

Supporting information

Solubility and Thermodynamic Analysis of Tenoxicam in Different Pure Solvents at Different Temperatures

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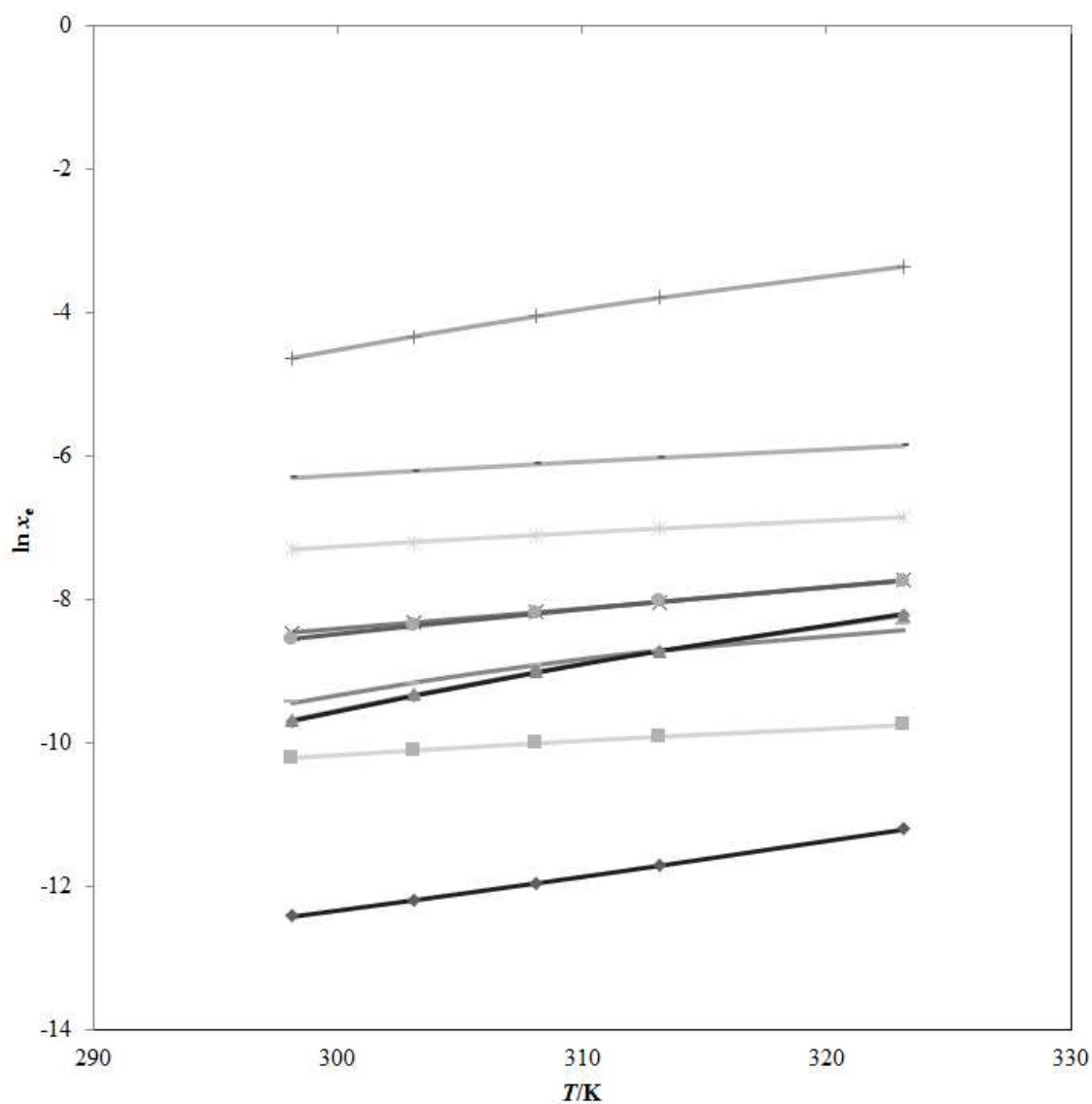


Figure S1 Correlation of experimental solubilities of crystalline TNX with the modified Apelblat model in different pure solvents at $T = (298.15 \text{ to } 323.15) \text{ K}$; ◆ water, ■ ethanol, ▲ IPA, × EG, ✱ EA, ● PG, + PEG-400, - Transcutol, — 1-butanol and ◆ 2-butanol (solid lines represent the calculated solubilities of TNX calculated by the modified Apelblat model)

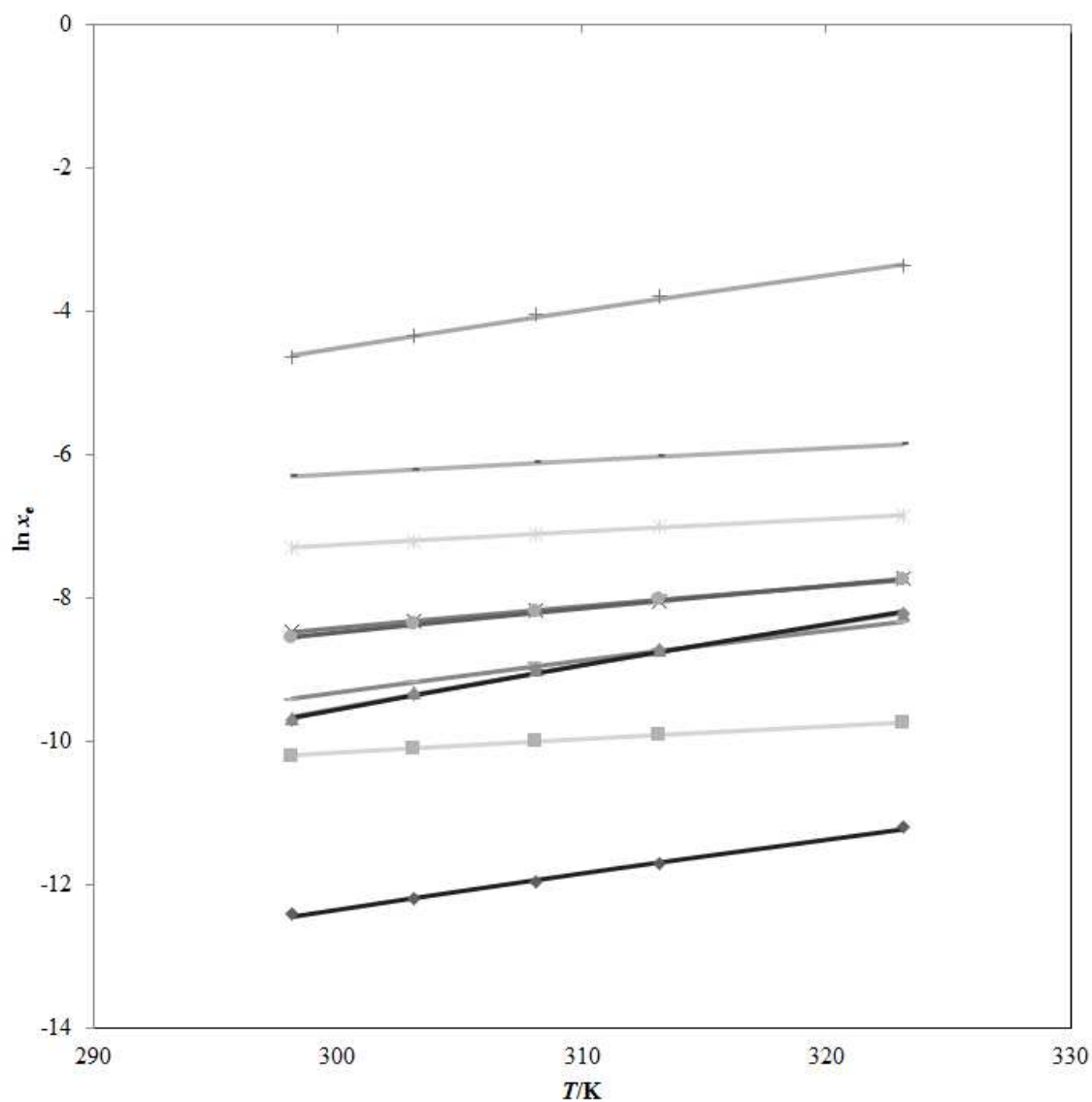


Figure S2 Correlation of experimental solubilities of crystalline LNX with Van't Hoff model in different pure solvents at $T = (298.15 \text{ to } 323.15) \text{ K}$; \blacklozenge water, \blacksquare ethanol, \blacktriangle IPA, \times EG, \ast EA, \bullet PG, $+$ PEG-400, $-$ Transcutol, $-$ 1-butanol and \blacklozenge 2-butanol (solid lines represent the calculated solubilities of TNX calculated by Van't Hoff model)

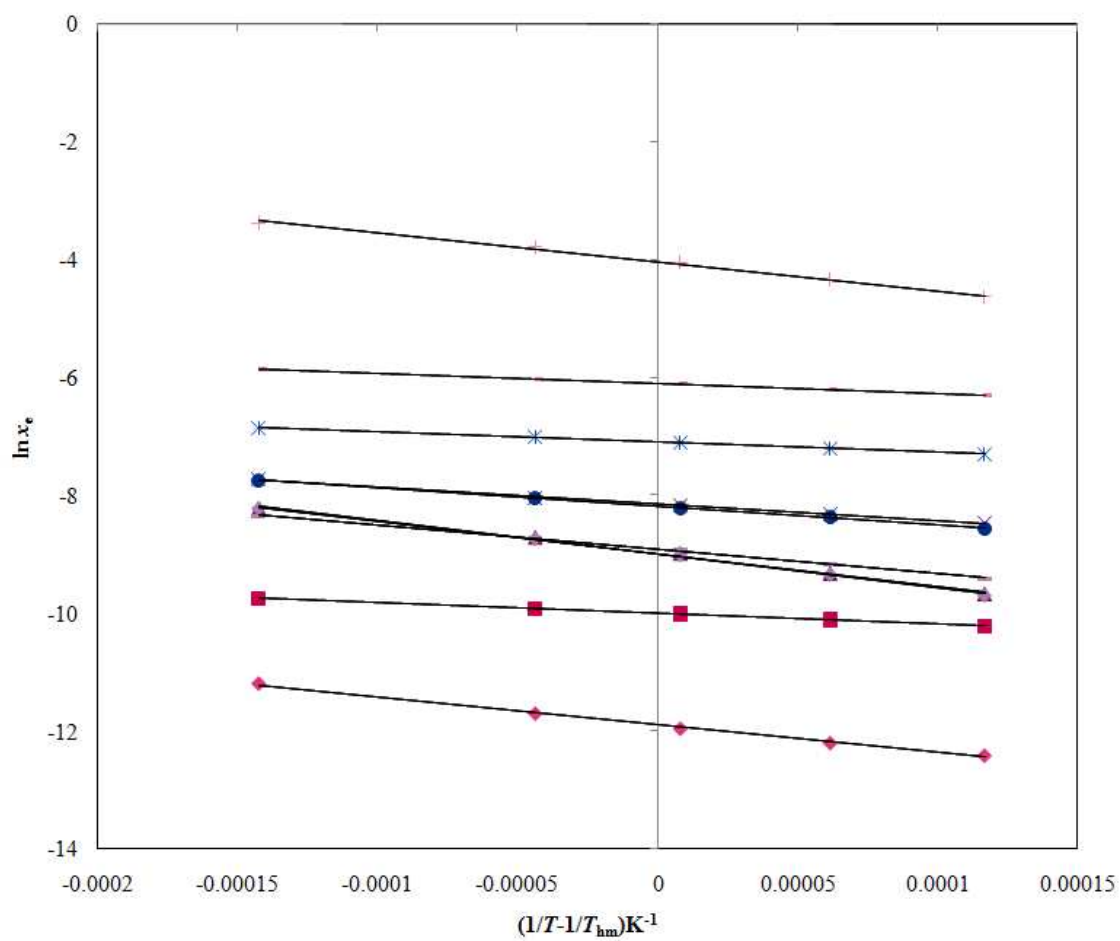


Figure S3 Van't Hoff plots for crystalline TNX in different pure solvents; \blacklozenge water, \blacksquare ethanol, \blacktriangle IPA, \times EG, \ast EA, \bullet PG, $+$ PEG-400, $-$ Transcutol, $-$ 1-butanol and \blacklozenge 2-butanol