

# Supporting Information

## Sb<sub>2</sub>O<sub>3</sub> Nanoparticles Anchored on Graphene Sheets via Alcohol Dissolution-Reprecipitation Method for Excellent Lithium Storage Properties

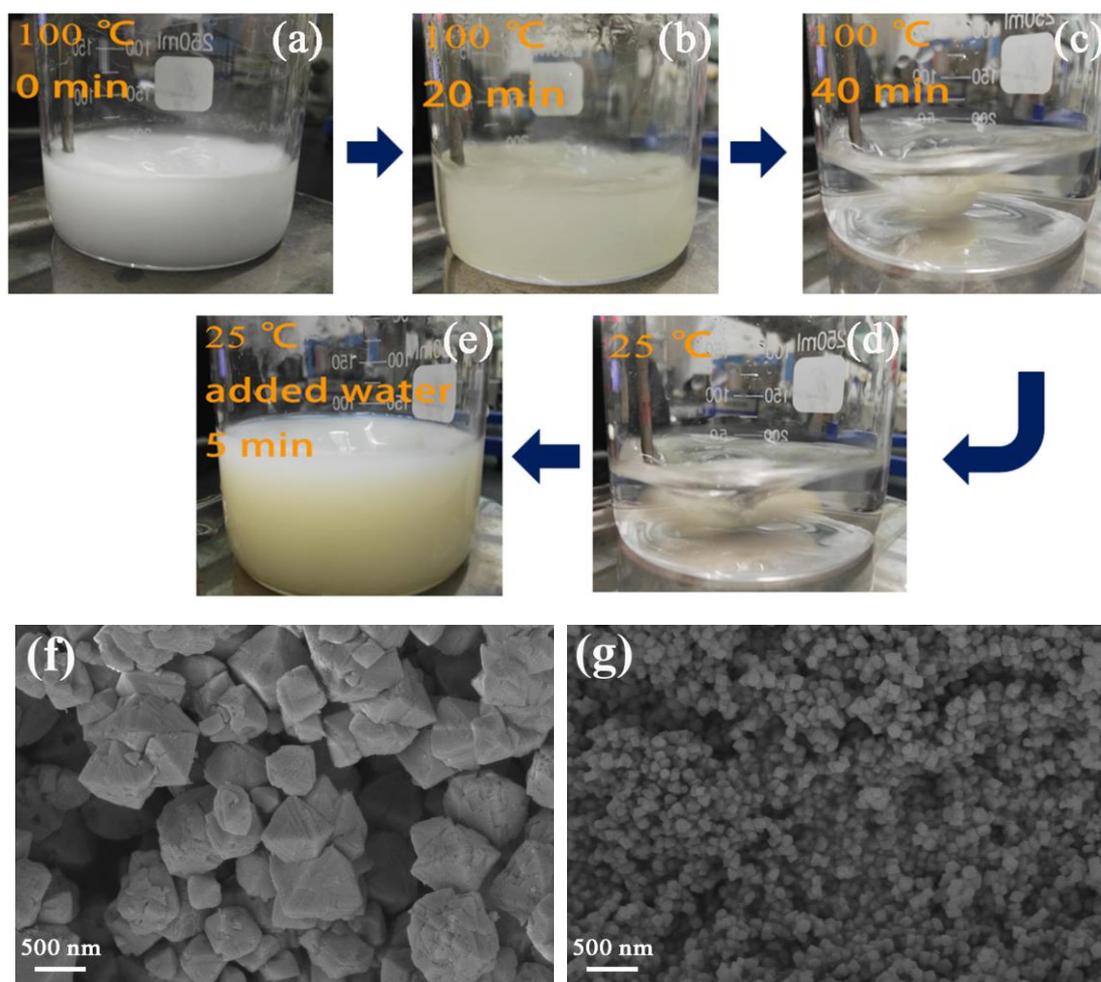
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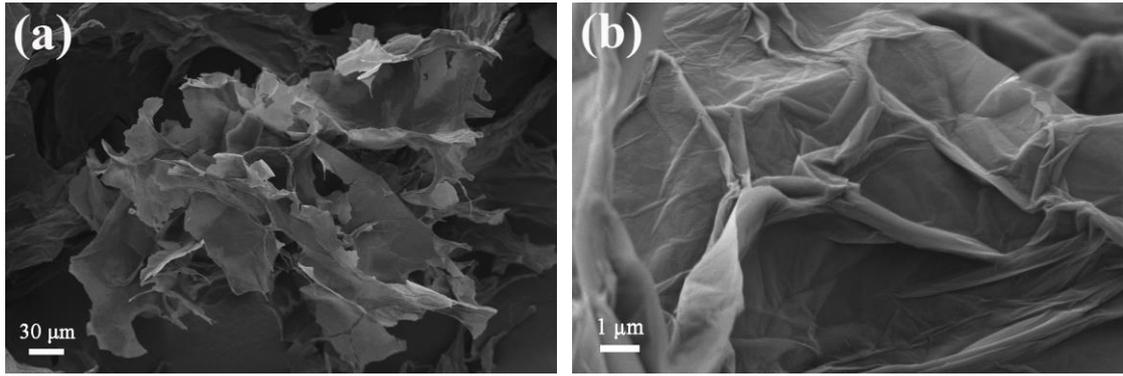
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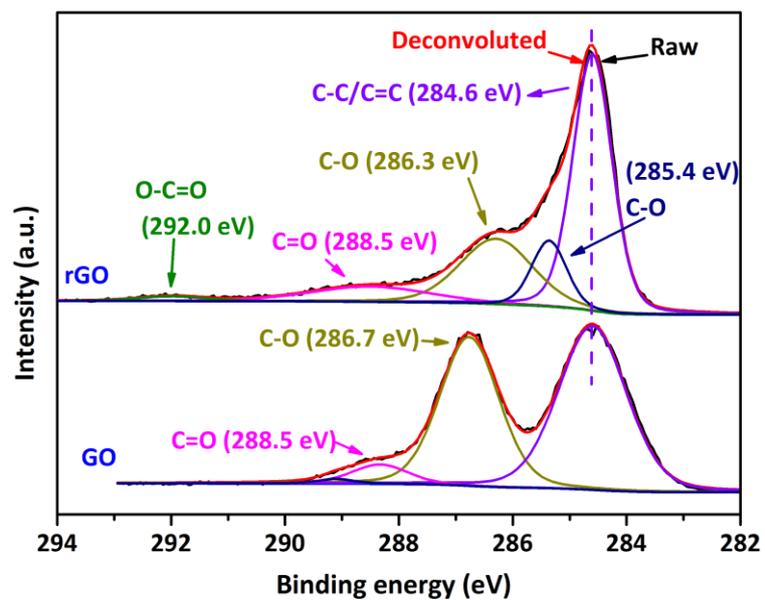
Ziqiang Lei, E-mail: leizq@nwnu.edu.cn, Tel/Fax: +86 931 7971261



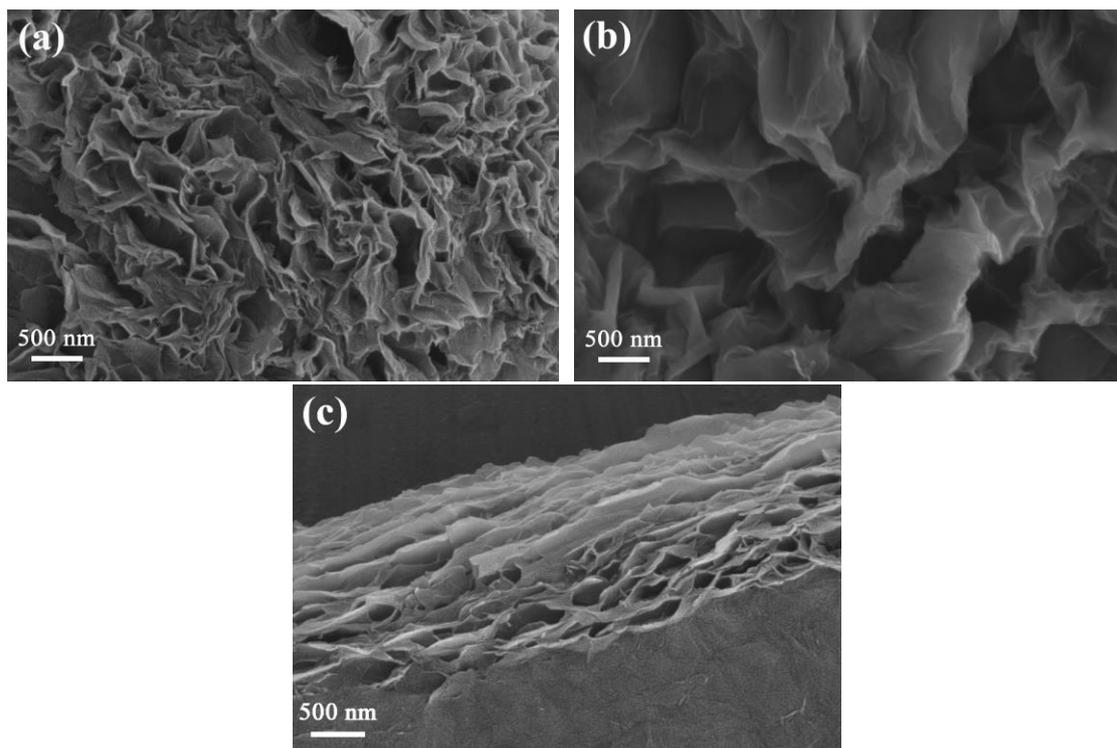
**Figure S1** Digital camera images for the procedure of dissolution and reprecipitation of  $\text{Sb}_2\text{O}_3$  in ethylene glycol (EG) solution.  $\text{Sb}_2\text{O}_3$  particles were dispersed by magnetic stirring at 100 °C to form a milk white suspension (a). As continuous stirring, the suspension gradually became transparent (b and c), due to the formation of ethylene glycol antimony (EG-Sb). After cooling to 25 °C, the solution maintained transparent on account of good chemical stability (d). The EG-Sb was hydrolyzed with addition of water and the  $\text{Sb}_2\text{O}_3$  was reprecipitation (e). SEM images for bulk  $\text{Sb}_2\text{O}_3$  obtained in its initial condition (f) and after the procedure of dissolution and reprecipitation in EG solution (g).



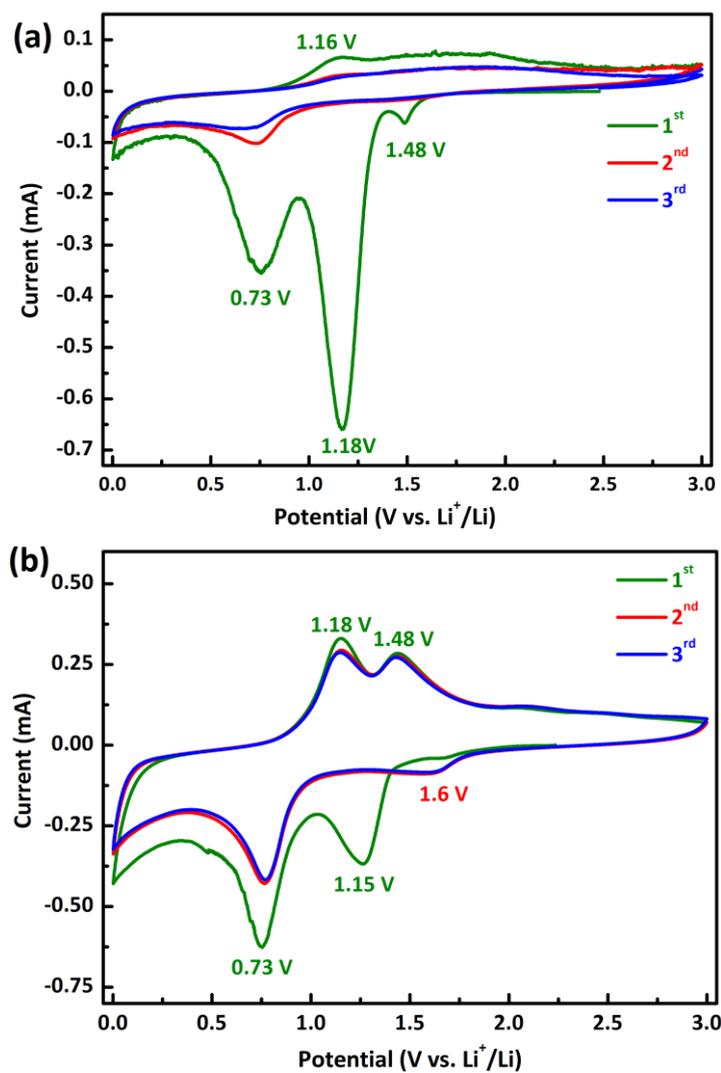
**Figure S2** SEM images with low (a) and high (b) magnification of the as-prepared GO.



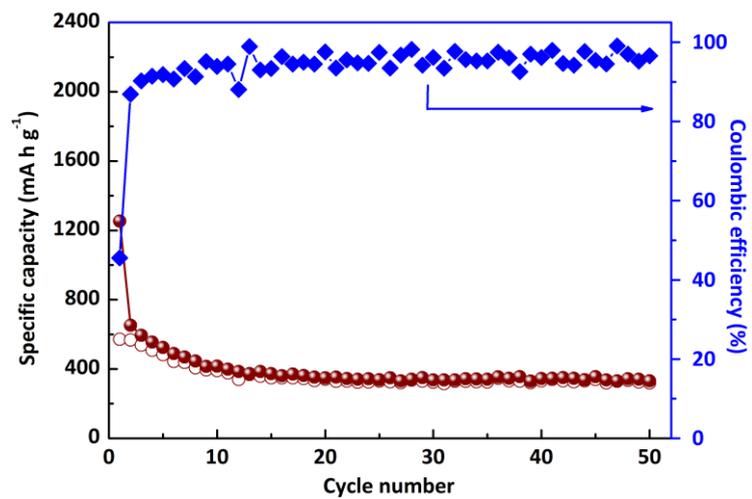
**Figure S3** High-resolution C1s XPS spectra of GO and rGO samples.



**Figure S4** SEM images for rGO obtained without the addition of L-AA (a), with a 24-hour stirring at room temperature before (b) and after (c) solvothermal treatment.



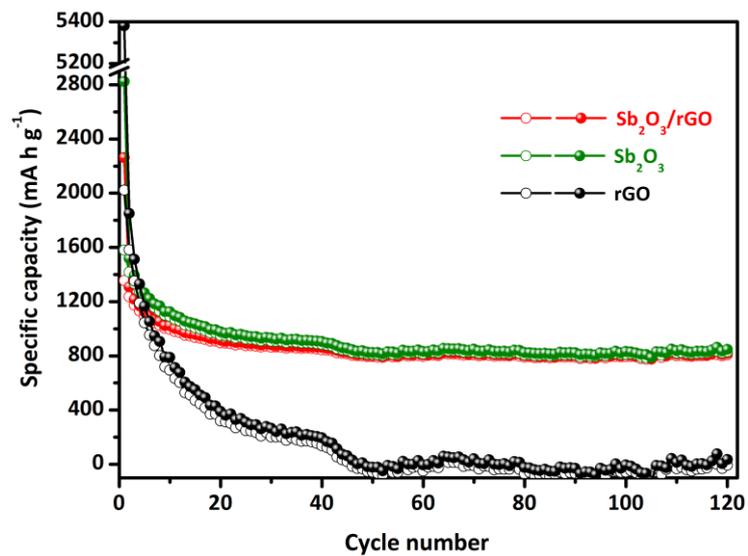
**Figure S5** Cyclic voltammograms of initial three cycles for bulk Sb<sub>2</sub>O<sub>3</sub> (a) and Sb<sub>2</sub>O<sub>3</sub>-rGO (b) electrodes at a scan rate of 0.2 mV s<sup>-1</sup>



**Figure S6** Cycling performance (left y-axis) and Coulombic efficiencies (right y-axis) of the bare rGO electrode at  $100 \text{ mA g}^{-1}$

**Table S1.** Summary of the  $\text{Sb}_x\text{O}_y$  based electrodes materials for LIB applications.

Materials	Initial reversible capacity ( $\text{mAh g}^{-1} / \text{A g}^{-1}$ )	ICE (%)	Capacity after (x) cycles ( $\text{mAh g}^{-1} / \text{cycles}$ )	Capacity at high rate ( $\text{mAh g}^{-1} / \text{A g}^{-1}$ )	Long-term cycling capacity after (x) cycles ( $\text{mAh g}^{-1} / \text{A g}^{-1} / \text{cycles}$ )	Potential (V)	Ref.
$\text{Sb}_2\text{O}_3$ thin films	794 / -	77.2	~750 / 70	-	-	0.01-3	1
$\text{Sb}_2\text{O}_3/\text{rGO}$	899 / 0.05	50.8	562 / 100	155 / 0.3	-	0.01-3	2
$\text{Sb}_6\text{O}_{13}/\text{rGO}$	1271 / 0.1	45.8	1109 / 140	201 / 3	430 / 0.5 / 300	0.01-3	3
$\text{Sb}_2\text{O}_4/\text{rGO}$	1170 / 0.1	53.8	798 / 200	320 / 3	428 / 0.55 / 500	0.01-3	4
hollow $\text{Sb}_2\text{O}_4$	727.1 / 0.1	67.1	700 / 50	370.9 / 2	415 / 1 / 100	0.01-3	5
$\text{Sb}_2\text{O}_3/\text{rGO}$	1355 / 0.1	60	808 / 120	188 / 5	525 / 0.6 / 700	0.001-3	This work



**Figure S7** Capacity contribution from  $Sb_2O_3$  or rGO during cycling in the  $Sb_2O_3/rGO$  nanocomposite, hypothesizing rGO contributed the fixed theoretical capacity of  $744 \text{ mA h g}^{-1}$  or  $Sb_2O_3$  contributed the fixed theoretical capacity of  $1109 \text{ mA h g}^{-1}$  during cycling, respectively.

## REFERENCES

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