Supporting Information

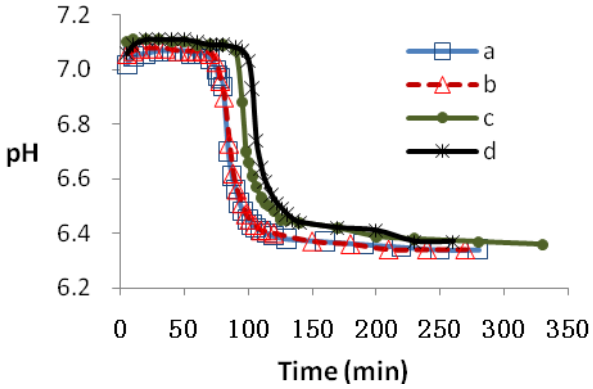
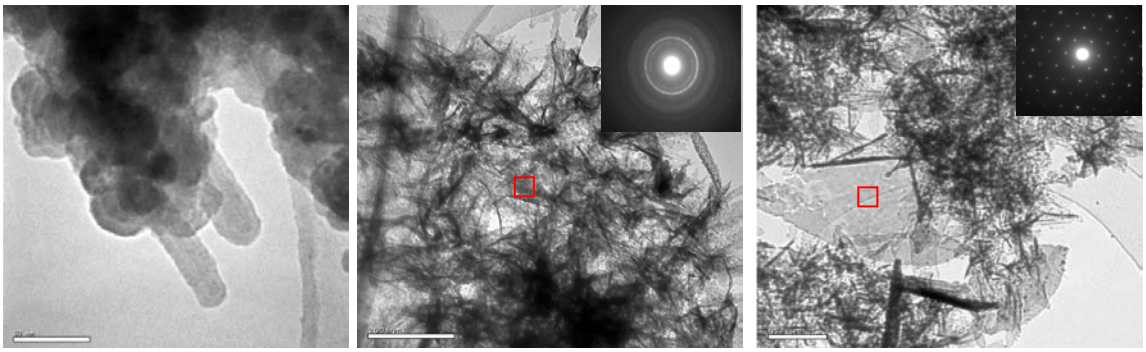


Figure S1. Influence of BTB on the precipitation kinetics of calcium phosphate. 25.0±0.1 °C. The reaction was initiated by mixing equal volumes of 8.00 mM $CaCl_2$ and 4.80 mM NaH_2PO_4, pH = 7.40. (a) and (b): In the absence of BTB; (c) and (d): BTB 30 μM.



(a) (b) (c)

Figure S2 Transmission electron micrograph of the solids isolated from calcium phosphate solution. Initial concentrations: Ca=4.00 mM, Pi=2.40 mM. (a) At 30 min from mixing the calcium and phosphate solution. The abrupt pH drop started at 85 min and concluded at 160 min. (b) At the completion of abrupt pH drop. (c) At 24 h after the completion. Samples in (b) and (c) were isolated from a separated run of experiment.