N-A ky lated 1,4 -Diazabicyclo[2.2.2]octane–Polyethylene Glycol Melt as Deep Eutectic Solvent for the Synthesis of Fisher Indoles and 1*H*-T etrazoles

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Thermal Study of DABCO-ILs (2, 4)

Figure-S1a shows overlap of thermogravimetric (TGA) graph and its first derivative (DTG) for compound **2**, and Figure-S1b shows overlap of TGA and DSC graphs. The compound remains stable on heating up to 200 °C. In the first step there is only a slight loss in weight of about 16% in the short temperature range of 200-220 °C. As can be seen from the DSC graph, this is an endothermic process and is probably due to the loss of water molecules. In the second step, also an endothermic process, there is a weight loss of 24% in the temperature range 220-260 °C, most likely due to loss of alkyl chain. From the DTG, the loss of water molecules is observed at 215 °C, and the first decomposition temperature is calculated to be 235 °C. The remaining weight loss (about 40%) occurs in the third and last gradual step in the temperature range 280-360 °C, the second decomposition temperature calculated from DTG is 315 °C.

Figure-S1c shows overlap of thermogravimetric (TGA) graph and its first derivative (DTG) for compound **4**, and Figure-S1d shows the overlap of TGA and DSC graphs. In the first step there is a weight loss of about 30% in the temperature range from ambient to 100 °C, the corresponding DSC data in this range suggests an endothermic process associated with this change. This weight loss is most likely due to loss of water molecules. After initial loss of water, the weight remains stable at heating up to 215 °C, after which a sharp decrease in weight (28%) is observed, this is the first decomposition temperature at 255 °C as indicated by the DTG. According to DSC, this weight loss is also an endothermic process and is probably due to loss of alkyl chains. Immediately after this step, a sharp weight loss of about 40% is observed. This corresponds to a second decomposition temperature of 300 °C as indicated by the DTG. Similar to previous step, this was also an endothermic process as can be seen from the DSC graph.

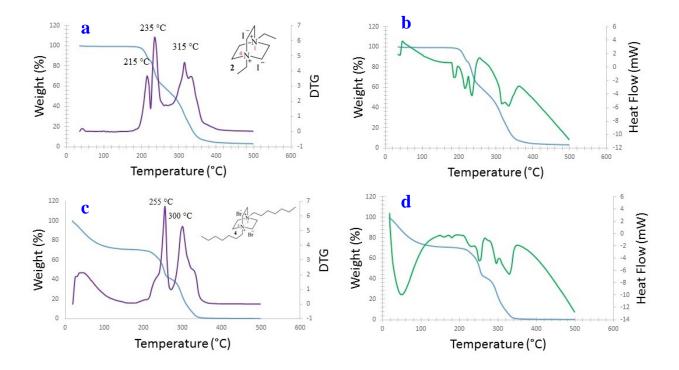


Figure-S1: a) Overlap of TGA (blue) and DTG (purple) graphs for compound **2**; b) Overlap of TGA (blue) and DSC (green) graphs for compound **2**; c) Overlap of TGA (blue) and DTG (purple) graphs for compound **4**; d) Overlap of TGA (blue) and DSC (green) graphs for compound **4**.

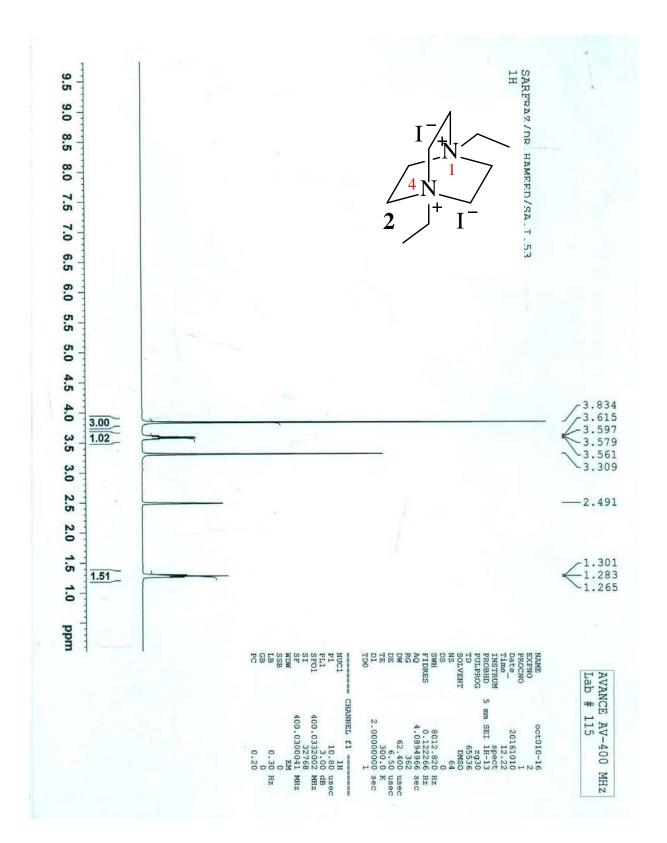


Figure-S2

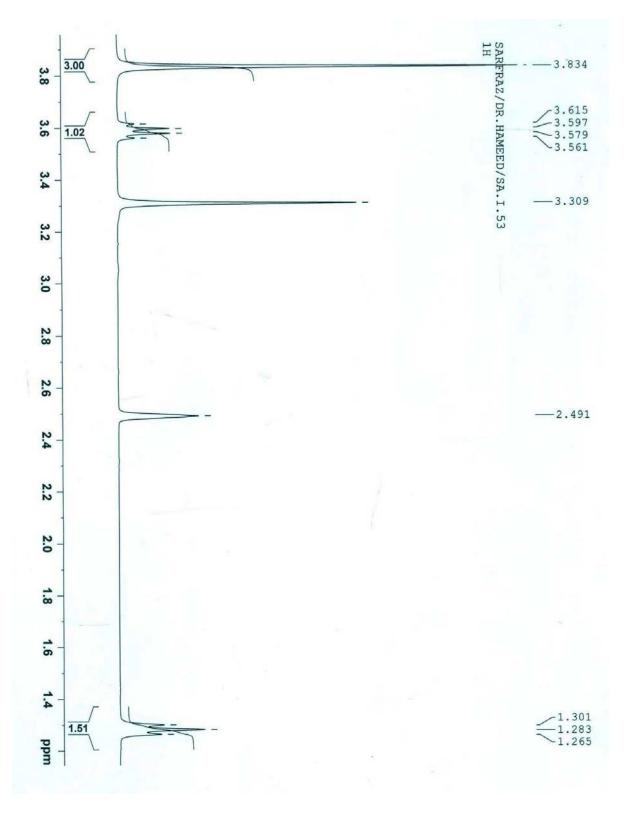


Figure-S2a

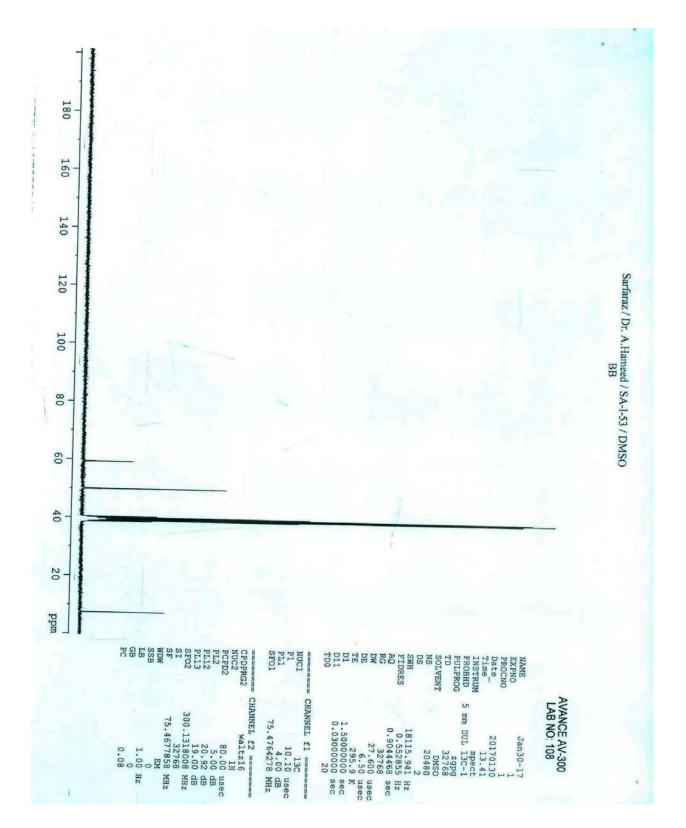


Figure-S3

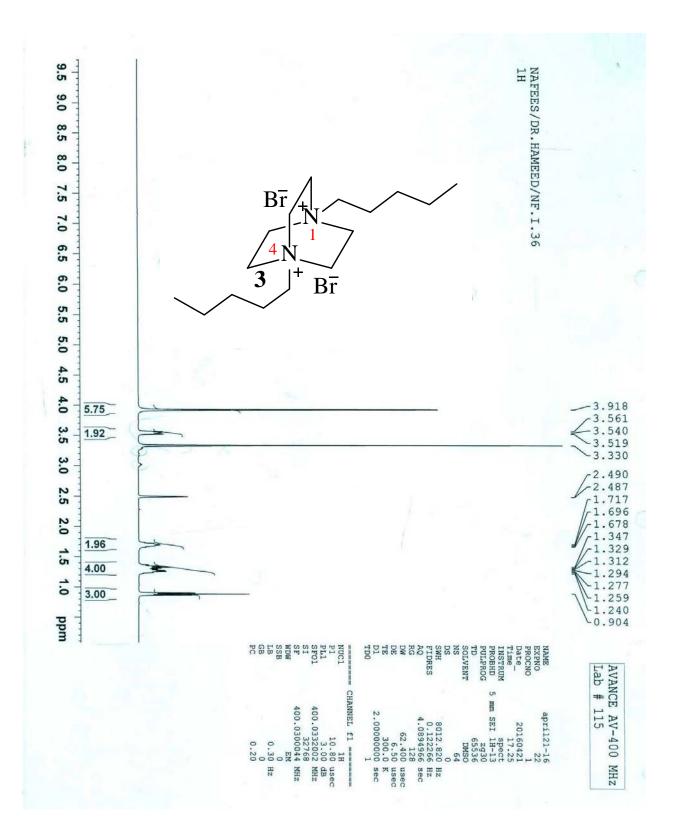


Figure-S4

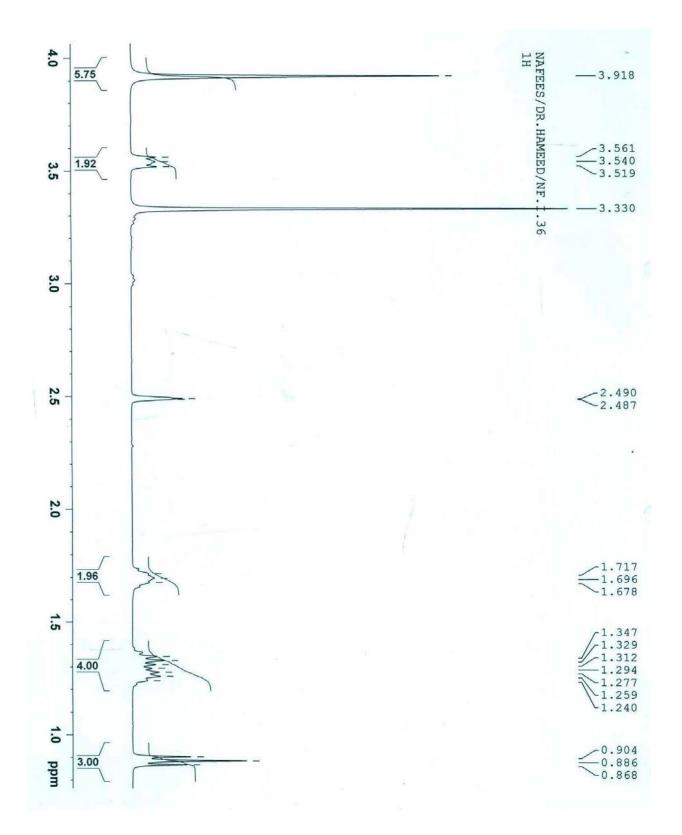


Figure-S4a

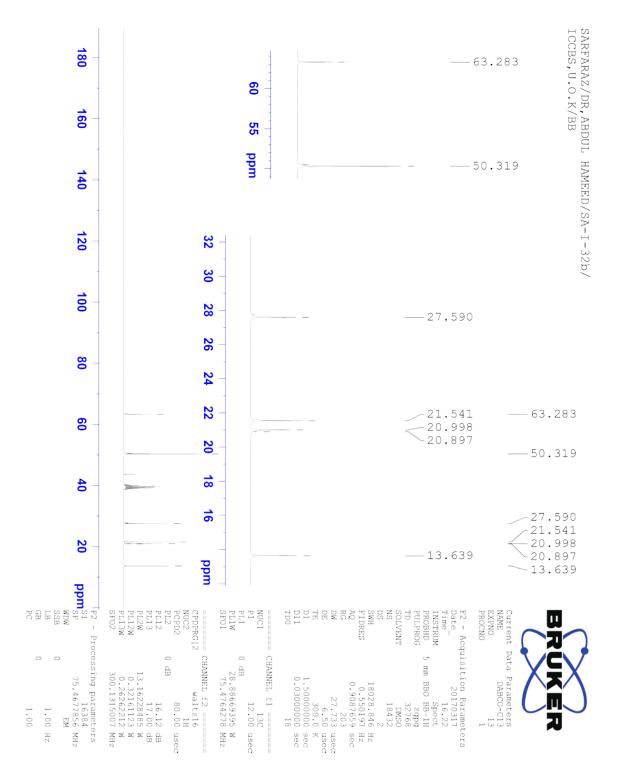


Figure-S5

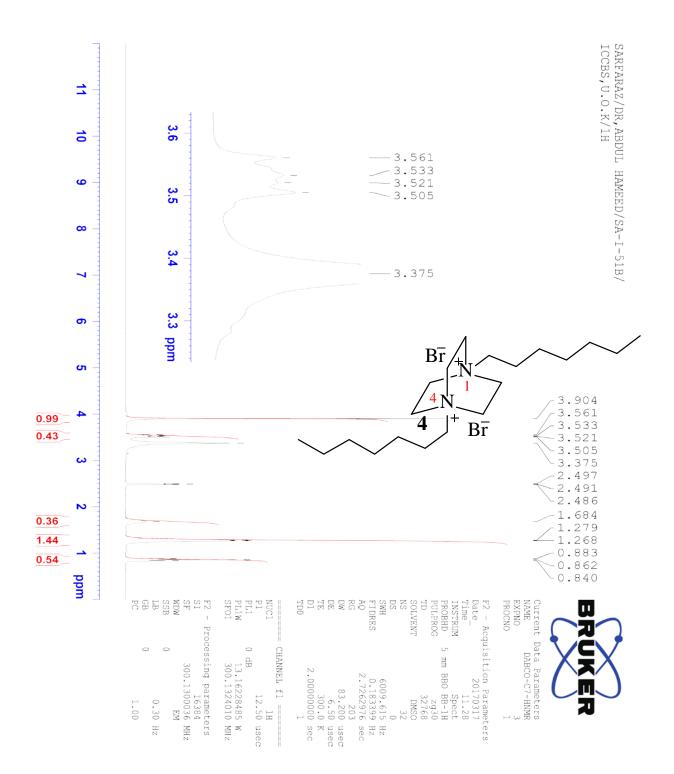


Figure-S6

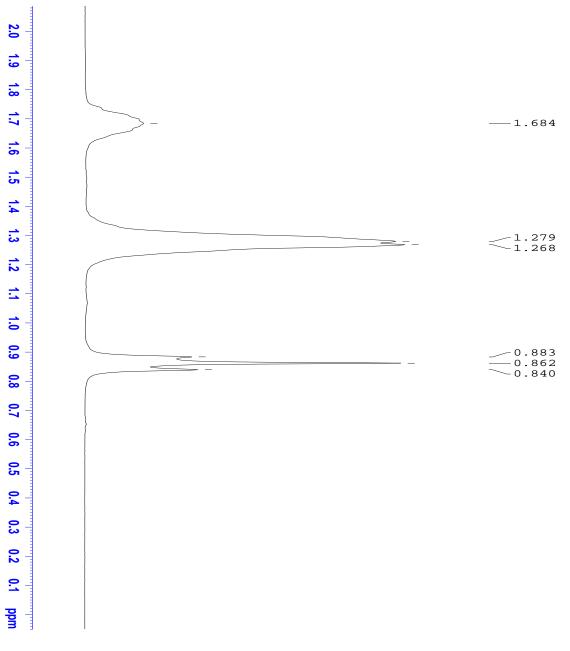


Figure-S6a

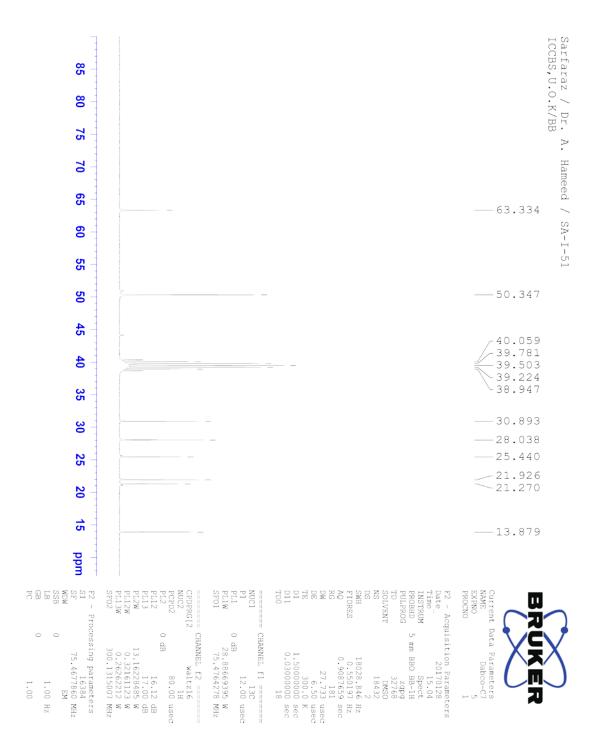


Figure-S7

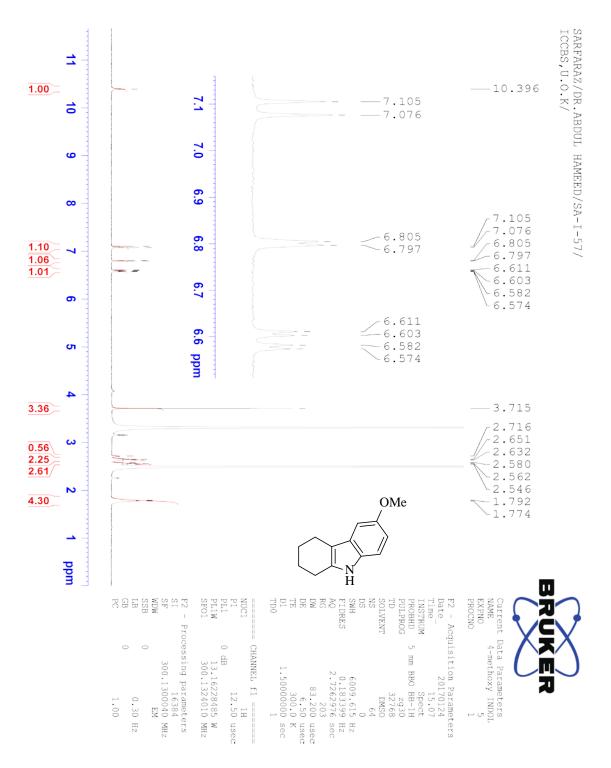


Figure-S8

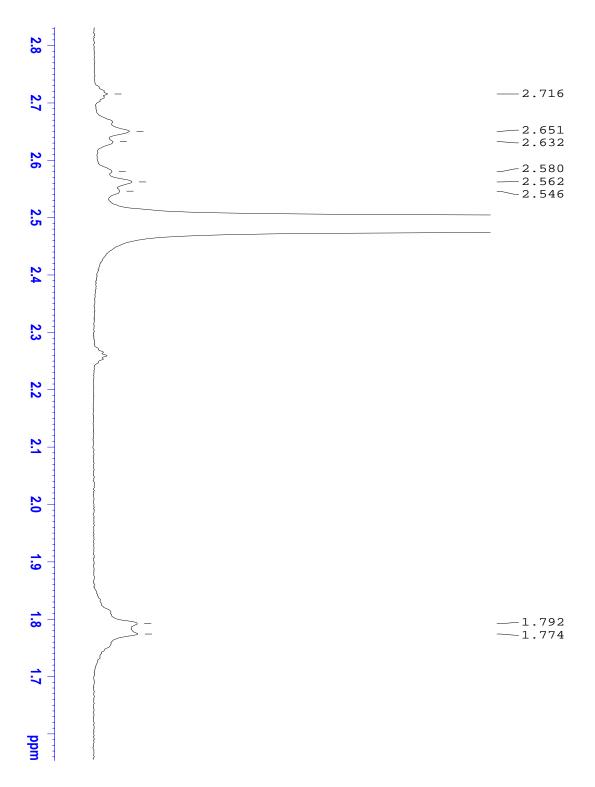


Figure-S8a

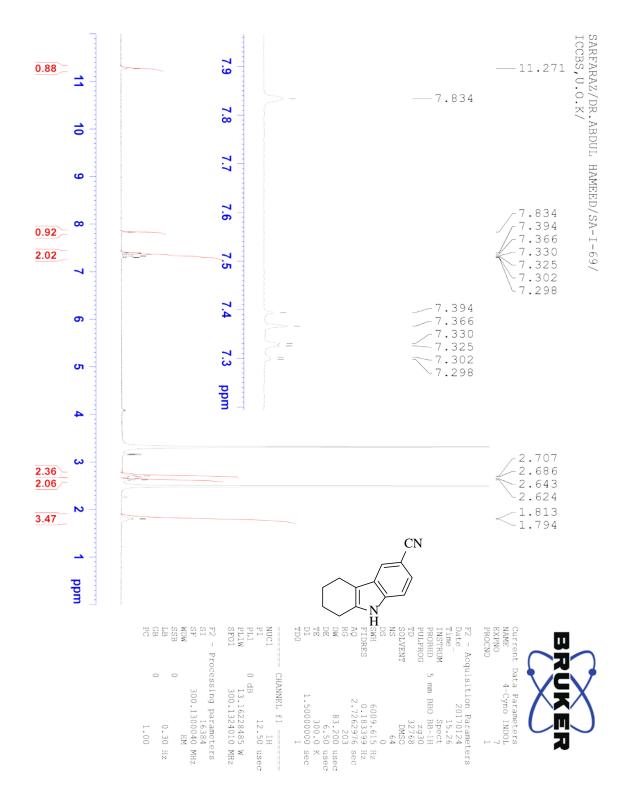


Figure-S9

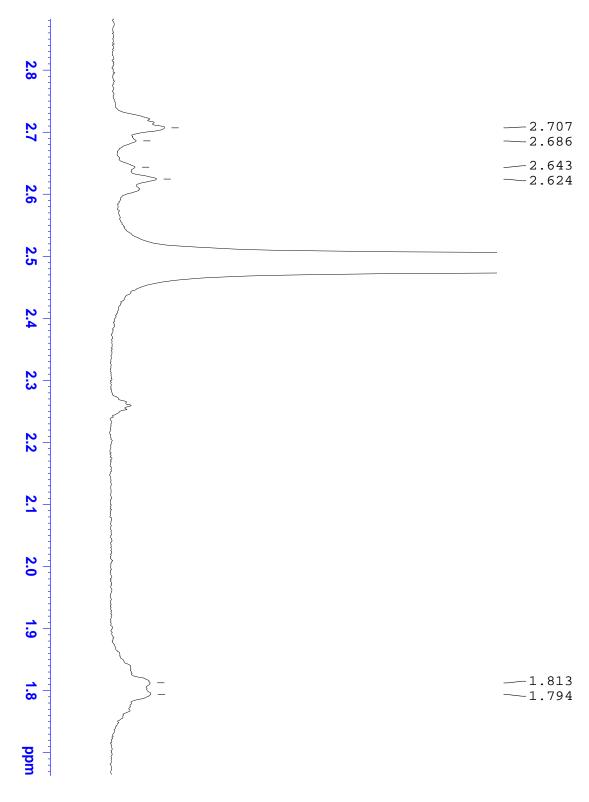


Figure-S9a

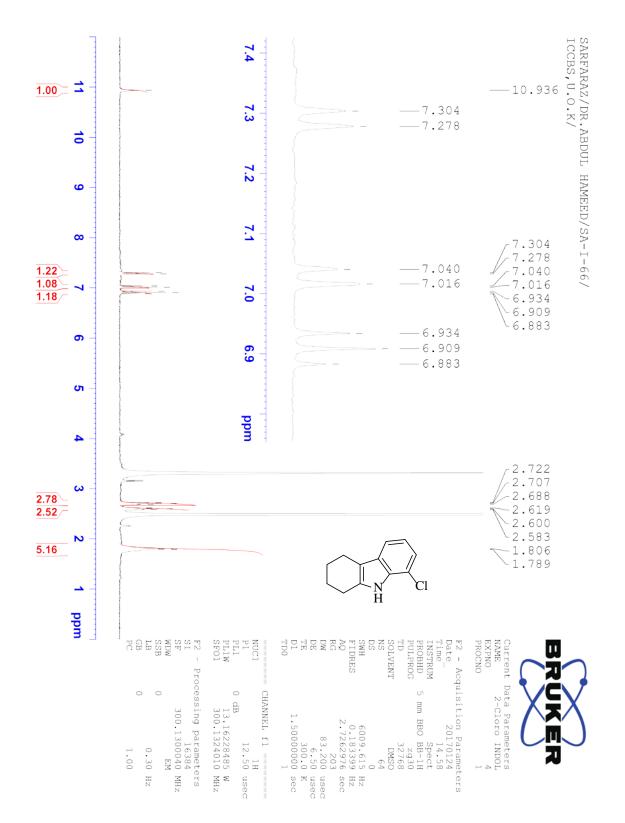


Figure-S10

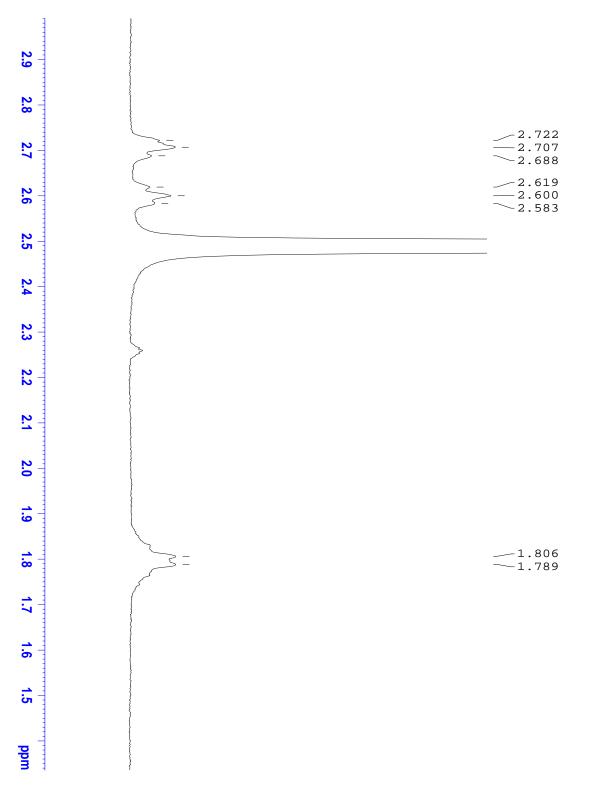


Figure-S10a

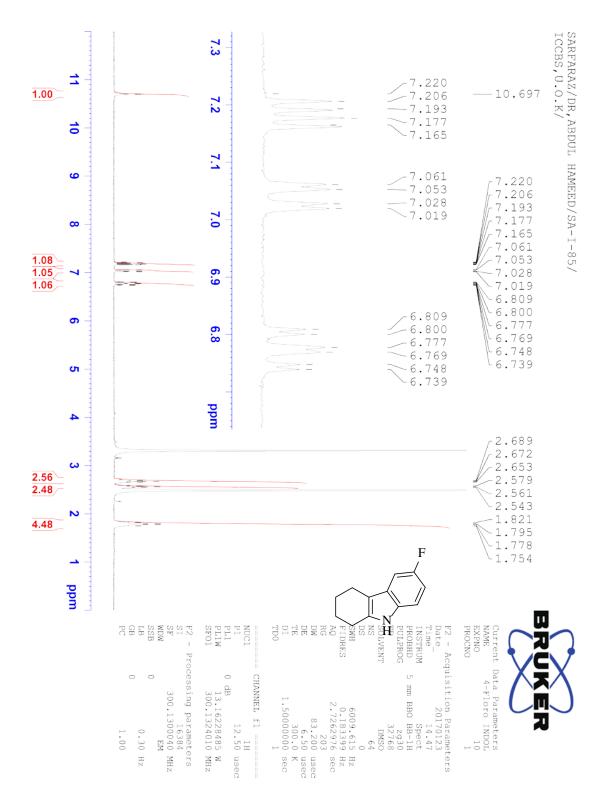


Figure-S11

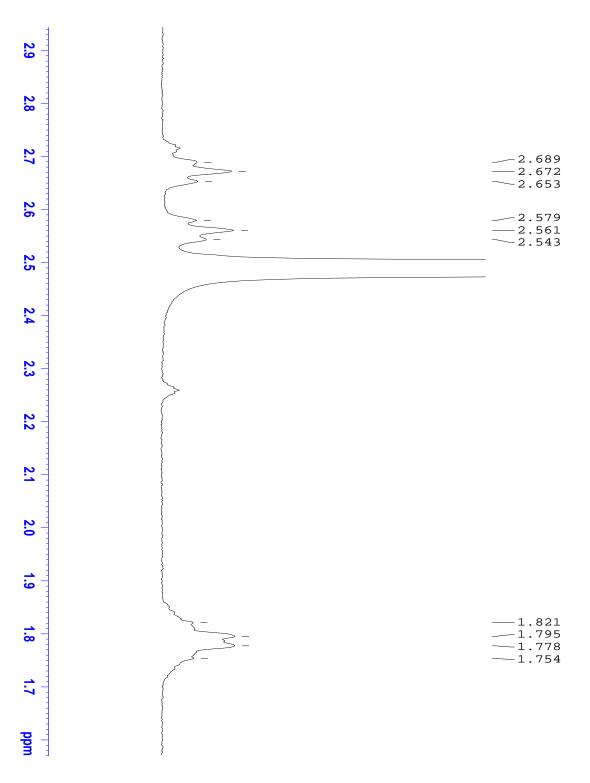


Figure-S11a

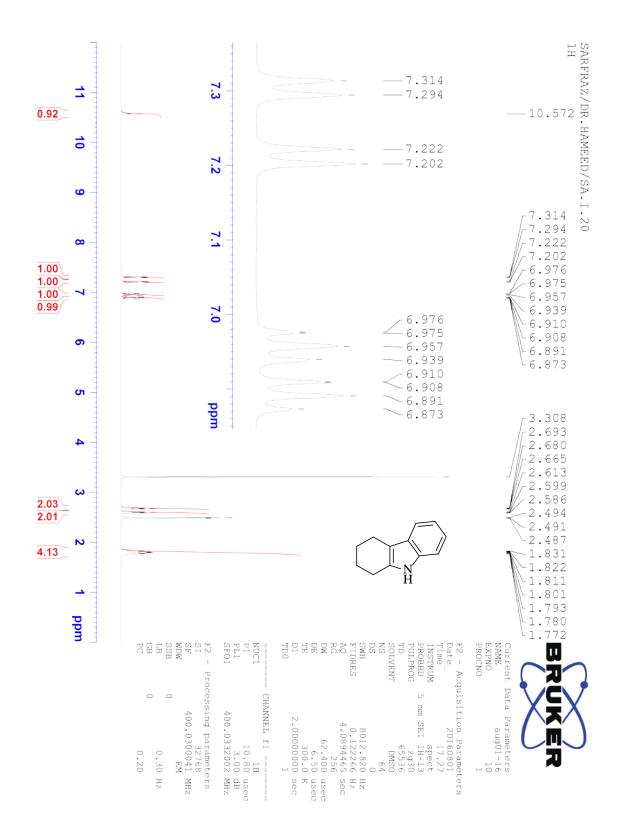


Figure-S12

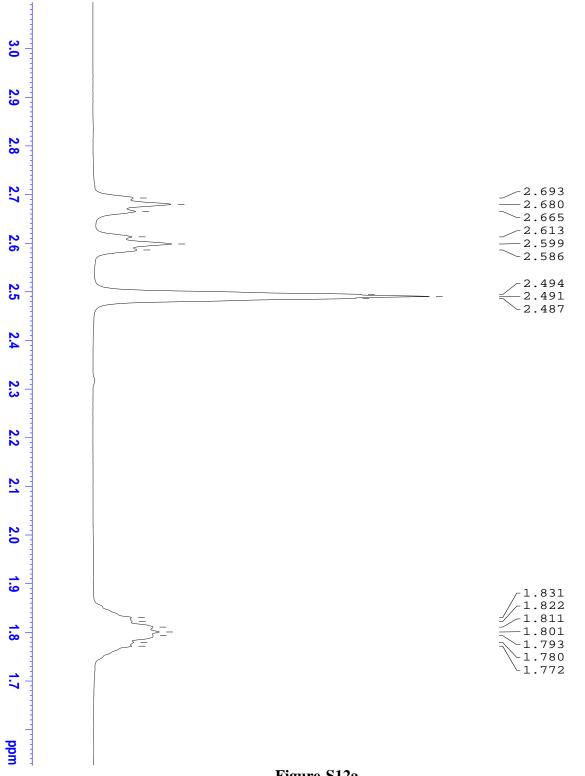


Figure-S12a

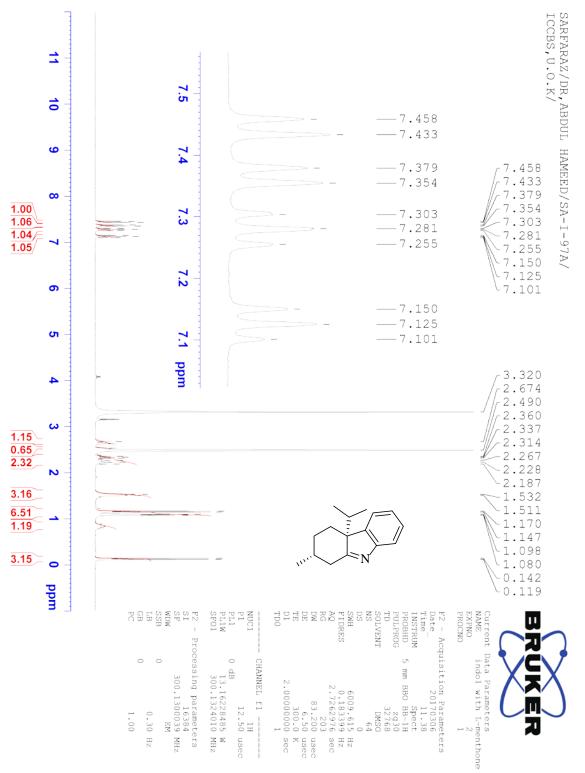
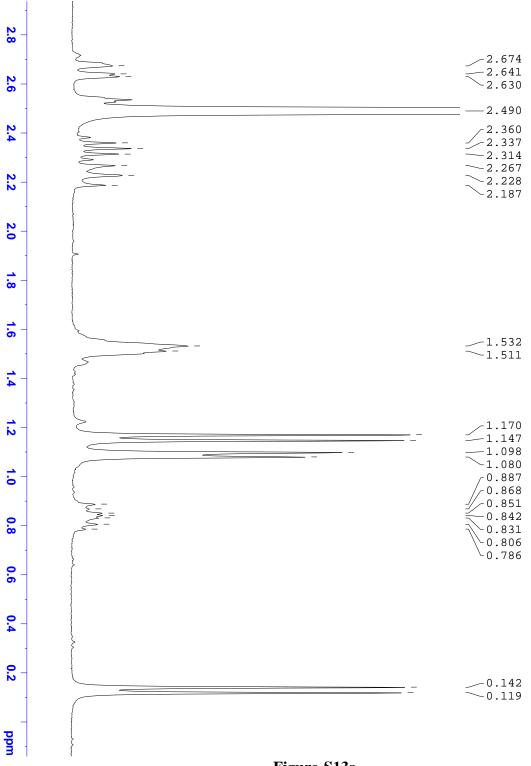
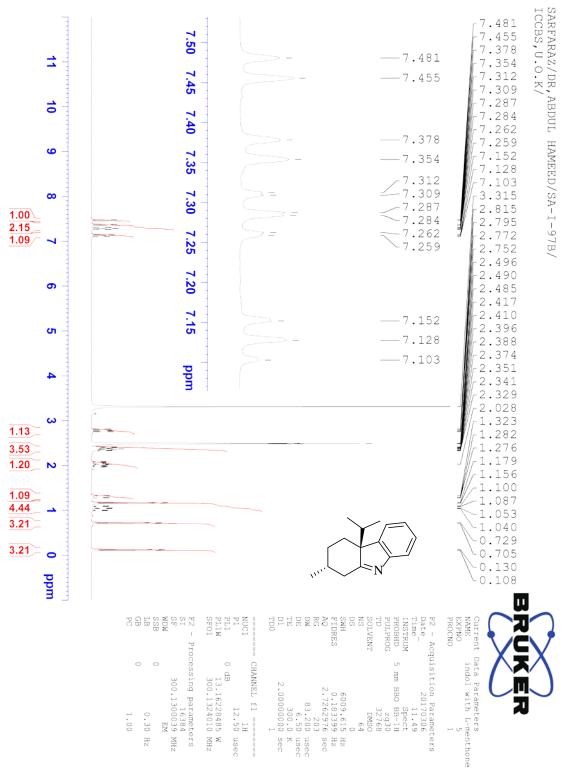


Figure-S13







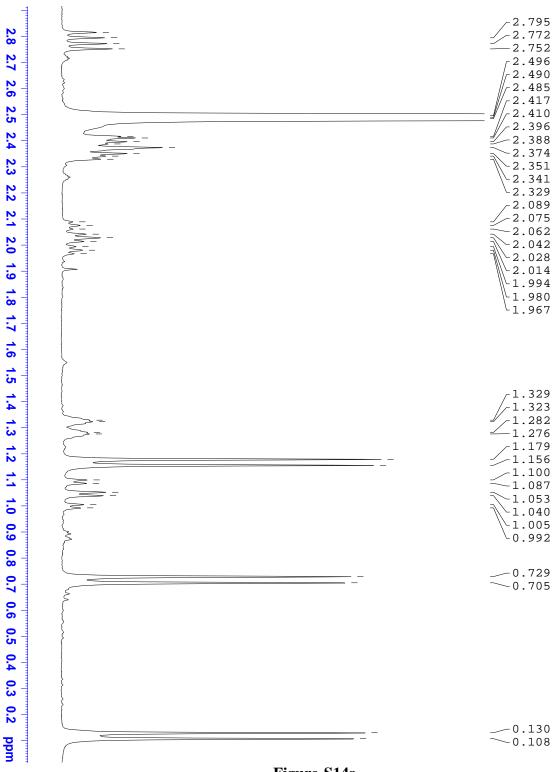


Figure-S14a

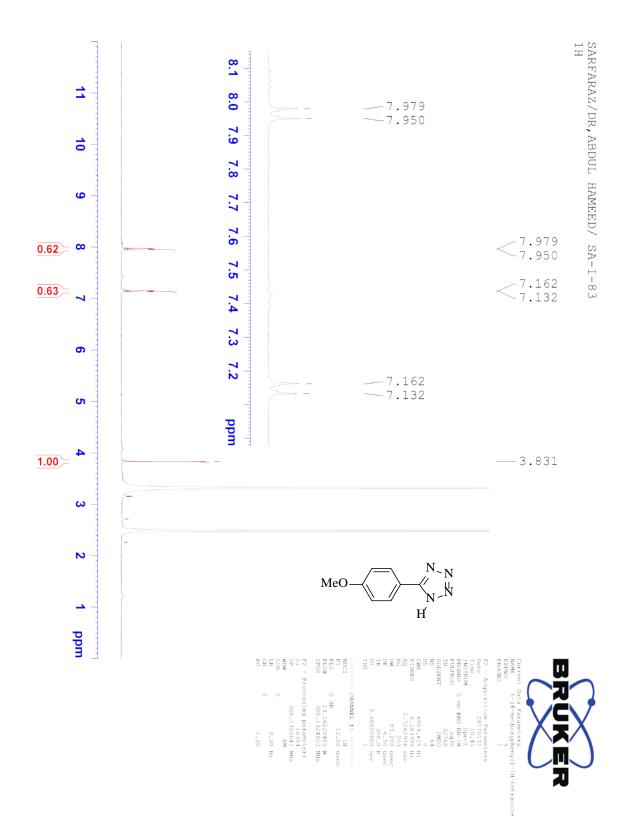


Figure-S15

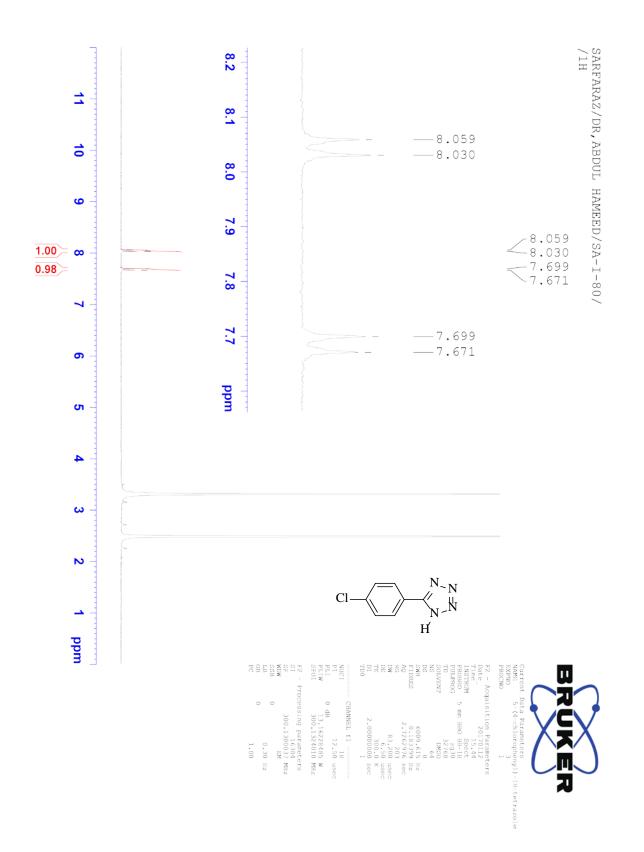


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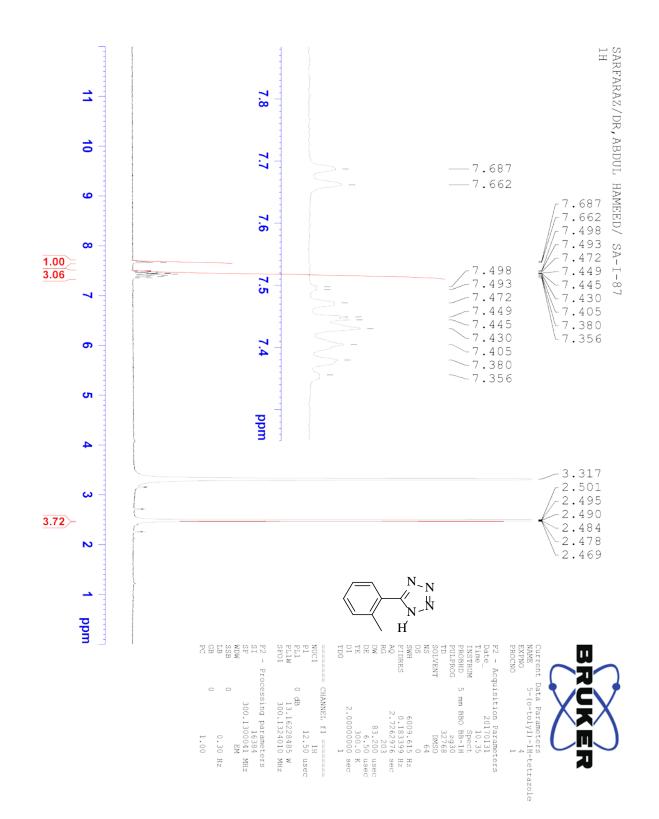
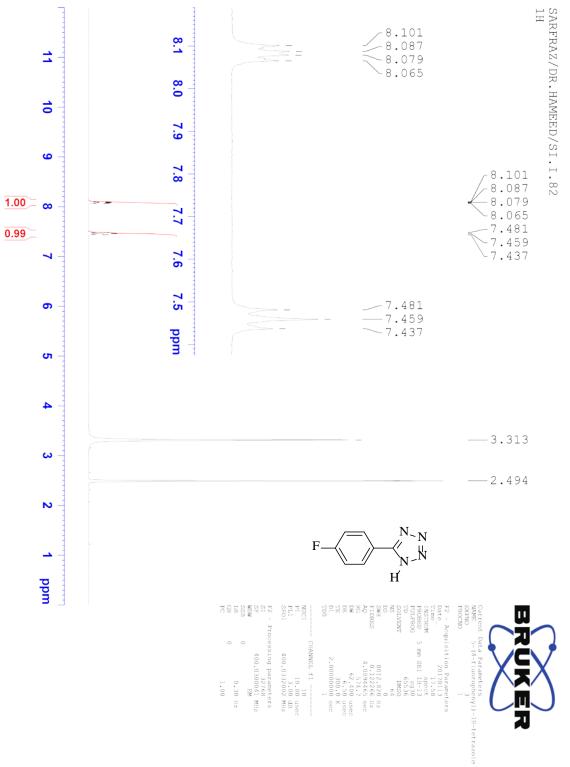


Figure-S17





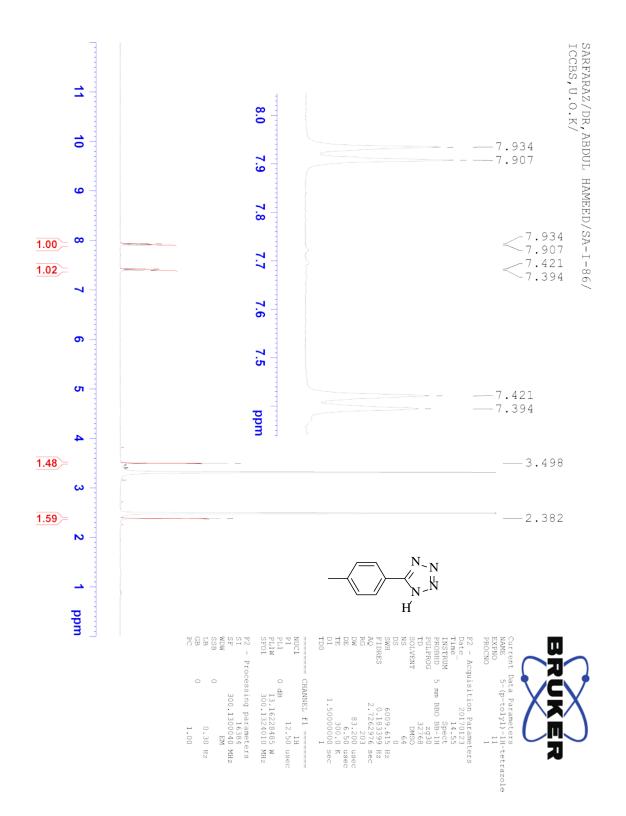


Figure-S19

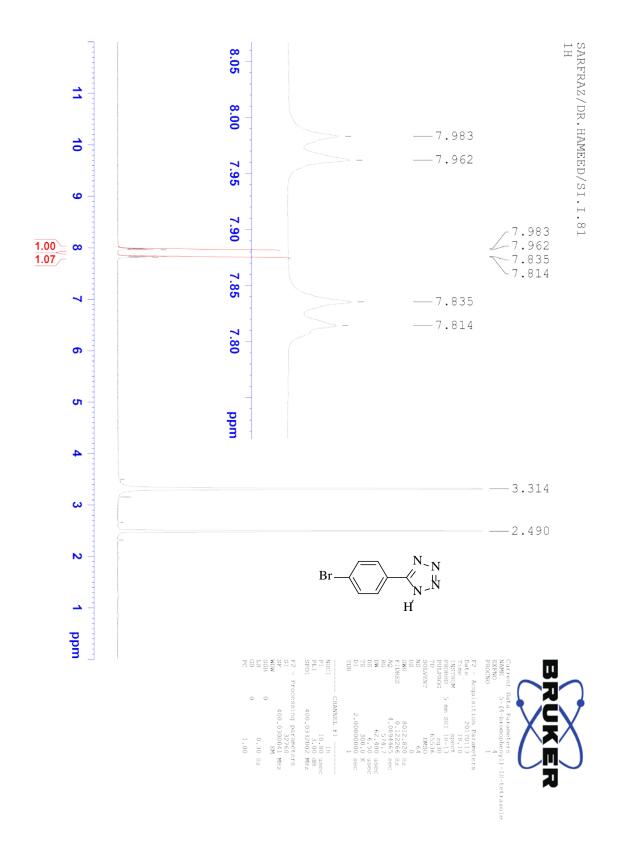
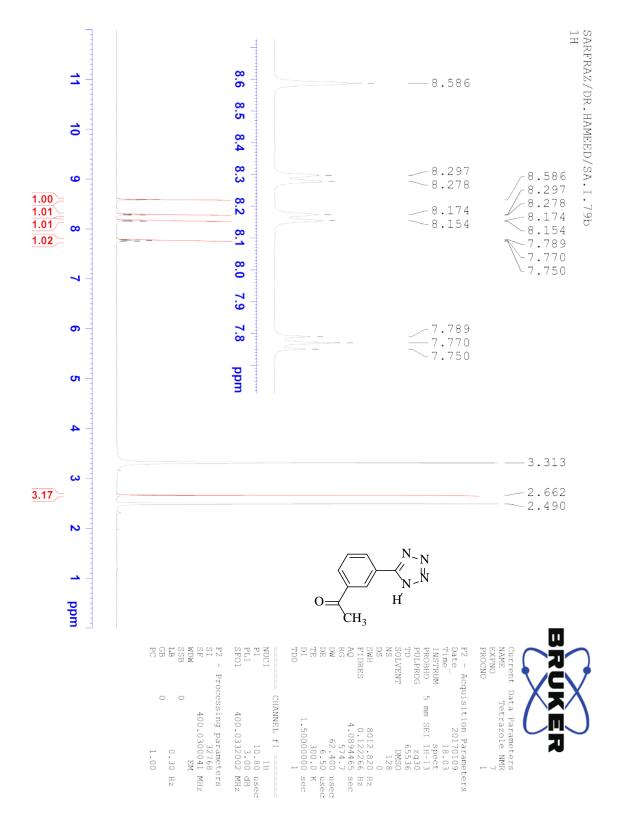


Figure-S20





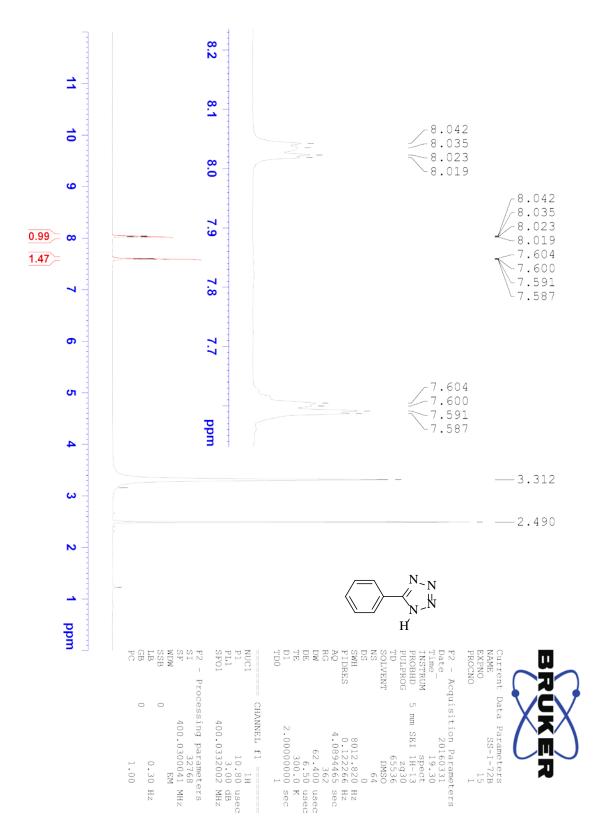


Figure-S22