Supporting Information

Time transient electrochemical monitoring of tetraalkylammonium polybromide solid particle formation: observation of ionic liquid-to-solid transition

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The tabulated Cartesian coordinates of the optimized geometries associated with Figure S21.

Synthesis and Characterization of QBrs and TBrs

Note S1. Synthetic mechanism for *N*-Methyl-*N*-ethyl pyrrolidinium bromide (MEPBr)

1-Methylpyrrolidine (8.5 g, 100 mmol), bromoethane (8.9 mL, 120 mmol) and ethyl acetate (20 mL) were added to a 100 mL round bottom flask. The mixture was stirred at room temperature for 6 h. The solid product was filtered, washed with ethyl acetate three times, and dried in a vacuum to yield a white solid (18.6 g, 96%). ¹H NMR (500 MHz, DMSO-*d*₆) δ 3.52 – 3.35 (m, 6H), 2.97 (d, *J* = 2.0 Hz, 3H), 2.07 (dd, *J* = 5.3, 4.0 Hz, 4H), 1.31 – 1.24 (m, 3H); ¹³C NMR (125 MHz, DMSO-*d*₆) δ 63.26, 58.63, 47.31, 21.49, 9.40; MS (EI) m/z = 114 (M⁺).

Note S2. Synthetic mechanism for N-Methyl-N-ethyl-morpholinium bromide (MEMBr)

4-Methylmorpholine (17.5 mL, 160 mmol), bromoethane (23.5 mL, 320 mmol), ethyl acetate (20 mL) were added to a 100 mL round bottom flask, and the reaction mixture refluxed at 40 °C

for 72 h. After it cooled to room temperature, the solid product was filtered, washed three times with ethyl acetate, and dried in a vacuum to yield a white solid (24.3 g, 72%). ¹H NMR (500 MHz, DMSO-*d*₆) δ 3.92 (t, *J* = 9.1 Hz, 4H), 3.52 (dd, *J* = 14.6, 7.3 Hz, 2H), 3.44 – 3.36 (m, 4H), 3.10 (d, *J* = 5.9 Hz, 3H), 1.25 (t, *J* = 7.3 Hz, 3H); ¹³C NMR (125 MHz, DMSO-*d*₆) δ 60.25, 59.70, 58.84, 45.79, 7.37; MS (EI) *m/z* = 130 (M⁺).

Note S3. Synthetic mechanism for 1-Ethylpyridinium bromide (EPyBr)



To a solution of pyridine (40.3 mL, 500 mmol) in ethyl acetate (40 mL), bromoethane (74 mL, 1.0 mol) was added dropwise in ice-bath. The mixture was stirred at 30 °C for 72 h. The solid product was filtered, washed three times with ethyl acetate, and dried in a vacuum to yield a white solid (59 g, 63%). ¹H NMR (500 MHz, DMSO-*d*₆) δ 9.11 (d, *J* = 5.8 Hz, 2H), 8.60 (t, *J* = 7.8 Hz, 1H), 8.16 (t, *J* = 6.9 Hz, 2H), 4.63 (q, *J* = 7.3 Hz, 2H), 1.54 (t, *J* = 7.3 Hz, 3H); ¹³C NMR (125 MHz, DMSO-*d*₆) δ 146.09, 145.25, 128.76, 57.02, 17.05; MS (EI) *m/z* = 108.1 (M⁺).

Note S4. Synthetic mechanism for Tetrapropylammonium bromide (TProABr)



Tripropylamine (15.0 mL, 80 mmol), 1-bromopropane (11.0 mL, 120 mmol), and ethanol (50 mL) were added to a 250 mL round bottom flask, and the reaction mixture was refluxed at 80 °C

for 48 h. After cooling to room temperature, the reaction mixture was concentrated to give a crude solid product. The crude product was washed with EtOAc and dried in a vacuum to yield a white solid (11.8 g, 55%). ¹H NMR (500 MHz, DMSO-*d*₆) δ 3.18 – 3.08 (m, 8H), 1.72 – 1.49 (m, 8H), 0.87 (t, *J* = 7.3 Hz, 12H); ¹³C NMR (125 MHz, DMSO-*d*₆) δ 59.75 (s), 15.31 (s), 11.00 (s); MS (EI) *m/z* = 186.2 (M⁺).

Note S5. Synthetic mechanism for Tetrapentylammonium bromide (TPABr)



[CAS No. 866-97-7]

1-Bromopentane (10 mL, 80 mmol), tripentylamine (46 mL 160 mmol), and ethanol (50 mL) were added to a 250 mL round bottom flask, and the reaction mixture was refluxed at 80 °C for 72 h. After cooling to room temperature, the reaction mixture was concentrated to give a crude solid product. The crude product was washed with EtOAc and dried in a vacuum to yield a white solid (13.4 g, 50%). ¹H NMR (500 MHz, DMSO-*d*₆) δ 3.23 – 3.10 (m, 8H), 1.66 – 1.48 (m, 8H), 1.48 – 1.15 (m, 16H), 0.87 (t, *J* = 7.2 Hz, 12H); ¹³C NMR (125 MHz, DMSO-*d*₆) δ 58.15 (s), 28.40 (s), 22.03 (s), 21.28 (s), 14.18 (s); MS (EI) *m/z* = 298.3 (M⁺).



Figure S1. The photograph of precipitated TBABr₃ on a Pt macro disk electrode with a radius of 1 mm after a potential of 1.5 V was applied for 1000 s in a 0.5 M H₂SO₄ aqueous solution with $C_{\text{TBABr}} = 50$ mM.



Figure S2. The Raman spectra measured from TBABr₃ formed electrochemically on a Pt macro disk electrode described in Figure S1 (black) and purchased from Sigma-Aldich (red).



Figure S3. The photographs of synthesized polybromides as a function of equiv. Br₂.



Figure S4. The Raman spectra obtained from TBr_{2n+1} , which were chemically synthesized by adding Br₂ to TBr aqueous solutions to have different ratios of $C_{Br_2(aq)}$ to $C_{Br^-(aq)}$.



Figure S5. 2D axial symmetric domain of the simulation for Figure 3.



Figure S6. The linear sweep voltammograms (LSVs, black) measured in 0.5 M H₂SO₄ aqueous solutions containing various concentrations of MEPBr (32, 42, 52, and 62 mM), and the corresponding simulation results (red) based on the *Cloud* model for the estimation of k_{et} -Br⁻/Br₂.



Figure S7. The LSVs (black) measured in 0.5 M H₂SO₄ aqueous solutions containing various concentrations of MEPBr (72, 82, 92, and 102 mM), and the corresponding simulation results (red) based on the *Cloud* model for the estimation of k_{et} -Br⁻/Br₂.



Figure S8. The LSVs (black) measured in 0.5 M H₂SO₄ aqueous solutions containing various concentrations of MEMBr (52, 62, 72, 82, 92, and 102 mM), and the corresponding simulation results (red) based on the *Cloud* model for the estimation of k_{et} -Br⁻/Br₂.



Figure S9. The LSVs (black) measured in 0.5 M H₂SO₄ aqueous solutions containing various concentrations of EPyBr (42, 52, 62, and 72 mM), and the corresponding simulation results (red) based on the *Cloud* model for the estimation of k_{et} -Br⁻/Br₂.



Figure S10. The LSVs (black) measured in 0.5 M H₂SO₄ aqueous solutions containing various concentrations of EPyBr (82, 92, and 102 mM), and the corresponding simulation results (red) based on the *Cloud* model for the estimation of k_{et} -Br⁻/Br₂.



Figure S11. The LSVs (black) measured in 0.5 M H₂SO₄ aqueous solutions containing various concentrations of TProABr (20, 30, 40, 50, and 60 mM), and the corresponding simulation results (red) based on the *Cloud* model for the estimation of k_{et} -Br⁻/Br₂.



Figure S12. The LSVs (black) measured in 0.5 M H₂SO₄ aqueous solutions containing various concentrations of TProABr (70, 80, 90, and 100 mM), and the corresponding simulation results (red) based on the *Cloud* model for the estimation of k_{et} -Br⁻/Br₂.



Figure S13. The LSVs (black) measured in 0.5 M H₂SO₄ aqueous solutions containing various concentrations of TBABr (10, 20, 30, 40, and 50 mM), and the corresponding simulation results (red) based on the *Cloud* model for the estimation of k_{et} -Br⁻/Br₂.



Figure S14. The LSVs (black) measured in 0.5 M H₂SO₄ aqueous solutions containing various concentrations of TBABr (60, 70, 80, 90, and 100 mM), and the corresponding simulation results (red) based on the *Cloud* model for the estimation of k_{et} -Br⁻/Br₂.



Figure S15. The LSVs (black) measured in 0.5 M H₂SO₄ aqueous solutions containing various concentrations of TPABr (10, 20, 30, 40, and 50 mM), and the corresponding simulation results (red) based on the *Cloud* model for the estimation of k_{et} -Br⁻/Br₂.



Figure S16. The LSVs (black) measured in 0.5 M H₂SO₄ aqueous solutions containing various concentrations of TPABr (60, 70, 80, and 90 mM), and the corresponding simulation results (red) based on the *Cloud* model for the estimation of k_{et} -Br⁻/Br₂.



Figure S17. The photographs of (a) TBABr and (b) TBABr₃ after the dynamic vapor sorption (DVS) analysis, which is depicted in (c); the graph describes change in mass (%) of TBABr (black) and TBABr₃ (red) powder as humidity changes (blue line) from 0 to 90 %.



Figure S18. The CA measured in 10 mM TBABr solution at 1.2 V for 300 s.



Figure S19. (a) Three dimensional, (b) the corresponding cross-sectional domain of the simulation, and (c) simulated, normalized steady-state voltammograms under the different conditions. IP adsorbed on different UME edge sites.



Figure S20. (a-h) The randomly chosen individual current spikes from a CA measured in a 0.5 M H_2SO_4 aqueous solution containing 50 mM TBABr at a constantly applied potential of 1.2 V for 60 s. The purpose of fitting the bulk electrolysis model to the individual current spikes is to estimate the corresponding radius of an adsorbed hemispherical *H*-TBABr₃ droplet.



Figure S21. DFT-optimized structures for the solvent-separated ion pairs of IL cations with $H \cdots Br$ distance in Å.

Tables

Table S1. Reactions, corresponding parameters, relevant time-dependent diffusion and chemical equations, and initial concentration of the chemical species using finite element analysis (Figure S5).

Reactions in aq.		Para	meters	
phase	$k_{\rm et}$ on Pt UME	ket on Cloud	$E_{ m eq}$	α
$Br \cdot + e^- \rightleftharpoons Br^-$	variable (cm/s)	0.1 (cm/s)	0.76 (V)	0.5
$2Br^{\cdot} \rightarrow Br_2$	$k_{f1} = 50$	$00 (M^{-1}s^{-1})$		
$Br_2 + e^- \rightleftharpoons Br_2^-$	0.1 (cm/s)	0.1 (cm/s)	0.72 V	0.5
H _{Cloud}		0.5	6 (V)	
$d_{\text{UME-Cloud}}$		var	riable	
	The relevant time-depe	endent diffusion equat	ions	
(1) $\frac{\partial C_{Br}}{\partial t} = D_{Br} \cdot \left[\frac{\partial^2 C_{I}}{\partial r} \right]$	$\frac{Br\cdot}{2} + \frac{1}{r}\frac{\partial C_{Br\cdot}}{\partial r} + \frac{\partial^2 C_{Br\cdot}}{\partial z^2} -$	$\frac{1}{2}k_{f1}C_{Br}.^2$		
(2) $\frac{\partial C_{Br}}{\partial t} = D_{Br} \left[\frac{\partial^2}{\partial t} \right]$	$\frac{C_{Br}}{\partial r^2} + \frac{1}{r} \frac{\partial C_{Br}}{\partial r} + \frac{\partial^2 C_{Br}}{\partial z^2} \right]$			
(3) $\frac{\partial C_{Br_2}}{\partial t} = D_{Br_2} \left[\frac{\partial^2 C_B}{\partial t} \right]$	$\frac{C_{Br_2}}{r^2} + \frac{1}{r} \frac{\partial C_{Br_2}}{\partial t} + \frac{\partial^2 C_{Br_2}}{\partial z^2} \right]$	$+\frac{1}{2}k_{f1}C_{Br}^{2}$		
(4) $\frac{\partial C_{Br_2}}{\partial t} = D_{Br_2}$	$\frac{\partial^2 C_{Br_2} \cdots}{\partial r^2} + \frac{1}{r} \frac{\partial C_{Br_2} \cdots}{\partial t} + \frac{\partial^2 C_{Br_2} \cdots}{\partial t} + \partial^2 C_{Br$	$\frac{\partial^2 C_{Br_2} - \cdot}{\partial z^2}$		
The in	itial condition, complet	ting the definition of the	he problem	

t = 0, all r, z; $C_{Br} = 0$, = variable, $C_{Br_2,Br_2} = 0$, $D_{Br,Br} = 1.58 \ge 10^{-5}$, $D_{Br_2} = 1.18 \ge 10^{-5}$, $D_{Br_2} = 1.00 \ge 10^{-5}$ cm²/s

Table S2. Reactions, corresponding parameters, relevant time-dependent diffusion and chemical equations, and initial concentrations of chemical species using finite element analysis (Figure 5).

Reactions in aq. Phase	Parameters
$1/2Br + e^- \Rightarrow Br^-$	$k_{et} = 0.1 \text{ cm/s}$ $E_{eq} 0.9 \text{ V}, \alpha = 0.5$

The relevant time-dependent diffusion equations

(1)
$$\frac{\partial C_{Br}}{\partial t} = D_{Br} \cdot \left[\frac{\partial^2 C_{Br}}{\partial r^2} + \frac{1}{r} \frac{\partial C_{Br}}{\partial r} + \frac{\partial^2 C_{Br}}{\partial z^2} \right]$$

(2)
$$\frac{\partial C_{Br}}{\partial t} = D_{Br} - \left[\frac{\partial^2 C_{Br}}{\partial r^2} + \frac{1}{r} \frac{\partial C_{Br}}{\partial r} + \frac{\partial^2 C_{Br}}{\partial z^2} \right]$$

The initial condition, completing the definition of the problem

t = 0, all r, z;
$$C_{Br} = 0, C_{Br} = 50 \times 10^{-3} M, D_{Br} = 1.58 \times 10^{-5}$$

IMED)]+ [Ɗ"]-			IMED)]+ [D _m]-		
	J∥[DI]						
28	3			30)		
~				~			
С	-1.101/04	-0.291425	1.066920	С	-1.957739	-1.799544	1.158100
Ν	-1.597483	0.271320	-0.232244	С	-1.521831	-0.342583	1.246288
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н	-1 603253	0.007710	1 889063	н	-0.497032	-0 253188	1 597526
и П	0.662321	2 2 2 0 / 2 /	1.546630	и П	-0.477032	0.280500	1.838762
11	-0.002321	-2.339434	1.340030	11	-2.1/101/	0.289309	1.050175
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п	-0.418/3/	-2./1/55/	-0.780288	П	-3.032870	-1.8851//	1.2/2009
Н	-2.166536	-2.605512	-0.903604	Н	-0./61488	-3.045530	-0.196939
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Н	0.061787	1.475468	-0.533246	Н	-0.629735	1.514853	-1.446394
Н	-3.434295	0.868562	0.609212	Н	0.317015	1.089247	-0.014866
Н	-3.393435	0.797525	-1.171726	Н	-3.398559	1.063199	0.018446
Н	-3.465148	-0.698175	-0.219078	Н	-2.855036	0.776464	-1.655891
Η	-0.914779	3.609261	0.150313	Н	-3.442759	-0.572763	-0.662336
Η	-2.494961	2.888375	0.416226	Н	-0.457041	3.372292	0.198279
Н	-1.127430	2.451226	1.454823	Н	-2.138269	2.919112	-0.040614
Н	2.183130	-1.768969	0.740017	Н	-1.272832	2.404567	1.418274
0	2.030256	-2.611669	1.213522	Н	2.067401	-1.305767	1.297639
Н	1.381952	-3.077747	0.679035	Н	0.901586	-2.283237	1.488642
				Br	2.849653	1.014498	1.353430
				Br	2.996272	-0.005320	-3.673083
IMEN	/[] ⁺ ∥ [Br] ⁻			IMEN	/] ⁺ ∥ [Br ₂] ⁻		
20))			31			
2,	, ,			51	L		
C	1 002360	1 473305	0 797060	0	3 415770	2 760470	0 227522
C	2 2 2 2 1 1 5	-1.4/3303	0.797000	C	2 005205	-2.700+70	0.227322
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IN C	-1.838390	-0.3/9228	0.300703		-3.223334	-0.353548	0.11/21/
C	-1.300049	-1./33207	0.092334	IN C	-1./09818	-0.393410	0.493505
C	-2.191069	-2.839282	0.708441	C	-1.226919	-1./42281	0.11804/
Ö	-3.564246	-2.717699	0.348/05	C	-2.070065	-2.861969	0.679237
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Н	-1.447706	0.454524	-1.357701	Н	-2.246379	-0.766402	2.523452
Н	-0.551141	-0.311095	2.150586	Н	-1.620939	1.595971	-0.082191
Н	-1.845449	0.915835	2.148815	Н	-1.234102	0.423351	-1.346287
Н	-0.330595	-1.826980	0.398573	Н	-0.564877	-0.313914	2.214370

Table S3. The tabulated Cartesian coordinates of the optimized geometries associated with Figure S21.

Н	-1.440989	-1.765492	-0.993353	Н	-1.854413	0.914471	2.121843
Н	-2.098950	-2.858696	1.793935	Н	-0.211823	-1.815081	0.490548
Н	-1.838148	-3.784633	0.310842	Н	-1.225133	-1.784389	-0.968537
Н	-4 096642	-1 452196	1 885935	Н	-2.052605	-2.875253	1 768399
н	-5 117096	-1 413072	0.446627	н	-1 674465	-3 801769	0.310719
и П	2 411022	0.220606	0.991027	и П	4 055058	1 522406	1 746249
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	1+ [D]-				1+ II [D.,]-		
	ן [םנ] ג				.j [D13] 5		
52	,			5.	<i>,</i>		
С	-1 162903	-1 849283	3 324085	C	-1 177666	-1 750357	3 387676
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