

SUPPORTING INFORMATION

For

Highly effective removal of pharmaceutical compounds from aqueous solution

by magnetic Zr-based MOFs composites

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$$\frac{t}{Q_t} = \frac{1}{k_2 Q_e^2} + \frac{1}{Q_e} t$$

where k_2 ($\text{g mg}^{-1}\text{min}^{-1}$) is the kinetic constant of pseudo-second-order adsorption. Q_e and Q_t (mg g^{-1}) are the amount of the target absorbed at equilibrium time e and any time t (min).

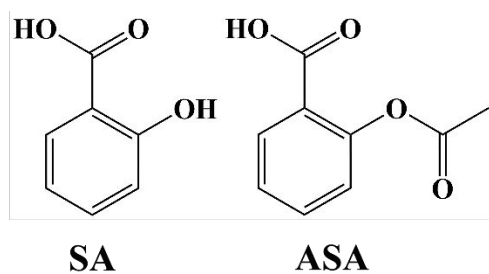


Figure S1. Structures of salicylic acid and acetylsalicylic acid.

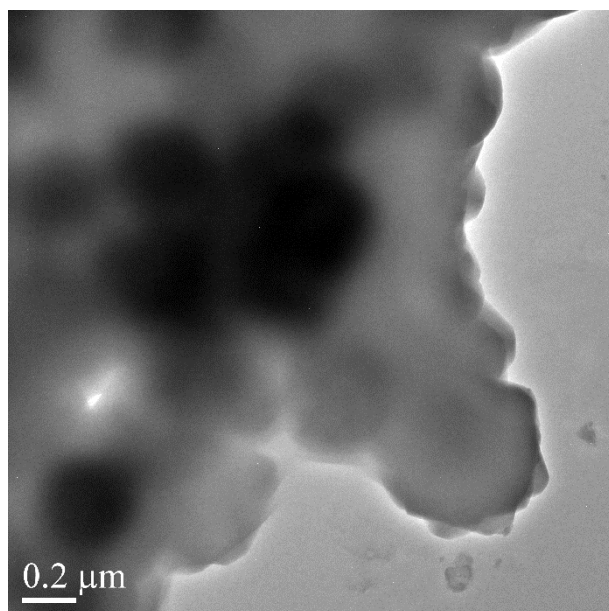


Figure S2. TEM images of $\text{Fe}_3\text{O}_4@\text{SiO}_2@\text{UiO-66-NH}_2$.

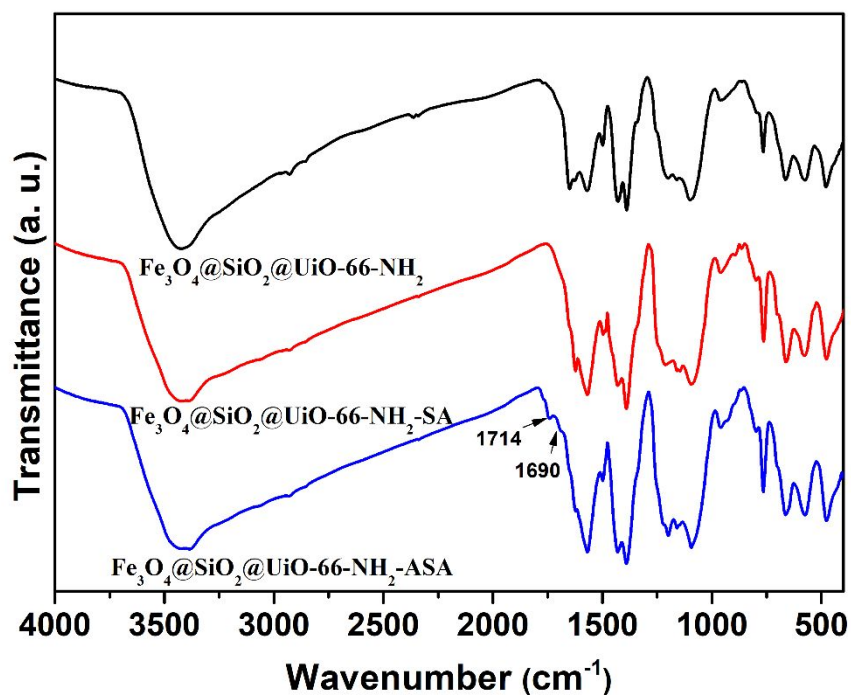


Figure S3. FTIR spectra of the magnetic UiO-66-NH₂ before and after adsorption of SA and ASA.

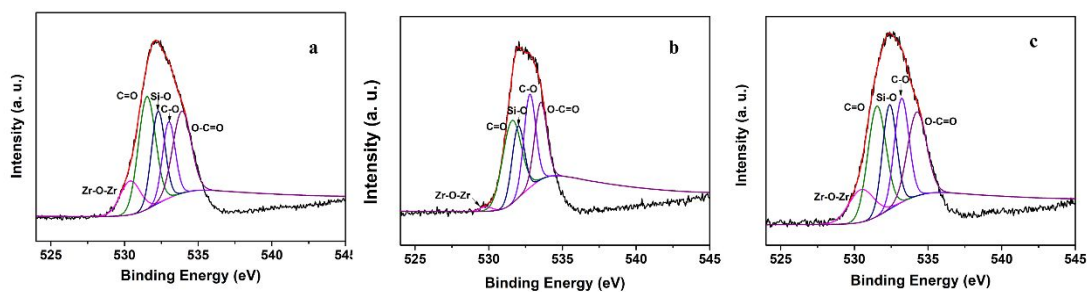


Figure S4. The O1s XPS spectra of the magnetic UiO-66-NH₂ before (a) and after adsorption of SA (b) and ASA (c).

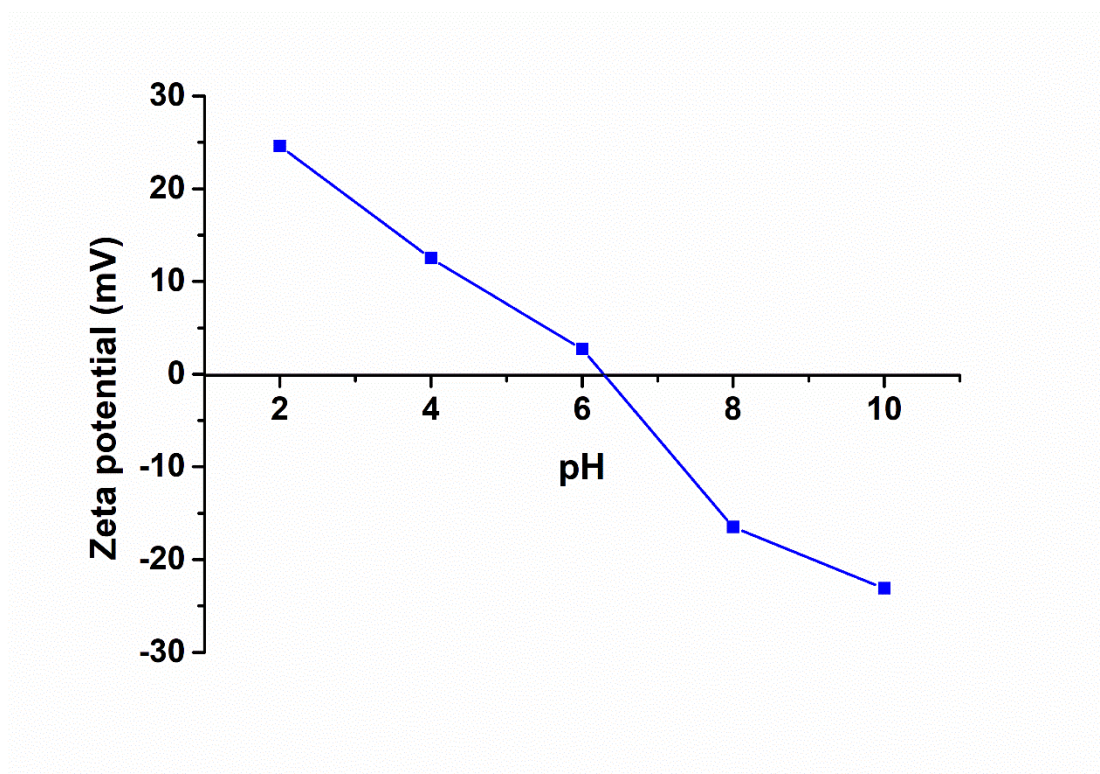


Figure S5. The zeta potential of the magnetic UiO-66-NH₂ as a function of pH.

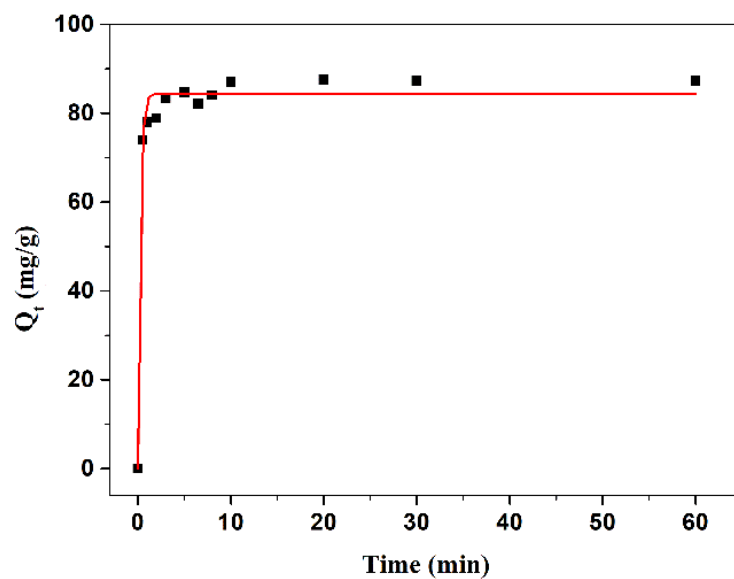


Figure S6. Pseudo-first-order kinetics of SA and adsorption on magnetic UiO-66-NH₂.

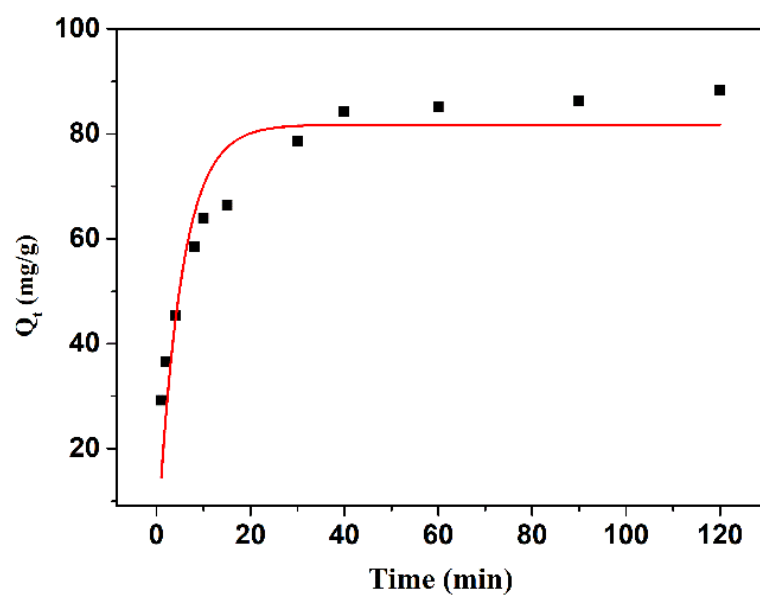


Figure. S7. Pseudo-first-order kinetics of ASA and adsorption on magnetic UiO-66-NH₂.