## **Supporting Information**

## Solvent and Anion Effects on the Electrochemistry of Manganese Dipyrrin-bisphenols

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**Figure S2.** Spin density plot of  $[(Ph)DPP^{\bullet}Mn^{III}]^+$  showing the distribution of spin over the metal center (3.91) and dipyrrin bisphenol framework (-0.91).



**Figure S3.** Plot of wavenumber (cm<sup>-1</sup>) for the major visible band of the neutral Mn<sup>III</sup> and singly reduced Mn<sup>II</sup> DPP complexes and the NIR band of the singly oxidized product  $vs \sigma$  in CH<sub>3</sub>CN containing 0.1 M TBAPF<sub>6</sub>.



**Figure S4.** Comparison of UV-vis spectral changes associated with the first metal-centered reduction (*i.e.*  $M^{III/II}$ ) of (Ar)DPPMn^{III} and (Ar)DPPCo^{III}(Py)\_2 in pyridine containing 0.1 M TBAPF\_6.



**Figure S5.** Plot of computationally determined LUMO energies vs the experimentally measured reduction potential for (Ph)<sub>3</sub>CorMn, (Ph)DPPMn and (Ph)<sub>4</sub>PorMnCl.



**Figure S6.** Plot of computationally determined HOMO energies vs the experimentally measured oxidation potential for (Ph)<sub>3</sub>CorMn, (Ph)DPPMn and (Ph)<sub>4</sub>PorMnCl.



Figure S7. LR-MS and HR-MS (ESI) mass spectra of (Ph)DPPMn<sup>III</sup> in MeOH.



Figure S8. LR-MS and HR-MS (ESI) mass spectra of (Mes)DPPMn<sup>III</sup> in MeOH.