

**TABLE 1S: Bond overlap populations,  $Q_{M-O}$ , between  $M^{z+}$  and O, effective charges,  $n$ , on  $M^{z+}$ , O and H obtained by the MO calculation of monohydrated ion clusters,  $M^{z+}\cdot H_2O$  and hydration numbers obtained experimentally,**

$N_{\text{hyd}}$ .

ion	$Q_{M-O}$	$n_M$	$n_O$	$n_H^a$	$N_{\text{hyd}}^b$
$Li^+$	0.385	+0.73	-0.30	+0.28	4.3
$Na^+$	0.273	+0.81	-0.37	+0.28	5.6
$K^+$	0.168	+0.89	-0.41	+0.26	5.5
$Rb^+$	0.147	+0.90	-0.43	+0.26	6.3 <sup>c</sup>
$Cs^+$	0.109	+0.92	-0.44	+0.26	7.0
$Ag^+$	0.311	+0.73	-0.31	+0.29	3.4
$Be^{2+}$	0.745	+1.15	-0.01	+0.43	4.0
$Mg^{2+}$	0.501	+1.48	-0.06	+0.29	6.0
$Ca^{2+}$	0.366	+1.66	-0.17	+0.25	6.3
$Sr^{2+}$	0.293	+1.73	-0.23	+0.25	8.1
$Ba^{2+}$	0.227	+1.79	-0.28	+0.24	8.1
$V^{2+}$	0.483	+1.39	+0.02	+0.30	6.0
$Cr^{2+}$	0.482	+1.30	+0.03	+0.33	6.0
$Mn^{2+}$	0.480	+1.43	-0.01	+0.29	6.0
$Fe^{2+}$	0.484	+1.36	+0.02	+0.31	6.0
$Co^{2+}$	0.497	+1.27	+0.03	+0.35	6.0
$Ni^{2+}$	0.486	+1.21	+0.05	+0.38	6.0
$Cu^{2+}$	0.450	+1.11	+0.08	+0.41	6.0
$Zn^{2+}$	0.488	+1.35	-0.03	+0.33	6.0

<sup>a</sup> Values for two H atoms in a  $H_2O$  were identical each other. <sup>b</sup> Mean of values reported in the reference 22 (at 25 °C). <sup>c</sup> The value assumed by the authors as a mean of values of  $K^+$  and  $Cs^+$ .

**TABLE 2S: Fundamental data used in the calculation of enthalpies due to the ligand field stabilization,  $H^\circ_{\text{LF}}$** 

$d^n$ <sup>a</sup>	ion	$N_t$ <sup>b</sup>	$N_e$ <sup>c</sup>	CFSE <sup>d</sup>	$10Dq$ <sup>e</sup>	$-H^\circ_{\text{LF}}$ <sup>f</sup>
		/ kJ mol <sup>-1</sup> / kJ mol <sup>-1</sup>				
divalent 3d elements						
$d^0$	$\text{Ca}^{2+}$	0	0	0Dq	0	0
$d^2$	$\text{Ti}^{2+}$	2	0	-8Dq	-	-
$d^3$	$\text{V}^{2+}$	3	0	-12Dq	14	169
$d^4$	$\text{Cr}^{2+}$	3	1	-6Dq	17	100
$d^5$	$\text{Mn}^{2+}$	3	2	0Dq	9	0
$d^6$	$\text{Fe}^{2+}$	4	2	-4Dq	12	48
$d^7$	$\text{Co}^{2+}$	5	2	-8Dq	12	96
$d^8$	$\text{Ni}^{2+}$	6	2	-12Dq	10	123
$d^9$	$\text{Cu}^{2+}$	6	3	-6Dq	16	93
$d^{10}$	$\text{Zn}^{2+}$	6	4	0Dq	0	0
trivalent 3d elements						
$d^0$	$\text{Sc}^{3+}$	0	0	0Dq	0	0
$d^1$	$\text{Ti}^{3+}$	1	0	-4Dq	24	97
$d^2$	$\text{V}^{3+}$	2	0	-8Dq	22	172
$d^3$	$\text{Cr}^{3+}$	3	0	-12Dq	21	253
$d^4$	$\text{Mn}^{3+}$	3	1	-6Dq	25	151
$d^5$	$\text{Fe}^{3+}$	3	2	0Dq	17	0
$d^{10}$	$\text{Ga}^{3+}$	6	4	0Dq	-	-

<sup>a</sup>  $d^n$  Electronic configuration of ion. <sup>b</sup> Number of electrons in  $t_{2g}$  orbital. <sup>c</sup> Number of electrons in  $e_g$  orbital. <sup>d</sup>

Stabilization estimated by the crystal field theory. <sup>e</sup>  $10Dq$  obtained experimentally.<sup>37</sup> <sup>f</sup> Ligand field stabilization enthalpy calculated by eq 5.