

Modeling the Reduction Kinetics of Munition Compounds by Humic Acids

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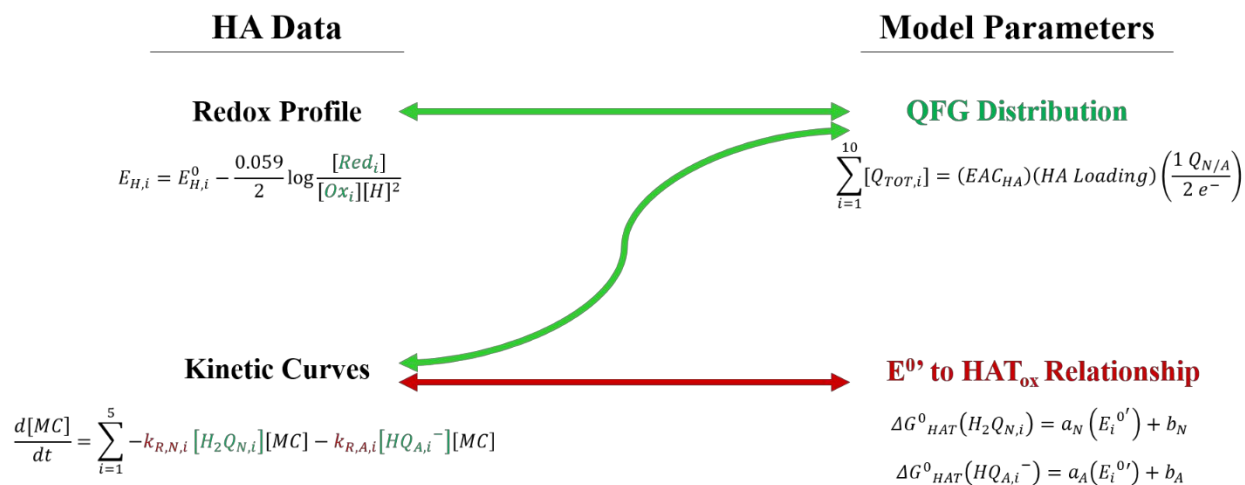
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HA Model Framework Description

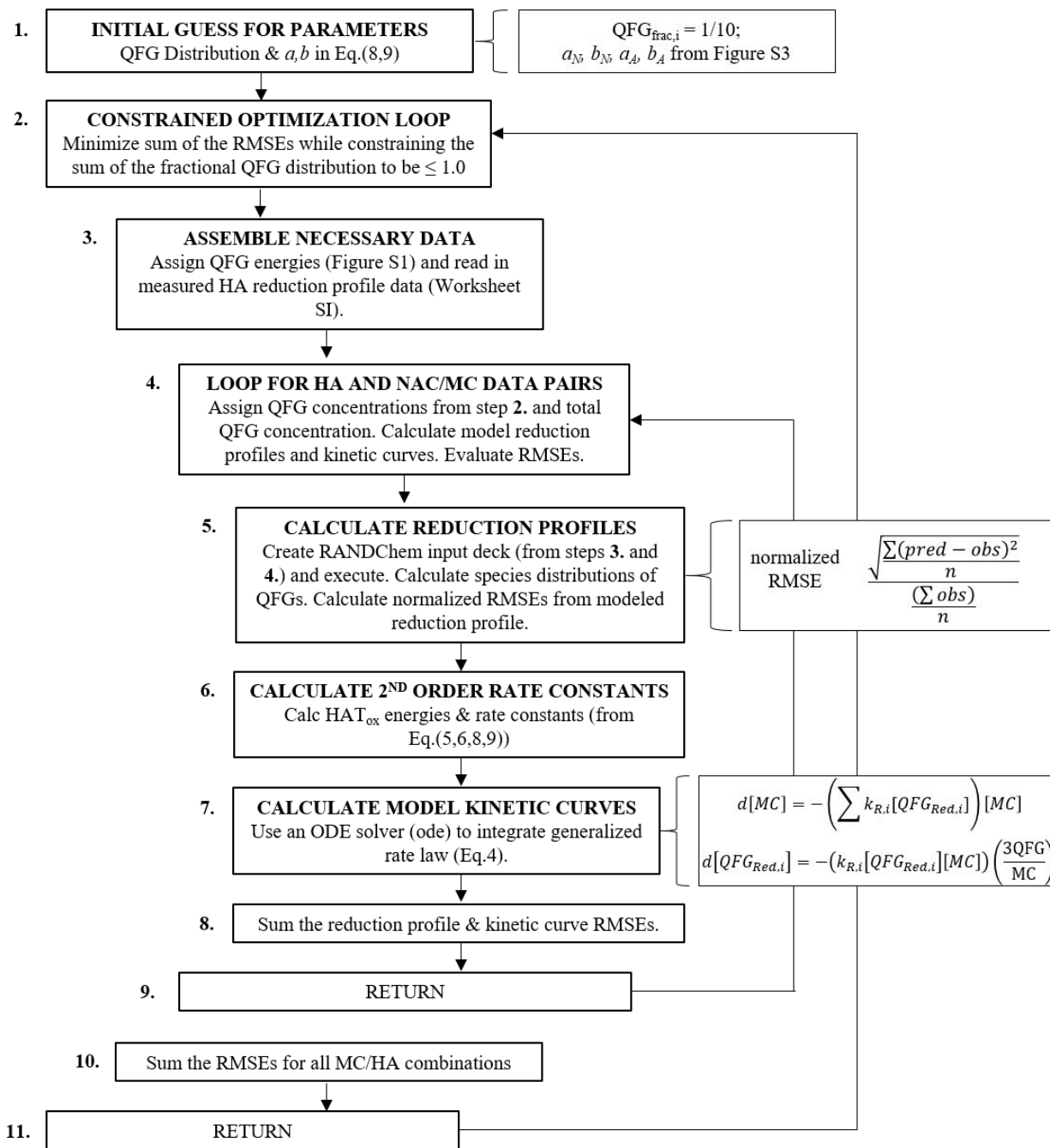
Scheme S1: Relationships between the HA data being optimized and the model's fitted parameters. Arrows indicate links between model parameters and the optimized data. Equations shown are predominant, but not exhaustive, for each segment.



The general framework for the humic acid (HA) model is shown in Scheme S1. The HA reduction profiles and NAC/MC kinetic curves serve as the thermodynamic and kinetic characteristics of a given HA source and are used to calibrate the quinone-like functional groups (QFGs) that comprise the HA model. The unique parameters being fit through this calibration are the QFG distribution and the formal reduction potential ($E^{0'}$) to HAT_{ox} energy relationships, as shown in Scheme S1.

The optimization procedure used for HA model calibration is shown in Scheme S2. The program was written and executed in R and utilizes an additional equilibrium speciation program (RANDChem), which was written and compiled in Fortran. The optimization program minimizes the root mean square errors (RMSEs) associated with both the HA reduction profiles and the NAC/MC kinetic curves concurrently through fitting the model parameters outlined in Scheme S1. It does this through a constrained optimization (constrOptim) which ensures that the sum of the QFG fractions does not exceed 1.0.

Scheme S2: Program schematic for optimization procedure. Written in R programming language with the use of RandChem equilibrium speciation program (Fortran).



Thermodynamic Description of Model Quinone-like Functional Groups

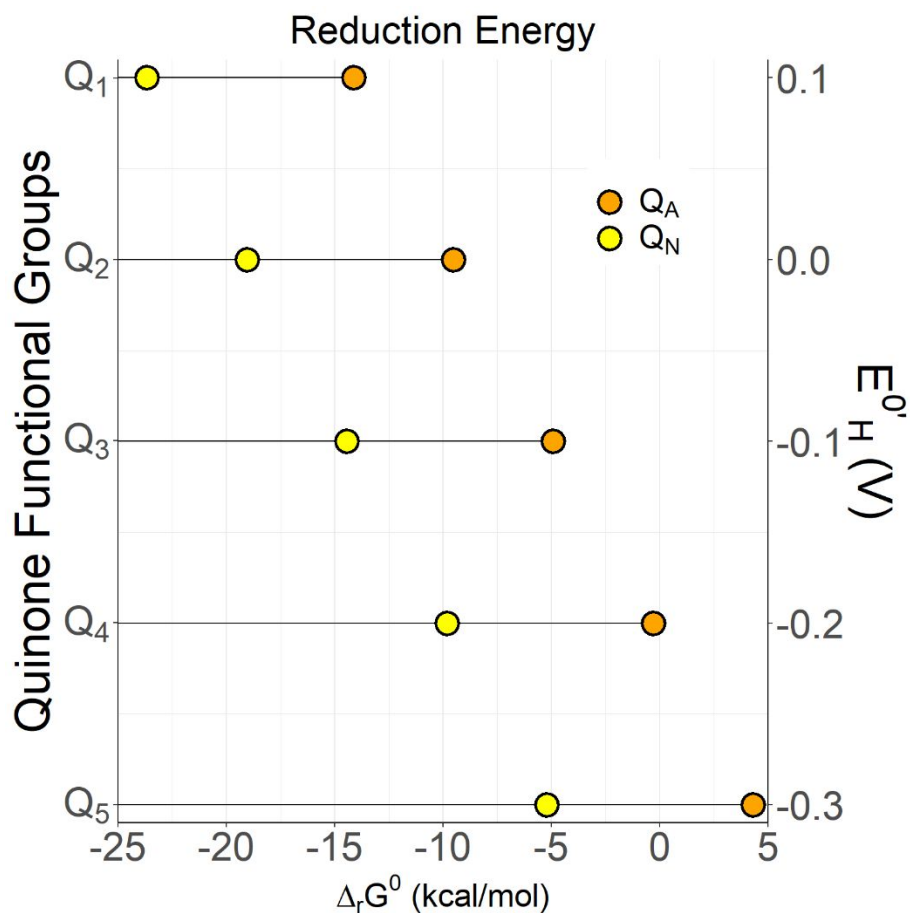


Figure S1: Energy Distribution of the ten model functional groups. Reduction energies were distributed evenly between the energy boundaries provided by the reduction profile. $E^0{}_H$ is the formal potential (pH = 7) relative to the standard hydrogen electrode. Q_N and Q_A are the neutral and anionic function groups respectively. $\Delta_r G^0$ corresponds to the standard state Gibbs free energy (pH = 0) of the reduction reactions for the QFGs (Eqs. 1, 2 in manuscript) and are used as their formation energies in the speciation calculations.

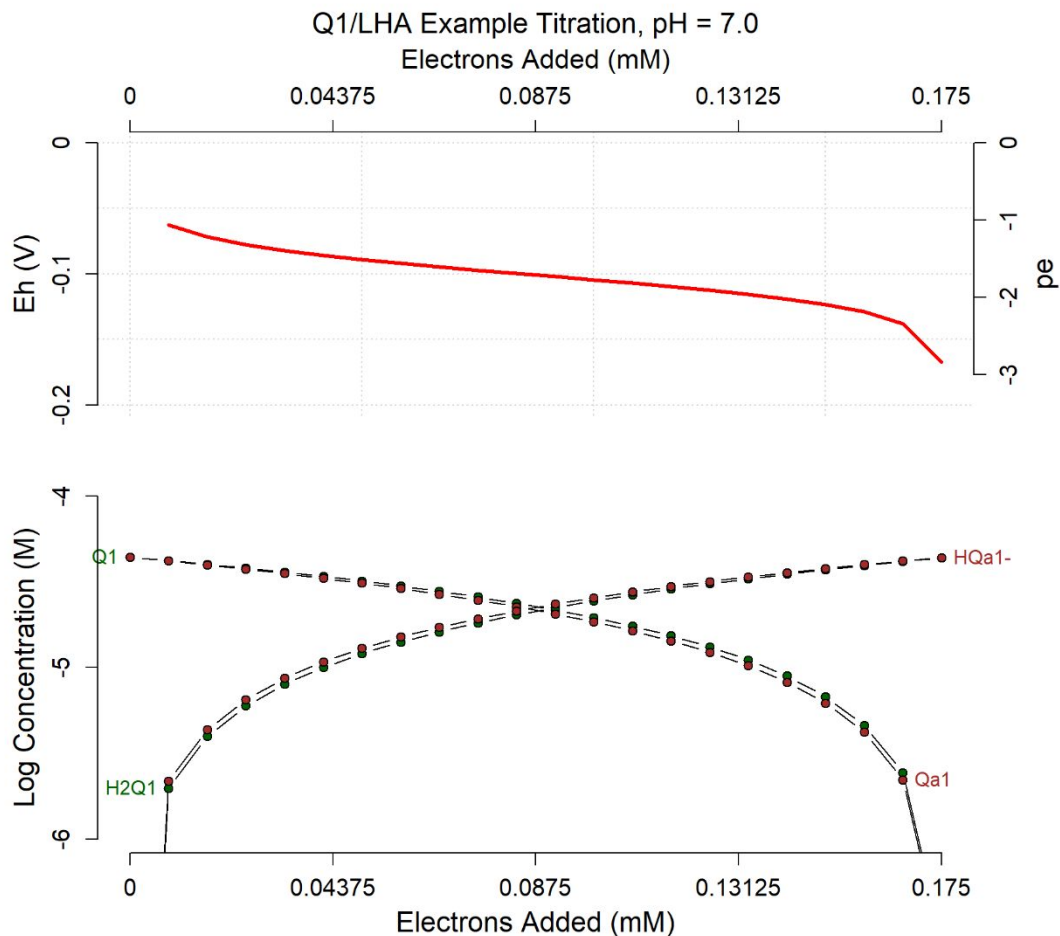


Figure S2: Results of example tableau in table 1. Model HA is comprised of one neutral (Q_{N1}) and one anionic (Q_{A1}) functional group at a 50/50 concentration ratio. Total functional group concentration is $88 \mu\text{M}$ as calculated from the EAC of LHA. The Q_{N1}/Q_{A1} reduction energies were selected as Q3 in Figure S1 simply to bisect the measured values. Top: Red line is model output. Bottom: Concentration of the oxidized and reduced form of the two functional groups as a function of electrons titrated, or reduction extent.

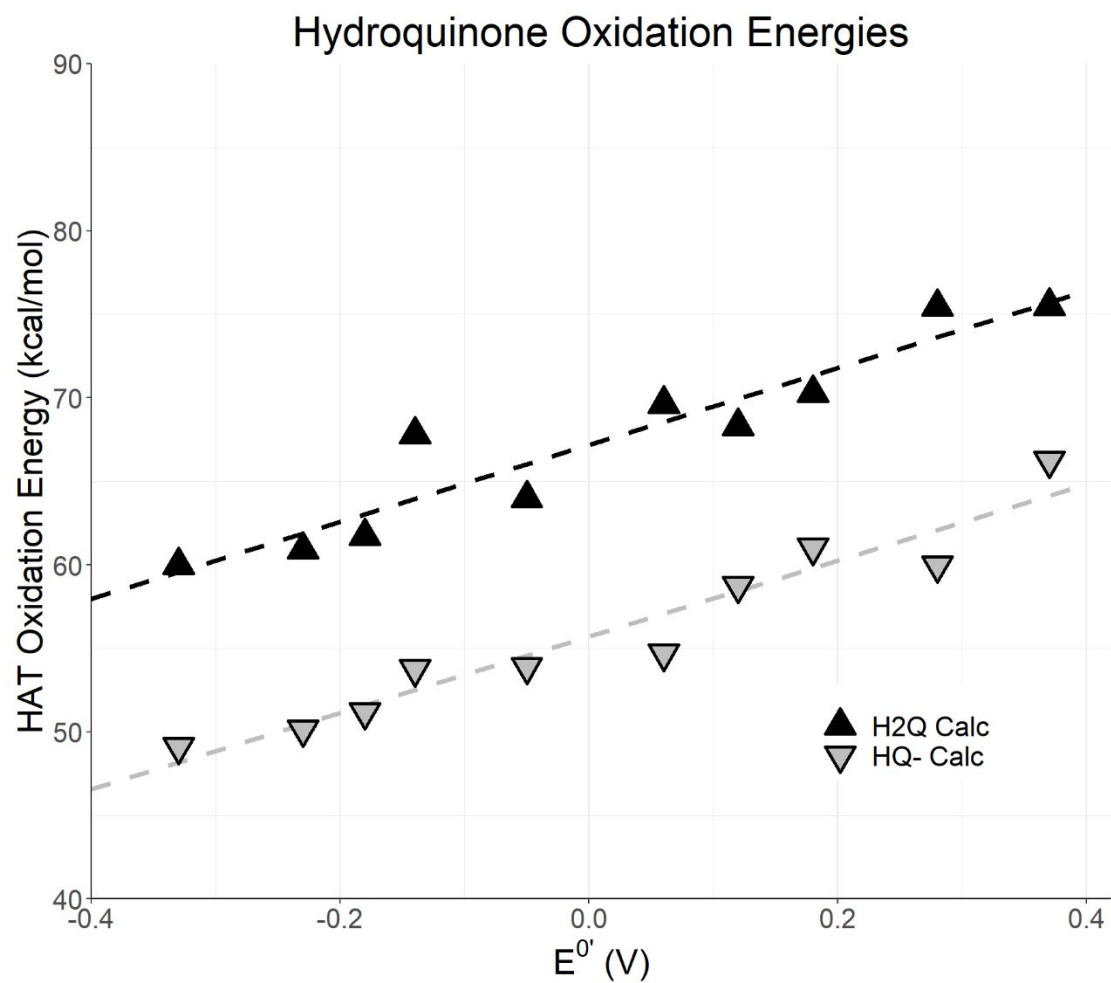


Figure S3: Calculated HAT oxidation energies as a function of reported formal potential for ten selected hydroquinones.

Functional Group Distribution and Energies

Table S1: Model fitted functional group distribution for the mean HA model.

Group	Composition (Fraction)
H ₂ Q _{N1}	0.043919
H ₂ Q _{N2}	0.01959
H ₂ Q _{N3}	0.111252
H ₂ Q _{N4}	0.410781
H ₂ Q _{N5}	0.147647
HQ _{A1} ⁻	0.002697
HQ _{A2} ⁻	0.089338
HQ _{A3} ⁻	0.045164
HQ _{A4} ⁻	0.085892
HQ _{A5} ⁻	0.043541

Table S2: Formal potentials and HAT oxidation energies for ten real hydroquinones and five model quinone-like functional groups. HAT energies for real hydroquinones are calculated quantum chemically, while the energies for the functional groups are calculated from Eq.8 with slopes and intercepts fitted through the mean HA model.

Quinone Functional Group		E ^{0'} (V)	HAT _{ox} Energies ^a	
			Neutral Form ^b (kcal/mol)	Anionic Form ^c (kcal/mol)
Existing Hydroquinones	1,2-benzoquinone	0.37	75.508	66.254
	1,4-benzoquinone	0.28	75.454	59.971
	2-hydroxy-1,4-benzoquinone	0.18	70.288	61.036
	1,4-naphthoquinone-2-sulfonate	0.12	68.317	58.766
	1,4-naphthoquinone	0.06	69.645	54.693
	5,8-dihydroxy-1,4-naphthoquinone	-0.05	63.989	53.904
	2-hydroxy-1,4-naphthoquinone	-0.14	67.818	53.765
	9,10-anthraquinone-2,6-disulfonate	-0.18	61.731	51.189
	9,10-anthraquinone-2-sulfonate	-0.23	60.922	50.163
	9,10-anthraquinone	-0.33	59.991	49.120
Model Functional Groups	Q ₁	0.10	84.220	73.371
	Q ₂	0.00	78.931	67.417
	Q ₃	-0.10	73.647	61.475
	Q ₄	-0.20	68.348	55.534
	Q ₅	-0.30	63.061	49.592

^a HAT_{ox} = a(E^{0'}) + b

^b a = 52.89; b = 78.98

^c a = 59.44; b = 67.45

Table S3: Conditions of NAC/MC reduction with HA experiments.

HA/NAC	pH	HA Loading (g/L)	EAC (mmol e ⁻ /g _{HA})
LHA/NTO	6.5, 8.0	0.5	1.7
LHA/DNAN	6.5	0.5	1.7
LHA/NB	7.5	0.1	1.7
ESHA/NTO	7.0	0.5	1.96
SRHA/NTO	7.0	0.5	0.8
PPHA/NTO	7.0	0.5	1.7

Reduction Profile of All HAs in Study

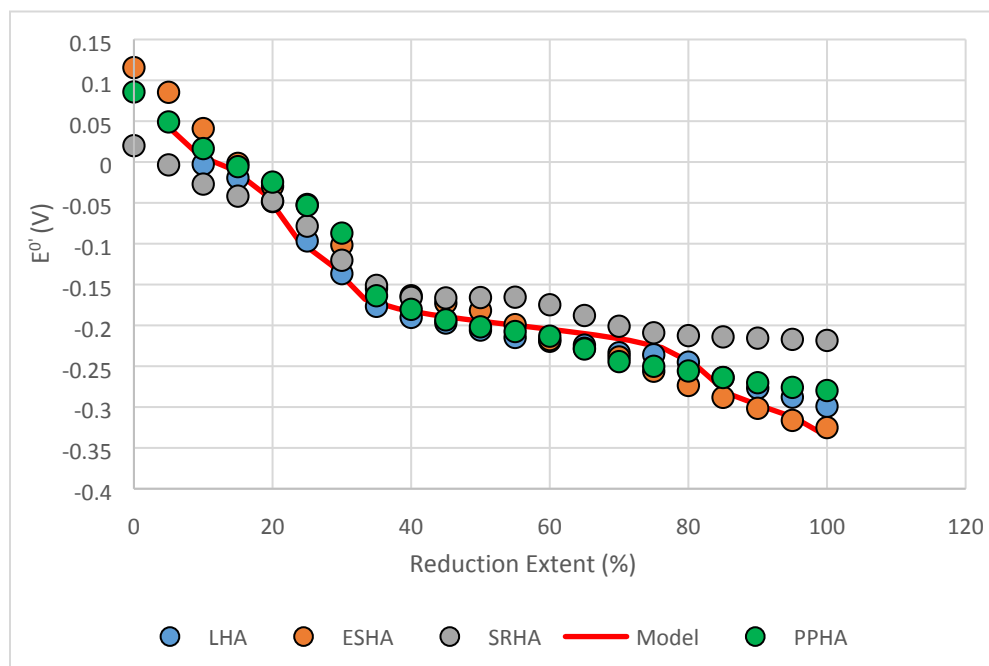


Figure S4: Measured reduction profiles for Leonardite Humic Acid (LHA), Eliot Soil Humic Acid (ESHA), Suwanee River Humic Acid (SRHA), and Pahokee Peat Humic Acid (PPHA), as well as the Mean HA model fit.

Independent ESHA/NTO Optimization

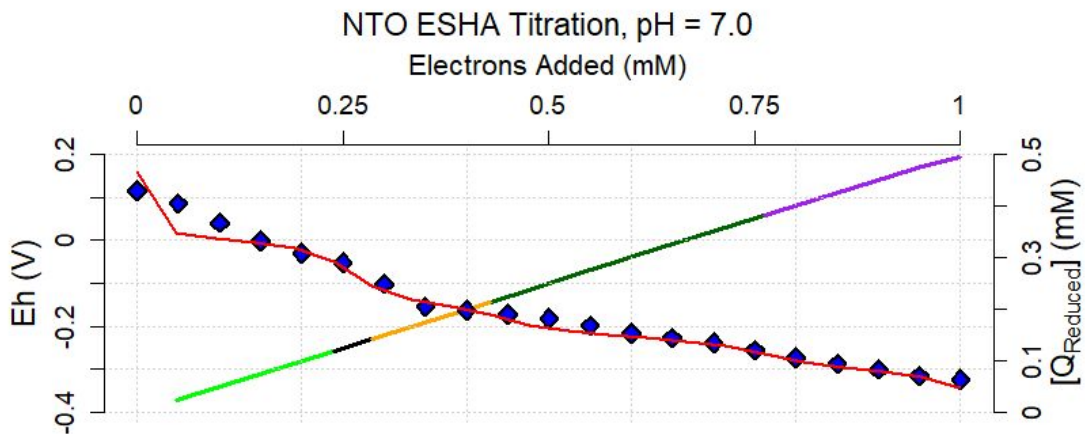


Figure S5: Reduction profile as potential vs. HA reduction extent for ESHA (blue diamonds) and the model fit (red line). Colored line on base indicates the active or redox controlling functional group as a function of reduction extent (right Y axis).

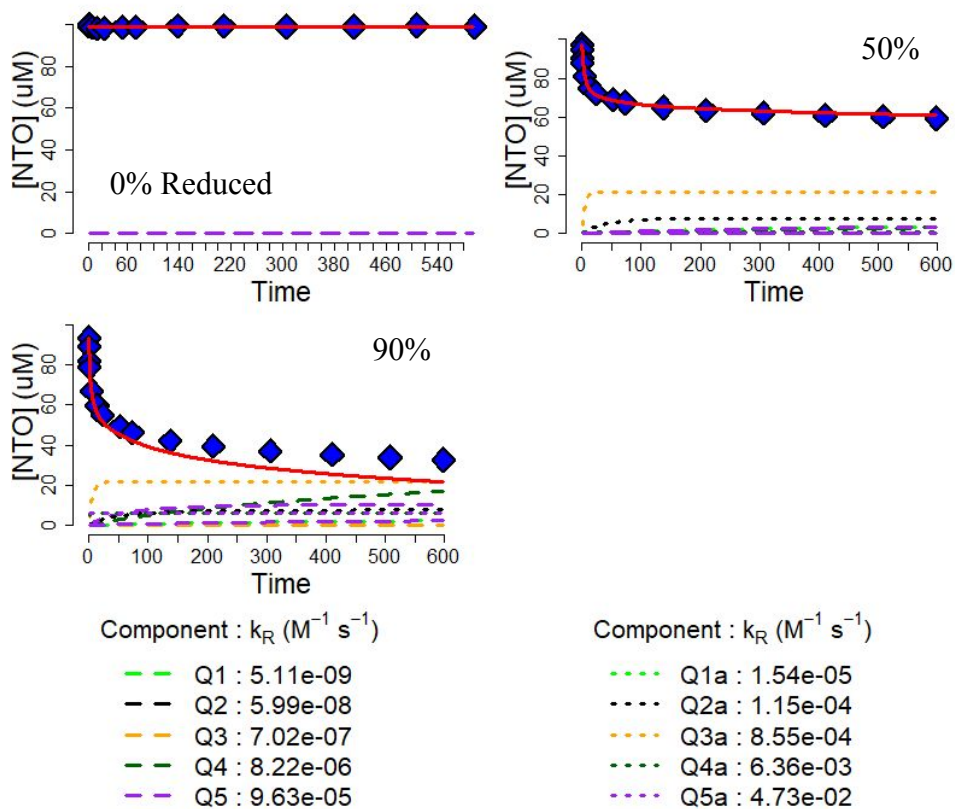


Figure S6: NTO concentration as a function of time for reduction with Eliot Soil Humic Acid (ESHA). Dashed and dotted lines correspond to the contribution of NTO loss from individual QFGs (dashed = neutral, dotted = anionic).

Example of NAC/MC Reduction kinetic Curves in Linear-Linear Scale

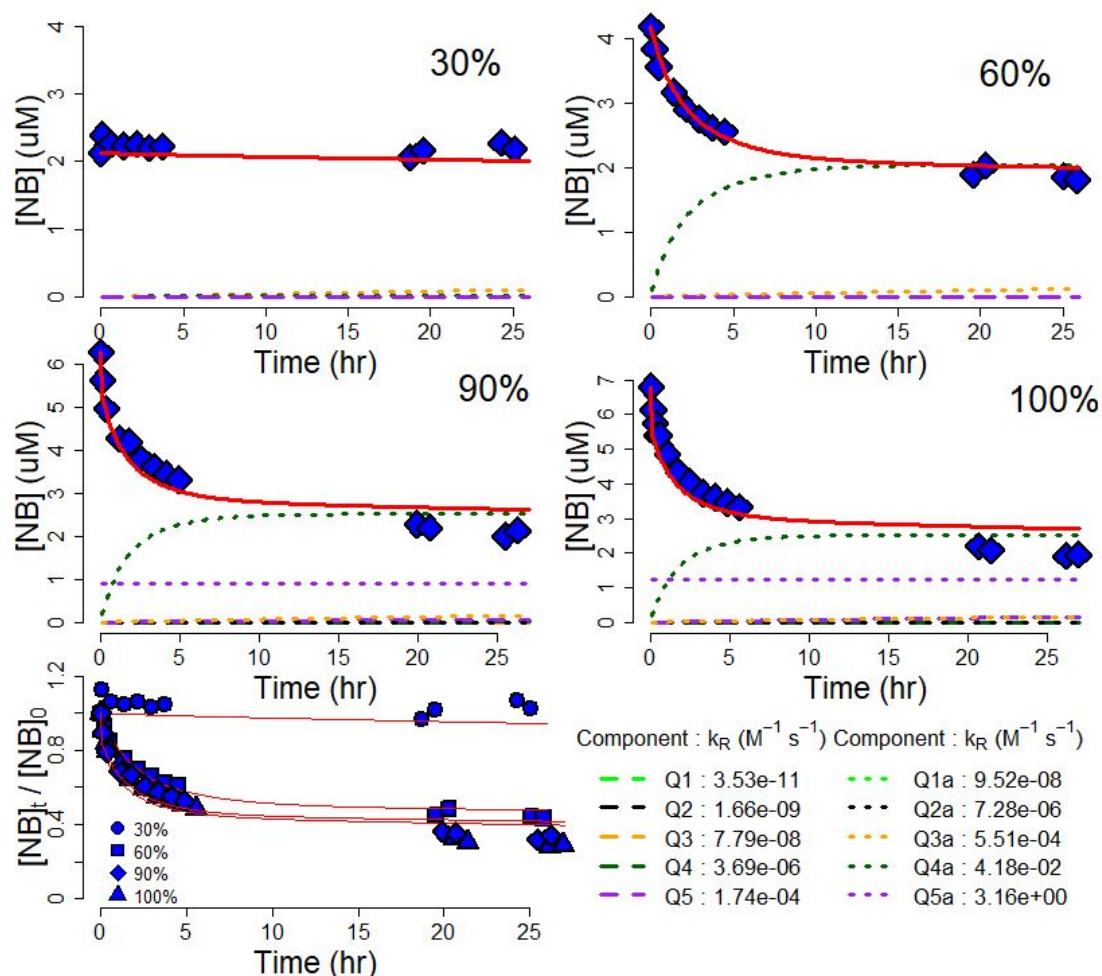


Figure S7: NB concentration as a function of time for reduction with Leonardite Humic Acid (LHA) at 30%, 60%, 90% and 100% reduced LHA. Same data presented in Figure 4, row A but plotted on a linear-linear scale. Dashed and dotted lines correspond to the contribution of NB loss from individual QFGs (dashed = neutral, dotted = anionic).

Example LHA Input Deck

LHA_DNAN_10Q
LITER
1.0
MULTIPLIERS

RT
25.
LIMIT

1000
ROWS

H2O	1			
	.000	.000	.000	.000
Hp	0			
	.000	.000	.000	.000
Em	0			
	.000	.000	.000	.000
Q1	0			
	.000	.000	.000	.000
Q2	0			
	.000	.000	.000	.000
Q3	0			
	.000	.000	.000	.000
Q4	0			
	.000	.000	.000	.000
Q5	0			
	.000	.000	.000	.000
Qd1	0			
	.000	.000	.000	.000
Qd2	0			
	.000	.000	.000	.000
Qd3	0			
	.000	.000	.000	.000
Qd4	0			
	.000	.000	.000	.000
Qd5	0			
	.000	.000	.000	.000

END

MATRIX

Aqueou	.00	.00	.00		.0	
H2O	-56.69	.00	.00	1.0H2O		.0
OHm	-37.59	.00	.00	1.0H2O	-1.0Hp	.0
Hp	.00	.00	.00	1.0Hp		.0
O2	3.92	.00	.00	2.0H2O	-4.0Hp	-4.0Em
H2	4.207	.00	.00	2.0Hp	2.0Em	
Q1	0.000	.00	.00	1.0Q7		
H2Q1	-12.63	.00	.00	1.0Q7	2.0Hp	2.0Em
Q2	0.000	.00	.00	1.0Qa		
H2Q2	-12.00	.00	.00	1.0Qa	2.0Hp	2.0Em
Q3	0.000	.00	.00	1.0Qb		
H2Q3	-11.25	.00	.00	1.0Qb	2.0Hp	2.0Em
Q4	0.000	.00	.00	1.0Q8		
H2Q4	-10.79	.00	.00	1.0Q8	2.0Hp	2.0Em
Q5	0.000	.00	.00	1.0Q9		
H2Q5	-8.482	.00	.00	1.0Q9	2.0Hp	2.0Em
Qd1	0.000	.00	.00	1.0Qd7		
H2Qd1	-12.63	.00	.00	1.0Qd7	1.0Hp	2.0Em
Qd2	0.000	.00	.00	1.0Qda		

H2Qd2	-12.00	.00	.00	1.0Qda	1.0Hp	2.0Em
Qd3	0.000	.00	.00	1.0Qdb		
H2Qd3	-11.25	.00	.00	1.0Qdb	1.0Hp	2.0Em
Qd4	0.000	.00	.00	1.0Qd8		
H2Qd4	-10.79	.00	.00	1.0Qd8	1.0Hp	2.0Em
Qd5	0.000	.00	.00	1.0Qd9		
H2Qd5	-8.482	.00	.00	1.0Qd9	1.0Hp	2.0Em
pH	.000	.00	.00			
HPLUS	.000	.00	.00	1.0Hp		

END

NOMESSAGES

RETURN

21	Aqueou		
H2O	1.000	1.000	1.0
Em	1.0e-9	8.5e-4	1.0
Q1	3.6e-5	3.6e-5	1.0
Q2	3.6e-5	3.6e-5	1.0
Q3	3.6e-5	3.6e-5	1.0
Q4	3.6e-5	3.6e-5	1.0
Q5	3.6e-5	3.6e-5	1.0
Qd1	3.6e-5	3.6e-5	1.0
Qd2	3.6e-5	3.6e-5	1.0
Qd3	3.6e-5	3.6e-5	1.0
Qd4	3.6e-5	3.6e-5	1.0
Qd5	3.6e-5	3.6e-5	1.0
HPLUS	1.0	1.0	1.0
pH	7.0	7.0	1.0

END

EXIT