Table S1. Redox potentials of Zn(TPP) and **1b-3b** in CH_2Cl_2.[a]

Porphyrin	Potential (V *vs.* SCE)[b]				
	$E^{ox}_{1/2}$ (2)	$E^{ox}_{1/2}$ (1)	$E^{red}_{1/2}$ (1)	$E^{red}_{1/2}$ (2)	$\Delta E_{1/2}$(1)
Zn(TPP)	1.12	0.80	-1.40	-1.84	2.20
Zn(TPP(CF_3)$_4$) (**1b**)	1.24	0.92	-0.58	-0.81	1.50
Zn(TPPBr$_4$) (**2b**)	1.10	0.93	-1.08	-1.30	2.01
Zn(TPP(CH_3)$_4$) (**3b**)	0.83	0.71	-1.50	-	2.21

[a] Experimental conditions: [porphyrin]: 1 mM; [$TBAPF_6$]: 0.1 M; scan rate: 0.05 V/s; reference electrode: Ag wire. Potentials were determined by the internal standard of ferrocene/ferrocenium redox couple (0.46 V *vs.* SCE; CH_2Cl_2).

[b] $E^{ox}_{1/2}$ (n)and $E^{red}_{1/2}$ (n) are n th oxidation and reduction potential respectively. $\Delta E_{1/2}(1) = E^{ox}_{1/2}(1) - E^{red}_{1/2}(1)$; HOMO-LUMO gap.