Supporting Information

Enhanced glyphosate removal by montmorillonite in the presence of Fe(III)

Zhong Ren[a,b], *Yuanhua Dong*[a,]*, *Yun Liu*[a,]*

[a]Key Laboratory of Soil Environment and Pollution Remediation, Institute of Soil Science, Chinese Academy of Sciences, Nanjing 210008, PR China

[b]Graduate University of Chinese Academy of Sciences, 100049 Beijing, China

* Corresponding author. Tel.: +86 25 86881317; +86 25 86881370; fax: +86 25 86881000. E-mail address: yhdong@issas.ac.cn (Yuanhua Dong); yliu@issas.ac.cn (Yun Liu).

Table S1. Main chemical compositions of MT.

Na_2O	MgO	Al_2O_3	SiO_2	K_2O	CaO	Fe_2O_3
0.09 %	5.24 %	16.64 %	58.19 %	0.10 %	2.83 %	4.29 %

Table S2. Deprotonation constants of PMG in Na(Cl) ionic medium (T = 25°C)

Reaction	Log10β		
	I = 0 M	I = 0.1 M	I = 0.6 M
H3L = H2L−+H+	−2.44	−2.25±0.006	−2.17±0.01
H3L = HL2−+2H+	−8.33	−7.70±0.01	−7.39±0.01
H3L = L3−+3H+	−19.14	−17.61±0.03	−17.35±0.05

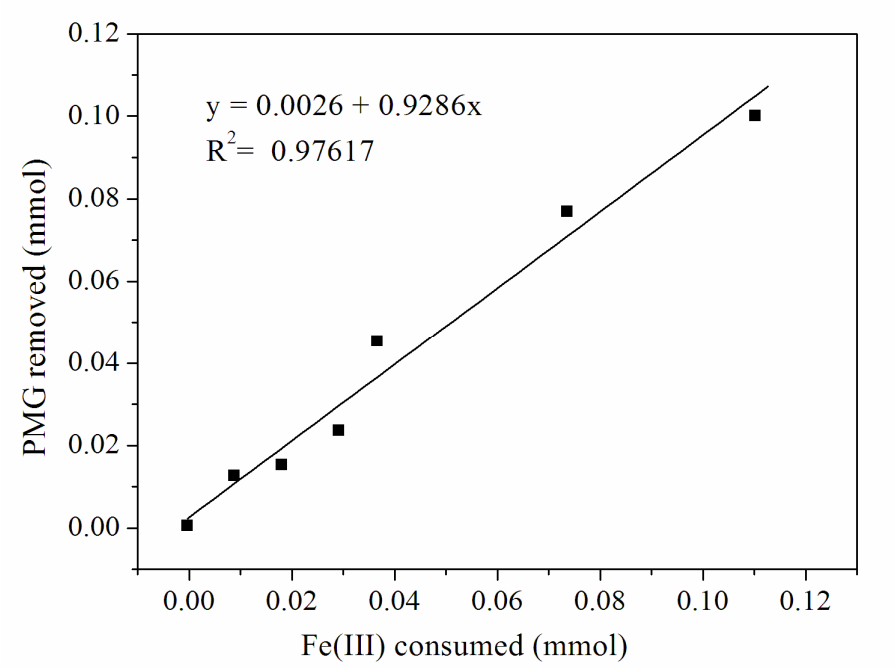

Figure S1**.** Linear fit of amount PMG removed versus consumed Fe(III) added.

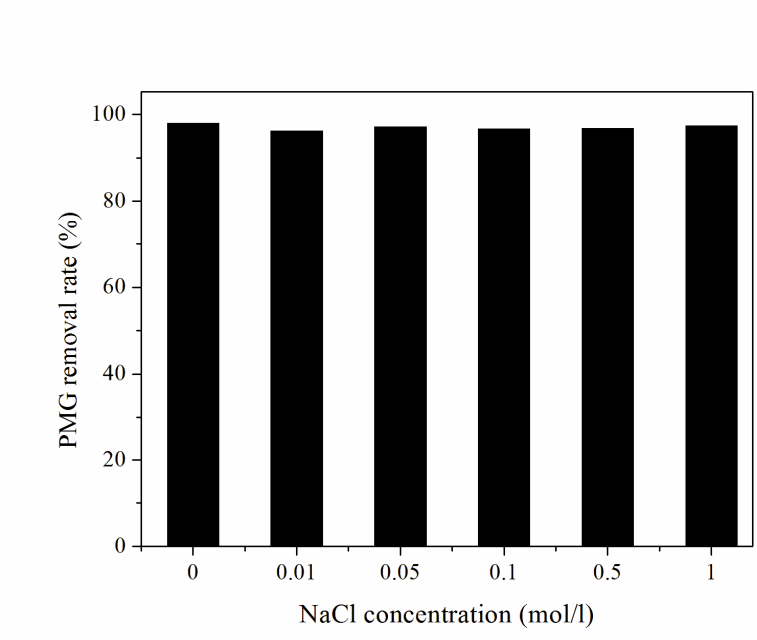

Figure S2**.** Effect of ion strength on removal of PMG by MT.

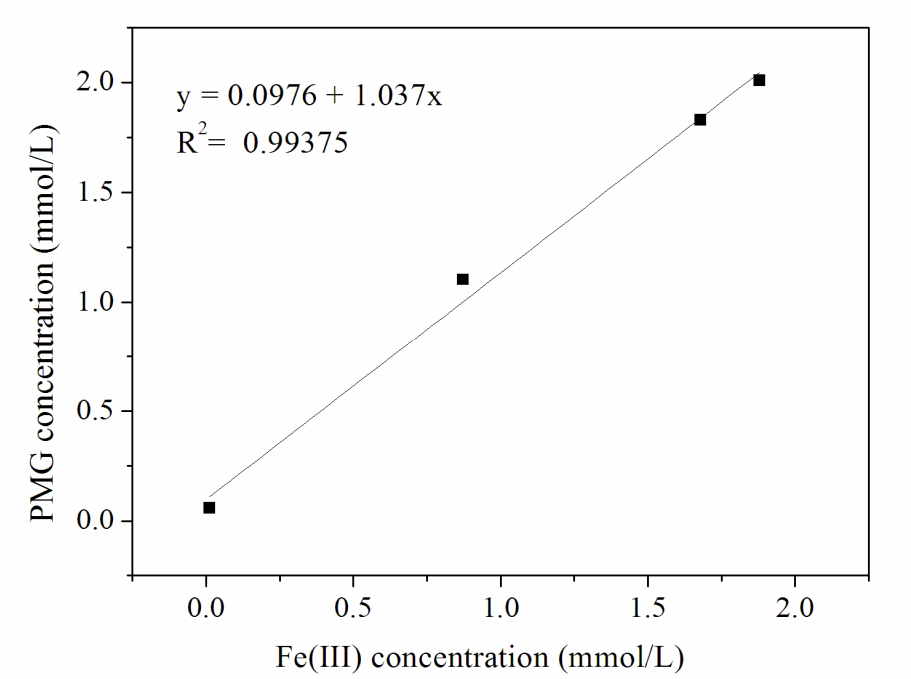

Figure S3**.** Relationship between PMG and Fe(III) concentration.

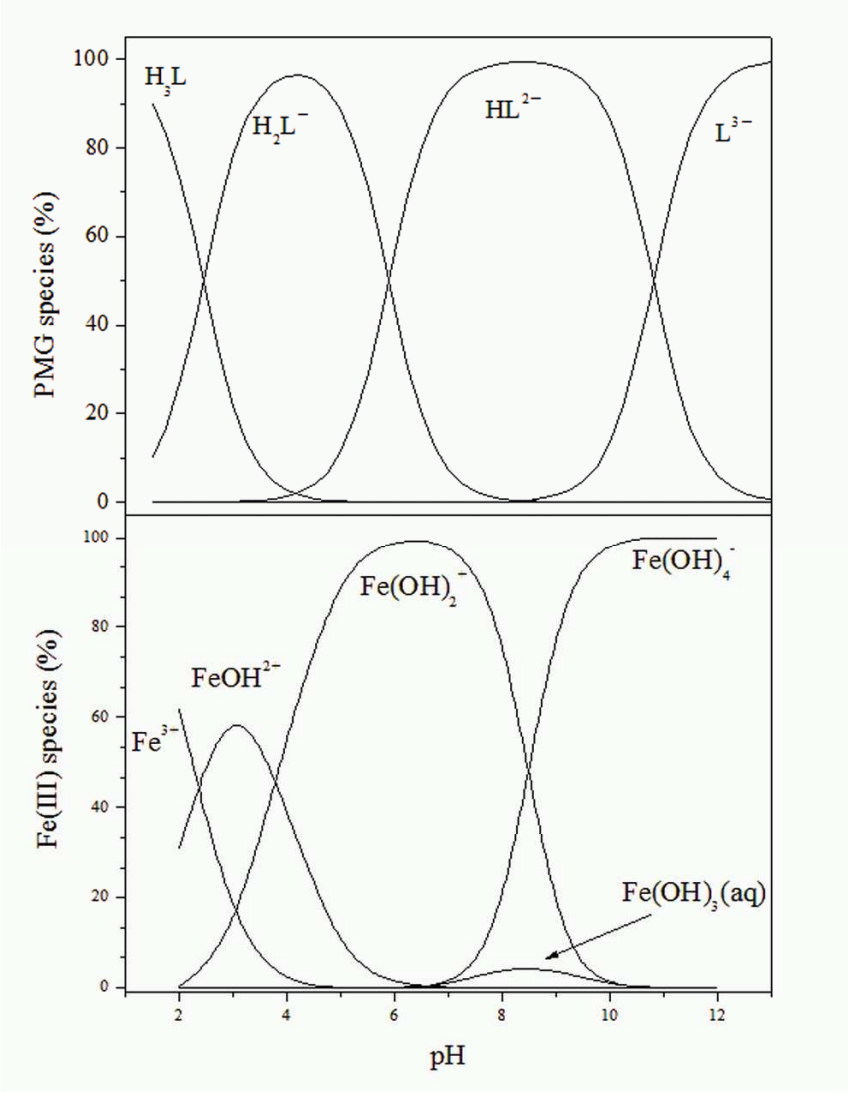

Figure S4**.** Speciation of Fe(III) and PMG in aqueous solution.