

Supporting Information Cover Sheet

Biosorption of Nonpolar Hydrophobic Organic Compounds to *Escherichia Coli*

Facilitated by Metal and Proton Surface Binding

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Submitted to *Environmental Science & Technology*

Prepared on *February 1, 2007*

Number of pages: 4

Number of figures: 3

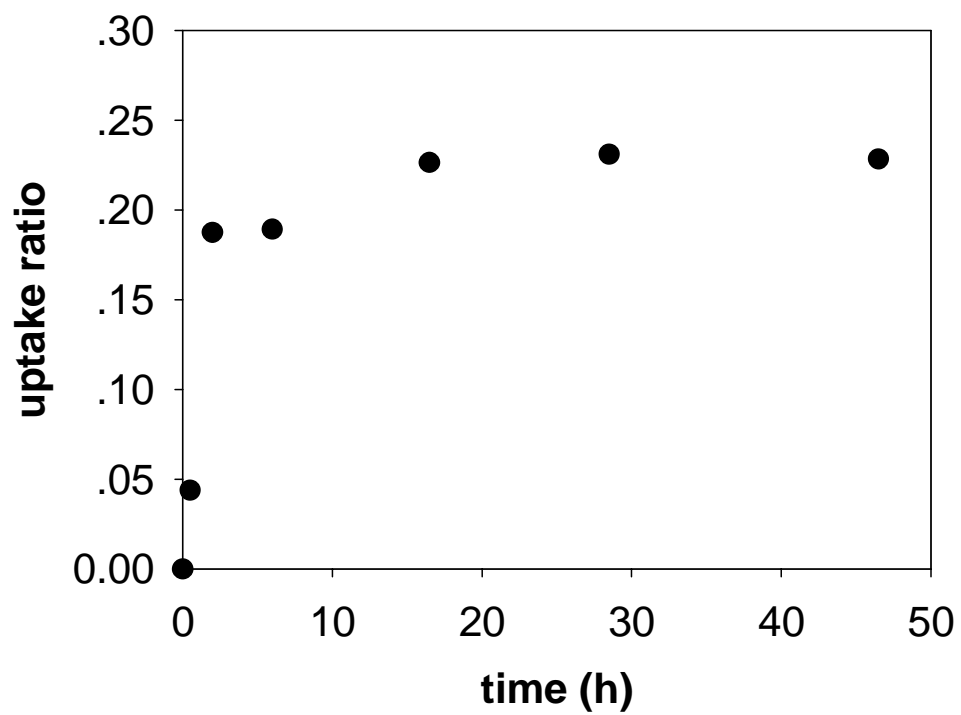


Figure S1. Sorption kinetics of PHEN to *E. coli* in 0.02 M NaNO₃, plotted as uptake ratio vs. time.

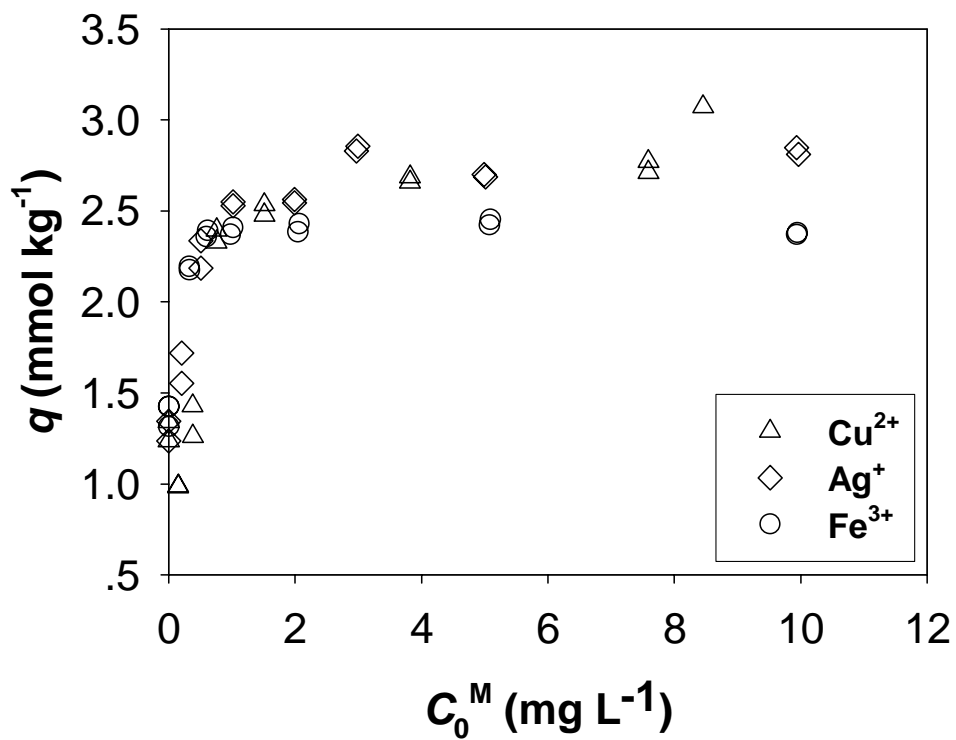


Figure S2. Sorbed concentration (q) vs. initial concentration of transition metal ion (C_0^M) at sorption equilibrium for PHEN to *E. coli* in 0.02 M NaNO₃. The equilibrium pH is 5.83±0.06 for Cu²⁺, 6.16±0.05 for Ag⁺, and 5.99±0.01 for Fe³⁺-containing samples (n = 12-14).

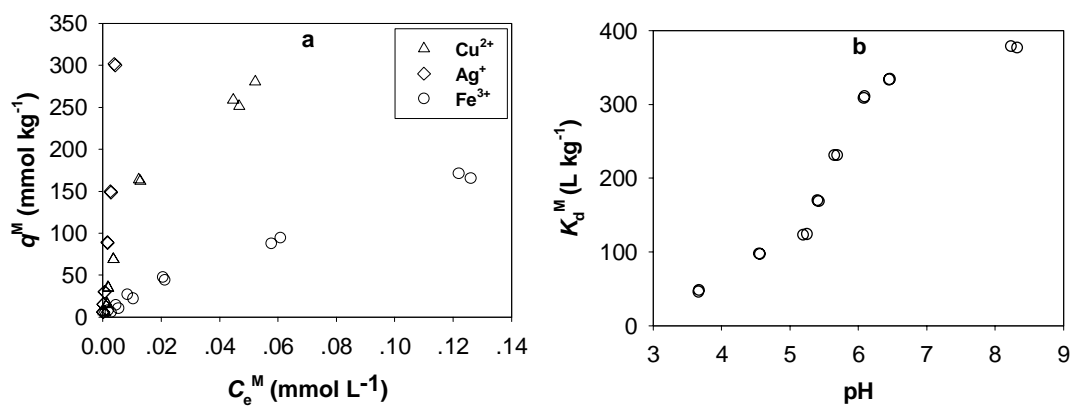


Figure S3. Sorption of heavy metals to *E. coli* in 0.02 M NaNO₃. (a) Sorbed concentration (q^M) vs. aqueous phase concentration (C_e^M) at sorption equilibrium. The equilibrium pH is 5.83 ± 0.06 for Cu²⁺, 6.16 ± 0.05 for Ag⁺, and 5.99 ± 0.01 for Fe³⁺-containing samples ($n = 12-14$). (b) Distribution coefficient (K_d^M) vs. pH for Cu²⁺.