

*Supporting Information*

**Tuning Task-Specific Ionic Liquids for the Extractive Desulfurization of Liquid Fuel**

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**Table S1** Names, abbreviations, sources, and densities of ionic liquids

#	Name	Abbreviation	Source	Density, g/mL (1 atm, 25 °C)
<b>Imidazolium-based ILs</b>				
1	1-butyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide	[BMIM][Tf <sub>2</sub> N]	Prepared by precipitation reaction <sup>1</sup> (see ESI <sup>2</sup> )	1.44 <sup>3</sup>
2	1-butyl-3-methylimidazolium bis(pentafluoroethylsulfonyl)imide	[BMIM][beti]	Prepared following ref. <sup>4</sup>	1.51 <sup>5</sup>
3	1-butyl-3-methylimidazolium tetrafluoroborate	[BMIM][BF <sub>4</sub> ]	Alfa Aesar (L19087), 98+%	1.20 <sup>6</sup>
4	1-butyl-3-methylimidazolium dicyanamide	[BMIM][dca]	Alfa Aesar (H59175), 97%	1.06 <sup>3</sup>
5	1-butyl-3-methylimidazolium trifluoromethanesulfonate	[BMIM][OTf]	Alfa Aesar (L19765), 99%	1.30 <sup>7</sup>
6	1-butyl-3-methylimidazolium thiocyanate	[BMIM][SCN]	Aldrich (724408), ≥95%	1.07 <sup>8</sup>
7	1-butyl-3-methylimidazolium nitrate	[BMIM][NO <sub>3</sub> ]	Prepared by anion-exchange (see ESI <sup>2</sup> )	1.16 <sup>9</sup>
8	1-butyl-3-methylimidazolium methylsulfate	[BMIM][MeSO <sub>4</sub> ]	Alfa Aesar (H27754), 99%	1.21 <sup>10</sup>
9	1-butyl-3-methylimidazolium dimethylphosphate	[BMIM][Me <sub>2</sub> PO <sub>4</sub> ]	Prepared following Refs. <sup>11, 12</sup>	1.16 <sup>13</sup>
10	1-butyl-3-methylimidazolium trifluoroacetate	[BMIM][CF <sub>3</sub> COO]	Prepared by anion-exchange (see ESI <sup>2</sup> )	1.21 <sup>14</sup>
11	1-butyl-3-methylimidazolium acetate	[BMIM][OAc]	Prepared by anion-exchange <sup>15</sup> (see ESI <sup>12</sup> )	1.05 <sup>16</sup>
12	1-butyl-3-methylimidazolium formate	[BMIM][HCOO]	Prepared by anion-exchange <sup>15</sup> (see ESI <sup>12</sup> )	0.94 <sup>a</sup>
13	1-butyl-3-methylimidazolium 1-butanesulfonate	[BMIM][BuSO <sub>3</sub> ]	Prepared following	1.26 <sup>a</sup>

			ref. <sup>4</sup>
14	1-butyl-3-methylimidazolium 2-(2-methoxyethoxy)ethyl sulfate	[BMIM][Me(OCH <sub>2</sub> CH <sub>2</sub> ) <sub>2</sub> SO <sub>4</sub> ]	Sigma-Aldrich (67421), ≥95.0%
15	1-ethyl-3-methylimidazolium hydrogensulfate	[EMIM][HSO <sub>4</sub> ]	Fluka (56486), ≥95%
16	1-ethyl-3-methylimidazolium acetate	[EMIM][OAc]	Prepared by anion-exchange <sup>15</sup> (see ESI <sup>12</sup> )
17	1-hexyl-3-methylimidazolium acetate	[HMIM][OAc]	Prepared by anion-exchange <sup>15</sup> (see ESI <sup>12</sup> )
18	1-hexyl-3-methylimidazolium hexafluorophosphate	[HMIM][PF <sub>6</sub> ]	Prepared following ref. <sup>4</sup>
19	1-methyl-3-octylimidazolium tetrafluoroborate	[OMIM][BF <sub>4</sub> ]	Prepared by anion-exchange (see ESI <sup>2</sup> )
20	1-benzyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide	[BzMIM][Tf <sub>2</sub> N]	Prepared by following ref. <sup>21</sup>
21	1-(cyclohexylmethyl)-3-methylimidazolium bis(trifluoromethylsulfonyl)imide	[ChxmMIm][Tf <sub>2</sub> N]	Prepared by following ref. <sup>21</sup>
<b>Pyridinium-based ILs</b>			
22	1-hexylpyridinium bis(trifluoromethylsulfonyl)imide	[C <sub>6</sub> Py][Tf <sub>2</sub> N]	Prepared following ref. <sup>23</sup>
23	1-hexylpyridinium thiocyanate	[C <sub>6</sub> Py][SCN]	Additional details provided below
24	1-hexylpyridinium dicyanamide	[C <sub>6</sub> Py][dca]	Prepared following ref. <sup>25</sup>
25	1-octylpyridinium bis(trifluoromethylsulfonyl)amide	[C <sub>8</sub> Py][Tf <sub>2</sub> N]	Prepared following ref. <sup>26</sup>
26	1-dodecylpyridinium bis(trifluoromethylsulfonyl)imide	[C <sub>12</sub> Py][Tf <sub>2</sub> N]	Prepared following ref. <sup>28</sup>
27	1-benzylpyridinium bis(trifluoromethylsulfonyl)imide	[BzPy][Tf <sub>2</sub> N]	Prepared by following ref. <sup>21</sup>
28	1-(cyclohexylmethyl)pyridinium bis(trifluoromethylsulfonyl)imide	[ChxmPy][Tf <sub>2</sub> N]	Prepared by following ref. <sup>21</sup>

***Ammonium-based ILs***

29	cholinium bis(trifluoromethylsulfonyl)imide	[Choline][Tf <sub>2</sub> N]	Prepared by precipitation reaction <sup>1</sup> (see ESI <sup>2</sup> )	1.53 <sup>a</sup>
30	tributylmethylammonium bis(trifluoromethylsulfonyl)imide	[N4441][Tf <sub>2</sub> N]	Prepared following ref. <sup>29</sup>	1.26 <sup>30</sup>
31	hexyltrimethylammonium bis(trifluoromethylsulfonyl)imide	[N6111][Tf <sub>2</sub> N]	Prepared following ref. <sup>29</sup>	1.31 <sup>29</sup>
32	methyltriocetylammnonium bis(trifluoromethylsulfonyl)imide	[N8881][Tf <sub>2</sub> N]	Prepared following ref. <sup>29</sup>	1.08 <sup>29</sup>
33	decytrimethylammonium bis(trifluoromethylsulfonyl)imide	[N10,111][Tf <sub>2</sub> N]	Prepared following ref. <sup>29</sup>	1.23 <sup>29</sup>
34	isopropyldimethylpropylammonium bis(trifluoromethylsulfonyl)imide	[C <sub>3</sub> Me <sub>2</sub> iPrN][Tf <sub>2</sub> N]	Prepared following ref. <sup>31</sup>	1.38 <sup>a</sup>
35	isopropyldimethyl(3-bromopropyl)ammonium bis(trifluoromethylsulfonyl)imide	[N11iPrC <sub>3</sub> Br][Tf <sub>2</sub> N]	Prepared following ref. <sup>31</sup>	1.47 <sup>a</sup>
36	( <i>n</i> -hexyl)isopropyldimethylammonium bis(trifluoromethylsulfonyl)imide	[C <sub>6</sub> Me <sub>2</sub> iPrN][Tf <sub>2</sub> N]	Prepared following ref. <sup>31</sup>	1.16 <sup>a</sup>
37	(3-chloro-2-hydroxypropyl)trimethylammonium bis(trifluoromethylsulfonyl)imide	[3Cl <sub>2</sub> OHPropN111][Tf <sub>2</sub> N]	Metathesis of commercial chloride salt (Sigma-Aldrich, 348287) using reported methods <sup>4</sup>	1.21 <sup>a</sup>
38	silver(I)/ <i>n</i> -propylamine bis(trifluoromethylsulfonyl)imide	[Ag( <i>n</i> -PrNH <sub>2</sub> ) <sub>2</sub> ][Tf <sub>2</sub> N]	Prepared following ref. <sup>32</sup>	1.46 <sup>a</sup>
39	silver(I)/isopropylamine bis(trifluoromethylsulfonyl)imide	[Ag( <i>i</i> -PrNH <sub>2</sub> ) <sub>2</sub> ][Tf <sub>2</sub> N]	Prepared following ref. <sup>32</sup>	1.61 <sup>a</sup>

***Pyrrolidinium-based ILs***

40	1-methyl-1-propylpyrrolidinium bis(trifluoromethylsulfonyl)imide	[C <sub>3</sub> MPyrr][Tf <sub>2</sub> N]	Prepared following ref. <sup>31</sup>	1.41 <sup>a</sup>
41	1-butyl-1-methylpyrrolidinium bis(trifluoromethylsulfonyl)imide	[C <sub>4</sub> MPyrr][Tf <sub>2</sub> N]	Prepared following ref. <sup>31</sup>	1.39 <sup>33</sup>
42	1-methyl-1-pentylpyrrolidinium	[C <sub>5</sub> MPyrr][Tf <sub>2</sub> N]	Prepared following	1.36 <sup>34</sup>

	bis(trifluoromethylsulfonyl)imide		ref. <sup>31</sup>	
43	1-hexyl-1-methylpyrrolidinium bis(trifluoromethylsulfonyl)imide	[C <sub>6</sub> MPyrr][Tf <sub>2</sub> N]	Prepared following ref. <sup>31</sup>	1.34 <sup>35</sup>
44	1-methyl-1-octylpyrrolidinium bis(trifluoromethylsulfonyl)imide	[C <sub>8</sub> MPyrr][Tf <sub>2</sub> N]	Prepared following ref. <sup>31</sup>	1.21 <sup>a</sup>
45	1-decyl-1-methylpyrrolidinium bis(trifluoromethylsulfonyl)imide	[C <sub>10</sub> MPyrr][Tf <sub>2</sub> N]	Prepared following ref. <sup>31</sup>	1.27 <sup>a</sup>
46	1-dodecyl-1-methylpyrrolidinium bis(trifluoromethylsulfonyl)imide	[C <sub>12</sub> MPyrr][beti]	Prepared following ref. <sup>31</sup>	— <sup>b</sup>
47	1-butyl-1-methylpyrrolidinium tetracyanoborate	[C <sub>4</sub> MPyrr][B(CN) <sub>4</sub> ] <sup>c</sup>	EMD Chemicals Inc., 99%	0.98 (20 °C), <sup>36</sup> 1.01 <sup>a</sup>
48	1-(cyclohexylmethyl)-1-methylpyrrolidinium bis(trifluoromethylsulfonyl)imide	[ChxmMPyrr][Tf <sub>2</sub> N]	Prepared by following ref. <sup>21</sup>	1.42 <sup>a</sup>
<b>Phosphonium-based ILs</b>				
49	trihexyltetradecylphosphonium bis(pentafluoroethylsulfonyl)imide	[P14,666][beti]	Prepared by ion exchange <sup>37</sup>	1.07 <sup>a</sup>
50	trihexyltetradecylphosphonium L-lactate	[P14,666][L-Lact]	Prepared by ion exchange <sup>37</sup>	— <sup>b</sup>
51	trihexyltetradecylphosphonium perfluoropentanoate	[P14,666][C <sub>4</sub> F <sub>9</sub> CO <sub>2</sub> ]	Prepared by ion exchange <sup>37</sup>	1.09 <sup>a</sup>
<b>Other types of ILs</b>				
52	1-butyl-1-methylpiperidinium bis(trifluoromethanesulfonyl)imide	[C <sub>4</sub> MPip][Tf <sub>2</sub> N]	Prepared following an earlier study <sup>4</sup>	1.38 <sup>38</sup>
53	Lithium(triglyme) bis(trifluoromethanesulfonyl)imide	[Li(G3)][Tf <sub>2</sub> N]	Prepared as reported previously <sup>39</sup>	1.51 <sup>a</sup>
54	1-octyl-1,4-diazabicyclo[2.2.2]octan-1-i um bis(trifluoromethylsulfonyl)imide	[DAB8][Tf <sub>2</sub> N]	Additional details provided below	— <sup>b</sup>
55	2,3,4,6,7,8-hexahydropyrrolo[1,2- <i>a</i> ]pyrimidinium acetate	[DBNH][OAc]	Prepared following an earlier study <sup>40</sup>	1.23 <sup>a</sup>
56	2,3,4,6,7,8-hexahydropyrrolo[1,2- <i>a</i> ]pyrimidinium trifluoroacetate	[DBNH][CF <sub>3</sub> COO]	Prepared following an earlier study <sup>40</sup>	1.30 <sup>a</sup>
57	<i>N,N,N',N'</i> -tetramethyl- <i>N'</i> -ethylguanidinium	[Me <sub>4</sub> Etguan][FAP]	EMD Chemicals	1.66 <sup>a</sup>

	tris(perfluoroethyl)trifluorophosphate	Inc., 98%	
<b>Deep eutectic solvents (DES)</b>			
58	Choline chloride/glycerol (1:2)	Choline chloride/glycerol (1:2)	Prepared following our earlier study <sup>41</sup> 1.18 (20 °C) <sup>42</sup>
59	Choline acetate/glycerol (1:1.5)	Choline acetate/glycerol (1:1.5)	Prepared following our earlier study <sup>41</sup> – <sup>b</sup>
<b>Alkoxy- or hydroxy-functionalized ILs</b>			
60	1-(2-hydroxyethyl)-3-methylimidazolium bis(trifluoromethanesulfonyl)imide	[HO-EMIM][Tf <sub>2</sub> N]	Prepared following ref. <sup>31</sup> 1.63 <sup>a</sup>
61	1-(2-methoxyethyl)-3-methylimidazolium bis(trifluoromethanesulfonyl)imide	[MeOCH <sub>2</sub> CH <sub>2</sub> -MIM][Tf <sub>2</sub> N]	Prepared following ref. <sup>31</sup> 1.59 <sup>a</sup>
62	1-ethyl-3-(2-(2-methoxyethoxy)ethoxy)ethylimidazolium bis(trifluoromethanesulfonyl)imide	[Me(OCH <sub>2</sub> CH <sub>2</sub> ) <sub>3</sub> -Et-Im][Tf <sub>2</sub> N]	Prepared in our earlier studies <sup>2, 12</sup> 1.20 <sup>a</sup>
63	1-butyl-3-(2-(2-methoxyethoxy)ethoxy)ethylimidazolium acetate	[Me(OCH <sub>2</sub> CH <sub>2</sub> ) <sub>3</sub> -Bu-Im][OAc]	Prepared in our earlier studies <sup>2, 12</sup> 1.15 <sup>a</sup>
64	1-ethyl-3-(2-(2-methoxypropyloxy)propyloxy)propylimidazolium acetate	[Me(OPr) <sub>3</sub> -Et-Im][OAc]	Prepared in our earlier studies <sup>2, 12</sup> 1.07 <sup>a</sup>
65	triethyl (2-(2-methoxyethoxy)ethoxy)ethylammonium acetate	[Me(OCH <sub>2</sub> CH <sub>2</sub> ) <sub>3</sub> -Et <sub>3</sub> N][OAc]	Prepared in our earlier studies <sup>2, 12</sup> 1.10 <sup>a</sup>
66	triethyl (2-(2-methoxyethoxy)ethoxy)ethylammonium bis(trifluoromethanesulfonyl)imide	[Me(OCH <sub>2</sub> CH <sub>2</sub> ) <sub>3</sub> -Et <sub>3</sub> N][Tf <sub>2</sub> N]	Prepared in our earlier studies <sup>2, 12</sup> 1.36 <sup>a</sup>
67	triethyl (2-(2-methoxyethoxy)ethoxy)ethylammonium formate	[Me(OCH <sub>2</sub> CH <sub>2</sub> ) <sub>3</sub> -Et <sub>3</sub> N][HCOO]	Prepared in our earlier studies <sup>2, 12</sup> 1.08 <sup>a</sup>
68	triethyl (2-(2-methoxyethoxy)ethoxy)ethylammonium bromide	[Me(OCH <sub>2</sub> CH <sub>2</sub> ) <sub>3</sub> -Et <sub>3</sub> N]Br	Prepared in our earlier studies <sup>2, 12</sup> 1.29 <sup>a</sup>
69	1-ethyl-3-(2-(2-methoxyethoxy)ethoxy)ethyl)piperidinium acetate	[Me(OCH <sub>2</sub> CH <sub>2</sub> ) <sub>3</sub> -Et-Pip][OAc]	Prepared in our earlier studies <sup>2, 12</sup> 1.08 <sup>a</sup>
70	1-ethyl-3-(2-(2-methoxyethoxy)ethoxy)ethyl)piperidinium bis(trifluoromethanesulfonyl)imide	[Me(OCH <sub>2</sub> CH <sub>2</sub> ) <sub>3</sub> -Et-Pip][Tf <sub>2</sub> N]	Prepared in our earlier studies <sup>2, 12</sup> 1.40 <sup>a</sup>

71	1-ethyl-3-(2-(2-methoxyethoxy)ethoxy)ethyl)piperidinium bromide	[Me(OCH <sub>2</sub> CH <sub>2</sub> ) <sub>3</sub> -Et-Pip]Br	Prepared in our earlier studies <sup>2, 12</sup>	1.23 <sup>a</sup>
72	(2-hydroxyethyl)dimethylammonium acetate	[Me <sub>2</sub> NH(CH <sub>2</sub> CH <sub>2</sub> OH)][OAc]	Bioniqs (York, UK), batch #07/10232	1.23 <sup>a</sup>
73	bis(2-methoxyethyl)ammonium acetate	[(CH <sub>3</sub> OCH <sub>2</sub> CH <sub>2</sub> ) <sub>2</sub> NH <sub>2</sub> ][OAc]	Bioniqs (York, UK), batch #07/10233	1.10 <sup>a</sup>

**Note:** <sup>a</sup> The density at 20 °C was estimated in this study following these procedures: a small vial containing IL was placed on a balance which was then zeroed; the vial was removed from the balance, and 50 µL of IL was withdrawn using a pipette; the vial was placed back to the balance to determine the mass loss due to IL removal; the density was calculated using the mass loss and the volume (50 µL). <sup>b</sup> The density cannot be determined by the method described above; in this case, the IL is too viscous and volumetric withdrawal using a micropipette is not reproducible. Therefore, the density was assumed to be 1.2 g mL<sup>-1</sup> for the purposes of calculating the Nernst sulfur partition coefficient. <sup>c</sup> This IL (**47**) tends to crystallize at room temperature over time; suggesting that its fluid state is a supercooled liquid.

#### Additional Synthetic Details:

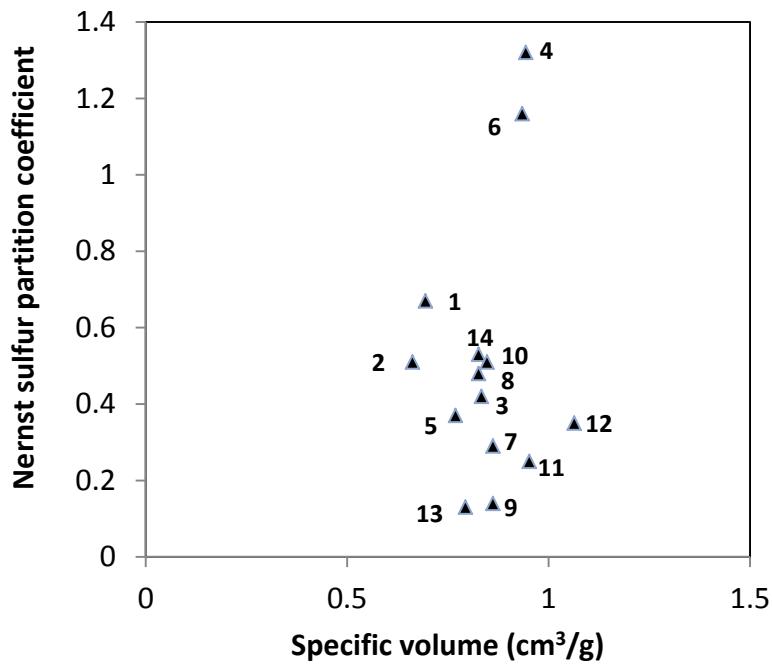
**1-hexylpyridinium thiocyanate, [C<sub>6</sub>Py][SCN]:** In a 100 mL round bottom flask, 5 g (20.5 mmol) of [C<sub>6</sub>Py][Br] was dissolved in 25 mL of dry acetone, followed by addition of 2.42 g (1.0 eq) of potassium thiocyanate (KSCN) which resulted in a white, milky suspension. After stirring the reaction for 2 days, the white suspension was allowed to settle overnight and then stored in a freezer for 24 h. The reaction was filtered via gravity filtration to remove NaBr. The filtrate was collected and the solvent removed by rotary evaporation to give a slightly yellow liquid.<sup>1</sup>H NMR (500 MHz, CDCl<sub>3</sub>): δ 9.57 (d, 2H), 8.52 (t, 1H), 8.14 (t, 2H), 4.99 (t, 2H), 2.03 (m, 2H), 1.33 (m, 2H), 1.27 (m, 4H), 0.82 (t, 3H).

**1-octyl-1,4-diazabicyclo[2.2.2]octan-1-ium bis(trifluoromethylsulfonyl)imide, [DAB8][Tf<sub>2</sub>N]:** In a 100 mL round bottom flask, 5 g (16.3 mmol) of [DAB8][Br] was dissolved in 25 mL of deionized water, followed by addition of 4.70 g (1.0 eq) of lithium bis(trifluoromethylsulfonyl)imide (LiTf<sub>2</sub>N) which resulted in the formation of a dense colorless bottom liquid phase. The reaction was washed with water five times and dried under vacuum for 2 days. <sup>1</sup>H NMR (500 MHz, D<sub>2</sub>O): δ 3.21 (m, 14 H), 1.70 (m, 2H), 1.30 (m, 10 H), 0.87 (t, 3H).

**Table S2** Miscibility between selected ILs and *n*-octane or *n*-dodecane

#	IL	<i>n</i> -octane	<i>n</i> -dodecane
4	[BMIM][dca]	Im	Im
6	[BMIM][SCN]	Im	Im
22	[C <sub>6</sub> Py][Tf <sub>2</sub> N]	Im	Im
23	[C <sub>6</sub> Py][SCN]	Im	Im
25	[C <sub>8</sub> Py][Tf <sub>2</sub> N]	Im	Im
26	[C <sub>12</sub> Py][Tf <sub>2</sub> N]	Im	Im
29	[Choline][Tf <sub>2</sub> N]	Im	M
30	[N4441][Tf <sub>2</sub> N]	Im	M
31	[N6111][Tf <sub>2</sub> N]	Im	M
32	[N8881][Tf <sub>2</sub> N]	Im	M
33	[N10,111][Tf <sub>2</sub> N]	Im	M
38	[Ag( <i>n</i> -PrNH <sub>2</sub> ) <sub>2</sub> ][Tf <sub>2</sub> N]	Im	M
39	[Ag( <i>i</i> -PrNH <sub>2</sub> ) <sub>2</sub> ][Tf <sub>2</sub> N]	Im	M
40	[C <sub>3</sub> MPyrr][Tf <sub>2</sub> N]	Im	M
41	[C <sub>4</sub> MPyrr][Tf <sub>2</sub> N]	Im	M
42	[C <sub>5</sub> MPyrr][Tf <sub>2</sub> N]	Im	M
43	[C <sub>6</sub> MPyrr][Tf <sub>2</sub> N]	Im	M
44	[C <sub>8</sub> MPyrr][Tf <sub>2</sub> N]	Im	M
45	[C <sub>10</sub> MPyrr][Tf <sub>2</sub> N]	Im	M
46	[C <sub>12</sub> MPyrr][beti]	Im	M
47	[C <sub>4</sub> MPyrr][B(CN) <sub>4</sub> ]	Im	Im
48	[ChxmMPyrr][Tf <sub>2</sub> N]	Im	Im
49	[P14,666][beti]	M	M
50	[P14,666][L-Lact]	M	M
51	[P14,666][C <sub>4</sub> F <sub>9</sub> CO <sub>2</sub> ]	M	M
64	[Me(OPr) <sub>3</sub> -Et-Im][OAc]	Im	Im
67	[Me(OCH <sub>2</sub> CH <sub>2</sub> ) <sub>3</sub> -Et <sub>3</sub> N][HCOO]	Im	Im
68	[Me(OCH <sub>2</sub> CH <sub>2</sub> ) <sub>3</sub> -Et <sub>3</sub> N]Br	Im	Im
73	[(CH <sub>3</sub> OCH <sub>2</sub> CH <sub>2</sub> ) <sub>2</sub> NH <sub>2</sub> ][OAc]	Im	Im

**Note:** Im = immiscible, M = miscible.



**Figure S1.** Correlation between the DBT partition coefficient and the specific volume of the ILs (specific volume = 1/density).

## References

1. Bonhote, P.; Dias, A.-P.; Michel, A.; Papageorgiou, N.; Kalyanasundaram, K.; Gratzel, M. Hydrophobic, highly conductive ambient-temperature molten salts. *Inorganic Chemistry* **1996**, *35*, 1168-1178.
2. Zhao, H.; Baker, G. A.; Song, Z.; Olubajo, O.; Crittle, T.; Peters, D. Designing enzyme-compatible ionic liquids that can dissolve carbohydrates. *Green Chem.* **2008**, *10*, 696-705.
3. de Castro, C. A. N.; Langa, E.; Morais, A. L.; Lopes, M. L. M.; Lourenço, M. J. V.; Santos, F. J. V.; Santos, M. S. C. S.; Lopes, J. N. C.; Veiga, H. I. M.; Macatrão, M.; Esperança, J. M. S. S.; Marques, C. S.; Rebelo, L. P. N.; Afonso, C. A. M. Studies on the density, heat capacity, surface tension and infinite dilution diffusion with the ionic liquids [C<sub>4</sub>mim][NTf<sub>2</sub>], [C<sub>4</sub>mim][dca], [C<sub>2</sub>mim][EtOSO<sub>3</sub>] and [Aliquat][dca]. *Fluid Phase Equilibria* **2010**, *294*, 157-179.
4. Burrell, A. K.; Sesto, R. E. D.; Baker, S. N.; McCleskey, T. M.; Baker, G. A. The large scale synthesis of pure imidazolium and pyrrolidinium ionic liquids. *Green Chemistry* **2007**, *9*, 449-454.
5. Tokuda, H.; Tsuzuki, S.; Abu Bin Hasan Susan, M.; Hayamizu, K.; Watanabe, M. How ionic are room-temperature ionic liquids? An indicator of the physicochemical properties. *Journal of Physical Chemistry B* **2006**, *110*, 19593-19600.
6. Afzal, W.; Liu, X.; Prausnitz, J. M. Solubilities of some gases in four imidazolium-based ionic liquids. *Journal of Chemical Thermodynamics* **2013**, *63*, 88-94.
7. Klomfar, J.; Součková, M.; Pátek, J. Temperature dependence measurements of the density at 0.1 MPa for 1-alkyl-3-methylimidazolium-based ionic liquids with the trifluoromethanesulfonate and tetrafluoroborate anion. *Journal of Chemical and Engineering Data* **2010**, *55*, 4054-4057.
8. Królikowska, M.; Hofman, T. Densities, isobaric expansivities and isothermal compressibilities of the thiocyanate-based ionic liquids at temperatures (298.15–338.15 K) and pressures up to 10 MPa. *Thermochimica Acta* **2012**, *530*, 1-6.
9. Mokhtarani, B.; Sharifi, A.; Mortaheb, H. R.; Mirzaei, M.; Mafi, M.; Sadeghian, F. Density and viscosity of 1-butyl-3-methylimidazolium nitrate with ethanol, 1-propanol, or 1-butanol at several temperatures. *Journal of Chemical Thermodynamics* **2009**, *41*, 1432-1438.
10. Matkowska, D.; Hofman, T. High-pressure volumetric properties of ionic liquids: 1-butyl-3-methylimidazolium tetrafluoroborate, [C<sub>4</sub>mim][BF<sub>4</sub>], 1-butyl-3-methylimidazolium methylsulfate [C<sub>4</sub>mim][MeSO<sub>4</sub>] and 1-ethyl-3-methylimidazolium ethylsulfate, [C<sub>2</sub>mim][EtSO<sub>4</sub>]. *Journal of Molecular Liquids* **2012**, *165*, 161-167.
11. Kuhlmann, E.; Himmeler, S.; Giebelhaus, H.; Wasserscheid, P. Imidazolium dialkylphosphates—a class of versatile, halogen-free and hydrolytically stable ionic liquids. *Green Chem.* **2007**, *9*, 233-242.
12. Tang, S.; Baker, G. A.; Ravula, S.; Jones, J. E.; Zhao, H. PEG-functionalized ionic liquids for cellulose dissolution and saccharification. *Green Chem.* **2012**, *14*, 2922-2932.
13. Gong, Y.; Shen, C.; Lu, Y.; Meng, H.; Li, C. Viscosity and density measurements for six binary mixtures of water (methanol or ethanol) with an ionic liquid ([BMIM][DMP] or [EMIM][DMP]) at atmospheric pressure in the temperature range of (293.15 to 333.15) K. *Journal of Chemical and Engineering Data* **2012**, *57*, 33-39.
14. Li, W.; Zhang, Z.; Han, B.; Hu, S.; Xie, Y.; Yang, G. Effect of water and organic solvents on the ionic dissociation of ionic liquids. *Journal of Physical Chemistry B* **2007**, *111*, 6452-6456.

15. Dinarès, I.; de Miguel, C. G.; Ibáñez, A.; Mesquida, N.; Alcalde, E. Imidazolium ionic liquids: A simple anion exchange protocol. *Green Chem.* **2009**, *11*, 1507-1510.
16. Safarov, J.; Geppert-Rybczyńska, M.; Kul, I.; Hassel, E. Thermophysical properties of 1-butyl-3-methylimidazolium acetate over a wide range of temperatures and pressures. *Fluid Phase Equilibria* **2014**, *383*, 144-155.
17. Chen, Y.; Mutelet, F.; Jaubert, J.-N. Solubility of CO<sub>2</sub> in 1-butyl-3-methylimidazolium diethylene-glycolmonomethylethersulfate and trihexyl(tetradecyl)phosphonium dodecylbenzenesulfonate. *Fluid Phase Equilibria* **2013**, *354*, 191-198.
18. Pinkert, A.; Ang, K. L.; Marsh, K. N.; Pang, S. Density, viscosity and electrical conductivity of protic alkanolammonium ionic liquids. *Physical Chemistry Chemical Physics* **2011**, *13*, 5136-5143.
19. Vakili-Nezhaad, G.; Vatani, M.; Asghari, M.; Ashour, I. Effect of temperature on the physical properties of 1-butyl-3-methylimidazolium based ionic liquids with thiocyanate and tetrafluoroborate anions, and 1-hexyl-3-methylimidazolium with tetrafluoroborate and hexafluorophosphate anions. *Journal of Chemical Thermodynamics* **2012**, *54*, 148-154.
20. Mokhtarani, B.; Mojtabedi, M. M.; Mortaheb, H. R.; Mafi, M.; Yazdani, F.; Sadeghian, F. Densities, refractive indices, and viscosities of the ionic liquids 1-methyl-3-octylimidazolium tetrafluoroborate and 1-methyl-3-butylimidazolium perchlorate and their binary mixtures with ethanol at several temperatures. *Journal of Chemical and Engineering Data* **2008**, *53*, 677-682.
21. Shirota, H.; Matsuzaki, H.; Ramati, S.; Wishart, J. F. Effects of Aromaticity in Cations and Their Functional Groups on the Low-Frequency Spectra and Physical Properties of Ionic Liquids. *The Journal of Physical Chemistry B* **2015**, *119*, 9173-9187.
22. Tao, R.; Tamas, G.; Xue, L.; Simon, S. L.; Quitevis, E. L. Thermophysical properties of imidazolium-based ionic liquids: The effect of aliphatic versus aromatic functionality. *Journal of Chemical and Engineering Data* **2014**, *59*, 2717-2724.
23. Liu, Q.-S.; Yang, M.; Li, P.-P.; Sun, S.-S.; Welz-Biermann, U.; Tan, Z.-C.; Zhang, Q.-G. Physicochemical Properties of Ionic Liquids [C<sub>3</sub>py][NTf<sub>2</sub>] and [C<sub>6</sub>py][NTf<sub>2</sub>]. *Journal of Chemical & Engineering Data* **2011**, *56*, 4094-4101.
24. Oliveira, F. S.; Freire, M. G.; Carvalho, P. J.; Coutinho, J. A. P.; Lopes, J. N. C.; Rebelo, L. P. N.; Marrucho, I. M. Structural and positional isomerism influence in the physical properties of pyridinium NTf<sub>2</sub>-based ionic liquids: Pure and water-saturated mixtures. *Journal of Chemical and Engineering Data* **2010**, *55*, 4514-4520.
25. Ma, X.-X.; Wei, J.; Guan, W.; Pan, Y.; Zheng, L.; Wu, Y.; Yang, J.-Z. Ionic parachor and its application to pyridinium-based ionic liquids of {[C<sub>n</sub>py][DCA]} (n = 2, 3, 4, 5, 6). *The Journal of Chemical Thermodynamics* **2015**, *89*, 51-59.
26. Papaiconomou, N.; Salminen, J.; Lee, J.-M.; Prausnitz, J. M. Physicochemical Properties of Hydrophobic Ionic Liquids Containing 1-Octylpyridinium, 1-Octyl-2-methylpyridinium, or 1-Octyl-4-methylpyridinium Cations. *Journal of Chemical & Engineering Data* **2007**, *52*, 833-840.
27. Yunus, N. M.; Mutualib, M. I. A.; Man, Z.; Bustam, M. A.; Murugesan, T. Thermophysical properties of 1-alkylpyridinium bis(trifluoromethylsulfonyl)imide ionic liquids. *Journal of Chemical Thermodynamics* **2010**, *42*, 491-495.
28. Burankova, T.; Reichert, E.; Fossog, V.; Hempelmann, R.; Embs, J. P. The dynamics of cations in pyridinium-based ionic liquids by means of quasielastic- and inelastic neutron scattering. *Journal of Molecular Liquids* **2014**, *192*, 199-207.

29. Kilaru, P.; Baker, G. A.; Scovazzo, P. Density and surface tension measurements of imidazolium-, quaternary phosphonium-, and ammonium-based room-temperature ionic liquids: Data and correlations. *Journal of Chemical and Engineering Data* **2007**, *52*, 2306-2314.
30. Bhattacharjee, A.; Luis, A.; Lopes-da-Silva, J. A.; Freire, M. G.; Carvalho, P. J.; Coutinho, J. A. P. Thermophysical properties of sulfonium- and ammonium-based ionic liquids. *Fluid Phase Equilibria* **2014**, *381*, 36-45.
31. Jin, H.; O'Hare, B.; Dong, J.; Arzhantsev, S.; Baker, G. A.; Wishart, J. F.; Benesi, A. J.; Maroncelli, M. Physical Properties of Ionic Liquids Consisting of the 1-Butyl-3-Methylimidazolium Cation with Various Anions and the Bis(trifluoromethylsulfonyl)imide Anion with Various Cations. *The Journal of Physical Chemistry B* **2008**, *112*, 81-92.
32. Wang, Y.; Hao, W.; Jacquemin, J.; Goodrich, P.; Atilhan, M.; Khraisheh, M.; Rooney, D.; Thompson, J. Enhancing liquid-pase oefin–paraffin separations using novel silver-based ionic liquids. *Journal of Chemical and Engineering Data* **2015**, *60*, 28-36.
33. Gacino, F. M.; Regueira, T.; Lugo, L.; Comunas, M. J. P.; Fernandez, J. Influence of molecular structure on densities and viscosities of several ionic liquids. *Journal of Chemical and Engineering Data* **2011**, *56*, 4984-4999.
34. Tagiuri, A.; Sumon, K. Z.; Henni, A. Solubility of carbon dioxide in three  $[\text{Tf}_2\text{N}]$  ionic liquids. *Fluid Phase Equilibria* **2014**, *380*, 39-47.
35. Řehák, K.; Morávek, P.; Strejc, M. Determination of mutual solubilities of ionic liquids and water. *Fluid Phase Equilibria* **2012**, *316*, 17-25.
36. Blahut, A.; Dohnal, V. Interactions of volatile organic compounds with the ionic liquids 1-butyl-1-methylpyrrolidinium tetracyanoborate and 1-butyl-1-methylpyrrolidinium bis(oxalato)borate. *Journal of Chemical Thermodynamics* **2013**, *57*, 344-354.
37. Page, P. M.; McCarty, T. A.; Baker, G. A.; Baker, S. N.; Bright, F. V. Comparison of Dansylated Aminopropyl Controlled Pore Glass Solvated by Molecular and Ionic Liquids. *Langmuir* **2007**, *23*, 843-849.
38. Bhattacharjee, A.; Carvalho, P. J.; Coutinho, J. A. P. The effect of the cation aromaticity upon the thermophysical properties of piperidinium- and pyridinium-based ionic liquids. *Fluid Phase Equilibria* **2014**, *375*, 80-88.
39. Ueno, K.; Yoshida, K.; Tsuchiya, M.; Tachikawa, N.; Dokko, K.; Watanabe, M. Glyme–Lithium Salt Equimolar Molten Mixtures: Concentrated Solutions or Solvate Ionic Liquids? *The Journal of Physical Chemistry B* **2012**, *116*, 11323-11331.
40. Parviainen, A.; King, A. W. T.; Mutikainen, I.; Hummel, M.; Selg, C.; Hauru, L. K. J.; Sixta, H.; Kilpeläinen, I. Predicting Cellulose Solvating Capabilities of Acid–Base Conjugate Ionic Liquids. *ChemSusChem* **2013**, *6*, 2161-2169.
41. Zhao, H.; Baker, G. A.; Holmes, S. New eutectic ionic liquids for lipase activation and enzymatic preparation of biodiesel. *Org. Biomol. Chem.* **2011**, *9*, 1908-1916.
42. Abbott, A. P.; Harris, R. C.; Ryder, K. S. Application of hole theory to define ionic liquids by their transport properties. *J. Phys. Chem. B* **2007**, *111*, 4910-4913.