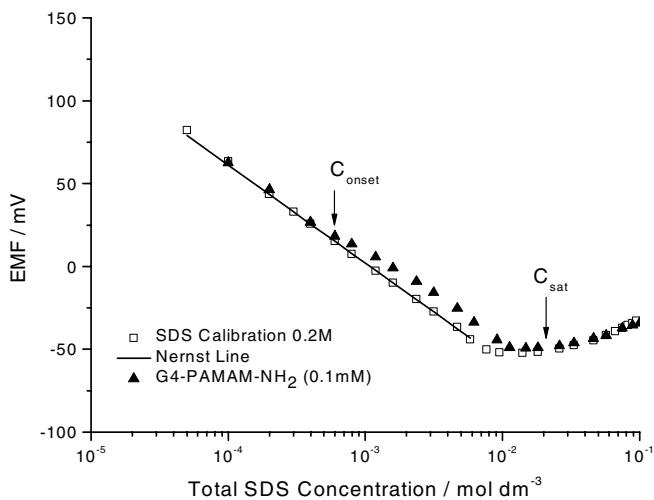
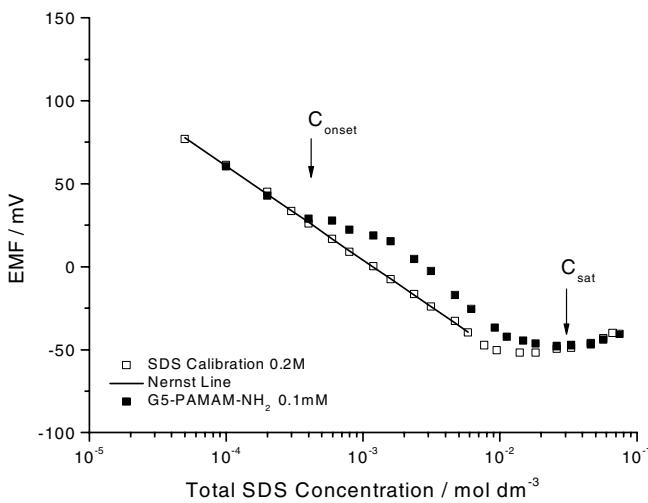


MS LA0494932 by Holzwarth *et al* additional Figures in supporting information:



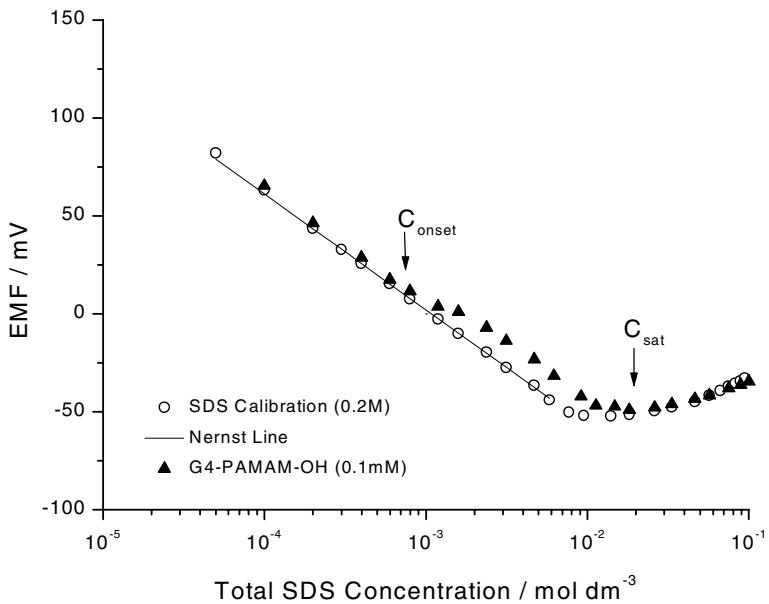
SI Figure A

Graph of the Electromotive Force of the DS⁻ selective electrode (relative to a bromide ion electrode) as a function of total SDS concentration for the (□) pure SDS aqueous solution and (■) SDS/0.1 × 10⁻³ mol dm⁻³ G5-PAMAM-NH₂ aqueous solution. Solutions are doped with NaBr 10⁻⁴ mol dm⁻³. T = 298K. C_{onset} and C_{sat} of binding; pH = 10.2.



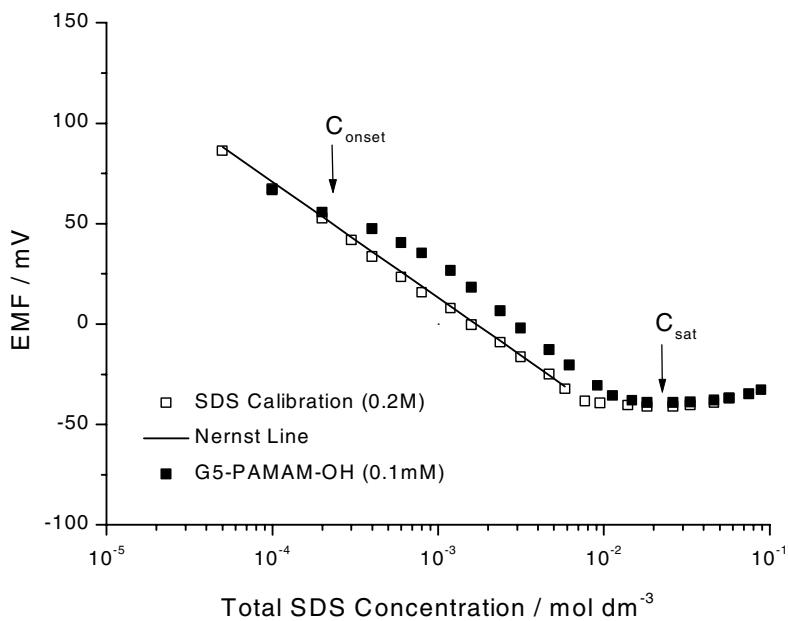
SI Figure B

Graph of the Electromotive Force of the DS⁻ selective electrode (relative to a bromide ion electrode) as a function of total SDS concentration for the (□) pure SDS aqueous solution and (■) SDS/0.1 × 10⁻³ mol dm⁻³ G5-PAMAM-NH₂ aqueous solution. Solutions are doped with NaBr 10⁻⁴ mol dm⁻³. T = 298K. C_{onset} and C_{sat} of binding; pH = 10.2.



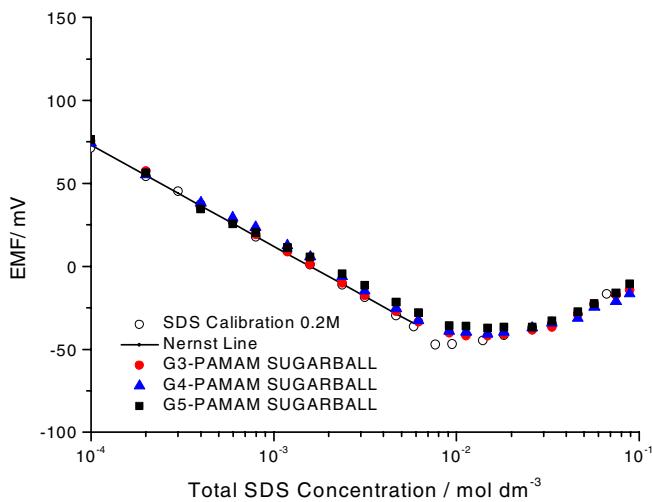
SI Figure C

Graph of the Electromotive Force of the DS⁻ selective electrode (relative to a bromide ion electrode) as a function of total SDS concentration for the (□) pure SDS aqueous solution and (▲) SDS/0.1 × 10⁻³ mol dm⁻³ G4-PAMAM-OH aqueous solution. Solutions are doped with NaBr 10⁻⁴ mol dm⁻³. T = 298K. C_{onset} and C_{sat} of binding; pH = 7.4.



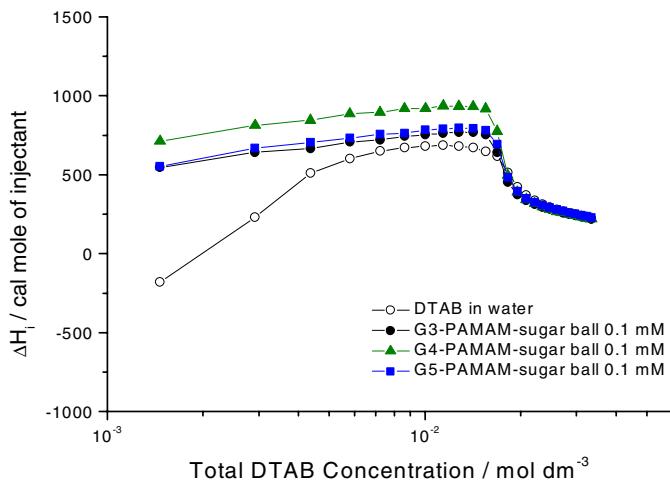
SI Figure D

Graph of the Electromotive Force of the DS⁻ selective electrode (relative to a bromide ion electrode) as a function of total SDS concentration for the (□) pure SDS aqueous solution and (■) SDS/0.1 × 10⁻³ mol dm⁻³ G5-PAMAM-OH aqueous solution. Solutions are doped with NaBr 10⁻⁴ mol dm⁻³. T = 298K. C_{onset} and C_{sat} of binding; pH = 7.4.



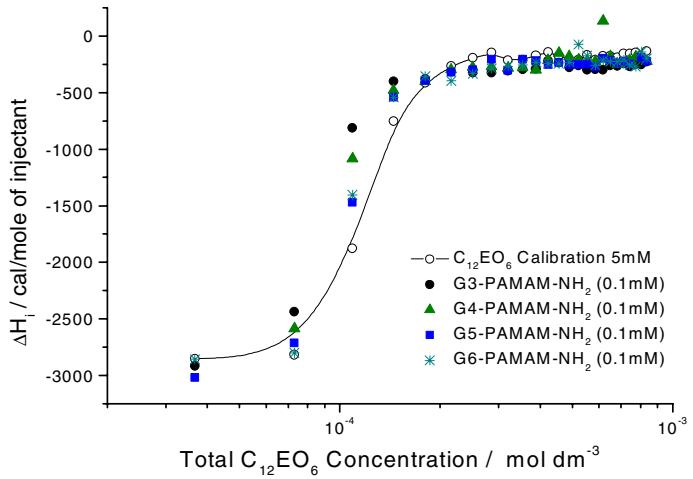
SI Figure E

Graph of the Electromotive Force of the DS⁻ selective electrode (relative to a bromide ion electrode) as a function of total SDS concentration for the (□) pure SDS aqueous solution and SDS/0.1 × 10⁻³ mol dm⁻³ Gx-PAMAM-Sugar ball aqueous solution. Solutions are doped with NaBr 10⁻⁴ mol dm⁻³. T = 298K. pH = 7.2.



SI Figure F

ITC graph of the enthalpy per injection as a function of total DTAB concentration for (o) pure DTAB aqueous solution, (●) DTAB /0.1 × 10⁻³ mol dm⁻³ G3-PAMAM-sugar ball aqueous solution, (▲) DTAB /0.1 × 10⁻³ mol dm⁻³ G4-PAMAM- sugar ball aqueous solution, (■) DTAB /0.1 × 10⁻³ mol dm⁻³ G5-PAMAM- sugar ball aqueous solution. T = 298K, some interactions at low DTAB concentrations; pH = 7.2.



SI Figure G1

ITC graph of the enthalpy per injection as a function of total $C_{12}\text{EO}_6$ concentration for (o) pure $C_{12}\text{EO}_6$ aqueous solution, (●) $C_{12}\text{EO}_6/0.1 \times 10^{-3}$ mol dm⁻³ G3-PAMAM-NH₂ aqueous solution, (▲) $C_{12}\text{EO}_6/0.1 \times 10^{-3}$ mol dm⁻³ G4-PAMAM-NH₂ aqueous solution, (■) $C_{12}\text{EO}_6/0.1 \times 10^{-3}$ mol dm⁻³ G5-PAMAM-NH₂ aqueous solution, (*) $C_{12}\text{EO}_6/0.1 \times 10^{-3}$ mol dm⁻³ G6-PAMAM-NH₂ aqueous solution. T = 298K, no interaction; pH = 10.2.

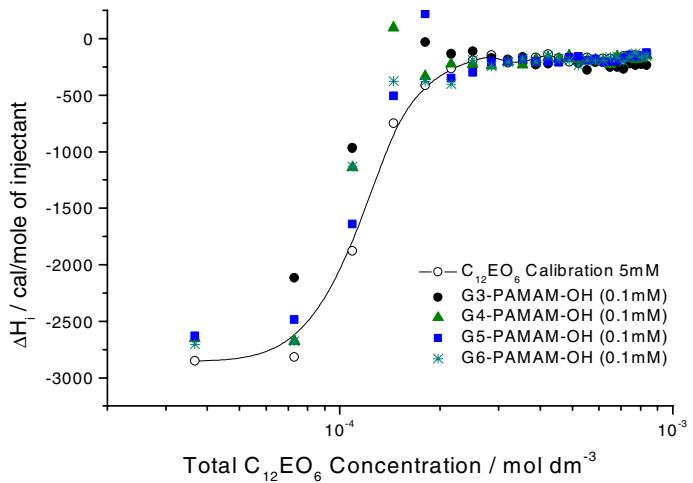


Figure G2

ITC graph of the enthalpy per injection as a function of total $C_{12}\text{EO}_6$ concentration for (o) pure $C_{12}\text{EO}_6$ aqueous solution, (●) $C_{12}\text{EO}_6/0.1 \times 10^{-3}$ mol dm⁻³ G3-PAMAM-OH aqueous solution, (▲) $C_{12}\text{EO}_6/0.1 \times 10^{-3}$ mol dm⁻³ G4-PAMAM-OH aqueous solution, (■) $C_{12}\text{EO}_6/0.1 \times 10^{-3}$ mol dm⁻³ G5-PAMAM-OH aqueous solution, (*) $C_{12}\text{EO}_6/0.1 \times 10^{-3}$ mol dm⁻³ G6-PAMAM-OH aqueous solution. T = 298K, very weak interaction; pH = 7.4.

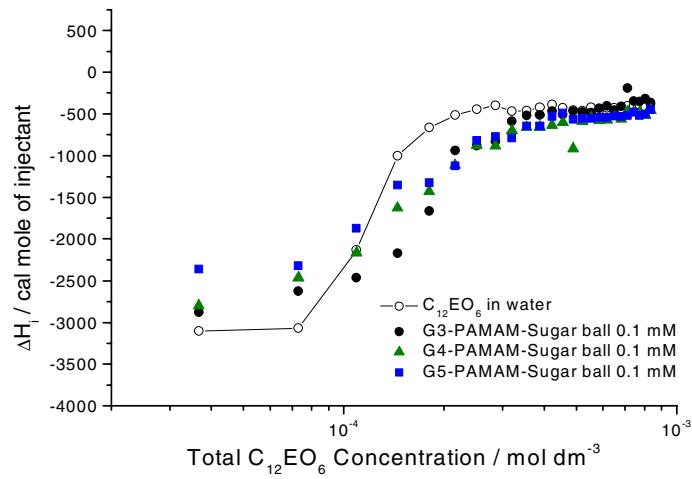
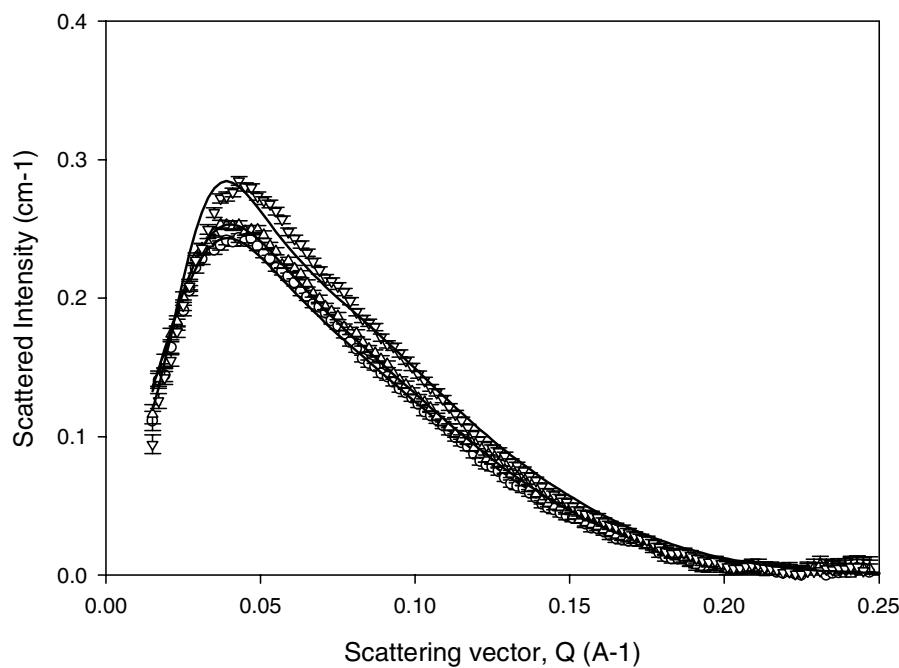


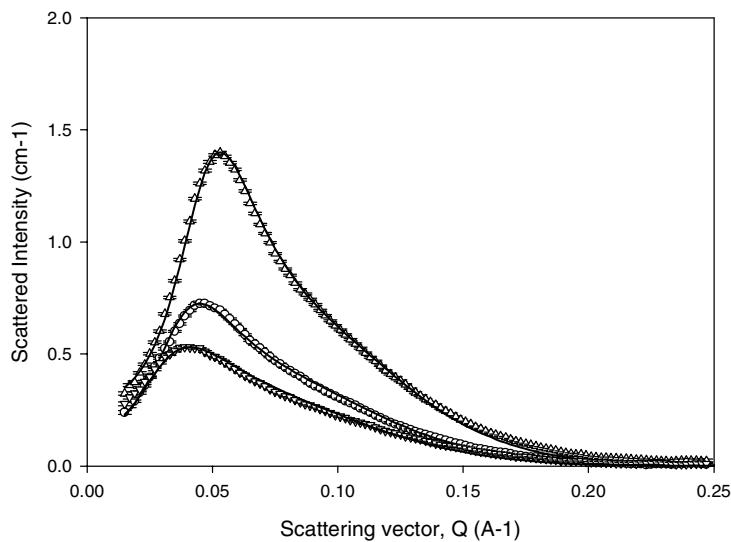
Figure G3

ITC graph of the enthalpy per injection as a function of total $C_{12}EO_6$ concentration for (o) pure $C_{12}EO_6$ aqueous solution, (●) $C_{12}EO_6/0.1 \times 10^{-3}$ mol dm^{-3} G3-PAMAM-sugar ball aqueous solution, (▲) $C_{12}EO_6/0.1 \times 10^{-3}$ mol dm^{-3} G4-PAMAM- sugar ball aqueous solution, (■) $C_{12}EO_6/0.1 \times 10^{-3}$ mol dm^{-3} G5-PAMAM- sugar ball aqueous solution. $T = 298K$, weak interaction. ; $pH = 7.2$.



SI Figure H

SANS scattered intensity for 0.1 mM G3-PAMAM-NH₂ / 12 mM SDS (o), 0.1mM (DDA core) G3-PAMAM-NH₂ / 18 mM SDS (∇), and 0.1mM G3-PAMAM-OH / 17 mM SDS (Δ) in D₂O. The solid lines are model fits as described in the text, using equations 1 and 2 and the parameters summarized in Table 3. DDA means that the core of this synthesized dendrimer is 1,12-diaminododecane. T = 298K.



SI Figure I

SANS scattered intensity for 0.35 mM (0.5% w/v) G5-PAMAM-NH₂ / 28 mM SDS (o), 0.1mM G5-PAMAM-NH₂ / 49 mM SDS (Δ), and 0.1mM G5-PAMAM-NH₂ / 23 mM SDS (∇) in D₂O. The solid lines are model fits as described in the text, using Equations 1 and 2 and the parameters summarized in Table 3. T = 298K.