Table A. Enthalpies of Transfer of Copolymers from Water to the Aqueous Surfactant Solutions at 298 K.

$m_{S,i}$ mol kg^{-1}	$\phi_S \cdot 10^6$ kg s^{-1}	$\phi_R \cdot 10^6$ kg s^{-1}	ΔH_t kJ mol^{-1}
	OPC/L64	$m_{R,i}$=10.0 mmol kg^{-1}	
0.6780	2.805	2.804	94.2
0.8660	2.905	2.822	107.4
1.0113	2.891	2.676	117.6
1.1608	2.891	2.774	122.8
1.2640	2.920	2.809	118.4
	OPC/F68	$m_{R,i}$=2.00 mmol kg^{-1}	
0.6780	2.834	2.819	166.8
0.8660	2.905	2.829	187.2
1.0113	2.861	2.743	188.8
1.2640	2.927	2.807	198.1
1.8117	2.948	2.759	196.0
	DeTAB/F68	$m_{R,i}$=1.97 mmol kg^{-1}	
0.44531	2.489	2.566	159.8
0.56139	2.483	2.587	158.1
0.60313	2.490	2.601	158.1
	DeTAB/F68	$m_{R,i}$=2.00 mmol kg^{-1}	
0.02001	2.781	2.609	8.0
0.05662	2.817	2.612	24.2
0.09014	2.739	2.609	46.0
0.1099	2.737	2.595	73.0
0.1502	2.764	2.582	120.7
0.1847	2.703	2.550	143.1
0.2301	2.711	2.549	154.3
0.2798	2.695	2.539	160.8
0.3302	2.713	2.583	166.6
0.4038	2.652	2.539	166.1
	DeTAB/F68	$m_{R,i}$=2.02 mmol kg^{-1}	
0.03980	2.552	2.599	13.4
0.07805	2.481	2.599	28.1
0.1224	2.416	2.600	86.7
0.3181	2.557	2.484	161.3

Table A (continued)

$m_{S,i}$ mol kg^{-1}	$\phi_S \cdot 10^6$ kg s^{-1}	$\phi_R \cdot 10^6$ kg s^{-1}	ΔH_t kJ mol^{-1}
	DeTAB/L64	$m_{R,i}$= 9.87 mmol kg^{-1}	
0.04070	4.877	3.732	6.7
0.08027	2.652	2.539	24.5
0.1203	2.848	2.673	49.5
0.1635	2.890	2.712	77.6
0.2086	2.743	3.722	83.8
0.2422	2.740	3.779	94.9
0.2795	2.730	3.625	103.8
	DeTAB/L64	$m_{R,i}$= 10.41 mmol kg^{-1}	
0.02407	2.844	2.566	1.8
0.3194	2.459	2.656	120.8
0.3602	2.510	2.692	124.2
0.3978	2.530	2.492	124.0
0.4416	3.549	2.965	124.7
0.4801	3.496	2.974	126.1
	DeTAB/L64	$m_{R,i}$= 9.89 mmol kg^{-1}	
0.01001	2.786	2.604	-2.0
0.06612	2.761	2.674	13.5
0.1040	2.751	2.711	39.2

Note that $m_{S,i}$ and $m_{R,i}$ stand for the initial concentrations of the surfactant and copolymer solutions, respectively.

Table B. Apparent molar volume of L64 in aqueous sodium decanoate solutions at 298 K.[a]

m_R	d_T	V_Φ
m_S = 0.06976	d_B =0.999040	
4.80	1.000435	2608.0
5.40	1.000609	2607.6
6.08	1.000814	2605.8
7.16	1.001134	2604.3
8.33	1.001476	2603.4
9.74	1.001898	2601.3
12.46	1.002687	2600.0
15.77	1.003654	2597.6
20.29	1.004952	2595.5
m_S= 0.1033	d_B=1.000013	
6.18	1.001546	2644.6
7.27	1.001809	2643.1
8.42	1.002137	2637.1
9.85	1.002522	2633.5
12.63	1.003289	2626.9
16.00	1.004206	2621.6
20.57	1.005443	2616.5

[a]Units are: concentrations, mol kg^{-1}; densities, g cm^{-3}; apparent molar volume, cm^3 mol^{-1}.