

SUPPORTING INFORMATION

Reactions of 1,3-diketones with a dipeptide isothiazolidin-3-one: toward agents that covalently capture oxidized protein tyrosine phosphatase 1B (PTP1B)

Kasi Viswanatharaju Ruddraraju[‡], Zachary D. Parsons[‡], Elizabeth M. Llufrío[‡], Natasha L.

Frost[‡], and Kent S. Gates^{‡,§,*}

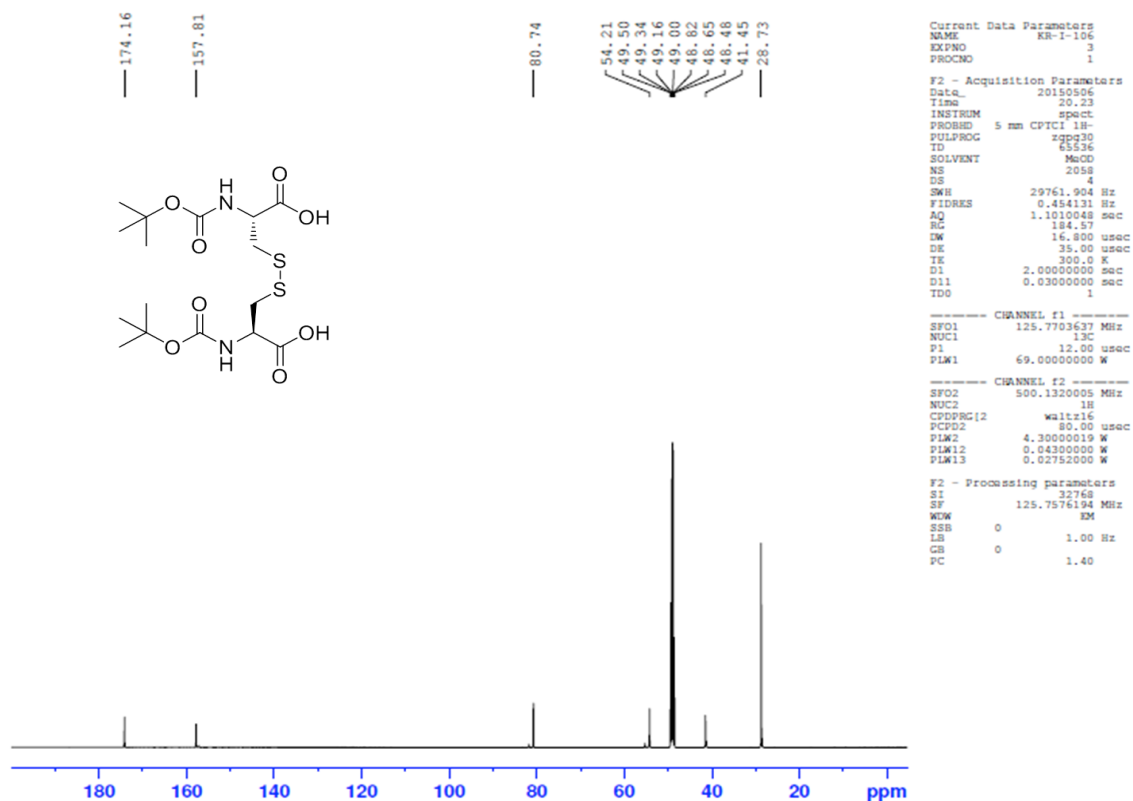
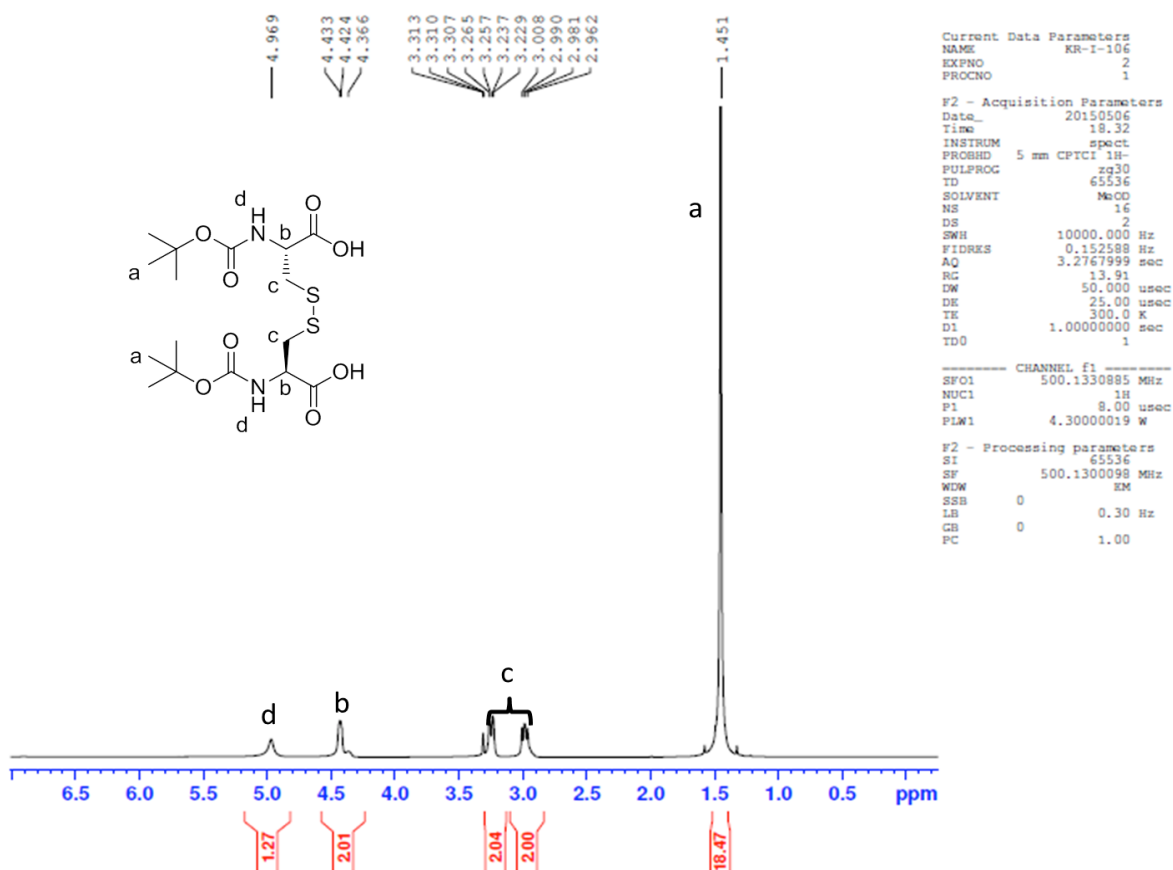
University of Missouri, Departments of [‡]Chemistry and [§]Biochemistry,
Columbia, MO 65211.

*Address correspondence to: gatesk@missouri.edu

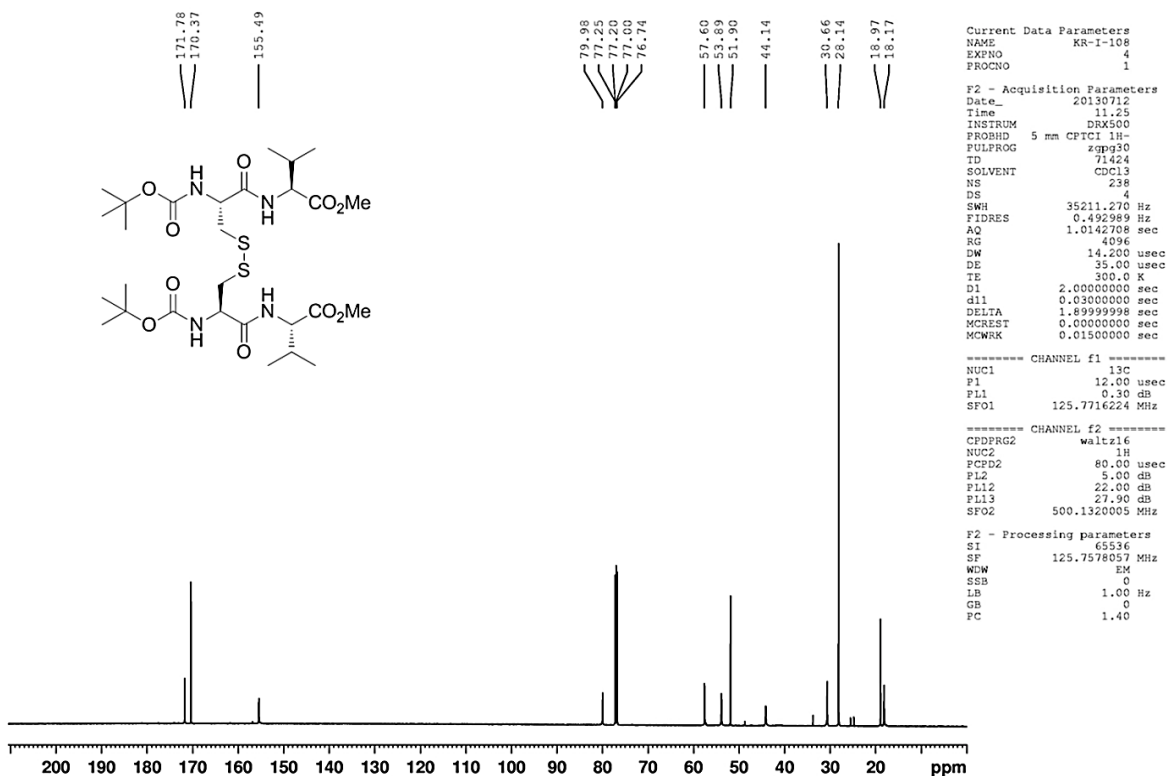
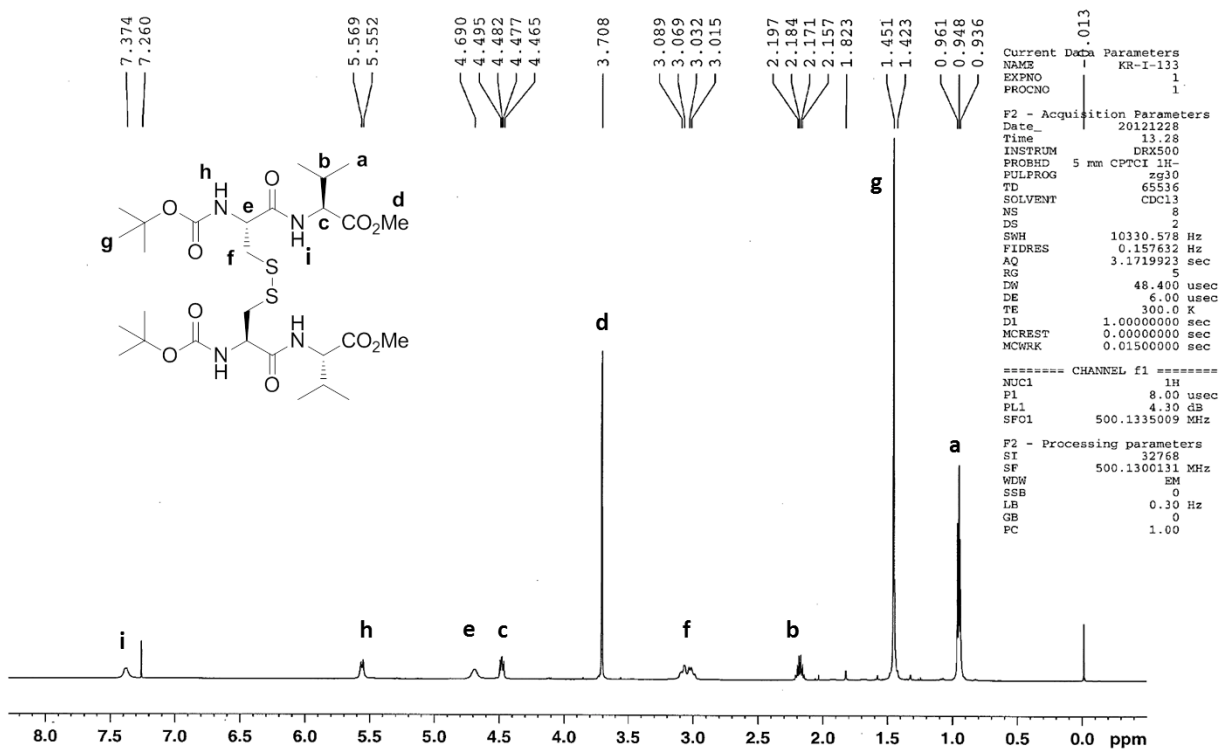
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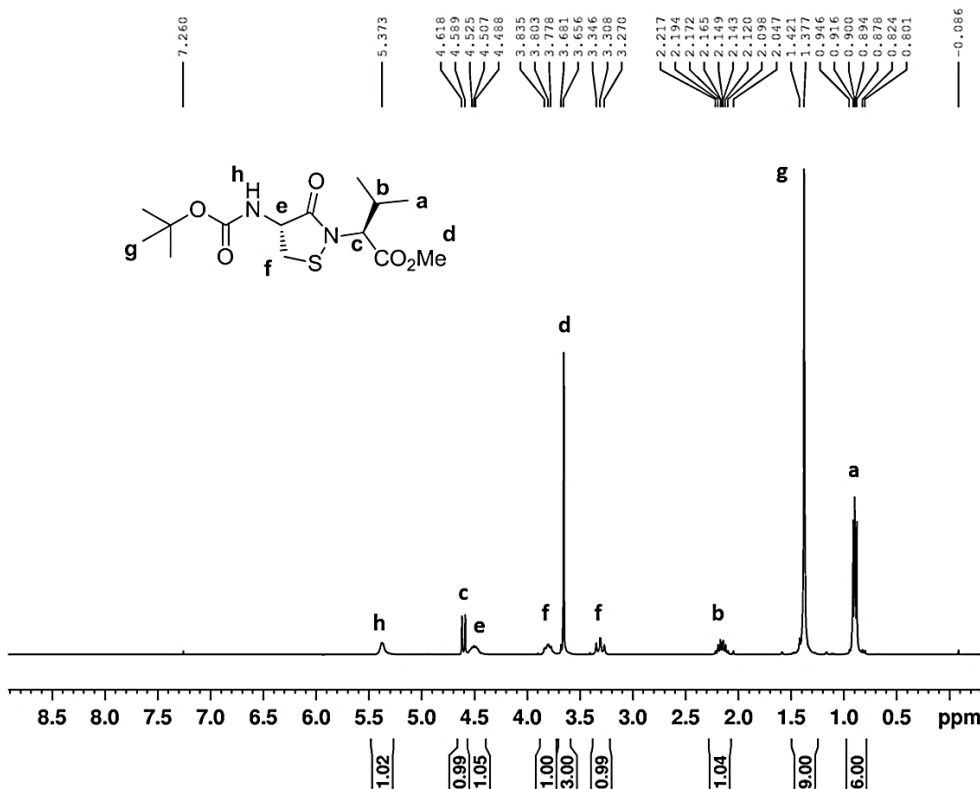
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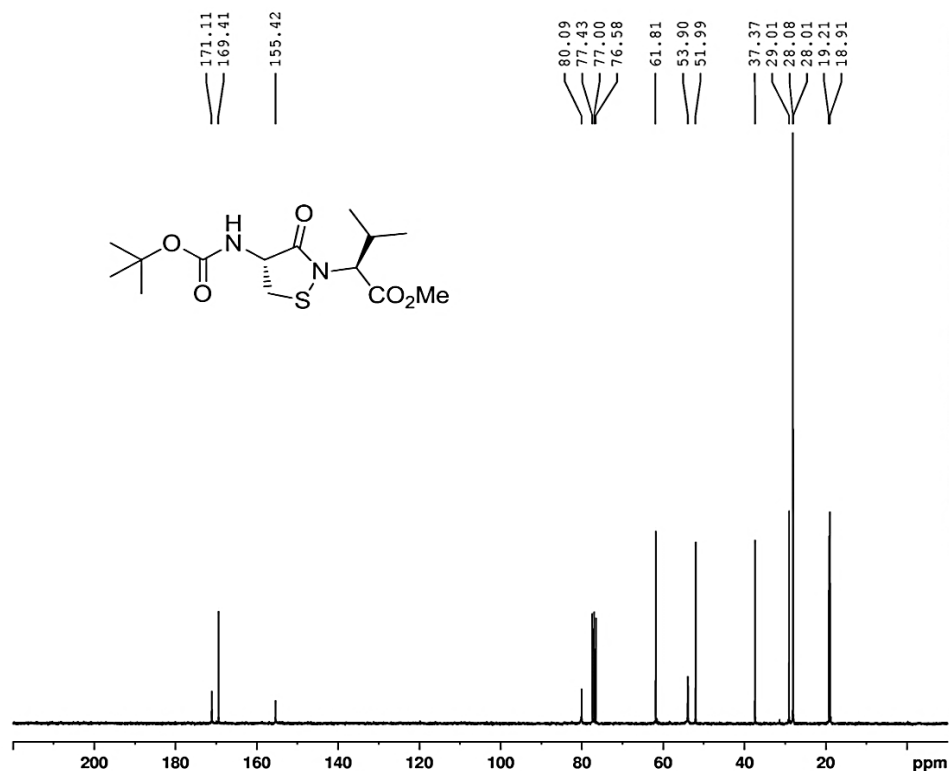


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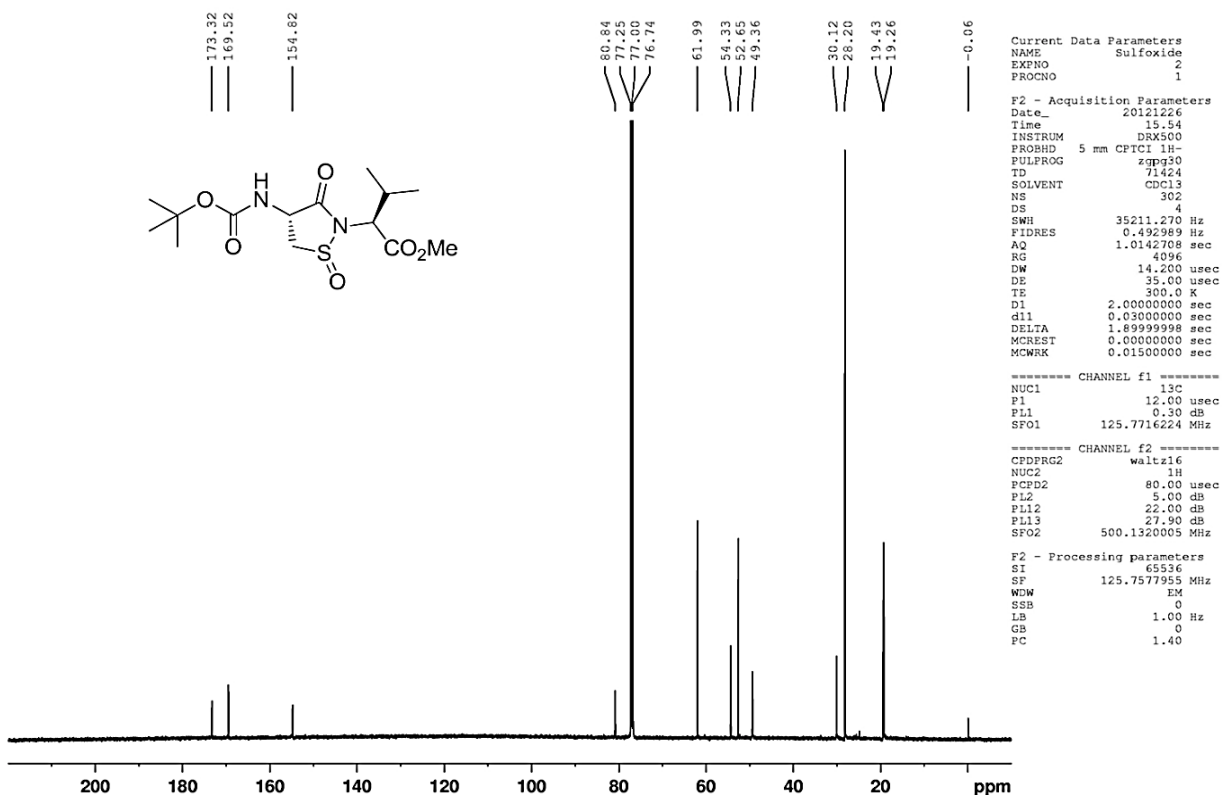
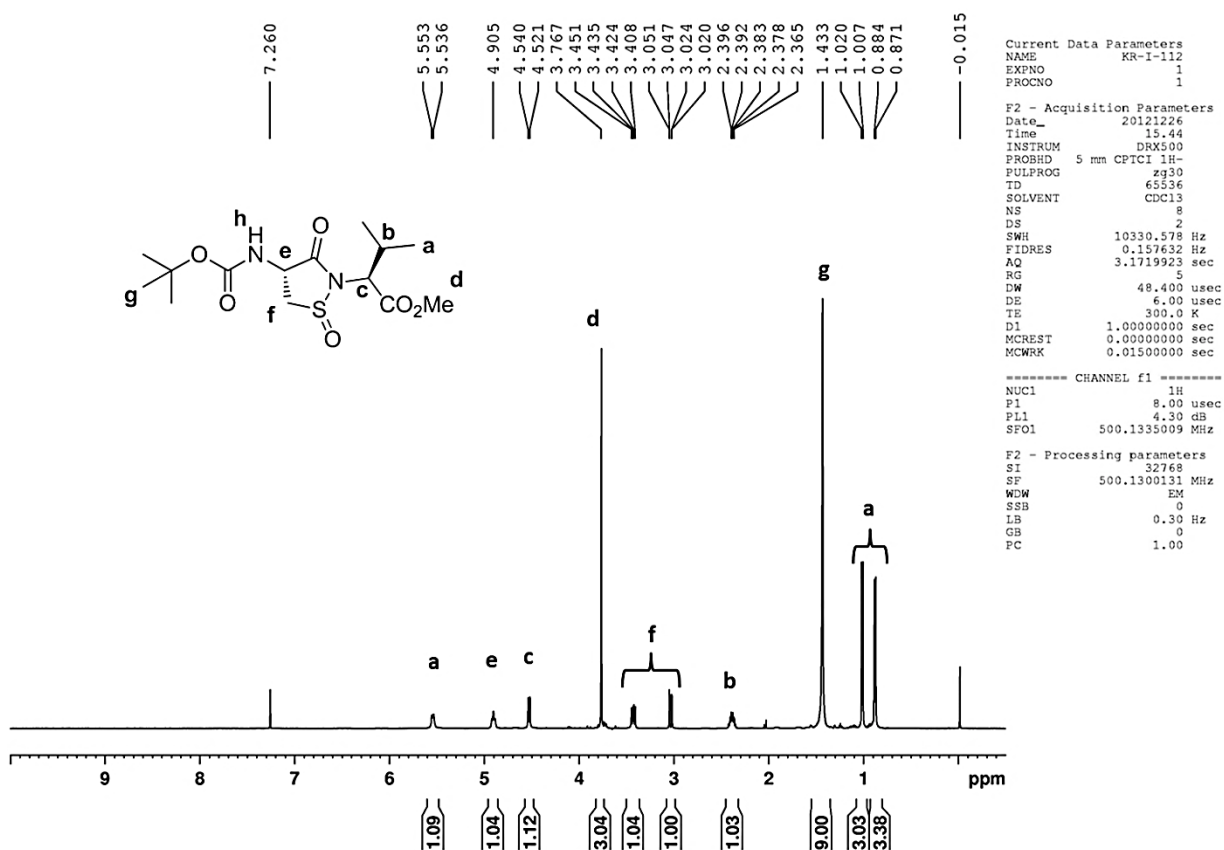
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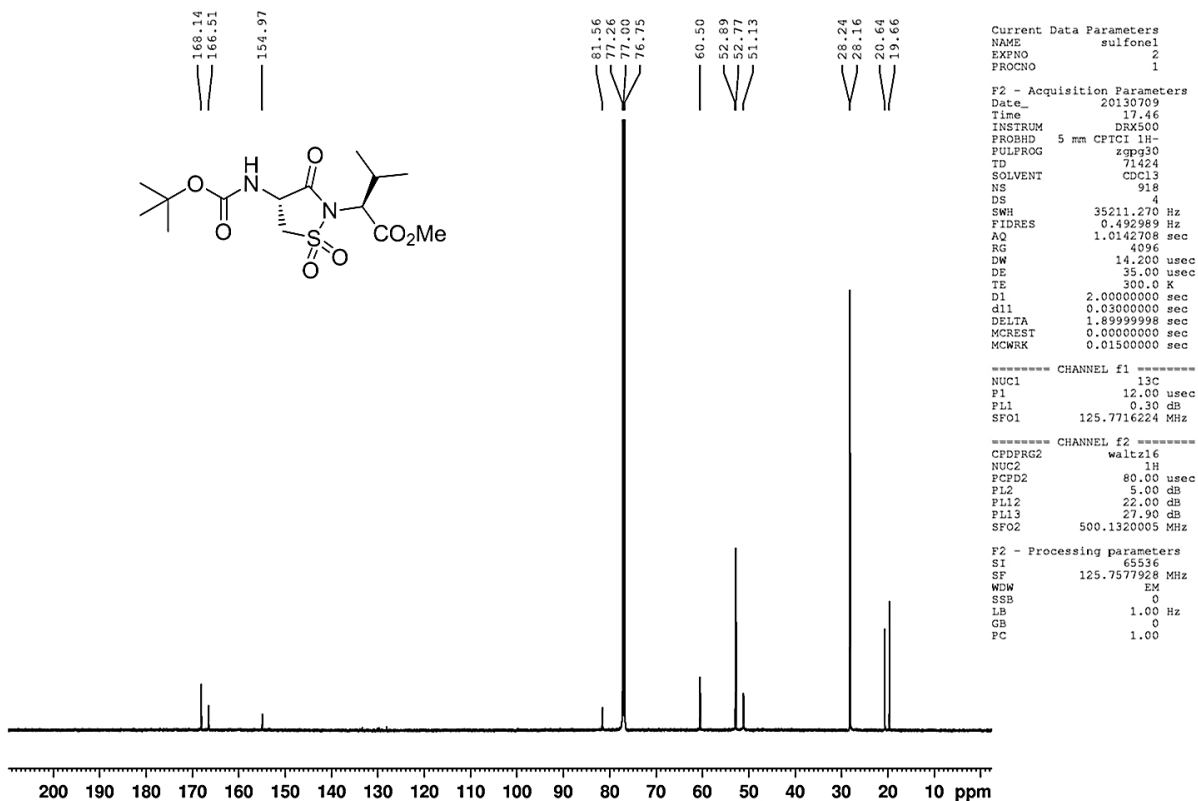
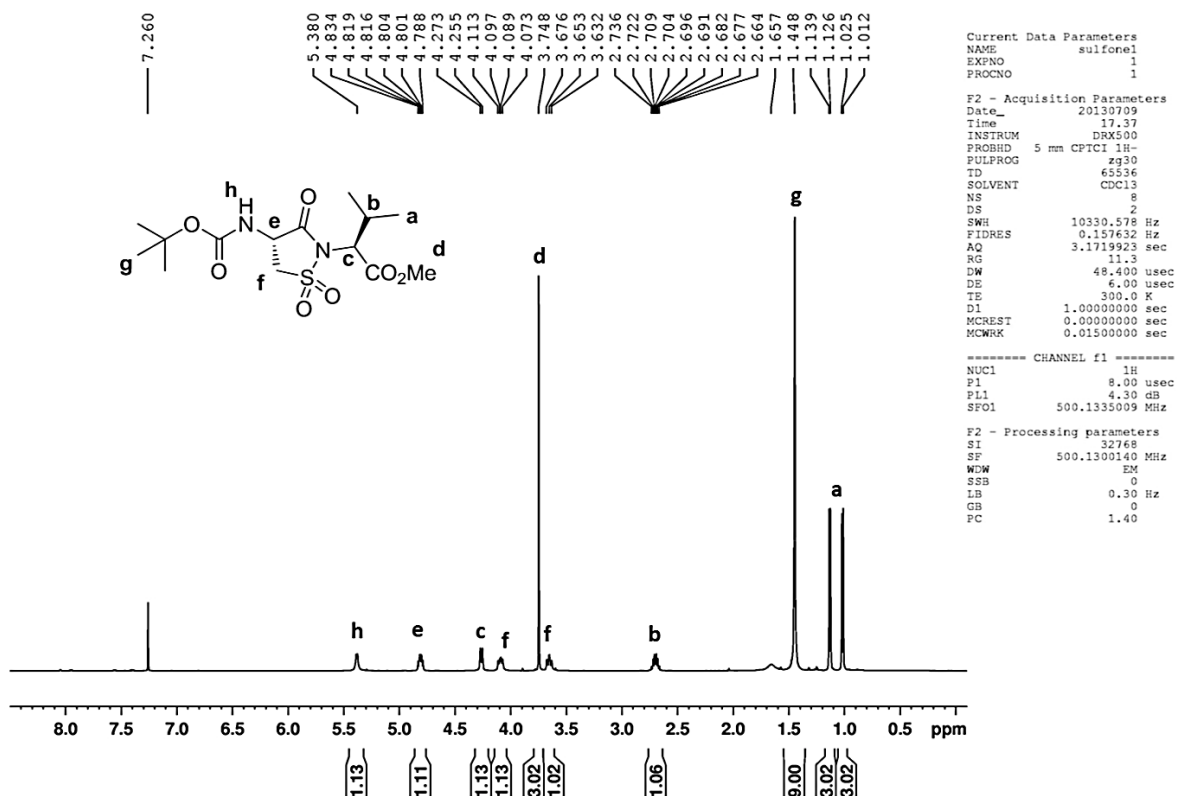
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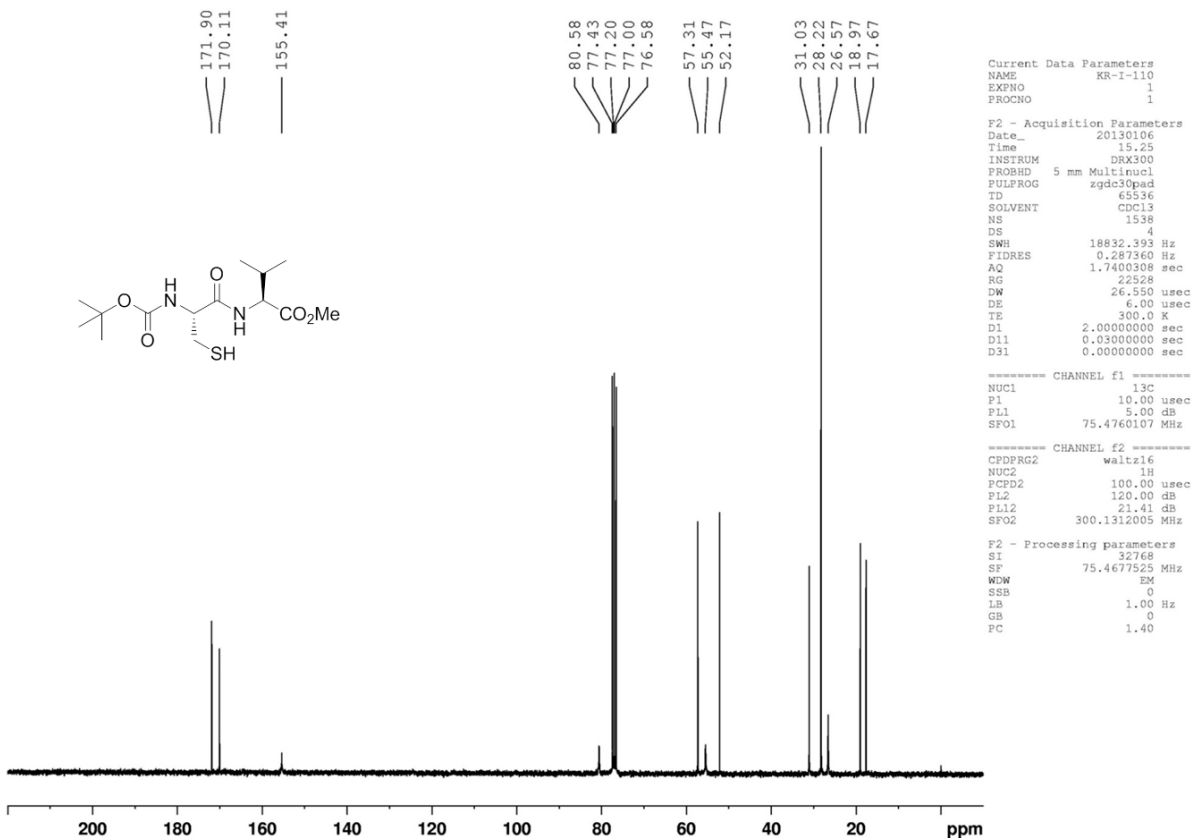
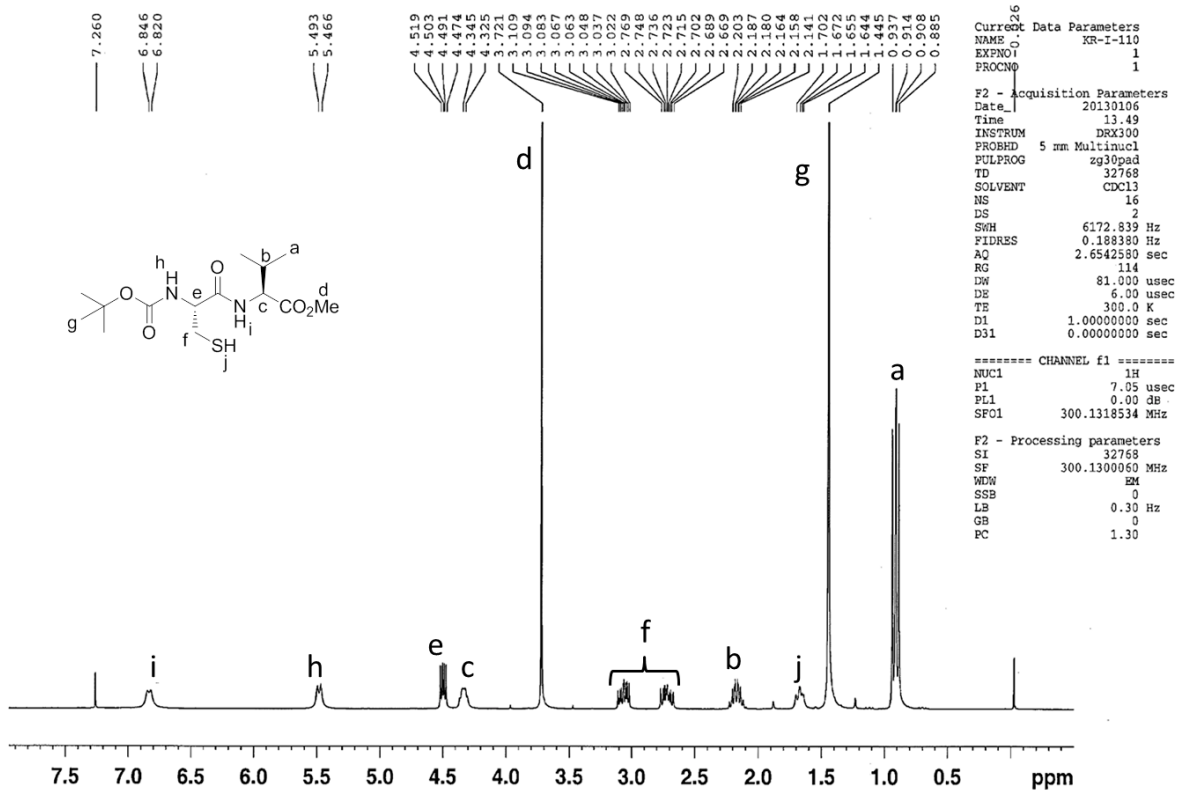
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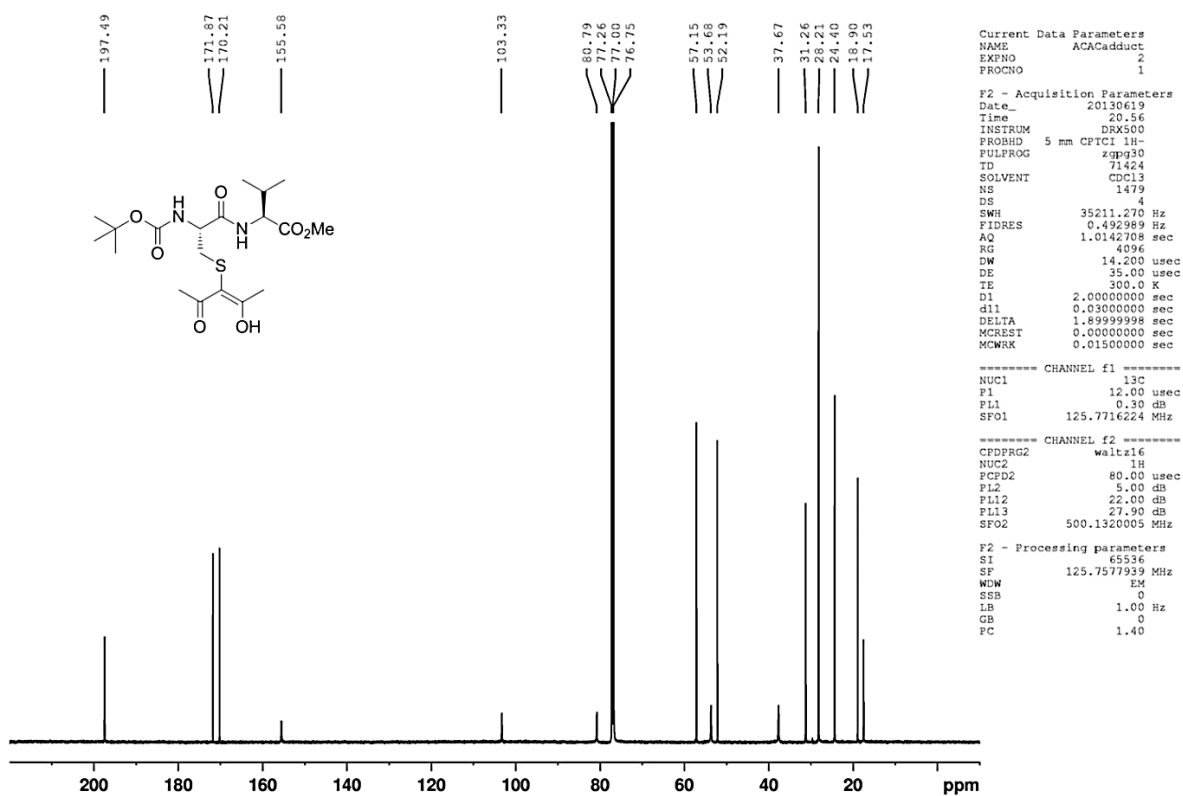
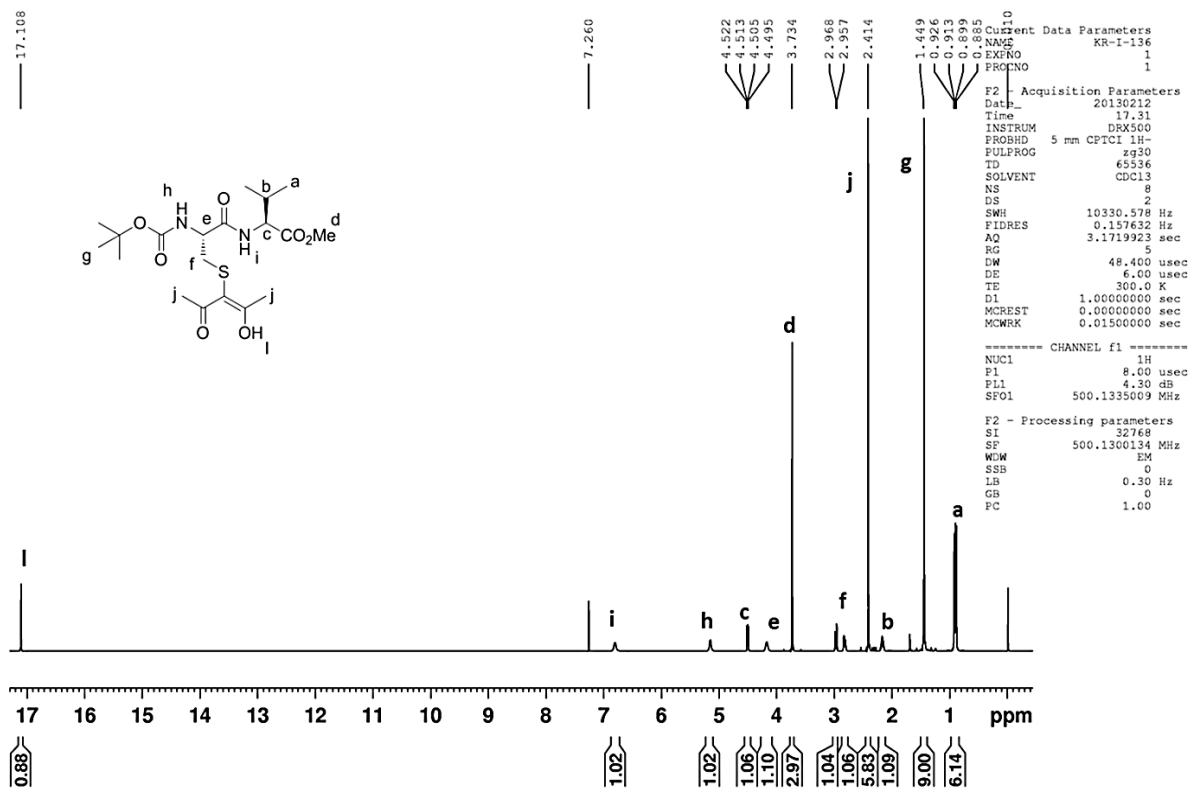
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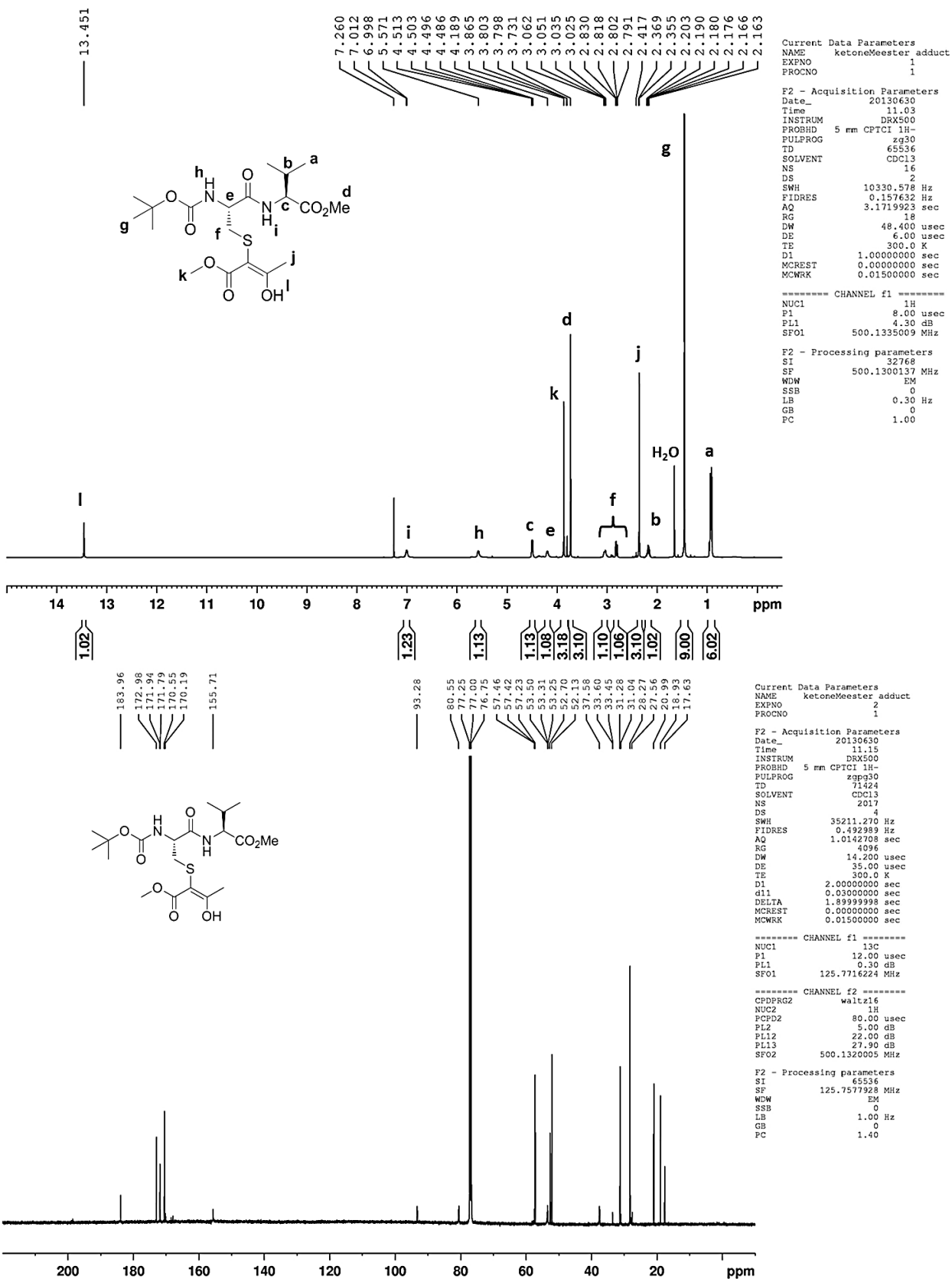
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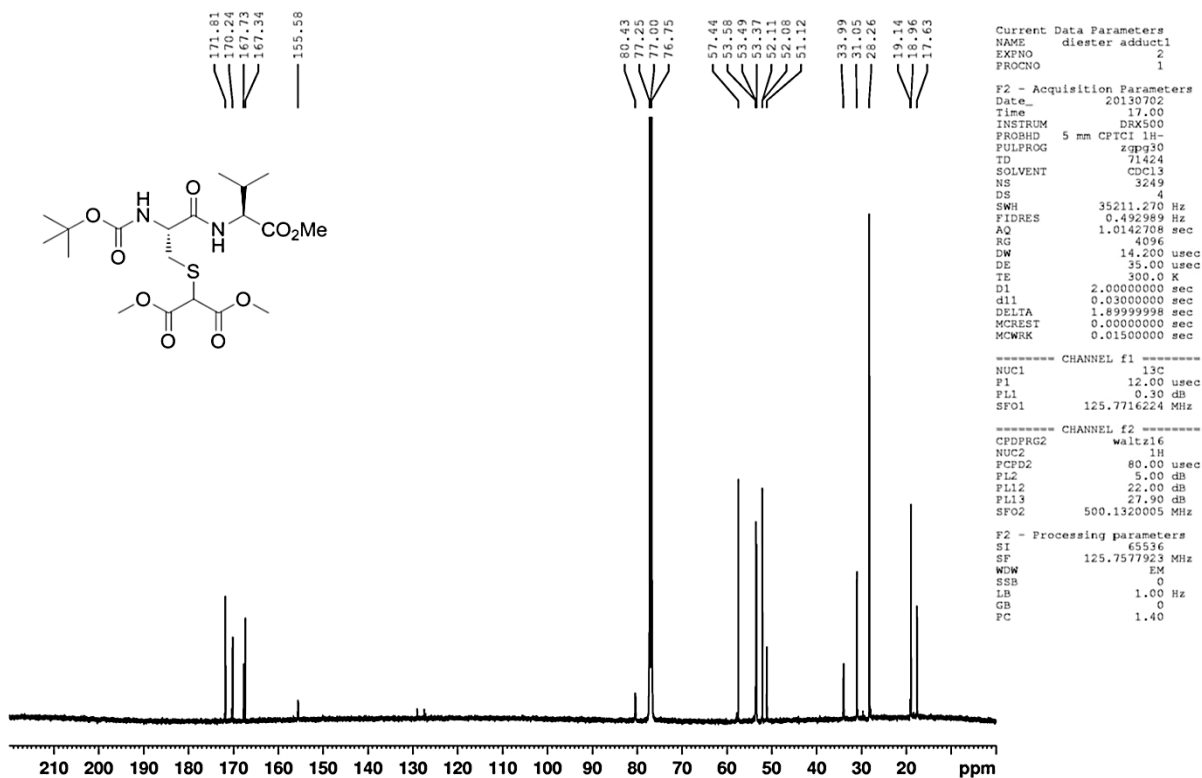
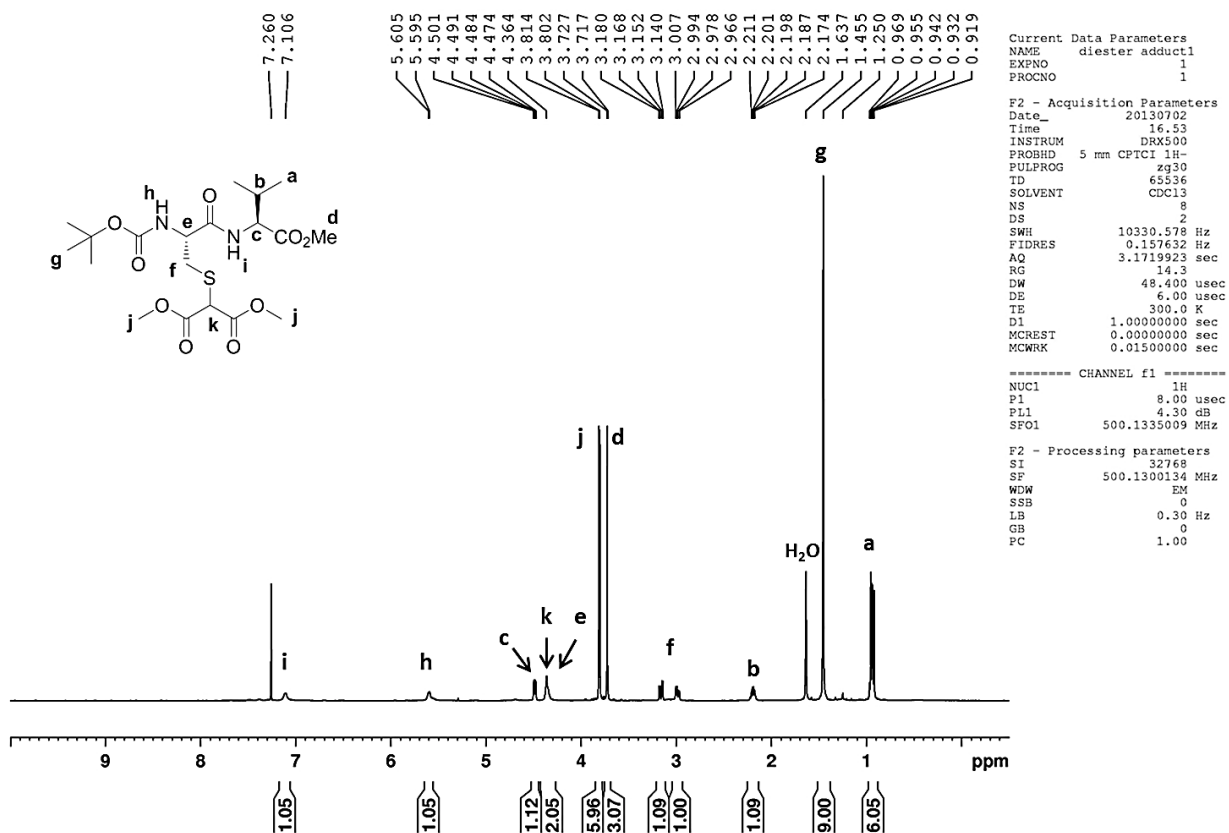
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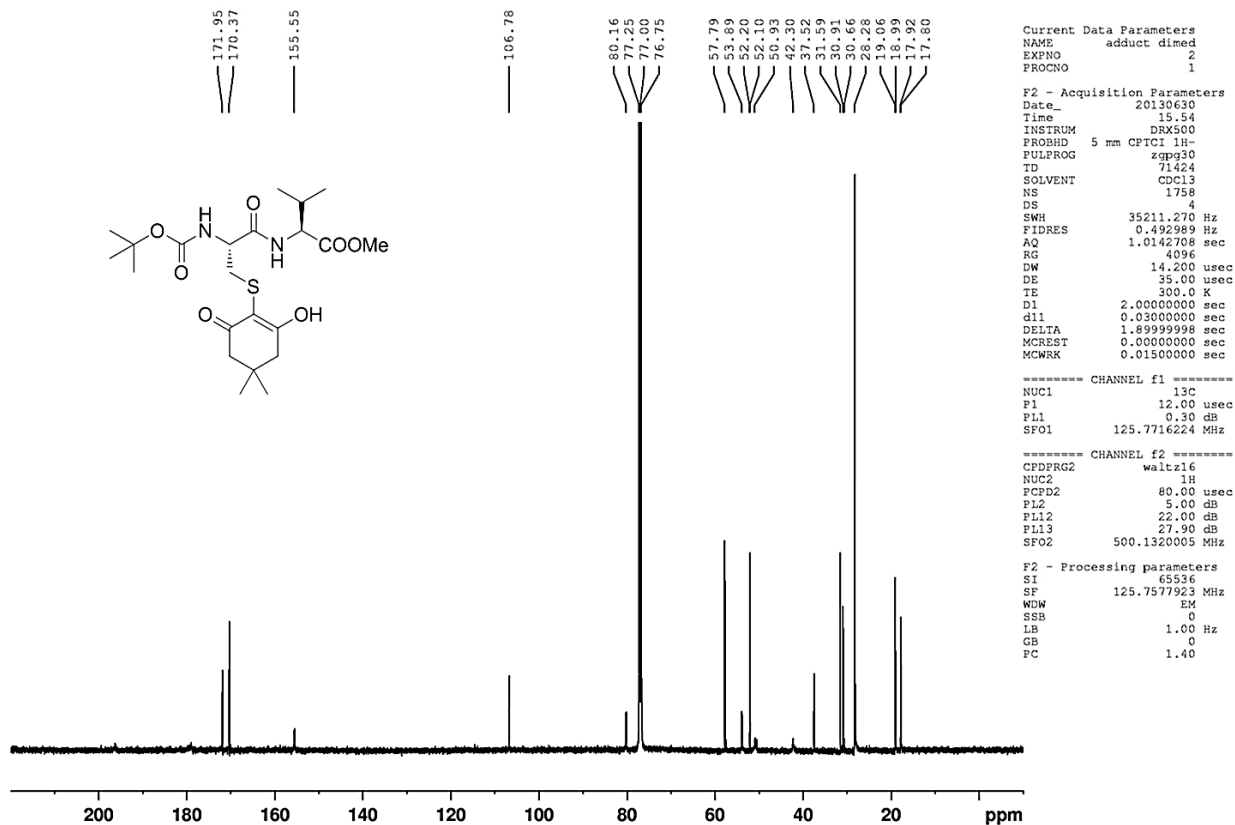
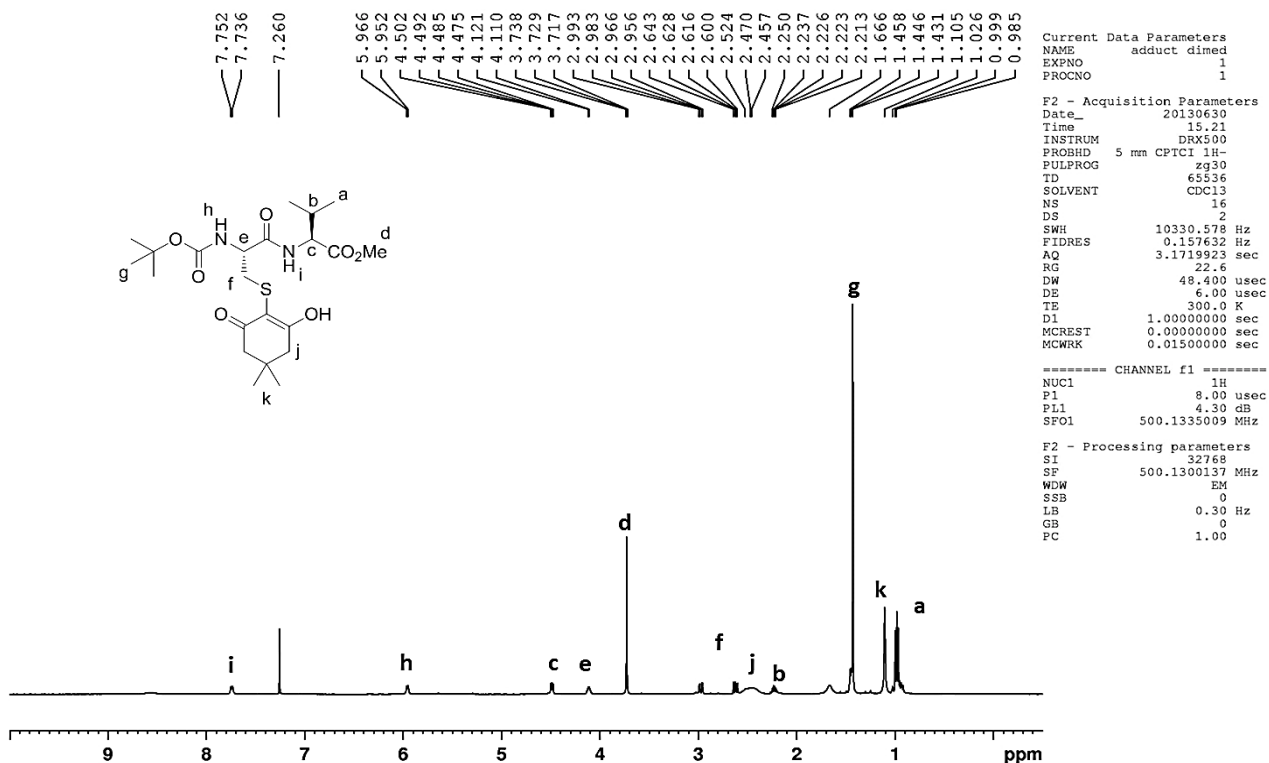
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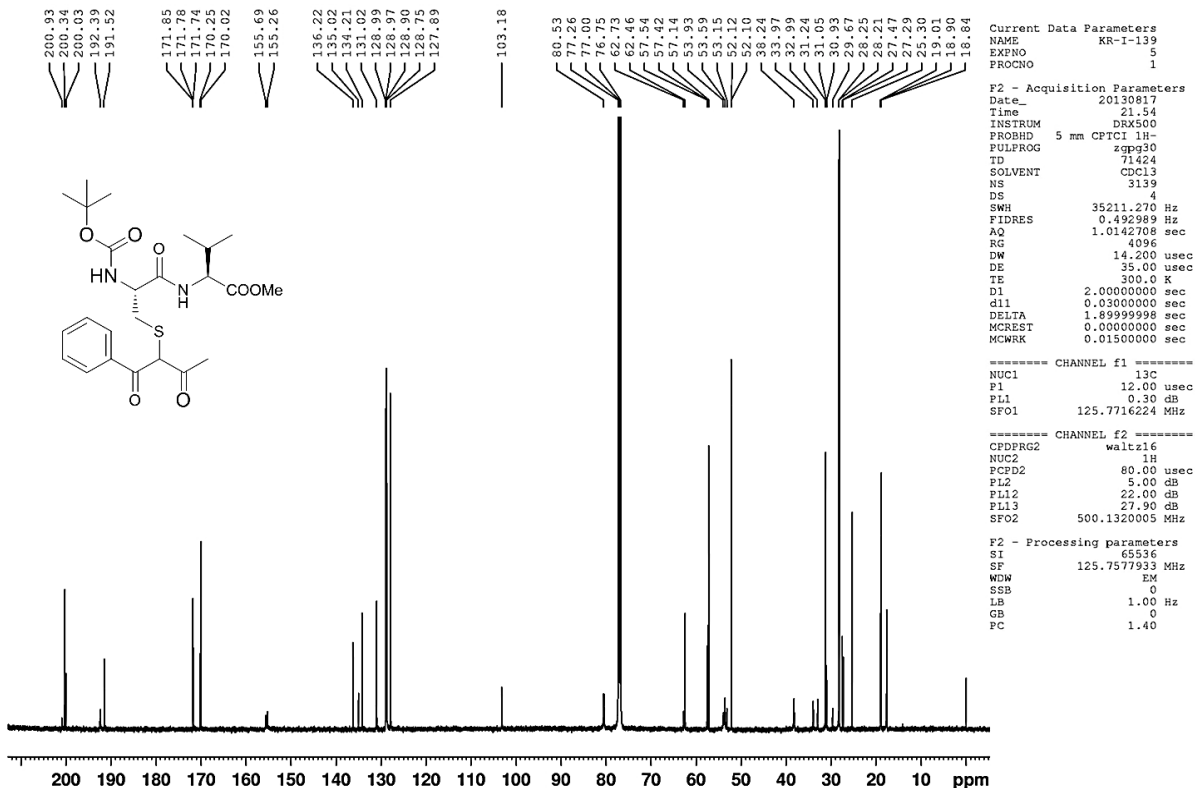
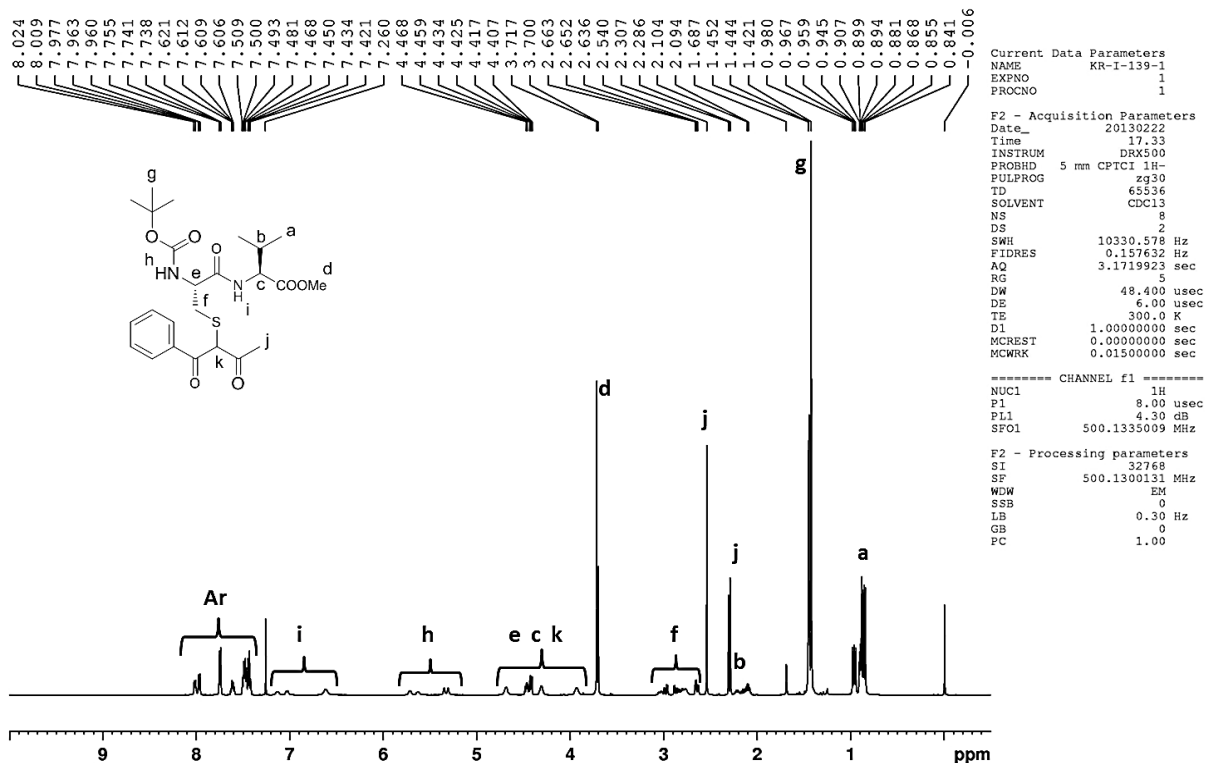
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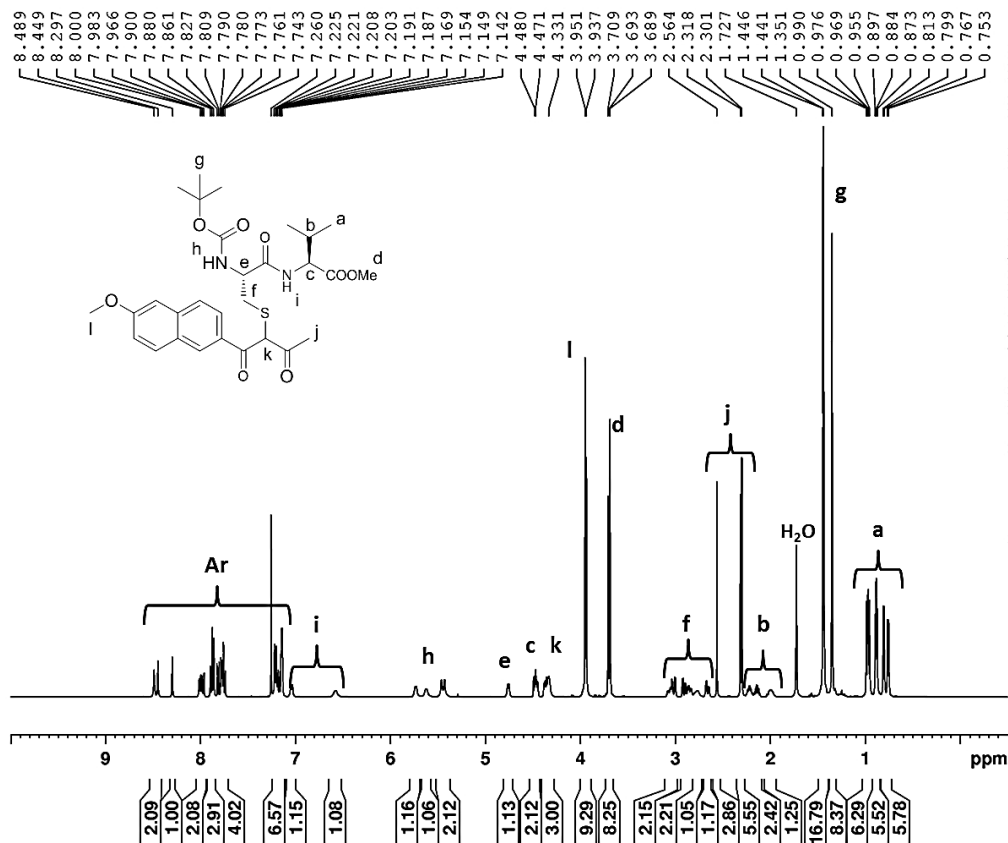
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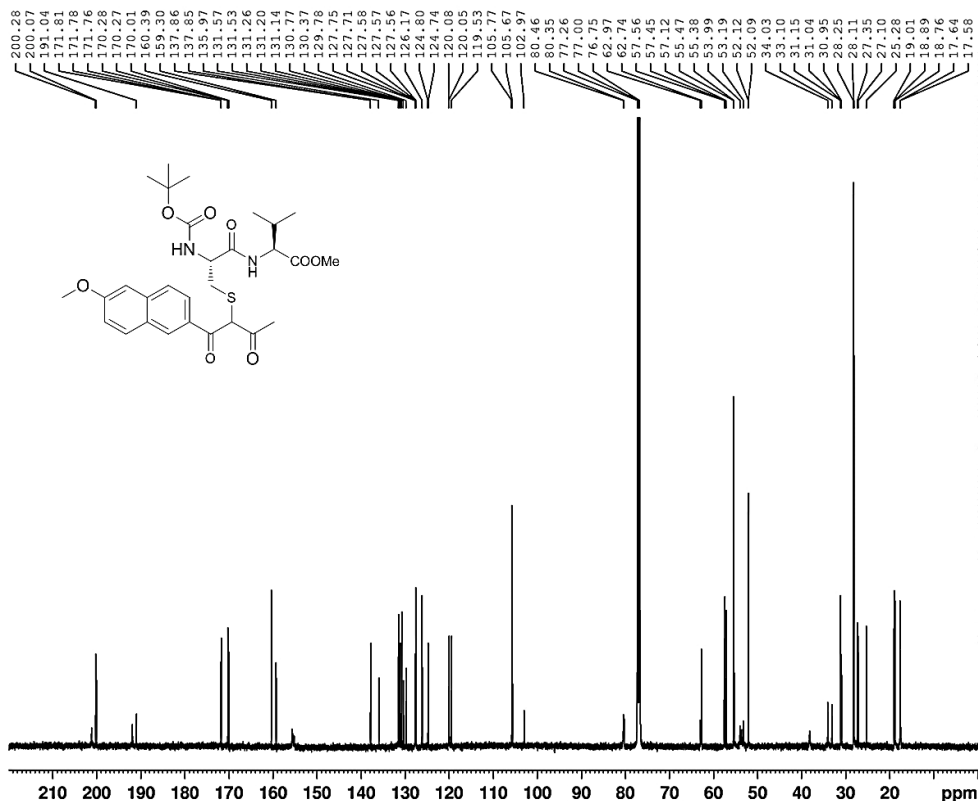


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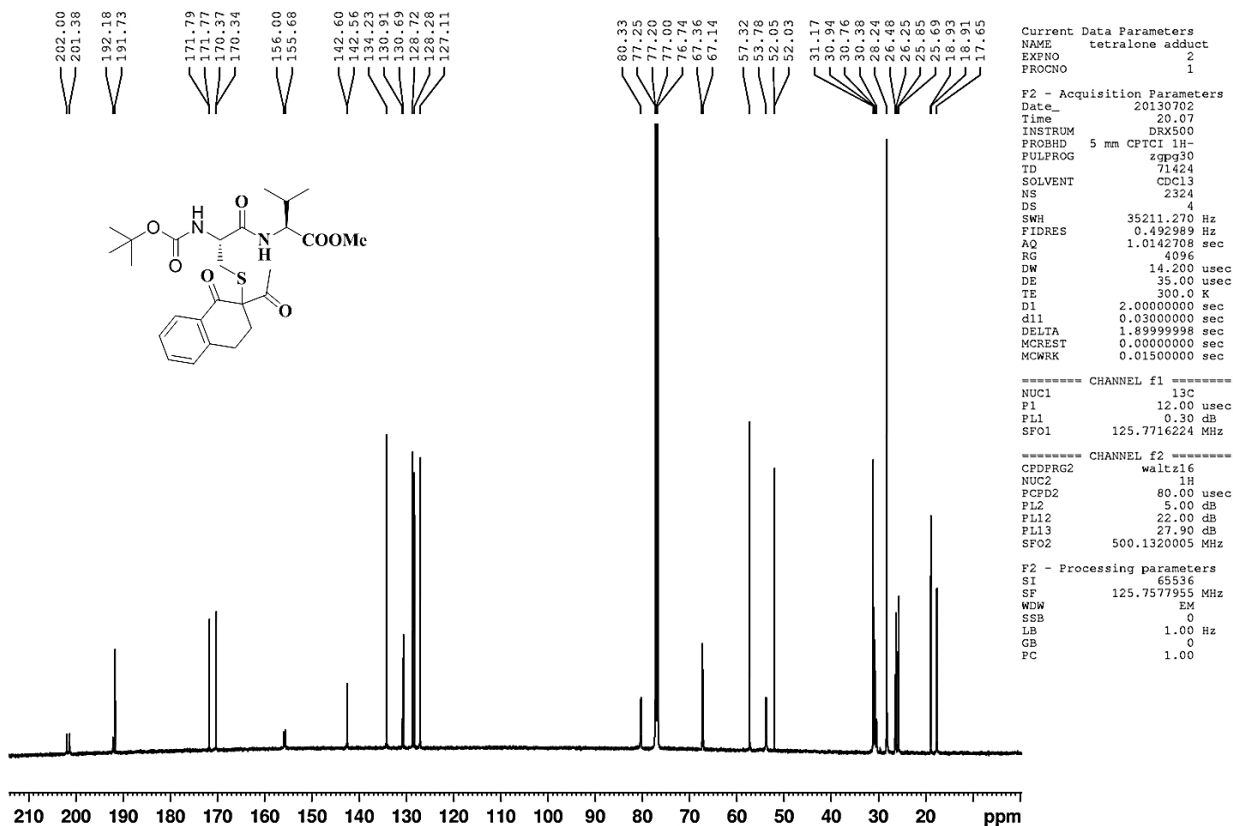
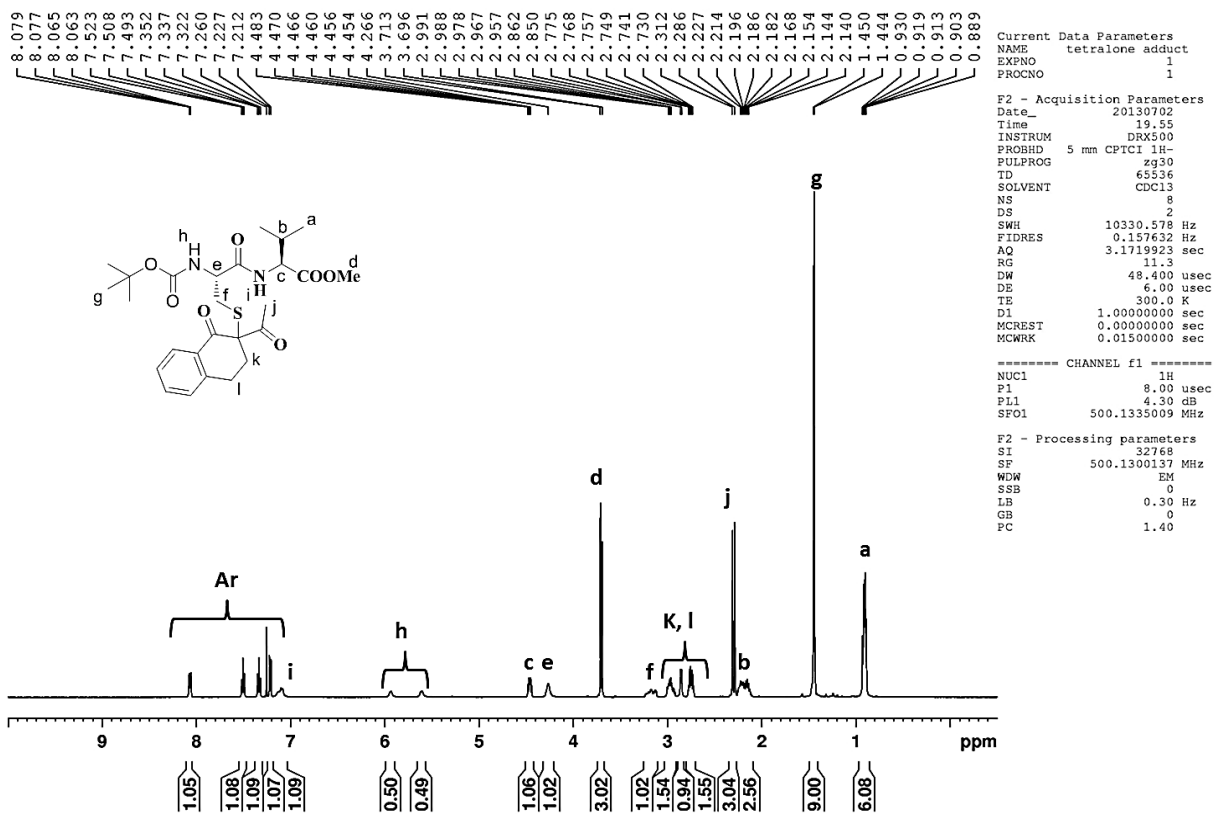
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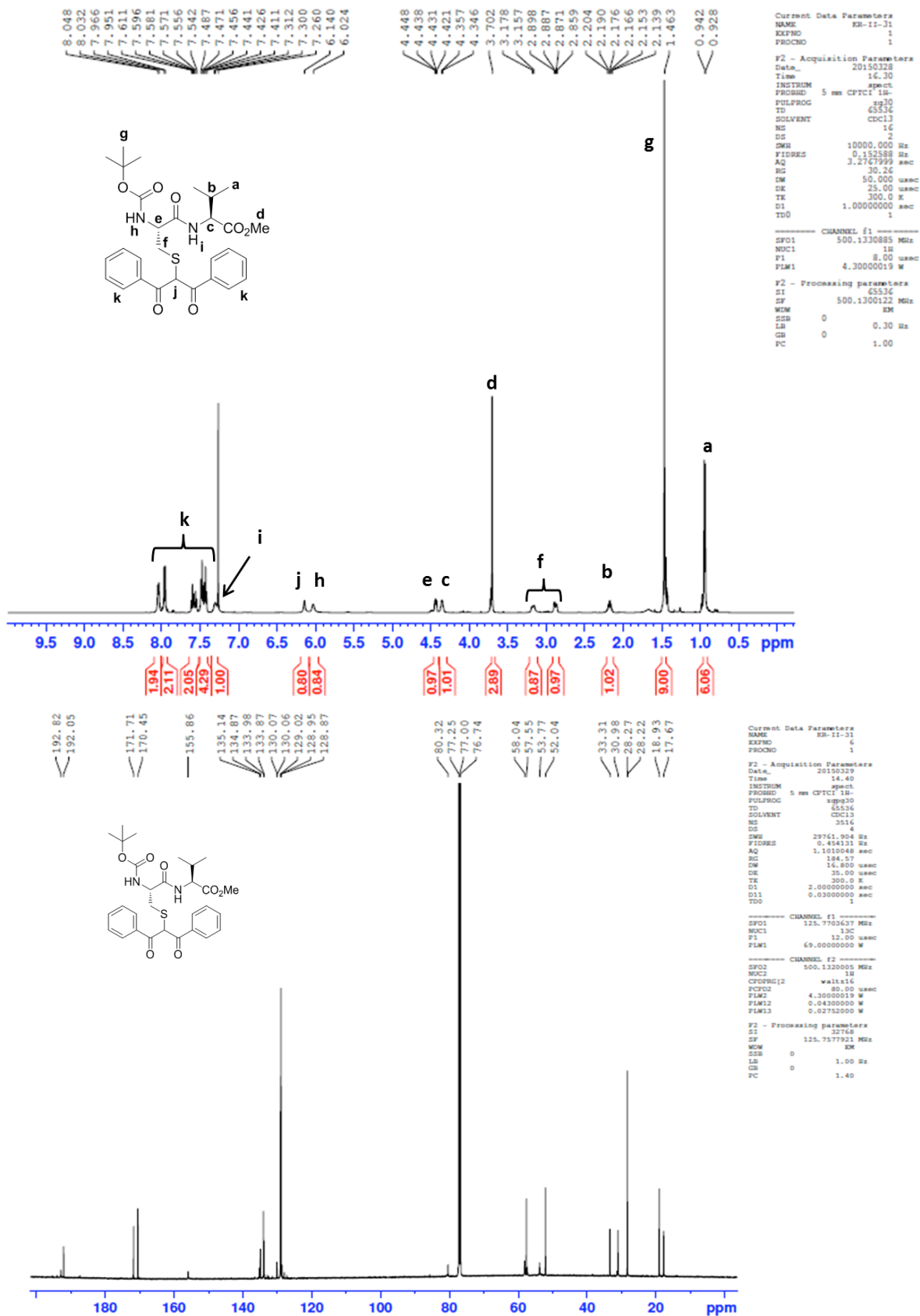
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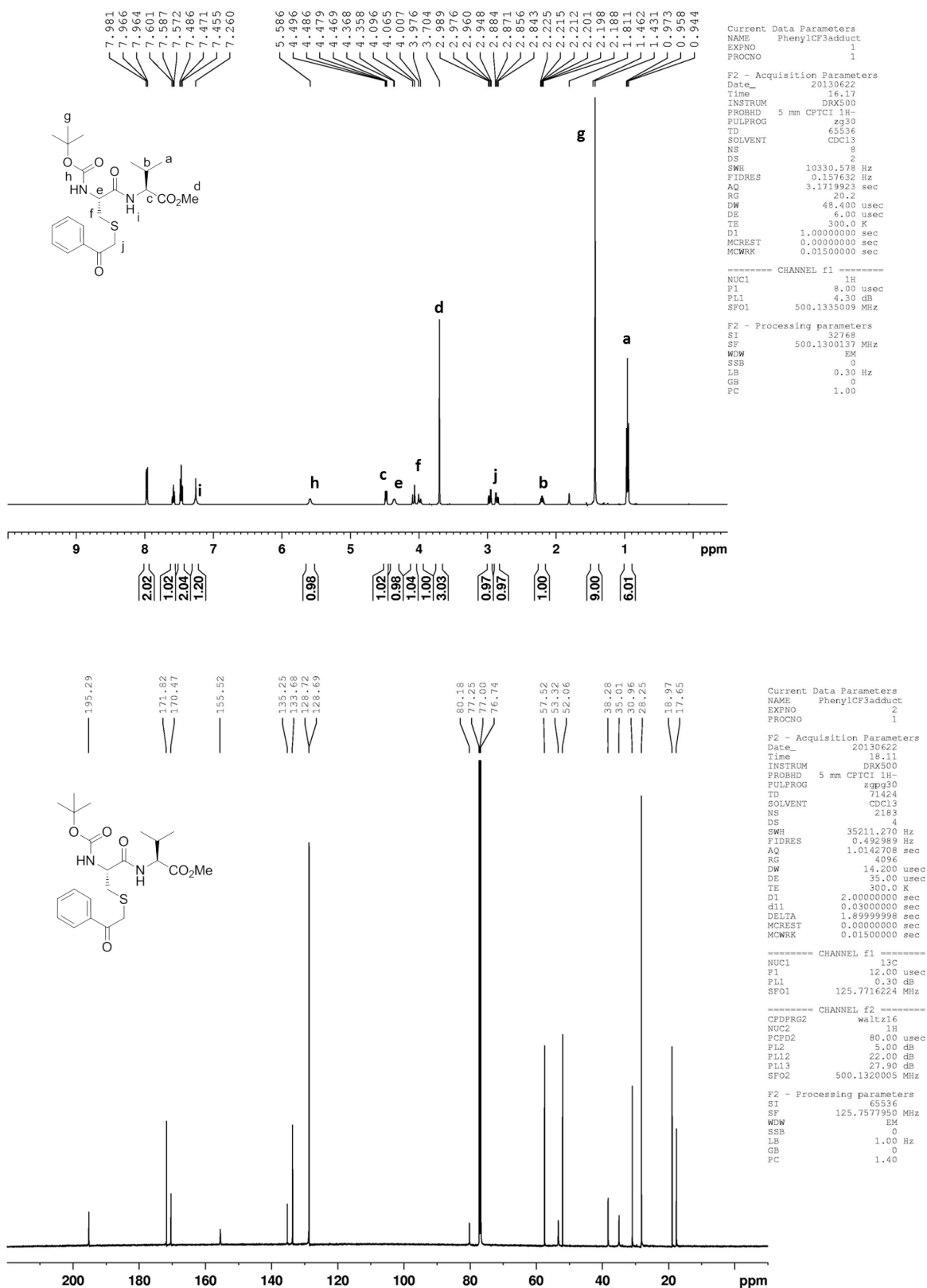
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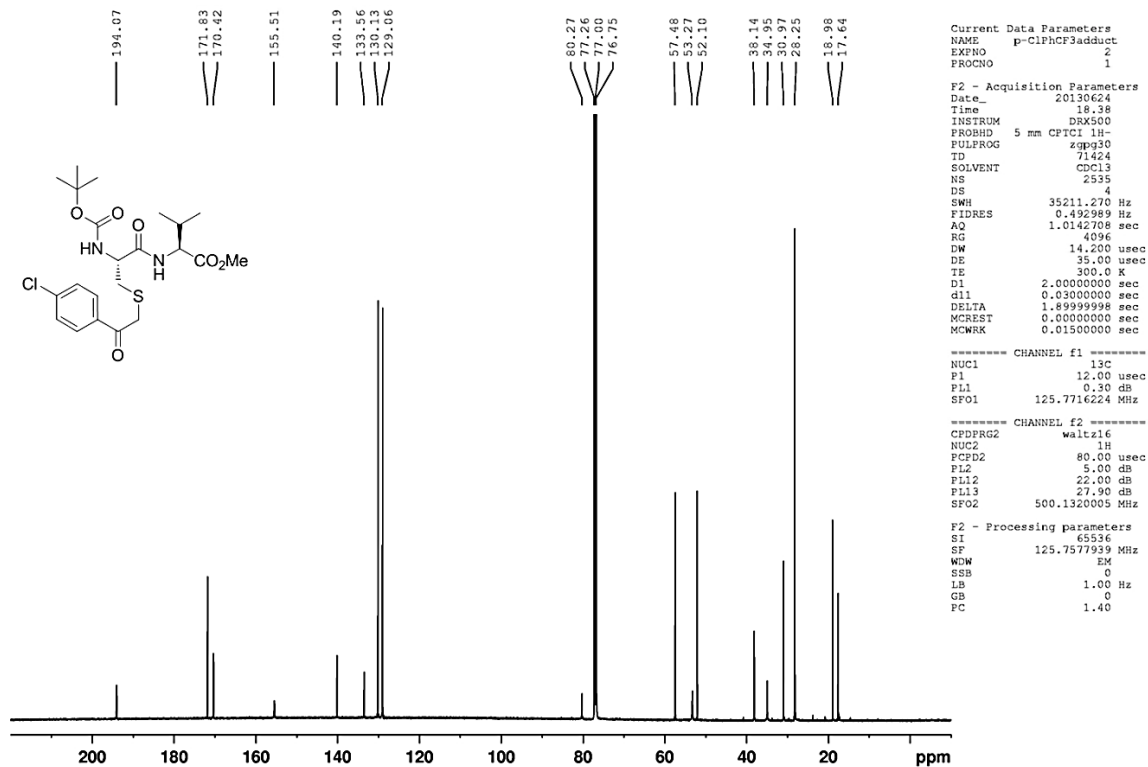
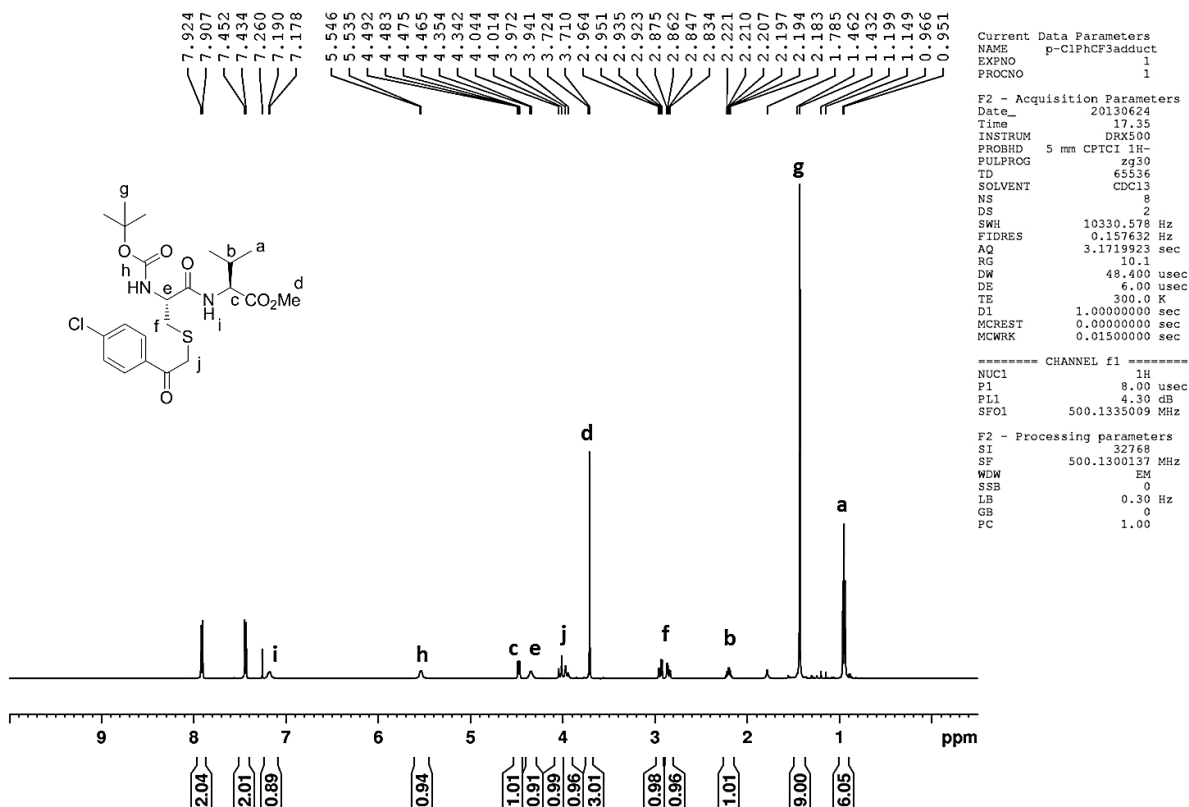
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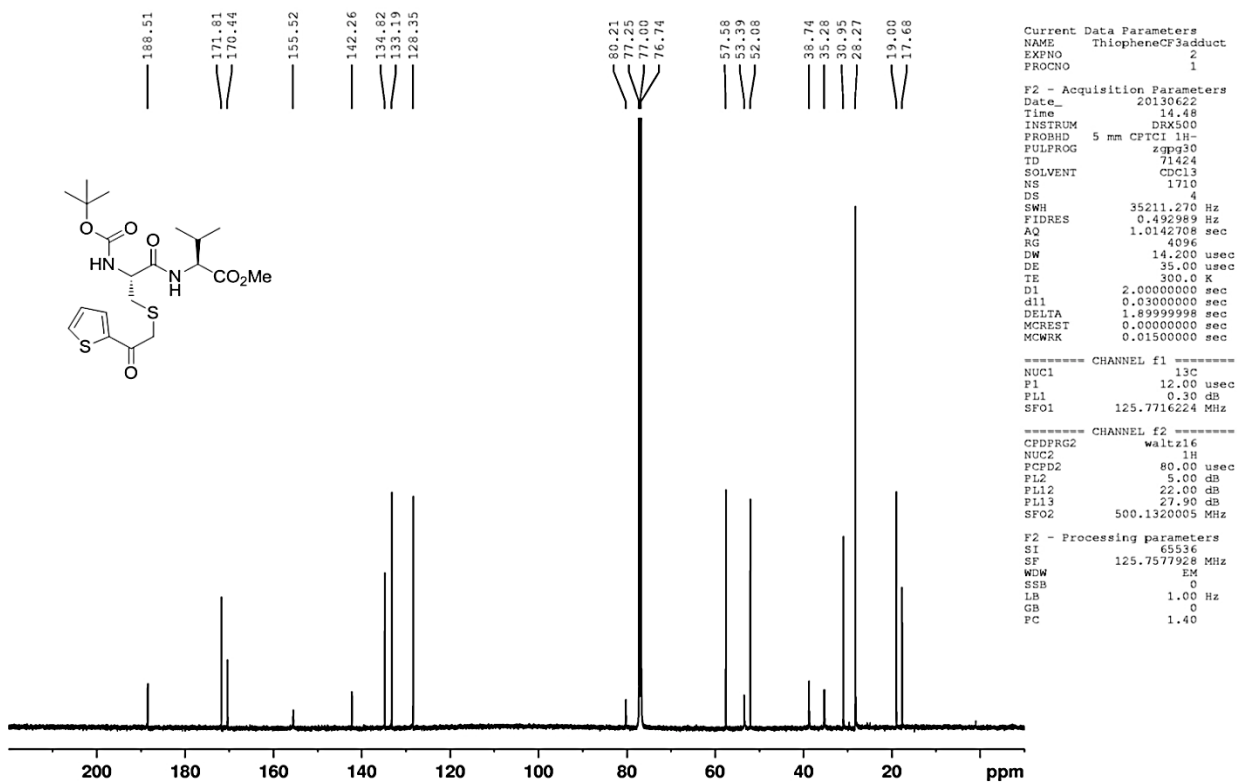
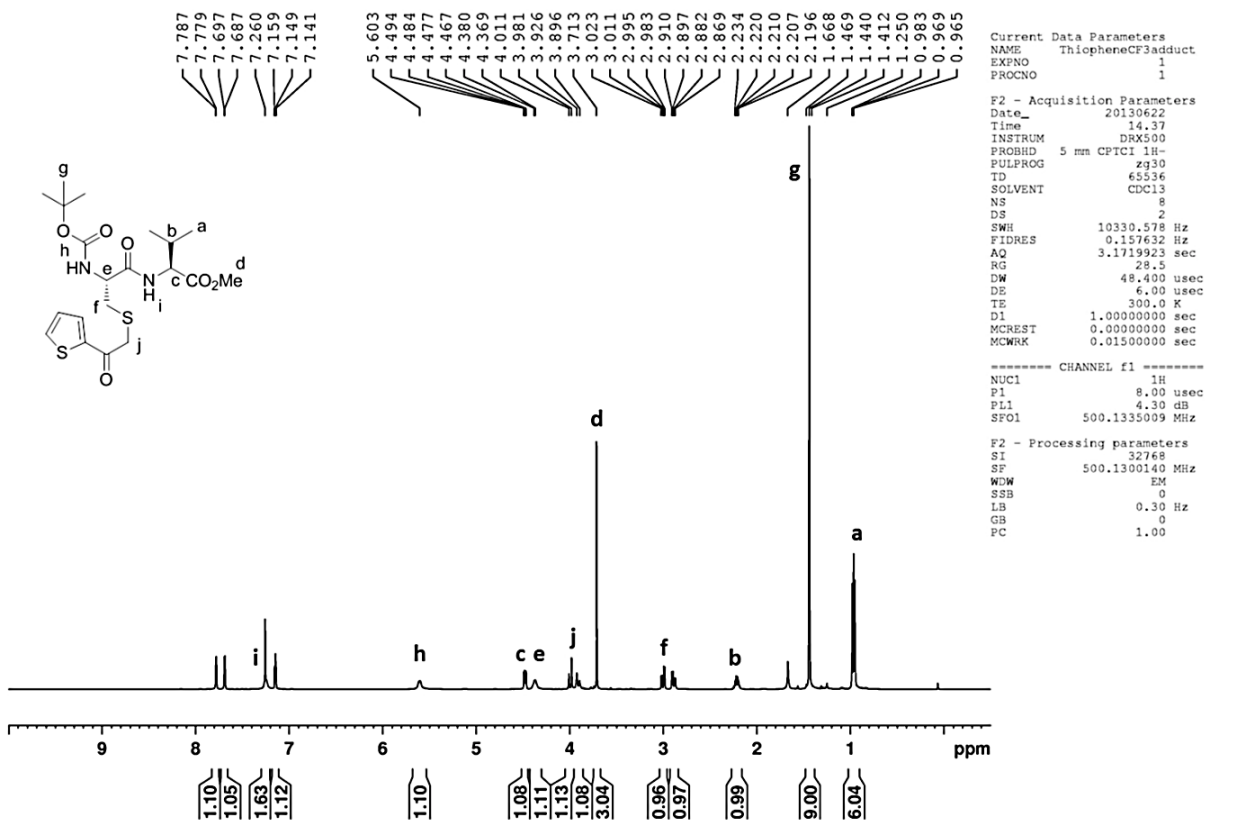
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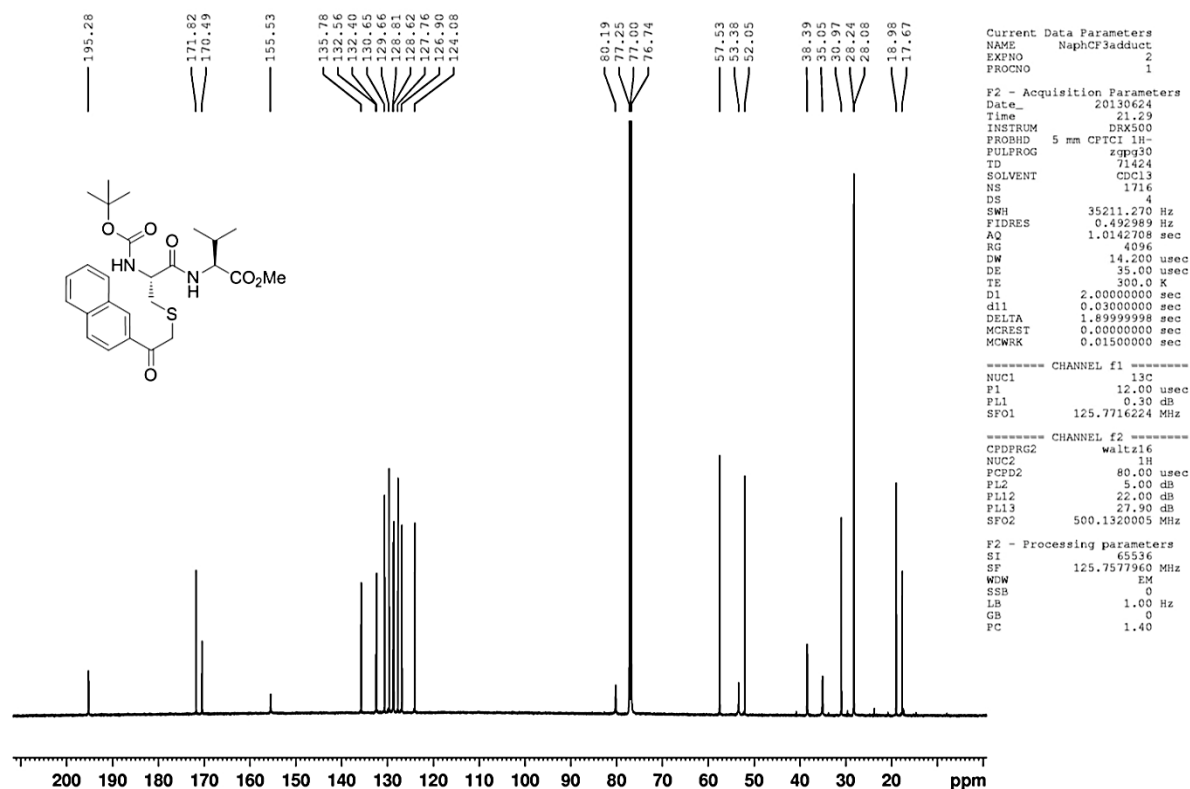
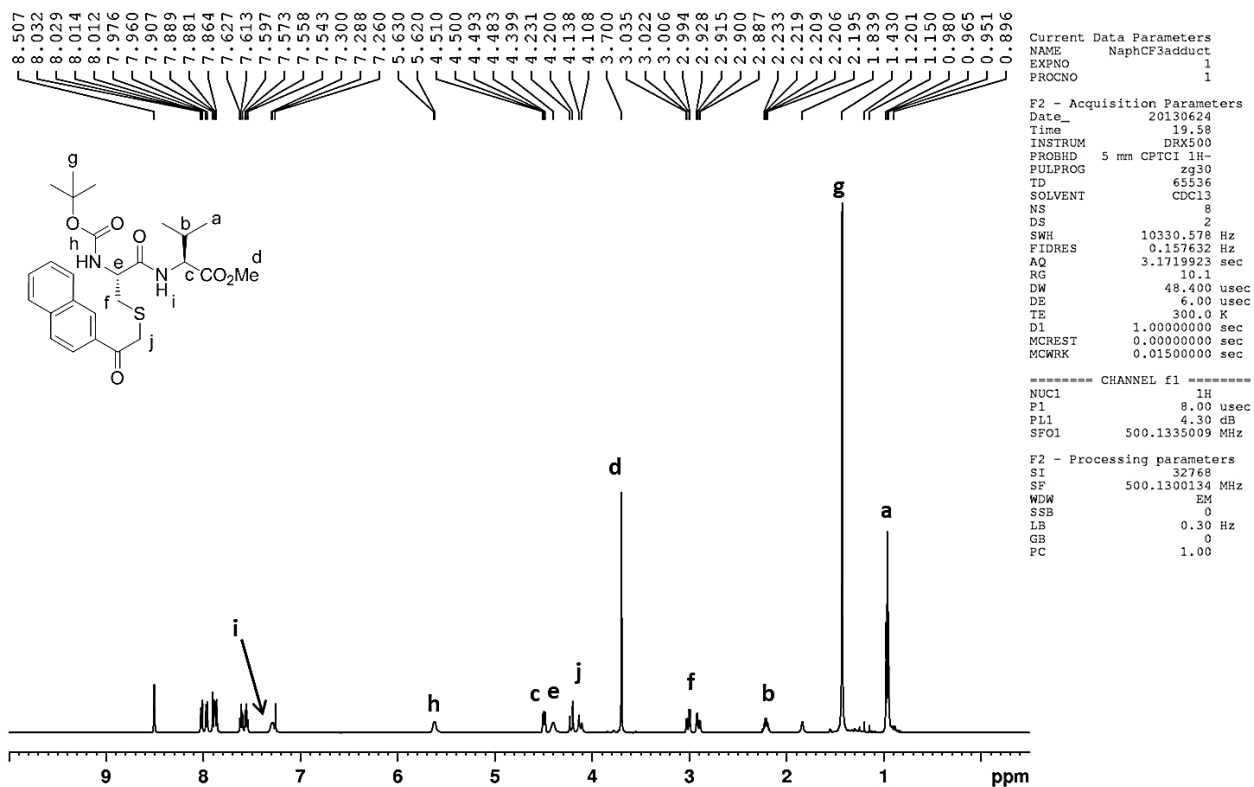
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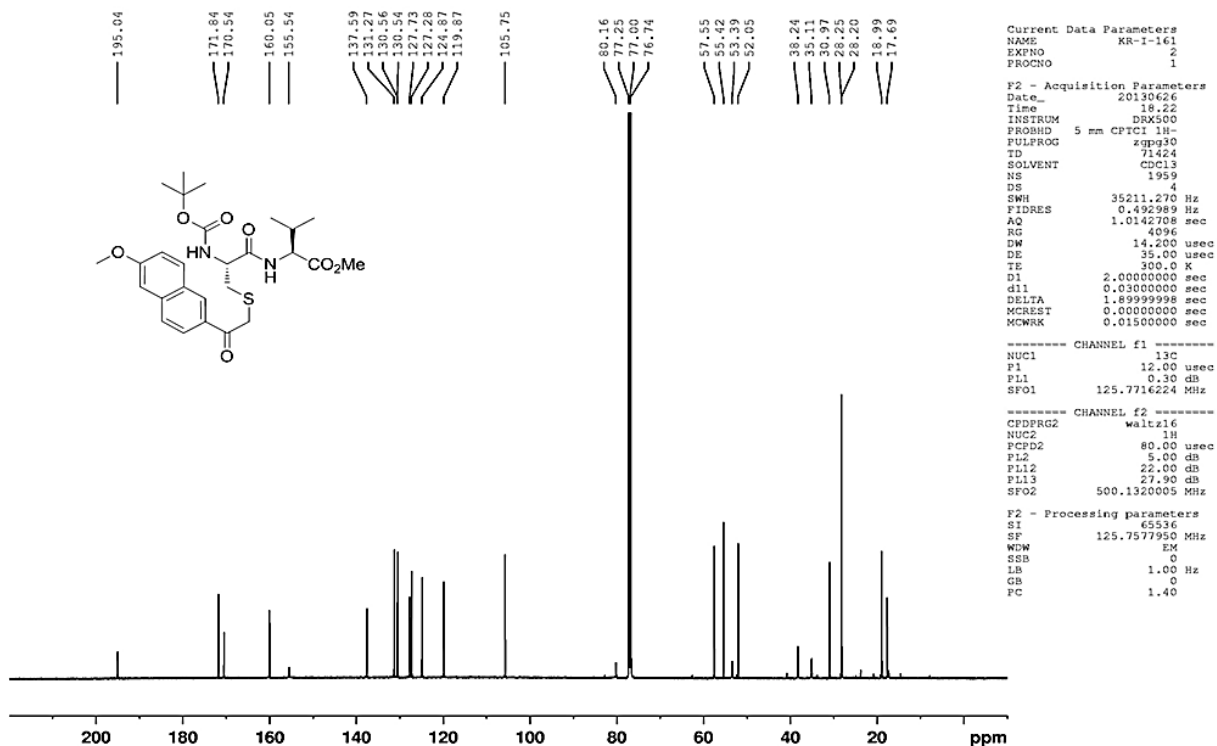
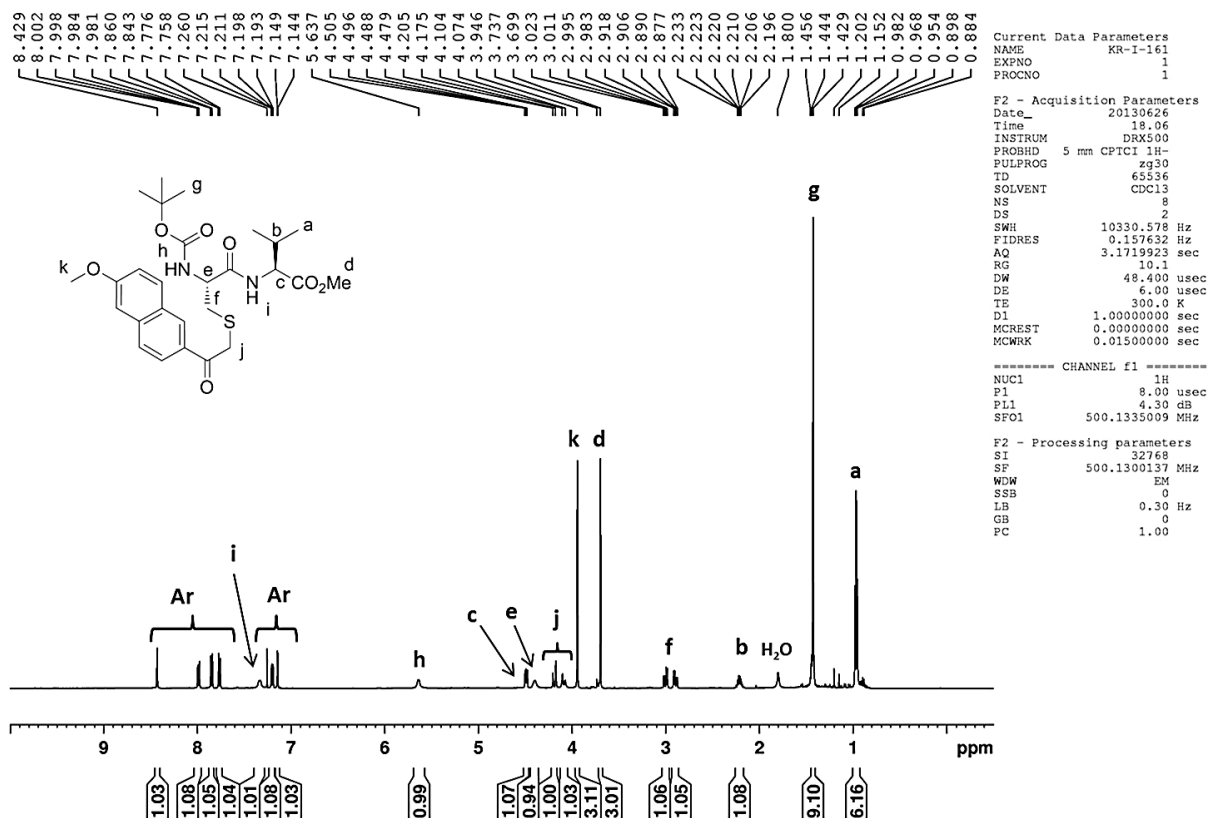
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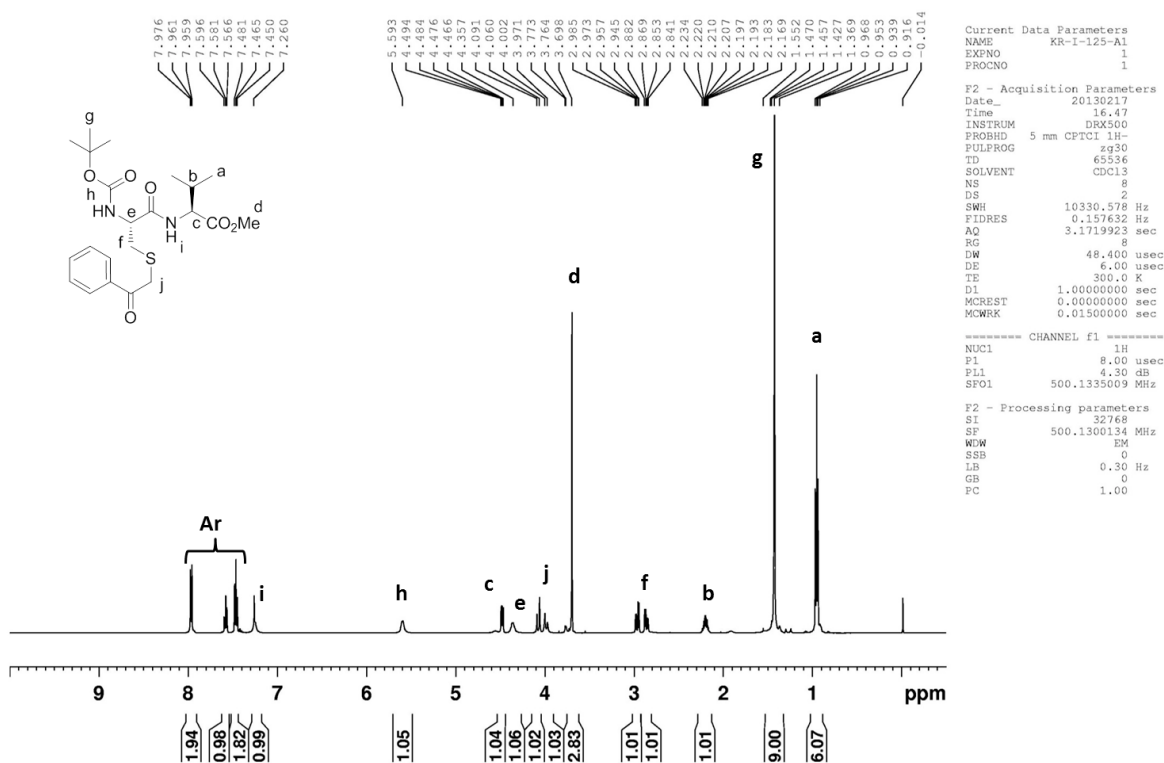
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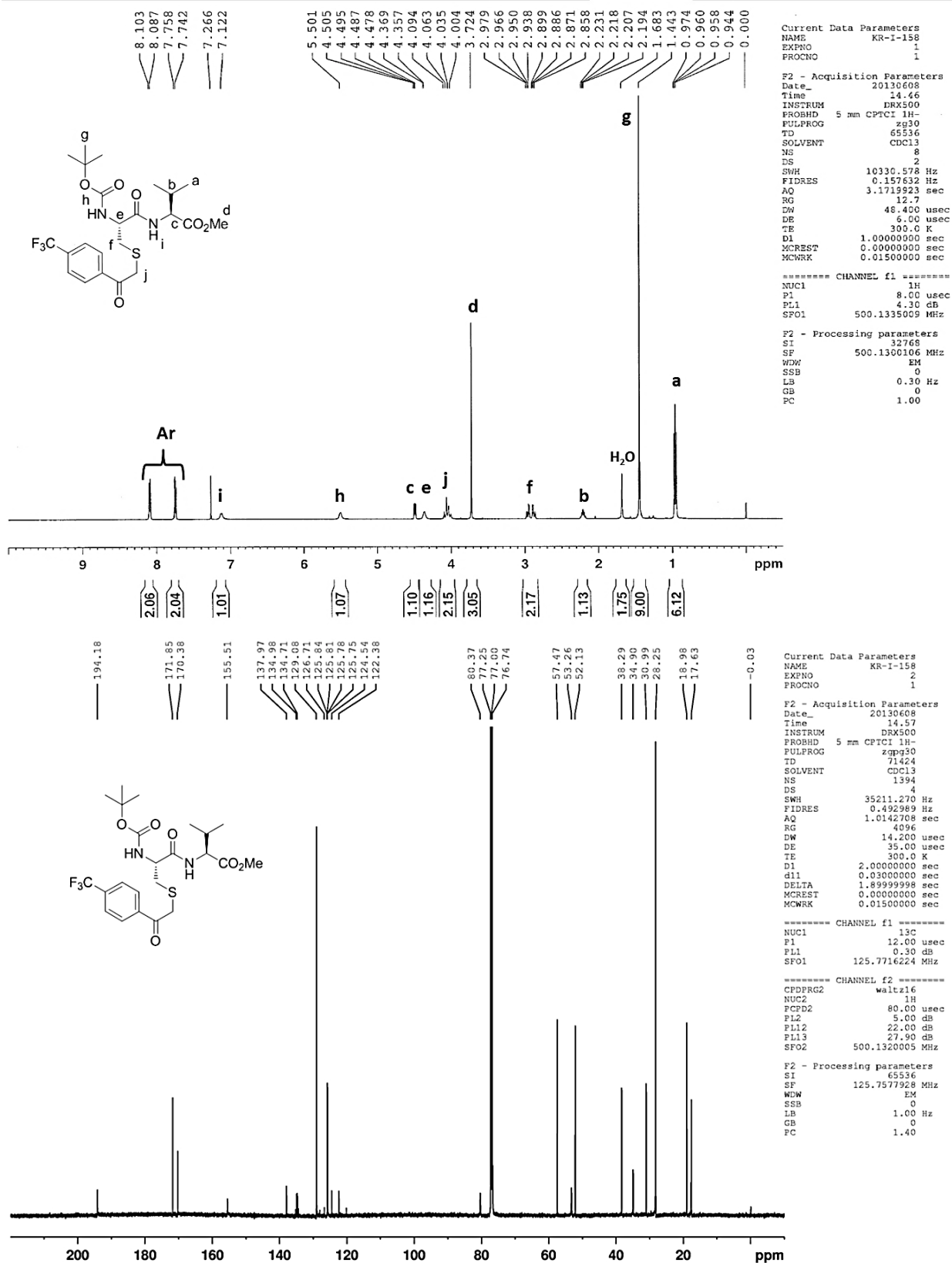
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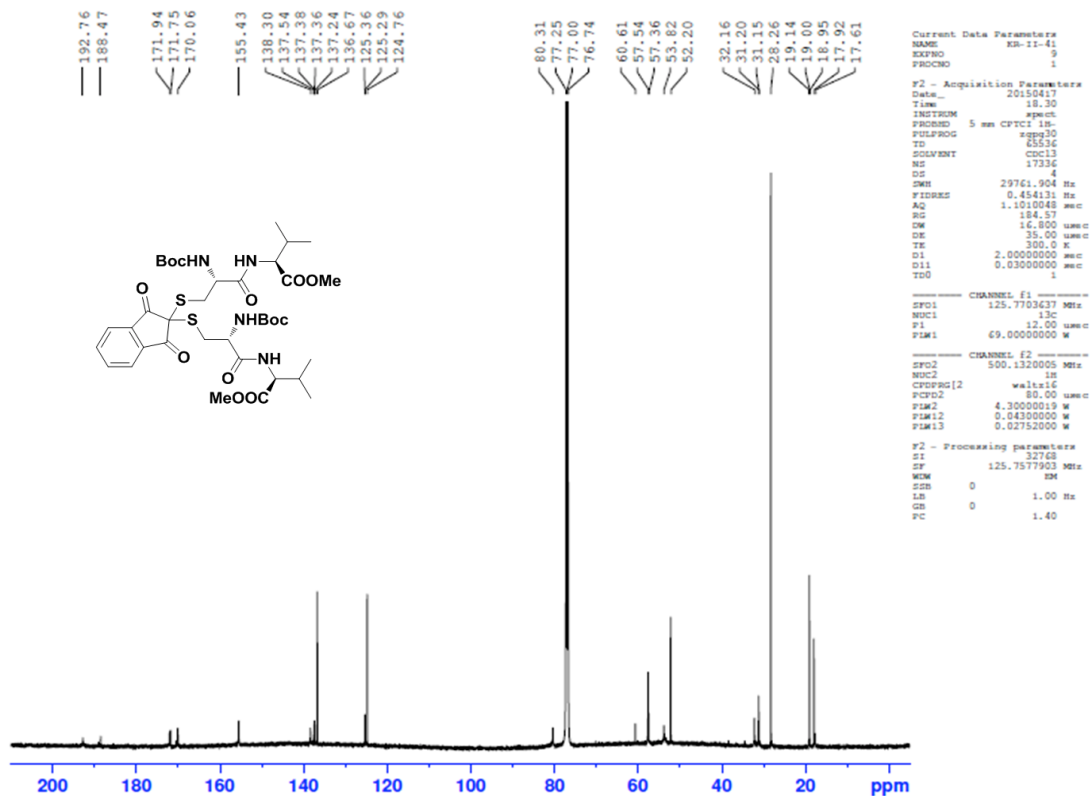
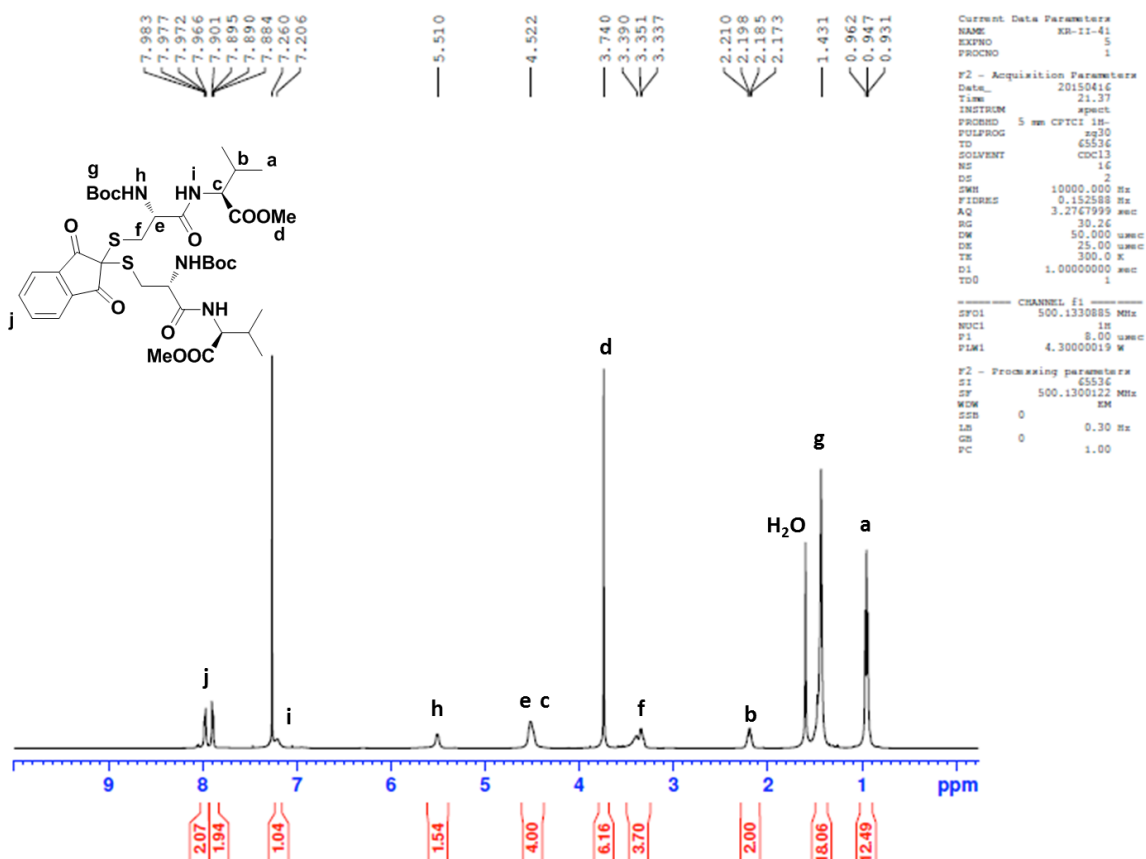
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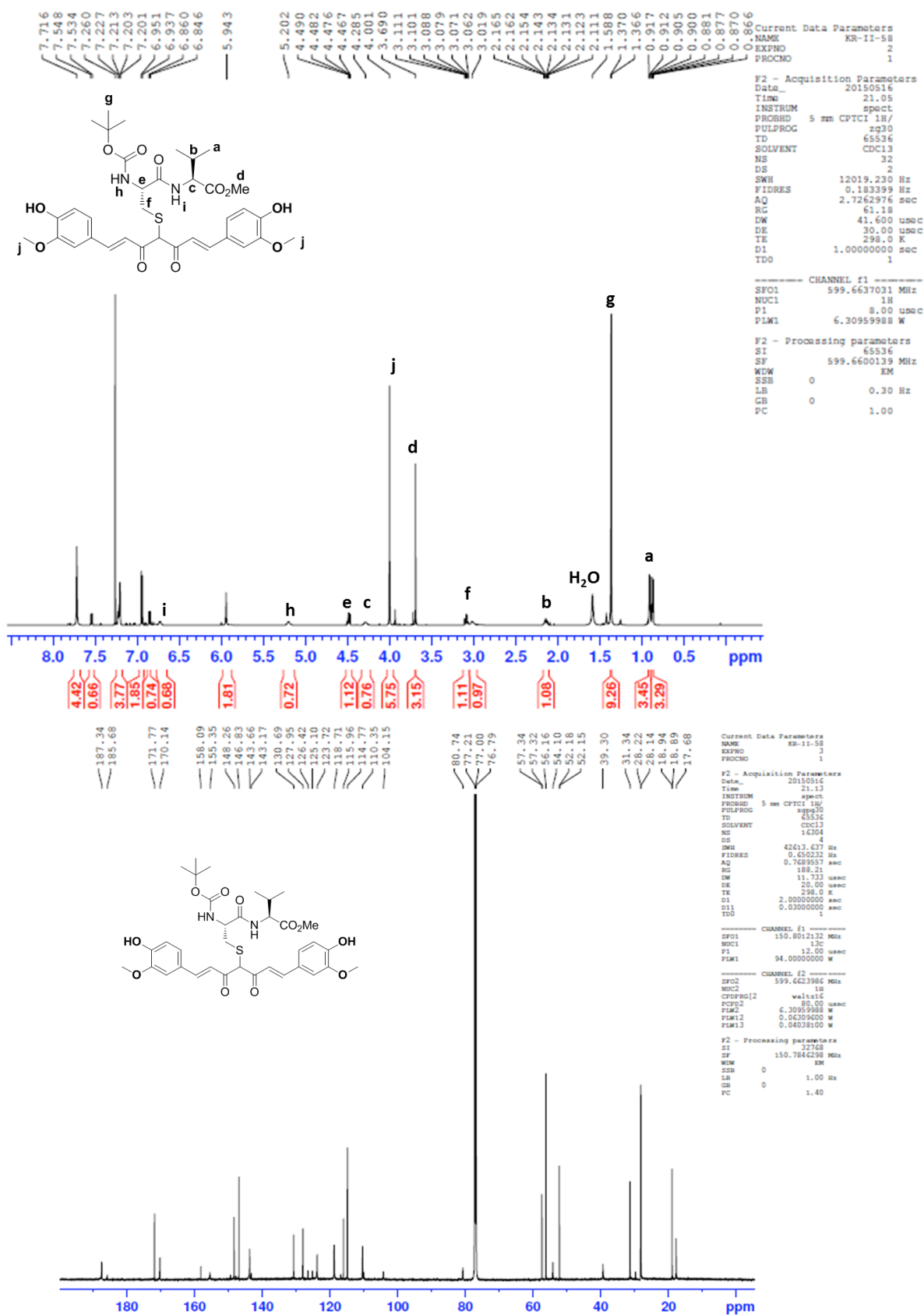
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Spectrophotometric Determination of Carbon Acid Ionization Constants. A buffer containing glycine, sodium acetate, and sodium phosphate (all 50 mM, final concentrations) was prepared in distilled, deionized H₂O. Portions of the buffer were prepared at pH 3 and pH 13, and at half pH unit intervals between. UV-vis measurements were made at room temperature (23 ± 2 °C). The approximate ionic strength of the buffer was 0.2 M (that due to buffer component contributions). Concentrated stock solutions of the carbon acids were prepared in DMSO (typically 1-10 mM) and stored at –20 °C until used. Because molar absorptivity may vary greatly as a function of molecular structure, final concentrations of carbon acids suitable for use in the spectrophotometer were determined separately for each molecule, but typically ranged from 50 – 250 µM.

In order to determine which carbon acid UV-Vis absorption maxima might be suitable for use in pK_a determination, we first examined the absorption spectra of a fixed concentration of carbon acid at the pH extremes (3 and 13), and looked for spectral changes as a function of pH. Generally, spectral peaks were blue-shifted and/or attenuated upon acidification (at pH 3), relative to those appearing at pH 13, consistent with anions being the relevant absorbing species. Having thus determined suitable concentrations of carbon acids and the associated wavelength(s) to monitor during titration experiments, fixed concentrations of each carbon acid were introduced to each buffer (pH 3 to 13), briefly allowed to equilibrate, and the pH-dependent spectral peak intensities recorded. In each case, the final composition of the mixture assayed in the cuvette was 4% DMSO in buffer (v/v; 40 µL carbon acid in DMSO, 960 µL buffer). Finally, absorption at λ_{max} for the wavelengths monitored was plotted as a function of pH, and the pK_a determined by fitting to the equation:

$$A_{pH} = A_{\min} + \frac{(A_{\max} - A_{\min})}{1 + 10^{pK_a - pH}}$$

where A_{min} , A_{max} , and A_{pH} are the absorbances at the pH minimum (pH 3), maximum pH (pH 13), and intermediate pH values, respectively, at the wavelength being monitored for a given carbon acid (Figure 4B). To validate both the solvent system employed (aqueous buffer components + 4% DMSO, $\mu = 0.2$ M) and the method used in data analysis, a set of benchmark carbon acids and heteroatom-based acids of known aqueous pK_a were examined, and the ionization constants determined in our hands matched excellently those reported in the literature, determined by various means.

Spectrophotometric determination of acid dissociation constants over a pH range 3-13.

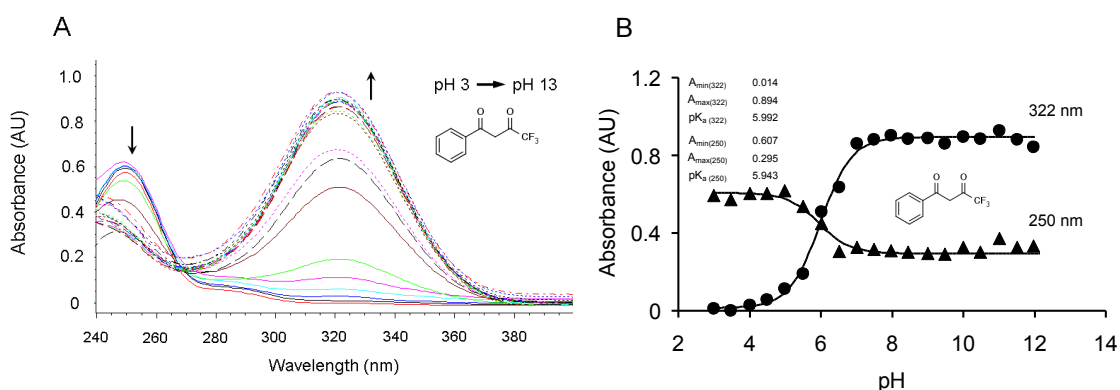
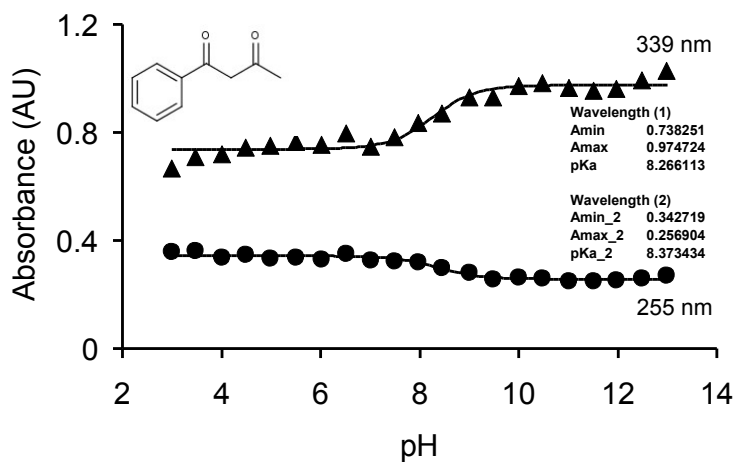
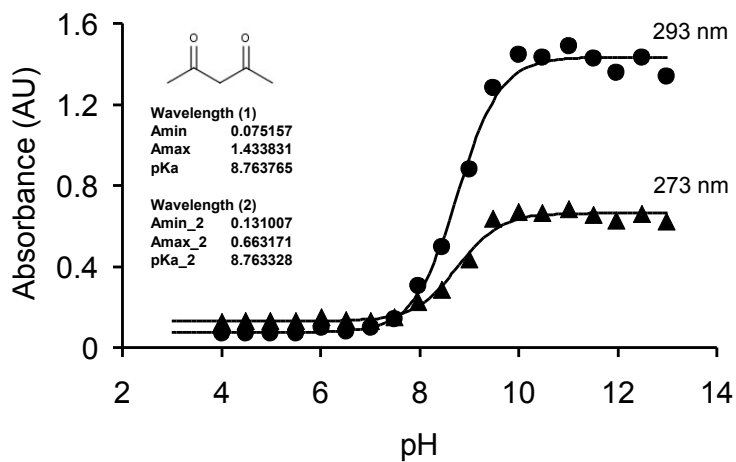
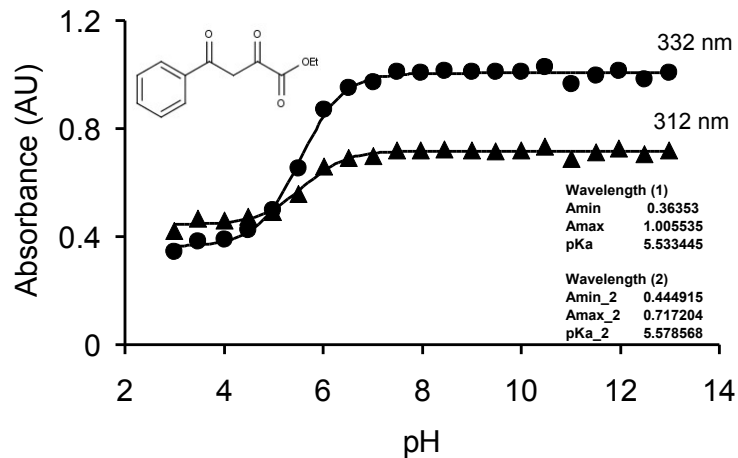
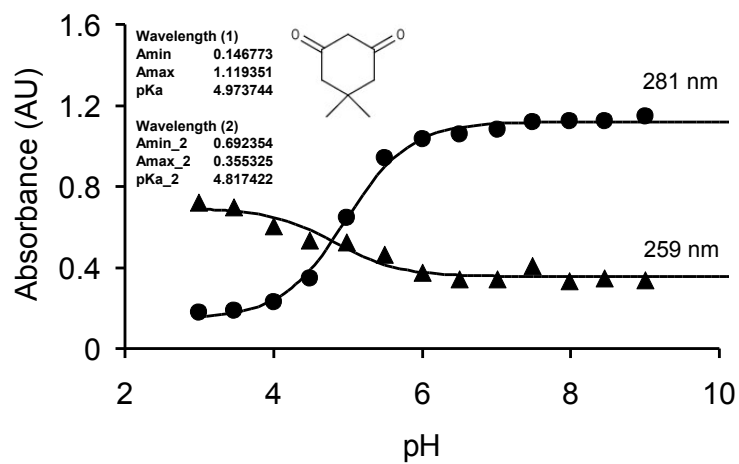
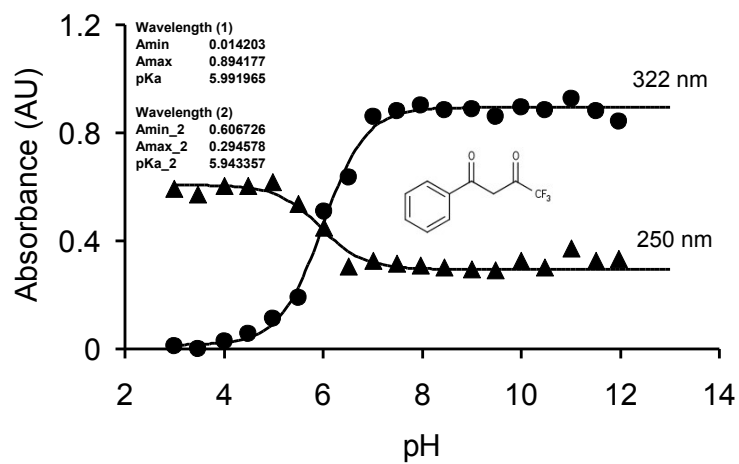
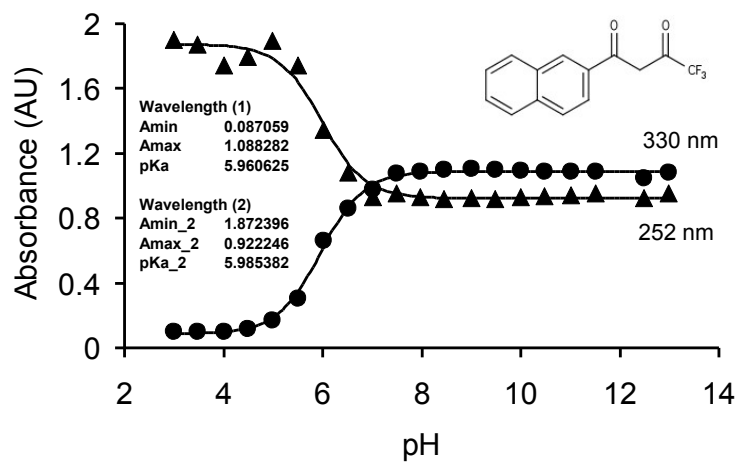
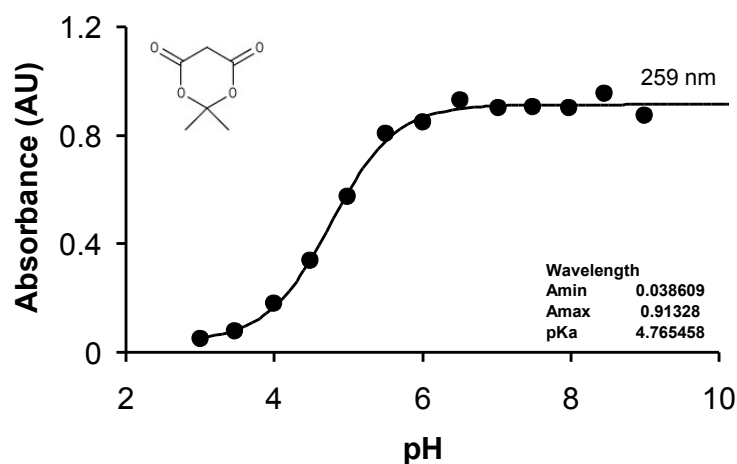


Figure S1. Spectrophotometric determination of carbon acid ionization constants. **A.** To 960 μ L samples of "universal buffer" pH 3 – 13 (see text) were added 40 μ L of 1.25 mM compound **H** in dmsO to final concentration 50 μ M. The solution was briefly allowed to equilibrate and the absorption spectrum versus buffer blank recorded. **B.** The absorbances of spectral peaks that varied as a function of pH were plotted versus the corresponding pH and the data fitted to afford the apparent pK_a .

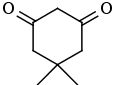
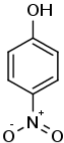
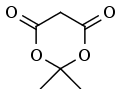
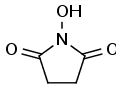
Carbon acid pK_a determination: titrimetric data and associated fits.







Benchmark set for pK_a determination methodology: all pK_a values determined in this work, listed below as " pK_a (det.)", were determined using the methodology detailed in the Methods section of the manuscript.

	pK_a (det.)	pK_a (lit.)		pK_a (det.)	pK_a (lit.)
	10.0	10.0 ^a		5.0	4.8 ^b
	6.9	7.2 ^a		4.8	4.9 ^b
	5.0	5.2 ^a		5.9	6.0 ^c
	4.5	4.6 ^a			

^a Harris, D. C. (2010) *Quantitative Chemical Analysis*, 8th ed.

^b Mofaddel, N., et al. *Anal. Biochem.* (2004), 380, 664-68

^c Cline, G. W. *J. Org. Chem.* (1988), 53, 3583-3586

Development of Assay Methodology for the Selective Detection of Sulfenyl Amide via Spectrophotometry. In order to follow the reaction of diketones with **4** spectrophotometrically, we first prepared the colorimetric reagent 2-nitro-5-sulfanylbzoic acid (TNB, $\text{Abs}_{410\text{nm}}$ in aqueous media). Synthesis of TNB was accomplished by reduction of Ellman's reagent (5,5'-dithiobis(2-nitrobenzoic acid)) with 8 equiv of cysteamine hydrochloride in water brought to pH 8 with NaHCO_3 . The reaction was stirred at room temperature for 2 h, acidified to pH 1 with 12 M HCl, and then extracted with ethyl acetate. The organic extract was dried over Na_2SO_4 , filtered, the solvent removed by rotary evaporation to afford pure TNB as an orange solid. NMR spectral data matched that previously reported in the literature.²

We determined that TNB reacts selectively and stoichiometrically (1:1) with **4**; that is, loss of absorbance at 410 nm from a solution of TNB was observed upon addition of sulfenyl amide to the solution, but not when disulfides such as 2-mercaptoethanol disulfide or thioethers such as **8a** were added (Supporting Information). Stoichiometric (1:1) reaction of TNB with **4** suggests the process proceeds to the mixed disulfide, but further reduction of the mixed disulfide to the cysteine dipeptide **7**, accompanied by formation of DTNB, does not occur. Failure of TNB to fully reduce **4** to **7** is in agreement with precedented thiol-disulfide reaction equilibria involving low- pK_a aryl thiols, wherein equilibrium heavily favors the free aryl thiol.³ Thus, we established that (a) the assay reading (Abs_{410}) changes solely as a function of concentration of **4** in solution, in 1:1 fashion, and (b) that instrument response is linearly dependent on analyte (**4**) concentration within the concentration regime considered. Accordingly, the assay employed a slight molar excess (1.25 equiv) of TNB over sulfenyl amide.

A concentrated stock of TNB (50 mM) in ethanol was prepared and stored at -20°C until use. During use, the solution was stored on ice. When handled this way, the stock thiol was found to be stable for weeks (no appreciable loss of Abs_{410} at a fixed concentration of TNB). *Note: DMSO was found to be unsuitable for storage of TNB, as the solution rapidly decolorized, presumably due to DMSO-mediated oxidation of TNB to colorless DTNB.*⁴ Stock solutions of all carbon nucleophiles and of the sulfenyl amide were prepared in DMSO and stored at -20°C before use.

Determination of Rate Constants for Reaction of 1,3-Diketo Nucleophiles with the Model Sulfenyl Amide. Approximately 10 min prior to the start of the assay, an aliquot (1–120 μL) of concentrated stock (250–500 mM) of carbon nucleophile in DMSO was added to the reaction buffer (1:1 methanol:buffer B – 50 mM Tris, 50 mM Bis-Tris, 10 mM DTPA, 100 mM NaOAc, pH 7.0, 11.88 mL total volume). To this solution was added additional DMSO, as needed, to bring the final volume of the reaction buffer to 12.0 mL total (1% DMSO v/v, final). This afforded final diketone nucleophile concentrations of 500 μM –5 mM. The 10 min pre-incubation period was intended to allow for ionization of the carbon acid to proceed toward equilibrium.⁵ Following the pre-incubation period, 2 mL of the mixture (now containing diketone), were removed and set aside for determination of initial (A_0) and endpoint (A_{∞}) measurements for the reaction.

Assay buffer solutions were prepared via addition of 10 μL of TNB (50 mM) in ethanol to 1990 μL of Buffer C (500 mM sodium acetate, 100 mM Bis-Tris, 75 mM DTPA, pH 5.0), affording 2 mL final volume and 250 μM TNB final concentration. The 2 mL of assay buffer solutions were subsequently divided into nine 200 μL aliquots and stored on ice until used; 7 aliquots were dedicated to reaction kinetics time point measurements, and the last 2 to endpoint determinations (A_0 and A_{∞} , *vide supra*).

To the remaining 10 mL of reaction buffer were added 10 μ L of the model sulfonyl amide (50 mM in DMSO), to give a final concentration of 50 μ M. Immediately upon introduction of the **4** to the reaction mixture, a timer was started and aliquots (800 μ L) were periodically removed from the reaction mixture and assayed for remaining (unreacted) **4** by dilution into ready-made 200 μ L aliquots of assay buffer containing TNB (*vide supra*). Reaction progress was typically assayed at 0.5, 1, 2, 3.5, 5, 8, and 11 min time points. The assay mixture, strongly buffered at pH 5, permitted simultaneous quenching of the enolate-**4** reaction and quantitative measurement of unreacted **4**.⁶ The quenched time point mixtures (1 mL total; 800 μ L reaction mixture, 200 μ L assay buffer + TNB) afforded 50 μ M TNB and 40 μ M **4** formal concentrations (final). Prior to measuring Abs₄₁₀ of these samples, they were allowed to thermally equilibrate by standing at room temperature for 5-10 min. The Abs₄₁₀ at time zero (A_0) and infinite time (A_∞) were determined separately by preparation of two independent endpoint samples; 800 μ L of the set-aside 2 mL of reaction mixture were added to 200 μ L assay buffer (done for two samples). Then, to the initial reading sample (A_0), 0.8 μ L of DMSO were added, and to the endpoint sample (A_∞) 0.8 μ L of **4** (50 mM) in DMSO were added. Use of these samples for determination of A_0 and A_∞ account for any contribution the diketone/enolate might have to Abs₄₁₀.

Plots of Abs₄₁₀ with respect to reaction time afforded monophasic rising curves, consistent with consumption of **4** by the carbon nucleophiles. Because the concentration of carbon nucleophile in the kinetics experiment was at least 10-fold higher than that of **4**, the data were analyzed using pseudo first-order kinetics by fitting to the equation:

$$A_t = [(A_0 - A_\infty) \times e^{-k_p \times t}] + A_\infty$$

where A_o , A_∞ , and A_t are the Abs₄₁₀ initially, finally, and intermediately during the reaction, and k_ψ is the pseudo first-order rate constant for reaction at a given concentration of carbon nucleophile. Note that the expression above is merely a rearranged form of the equation that describes any (pseudo) first-order process:

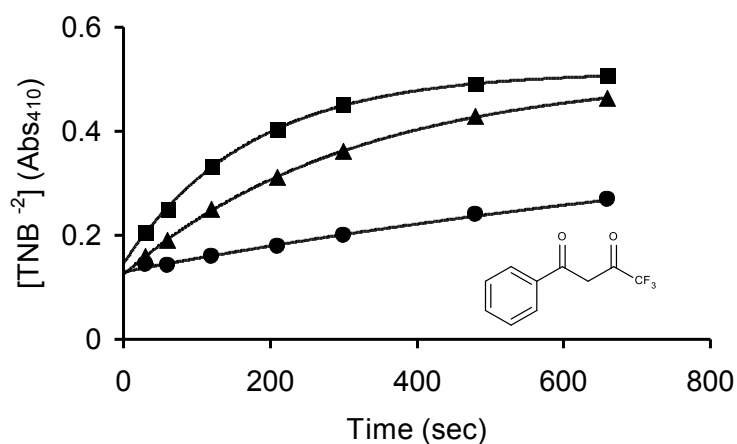
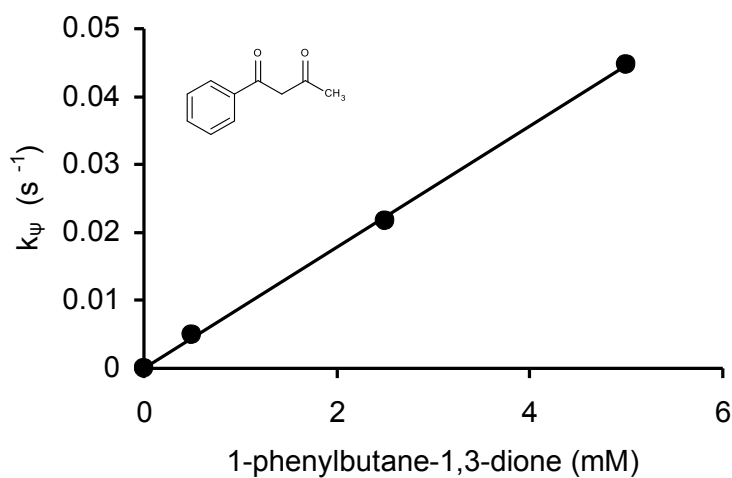
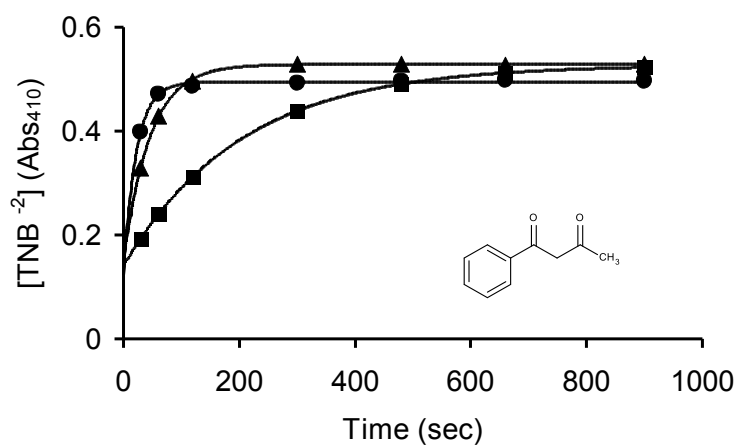
$$\frac{A_t - A_\infty}{A_o - A_\infty} = e^{-k^*t}$$

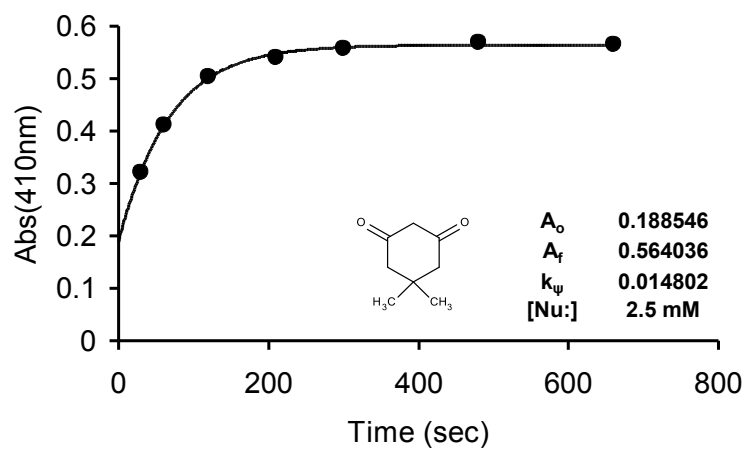
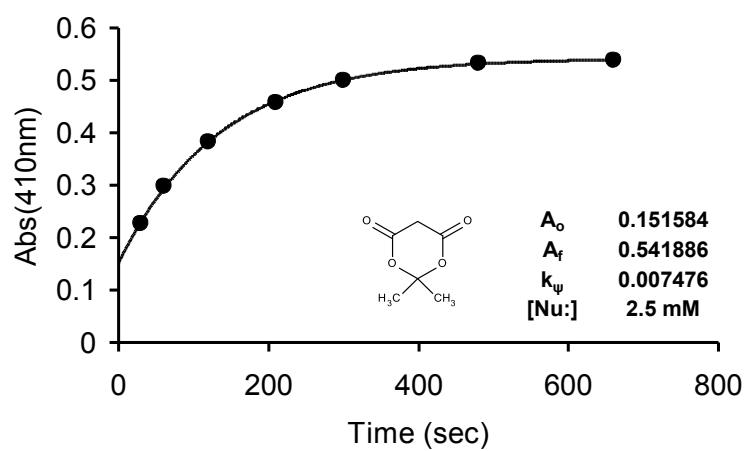
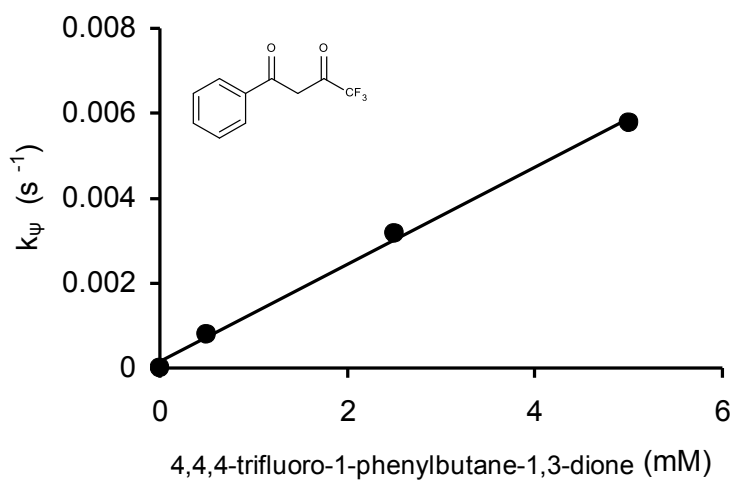
Indeed, first-order kinetic models fit the data quite well, and plotting the observed pseudo first-order rate constants versus concentration of the carbon nucleophile in excess afforded a straight line passing through the origin. This result is consistent with a simple bimolecular process being involved in the rate-determining step, described by rate law:

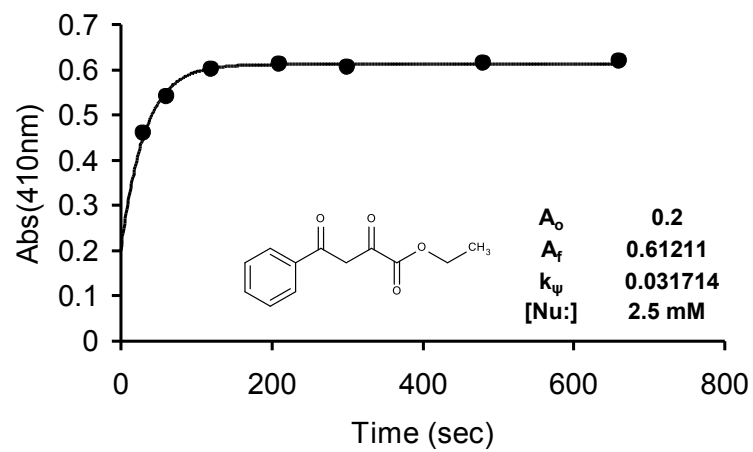
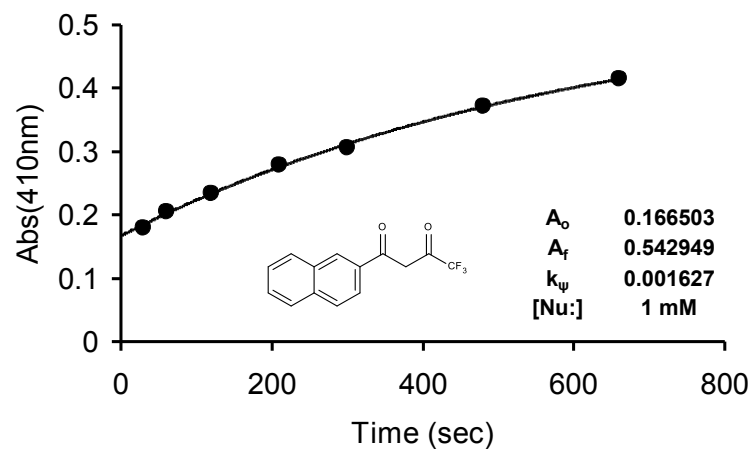
$$Rate = -\frac{d[SA]}{dt} = +\frac{d[P]}{dt} = k_{obs}[SA][Nu]$$

where k_{obs} is the second-order rate constant (also referred to as k_{trap} herein), and [SA], [P], and [Nu] are the concentrations of **4**, product thioether, and enolate nucleophile, respectively.

Kinetics data: determination of rate constants for reaction of carbon nucleophile with model sulfenyl amide *via* TNB assay methodology, associated fitting to kinetic data, and determination of reaction order with respect to various carbon nucleophiles.







TNB assay methodology development and "calibration": determination of instrument response, reaction selectivity, and stoichiometry of the assay.

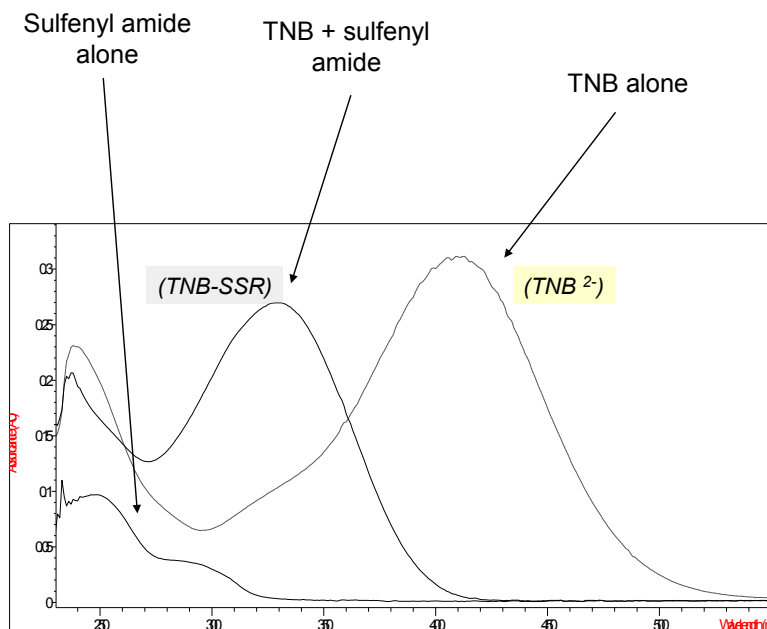


Figure S2. An equimolar concentration of the model sulfenyl amide to the TNB^{2-} ion under the TNB assay conditions completely converts TNB^{2-} to TNB-SSR . To 1:1 MeOH : pH 5 Buffer C, 50 μM sulfenyl amide or 50 μM TNB^{2-} were added (leftmost and rightmost traces/peaks, respectively). The absorption spectra reveal that TNB^{2-} absorbs strongly at 410 nm in the assay mixture and that the sulfenyl amide contributes no appreciable absorbance at this wavelength. When 50 μM sulfenyl amide was added to an assay mixture containing 50 μM TNB^{2-} , immediate loss in yellow color of the solution (due to A_{410} of the TNB^{2-} ion) was observed; the absorption spectrum of the resulting mixture reveals strong absorbance at 325 nm, consistent with TNB -disulfide formation. Because of the reaction stoichiometry, it may be safely assumed the disulfide is a mixed disulfide of the TNB thiol and sulfenyl amide sulfur. Importantly, the mixed disulfide does not contribute to A_{410} .

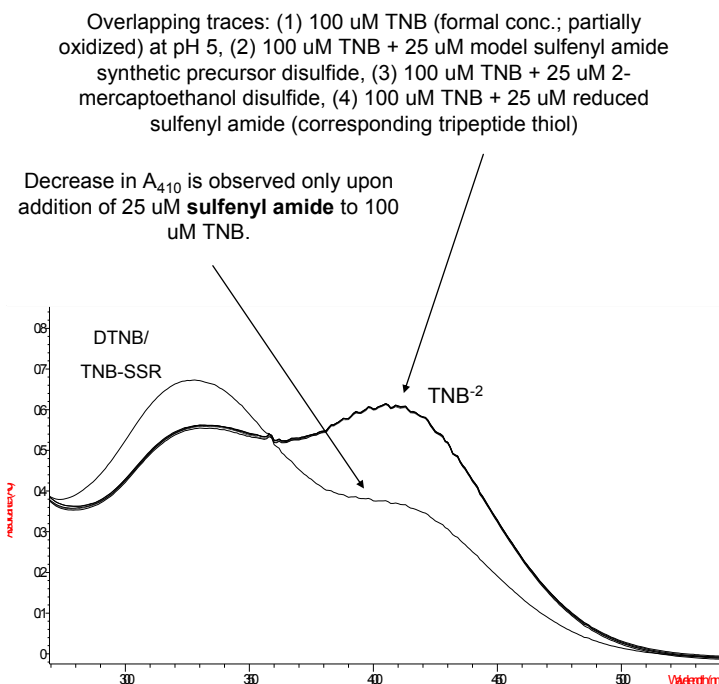


Figure S3. Determination of selectivity of the TNB assay to other potentially-interfering redox-active species. Oxidation of TNB^{-2} (to TNB-SSR or DTNB) was observed **only** upon addition of sulfenyl amide to the mixture; neither the disulfide synthetic precursor to the sulfenyl amide (see synthetic route), nor β -mercaptoethanol disulfide effected conversion of TNB^{-2} to the corresponding mixed disulfide (final concentrations indicated in Figure; all spectra collected in 1:1 MeOH : pH 5 Buffer C). Additionally, addition of thiol agent (the thiol corresponding to fully-reduced sulfenyl amide) did not convert any DTNB present to the TNB^{-2} ion. Finally, only the addition of sulfenyl amide (25 μM) afforded a decrease in A_{410} , corresponding to the conversion of TNB^{-2} to the TNB-model sulfenyl amide mixed disulfide ("TNB-SSR").

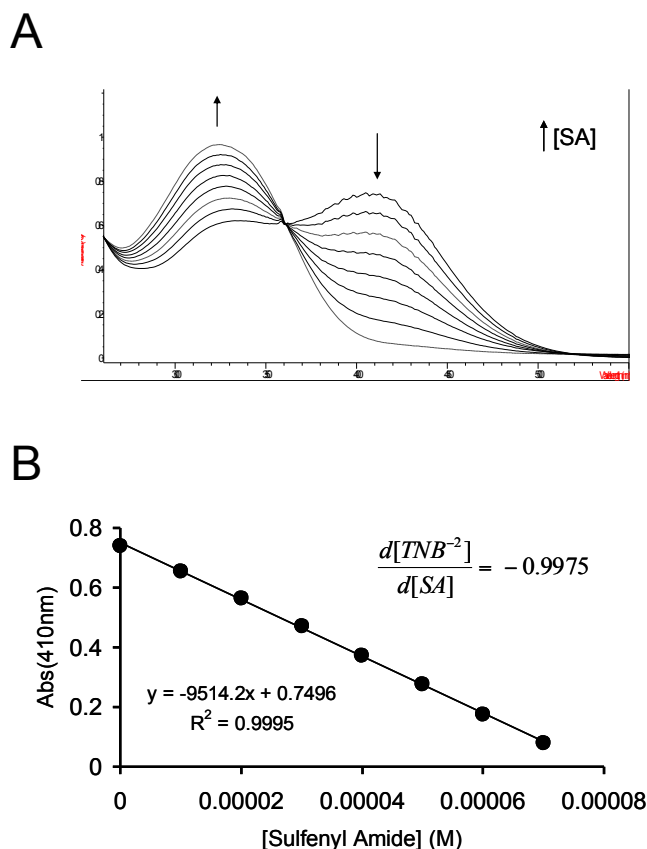
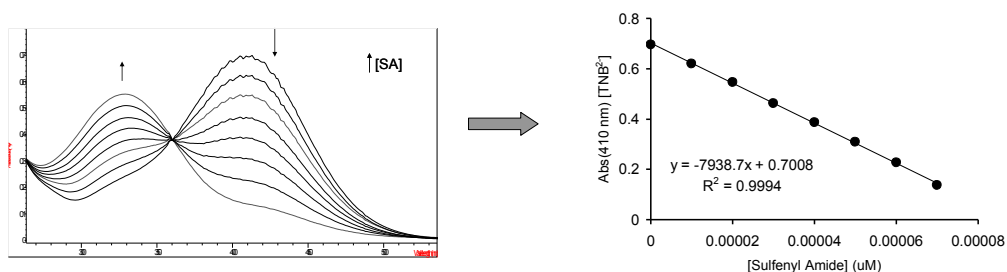


Figure S4. Determination of reaction stoichiometry between the TNB^{2-} ion and the model sulfenyl amide under the assay conditions employed (1:1 MeOH : Buffer C, pH 5.0). **A.** To a 1 mL solution of 1:1 MeOH : Buffer C containing 100 μM TNB^{2-} (formal conc.; partially oxidized by air) were added aliquots (2 μL) of 5 mM sulfenyl amide in dmsO, from 0 μM to 70 μM , final. As sulfenyl amide was titrated in, A_{410} smoothly decreased with concomitant increase in A_{325} , corresponding to conversion of TNB^{2-} to the TNB-model sulfenyl amide mixed disulfide ("TNB-SSR"). **B.** A plot of the absorbances at 410 nm versus molar concentration of the sulfenyl amide titrated in affords excellent linear correlation with a slope of -9514. In a separate experiment, wherein TNB^{2-} was maintained in fully reduced form by the presence of an excess (2 equiv) of known thiol-reducing agent TCEP (tris(2-carboxyethyl)phosphine), the approximate extinction coefficient of the TNB^{2-} ion under these conditions was found to be $\sim 9538 M^{-1} cm^{-1}$. Thus, using the established extinction coefficient and slope of the line relating $[TNB^{2-}]$ to molar concentration of sulfenyl amide titrated in, we find that the reaction proceeds stoichiometrically (1:1, as $d[TNB^{2-}]/d[SA] \approx 1.0$, inset in panel **B**). Furthermore, because instrument response is linear with respect to $[SA]$, monitoring A_{410} is a suitable means by which to track $[SA]$ in kinetics experiments under these conditions.

A



B

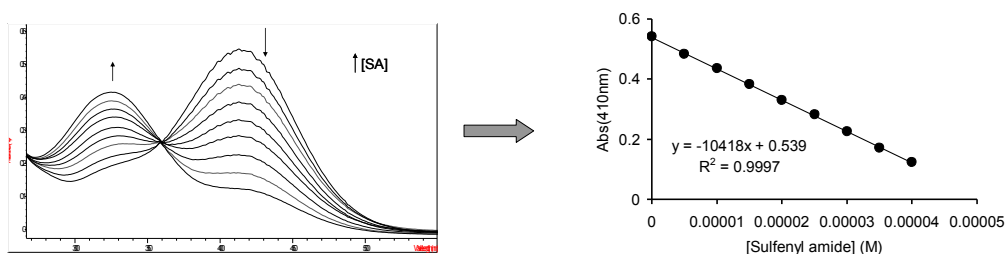
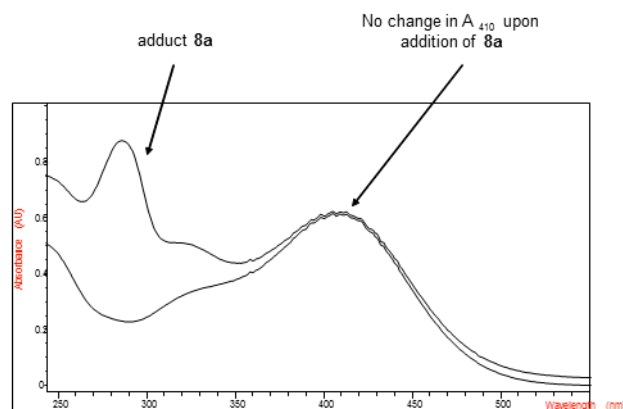


Figure S5. The "TNB assay" is robust with respect to presence or absence of organic cosolvent (methanol) and pH of the assay mixture. Essentially identical methodology to that described above in sulfenyl amide titrations was used here to determine instrument response to increasing concentrations of secondary analyte (sulfenyl amide). **A.** Increasing concentrations of sulfenyl amide were titrated into a solution of TNB⁻² in pure aqueous buffer, pH 5.0 (Buffer C). Excellent linear dependence of instrument response on molar concentration is observed in the absence of organic cosolvent (methanol). **B.** Again, excellent linear dependence of A_{410} on molar concentration of sulfenyl amide is observed when the assay solution pH is changed (1:1 MeOH : Buffer B (see text)). Note that the extinction coefficient is undoubtedly larger at pH 7, relative to that at pH 5 (due to an increased fraction of TNB thiolate at higher pH); hence the increased negative slope at pH 7 relative to that at pH 5.

A



B

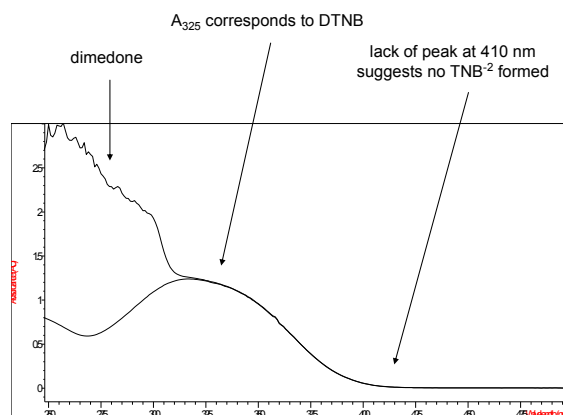


Figure S6. Thioether products and carbon nucleophiles do not interfere with the TNB assay. **A.** To confirm the sulfide products formed during kinetics experiments would not react with the TNB^{2-} ion, to a solution of $100\ \mu\text{M}$ TNB^{2-} in 1:1 MeOH : Buffer C ("standard assay conditions") was added product **8a** to a final concentration of $100\ \mu\text{M}$ and allowed to stand for 5 minutes. No change in A_{410} resulted from introduction of the sulfide product. **B.** Because kinetics assays were performed under pseudo first-order conditions (large molar excess of carbon nucleophile to sulfenyl amide), confirmation of lack of cross-reactivity was crucial. Namely, in principle, carbon nucleophiles could attack electrophilic disulfides (such as TNB-SSR or DTNB), releasing chromaphoric TNB^{2-} . However, when dimedone (**D**, $1\ \text{mM}$ final concentration) was added to a solution of $100\ \mu\text{M}$ DTNB (the disulfide of TNB^{2-}) in 1:1 MeOH : Buffer C and allowed to stand for 2 minutes, no release of TNB^{2-} (at A_{410}) was observed, confirming the "inertness" of excess nucleophilic carbon acid to the TNB assay under the conditions employed.

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