SUPPORTING MATERIAL

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S2 **Table S1.** The effect of [BrO-] and pH on k_{obs} s^{-1} for the reaction of NPDEPN with BrO^-, from complex **1** in 1.0 M KCl at 25 °C.

S3 **Table S2** Values of k_{obs} (s^{-1}) for the reaction of NPDEPN with BrO^- in 1.0 M KCl at 25 °C.

S4 **Table S3** Values of k_{obs} (s^{-1}) for the reaction of NPDEP with BrO^- in 1.0 M KCl at 25 °C.

S5 **Table S4** Values of k_{obs} (s^{-1}) for the reaction of NPOTos with BrO^- in 1.0 M KCl at 25 °C.

Table S1. The effect of [BrO-] and pH on k_{obs} s^{-1} for the reaction of NPDEPN with BrO^-, from complex **1** in 1.0 M KCl at 25 °C.

[BrO⁻], M [a]	10^3 k_{obs}, s^{-1}	pH [b]	k_{obs} 10^3, s^{-1}
0.00375	0.82	11.00	1.52
0.0055	1.10	11.35	1.60
0.0075	1.30	11.52	1.85
0.0092	1.39	11.76	2.10
0.011	1.85	12.02	2.90
0.015	2.14		
0.0184	2.69		
0.022	3.26		

[a] at pH = 11.35-11.39; [b] at [BrO-] = 0.0101 M.

Table S2 Values of k_{obs} (s^{-1}) for the reaction of NPDEPN with BrO^- in 1.0 M KCl at 25 °C.

$[BrO^-]$, M [a]	10^3 k_{obs}, s^{-1}	$[BrO^-]$, M [b]	10^3 k_{obs}, s^{-1}	$[BrO^-]$, M [c]	10^3 k_{obs}, s^{-1}	$[BrO^-]$, M [d]	10^3 k_{obs}, s^{-1}
0.0025	0.75	0.0026	0.48	0.0025	0.58	0.00325	0.63
0.0039	0.89	0.0026	0.54	0.005	0.90	0.00425	0.74
0.005	1.05	0.0035	0.60	0.009	1.30	0.0065	0.97
0.00778	1.48	0.0052	0.81	0.010	1.44	0.0085	1.23
0.009	1.72	0.007	1.14	0.018	2.40	0.013	1.65
0.01	1.80	0.014	1.85	0.02	2.60	0.017	2.24
0.0155	2.62					0.026	3.32
0.018	3.10					0.034	4.20
0.02	3.37						
0.031	5.21						

[a] BrO^- from Br_2 in water, pH=11.30-11.35; [b] BrO^- from complex **2**, pH=11.15-11.20; [c] BrO^- from complex **3**, pH=11.30-11.35; [d] BrO^- from complex **4**, pH=11.30-11.35

Table S3 Values of k_{obs} (s^{-1}) for the reaction of NPDEP with BrO^- in 1.0 M KCl at 25 °C.

$[BrO^-]$, M [a]	10^4 k_{obs}, s^{-1}	$[BrO^-]$, M [b]	10^4 k_{obs}, s^{-1}	$[BrO^-]$, M [c]	10^4 k_{obs}, s^{-1}	$[BrO^-]$, M [d]	10^4 k_{obs}, s^{-1}
0.0044	0.60	0.0063	1.21	0.0107	1.20	0.008	0.75
0.0044	0.75	0.0052	1.01	0.007	1.10	0.0105	0.86
0.0052	0.75	0.0104	1.49	0.014	1.84	0.015	1.15
0.0052	0.70	0.0089	1.40	0.0215	3.07	0.016	1.90
0.0088	1.12	0.0125	2.08	0.028	3.23	0.021	2.40
0.0105	1.30	0.0178	2.25	0.03	3.40	0.03	3.30
0.0175	2.00	0.0208	2.55	0.035	4.30	0.042	4.45
0.021	2.36	0.0250	3.45	0.043	5.04		
0.035	3.68	0.0356	3.99				
0.042	4.45	0.0415	4.65				

[a] BrO^- from Br_2 in water , pH=11.30-11.35; [b] BrO^- from complex **1**, pH=11.35-11.39; [c] BrO^- from complex **3**, pH=11.70-11.75; [d] BrO^- from complex **4**, pH=11.70-11.75.

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Table S4 Values of k_{obs} (s^{-1}) for the reaction of NPOTos with BrO^- in 1.0 M KCl at 25 °C.

$[BrO^-]$, M [a]	10^4 k_{obs}, s^{-1}	$[BrO^-]$, M [b]	10^4 k_{obs}, s^{-1}	$[BrO^-]$, M [c]	10^4 k_{obs}, s^{-1}	$[BrO^-]$, M [d]	10^4 k_{obs}, s^{-1}	$[BrO^-]$, M [e]	10^3 k_{obs}, s^{-1}
0.0021	1.99	0.0055	1.40	0.007	1.48	0.0105	2.36	0.0105	2.67
0.0021	2.22	0.0075	1.70	0.0051	1.07	0.015	3.05	0.0125	2.90
0.0042	2.34	0.0085	1.94	0.0092	1.83	0.018	3.93	0.014	3.10
0.0085	3.08	0.011	2.36	0.011	2.00	0.021	3.96	0.019	3.97
0.0107	3.45	0.015	2.99	0.014	2.40	0.03	5.50	0.021	4.69
0.016	4.30	0.017	3.30	0.0185	2.95	0.036	6.96	0.025	4.80
0.017	4.59	0.022	4.25	0.022	3.66	0.042	6.90	0.028	5.24
0.0209	5.30	0.034	6.20	0.028	4.80			0.038	6.86
0.026	6.28			0.037	5.85			0.042	7.42
0.034	7.32			0.044	6.96				

[a] BrO^- from Br_2 in water, pH=12.28-12.29; [b] BrO^- from Br_2 in water, pH=11.72-11.75; [c] BrO^- from complex **1**, pH=11.60-11.65; [d] BrO^- from complex **3**, pH=12.0-12.05; [e] BrO^- from complex **4**, pH=12.10-12.15.