Supporting information

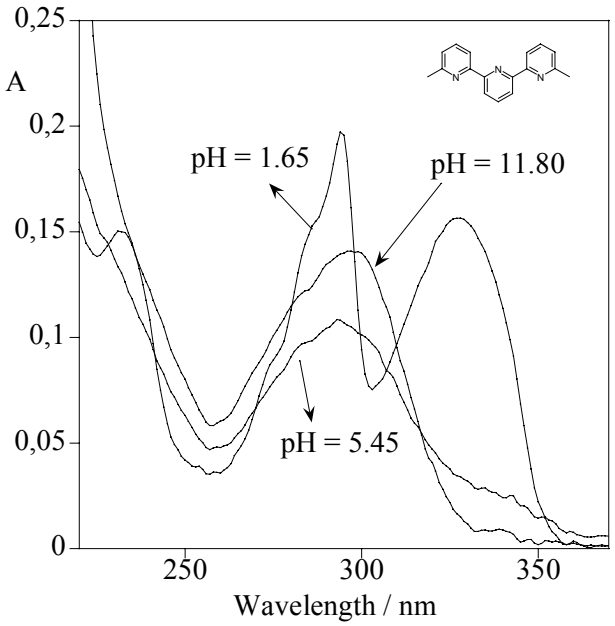

Figure S1. pH dependence of the absorption spectra of Terpyridine in aqueous solution [L]= $1.95.10^{-5}$

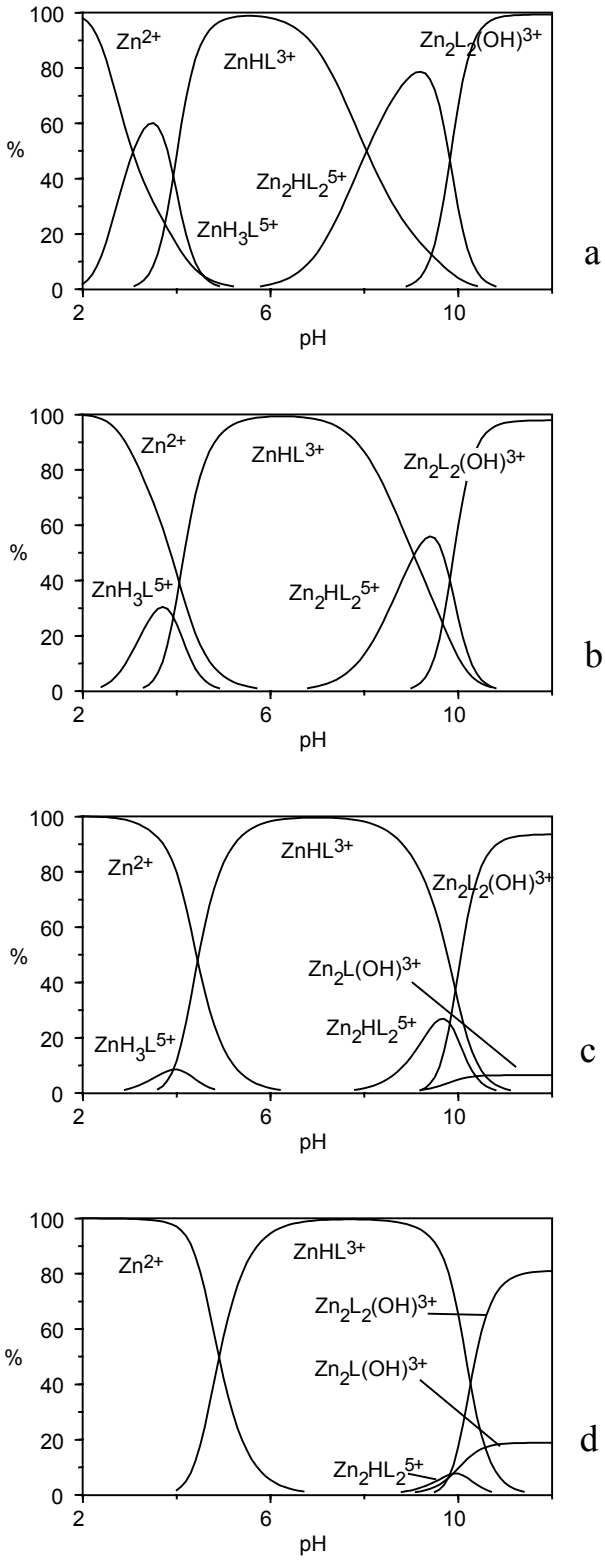

Figure S2. Distribution diagrams of the species for the systems **L**\Zn(II) at different concentrations in NMe_4Cl 0.1 mol dm^{-3} at 298.1 K; [**L**] = $[Zn^{2+}] = 1 \cdot 10^{-2}$ M (a), [**L**] = $[Zn^{2+}] = 1 \cdot 10^{-3}$ M (b), [**L**] = $[Zn^{2+}] = 1 \cdot 10^{-4}$ M (c), [**L**] = $[Zn^{2+}] = 1 \cdot 10^{-5}$ M (d).

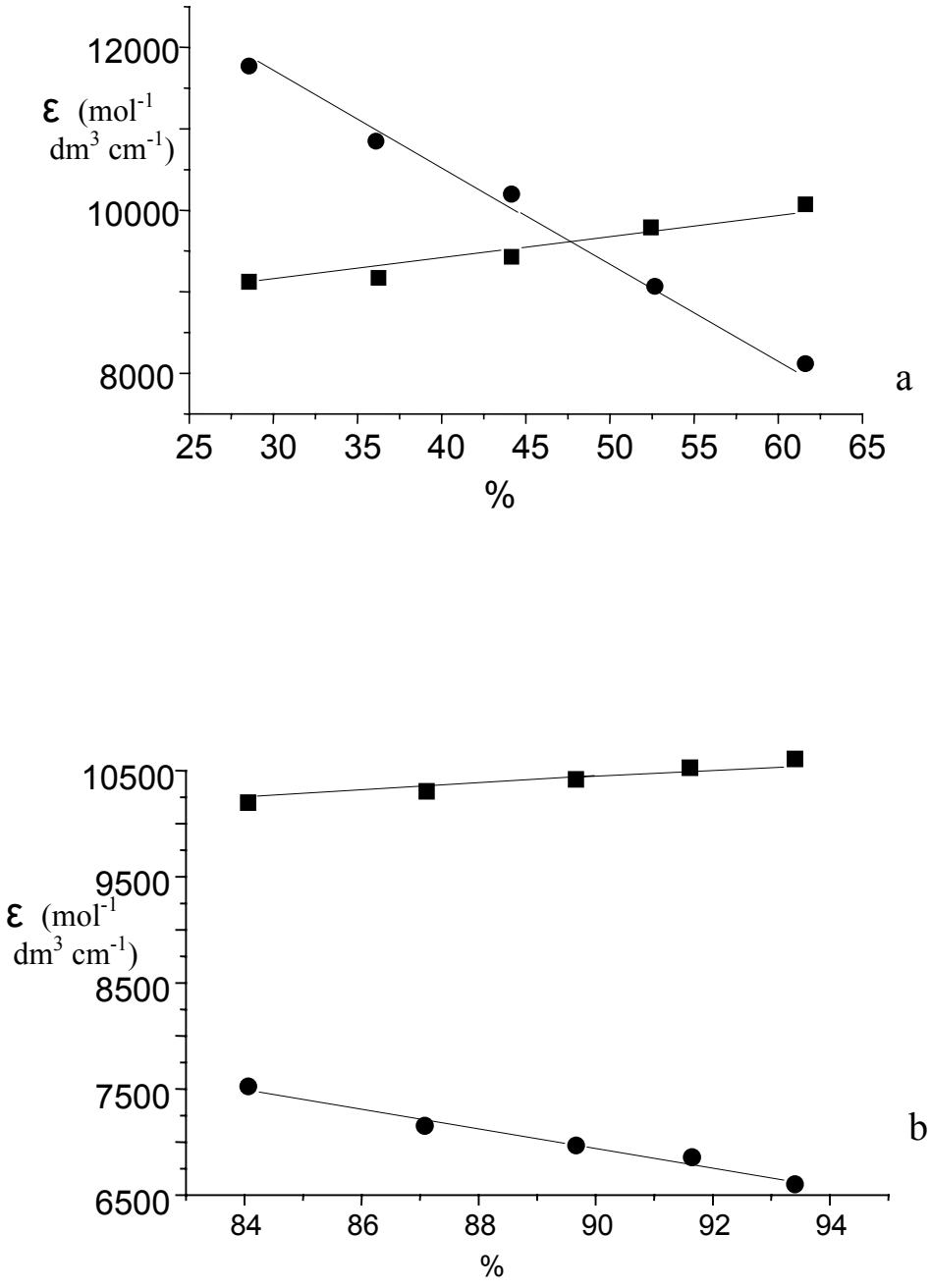

Figure S3. ε values at 310 nm (■) and 335 nm (●) measured on solutions with different concentrations of the Zn(II) complex with **L** at pH 9.9 (a) and 11.0 (b) as a function of the overall percentage of the dimeric species ($[Zn_2\mathbf{L}_2H]^{5+}$ + $[Zn_2\mathbf{L}_2(OH)]^{3+}$) calculated on the basis of the potentiometric results.