

Supporting Information

Combustion Mechanisms and Kinetics of Fuel Additive: A ReaxFF Molecular Simulation

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The force field used in this work

Reactive MD-force field: Combustion C/H/O force field - September 7, 2016

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39      ! Number of general parameters
50.0000 !p(boc1)
 9.5469 !p(boc2)
26.5405 !p(coa2)
 0.6863 !p(trip4)
 2.7295 !p(trip3)
70.0000 !kc2
 1.0588 !p(ovun6)
 4.1262 !p(trip2)
12.1176 !p(ovun7)
13.3056 !p(ovun8)
-68.9784 !p(trip1)
 0.0000 !Lower Taper-radius (swa)
10.0000 !Upper Taper-radius (swb)
 0.0000 !not used
33.8667 !p(val7)
 6.0891 !p(lp1)
 1.0563 !p(val9)
 2.0384 !p(val10)
 6.1431 !not used
 6.9290 !p(pen2)
 0.3989 !p(pen3)
 3.9954 !p(pen4)
 0.0000 !not used
 5.7796 !p(tor2)
10.0000 !p(tor3)
 1.9487 !p(tor4)
 0.0000 !not used
 2.1645 !p(cot2)
 1.5591 !p(vdW1)
 0.1000 !Cutoff for bond order*100 (cutoff)
 2.1365 !p(coa4)
 0.6991 !p(ovun4)
50.0000 !p(ovun3)
 1.8512 !p(val8)
 0.0000 !not used
 0.0000 !not used
 0.0000 !not used
 1.0000 !not used
 2.6962 !p(coa3)
3      ! Nr of atoms; atomID;ro(sigma); Val;atom mass;Rvdw;Dij;gamma
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alfa;gamma(w);Val(angle);p(ovun5);n.u.;chiEEM;etaEEM;n.u.
 ro(pipi);p(lp2);Heat increment;p(boc4);p(boc3);p(boc5);n.u.;n.u.
 p(ovun2);p(val3);n.u.;Val(boc);p(val5);n.u.;n.u.;n.u.

C		1.3674	4.0000	12.0000	2.0453	0.1444	0.9500	1.1706	4.0000
		9.0000	1.5000	4.0000	27.5134	79.5548	5.0191	7.0500	0.0000
		1.1168	0.0000	181.0000	14.2732	24.4406	6.7313	0.8563	0.0000
		-4.1021	5.0000	1.0564	4.0000	2.9663	0.0000	0.0000	0.0000
H		0.9479	1.0000	1.0080	1.1364	0.0232	0.9900	-0.1000	1.0000
		9.0643	4.7746	1.0000	0.0000	121.1250	4.7757	9.7732	1.0000
		-0.1000	0.0000	62.4879	2.5194	2.3785	0.2223	1.0698	0.0000
		-15.7683	2.1488	1.0338	1.0000	2.8793	0.0000	0.0000	0.0000
O		1.1939	2.0000	15.9990	1.9289	0.1201	0.9900	1.0981	6.0000
		10.4842	8.2916	4.0000	28.8967	116.0768	7.9703	7.0500	2.0000
		1.0479	20.0000	60.8726	10.0338	2.2024	0.9942	0.9745	0.0000
		-3.6141	2.7025	1.0493	4.0000	2.9225	0.0000	0.0000	0.0000

6 ! Nr of bonds; at1;at2;De(sigma);De(pi);De(pipi);p(be1);p(b
 p(be2);p(bo3);p(bo4);n.u.;p(bo1);p(bo2)

1	1	80.8865	107.9944	52.0636	0.5218	-0.3636	1.0000	34.9876	0.7769
		6.1244	-0.1693	8.0804	1.0000	-0.0586	8.1850	1.0000	0.0000
1	2	179.5195	0.0000	0.0000	-0.5242	0.0000	1.0000	6.0000	0.7187
		5.4740	1.0000	0.0000	1.0000	-0.1144	6.7029	0.0000	0.0000
2	2	113.9232	0.0000	0.0000	-0.5971	0.0000	1.0000	6.0000	0.9093
		1.7152	1.0000	0.0000	1.0000	-0.0450	6.0710	0.0000	0.0000
1	3	136.4945	164.1201	5.5000	-0.9159	-0.1075	1.0000	10.6519	0.8644
		0.6858	-0.4602	9.5754	1.0000	-0.1745	4.5987	0.0000	0.0000
3	3	148.0798	155.2406	20.1160	-1.0000	-0.1254	1.0000	33.0027	0.7790
		0.7673	-0.1697	7.0028	1.0000	-0.1300	5.1959	1.0000	0.0000
2	3	169.1351	0.0000	0.0000	-0.8810	0.0000	1.0000	6.0000	0.5757
		1.5482	1.0000	0.0000	1.0000	-0.1788	4.6622	0.0000	0.0000

3 ! Nr of off-diagonal terms. at1;at2;Dij;RvdW;alfa;ro(sigma);r

1	2	0.1253	1.5717	9.9736	1.2057	-1.0000	-1.0000
2	3	0.1125	1.6311	8.7528	1.0929	-1.0000	-1.0000
1	3	0.0953	1.7397	8.8986	1.4256	1.1067	1.1265

18 ! Nr of angles. at1;at2;at3;Thetao,o;p(val1);p(val2);p(coal1);

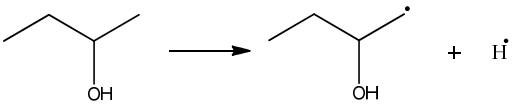
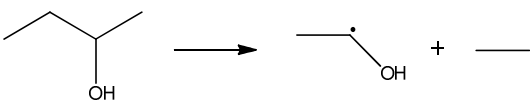
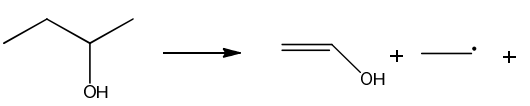
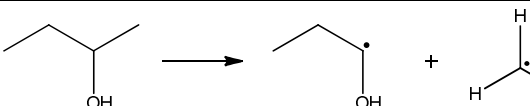
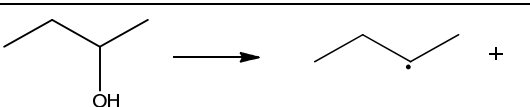
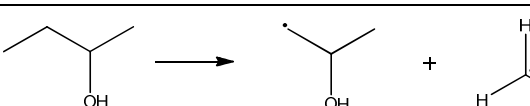
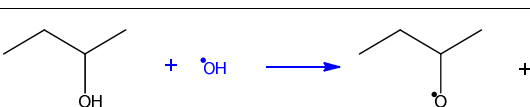
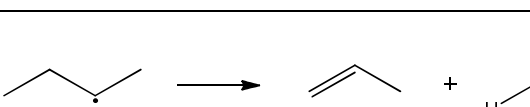
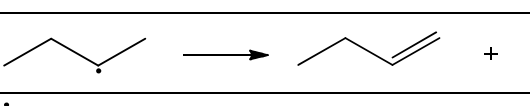
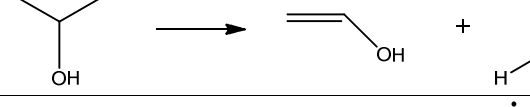
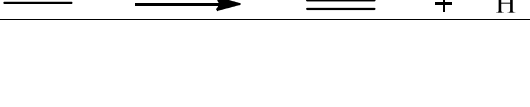
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1	1	2	68.0572	9.9461	4.7000	0.0000	0.4566	0.0000	1.8532
2	1	2	65.6815	35.0000	1.8622	0.0000	0.0490	0.0000	1.0937
1	2	2	0.0000	4.0000	7.2043	0.0000	0.0000	0.0000	1.0728
1	2	1	0.0000	3.4110	7.7350	0.0000	0.0000	0.0000	1.0400
2	2	2	0.0000	30.0000	5.6235	0.0000	0.0000	0.0000	1.0400
1	1	3	78.3624	13.0773	9.0480	0.0000	0.1270	52.1129	2.3964
3	1	3	76.7101	24.3833	5.8613	-21.8559	2.6395	-32.6534	3.6179
2	1	3	79.1288	30.0000	1.4632	0.0000	0.2065	0.0000	2.0000
1	3	1	80.7352	16.4130	4.9987	0.0000	0.0843	0.0000	1.0137

1	3	3	85.4436	14.4937	3.9928	0.0000	1.4350	44.5320	1.1348	
3	3	3	89.9282	32.1199	2.7181	0.0000	0.3323	57.6122	1.0000	
1	3	2	82.9640	32.4874	0.8777	0.0000	0.9627	0.0000	1.0010	
2	3	3	85.7838	17.3139	1.9157	0.0000	3.6306	0.0000	2.1858	
2	3	2	84.2527	33.1226	0.6730	0.0000	0.7238	0.0000	2.4348	
1	2	3	0.0000	14.4588	3.1507	0.0000	3.4571	0.0000	1.0149	
3	2	3	0.0000	0.9696	3.6303	0.0000	0.0000	0.0000	1.6987	
2	2	3	0.0000	0.5797	1.9739	0.0000	0.0000	0.0000	2.4494	
21	! Nr of torsions. at1;at2;at3;at4;;V1;V2;V3;p(tor1);p(cot1);n									
1	1	1	1	2.0474	32.6719	0.5282	-9.0000	-2.6449	0.0000	0.0000
1	1	1	2	1.6328	78.4995	-0.1514	-6.9161	-2.9986	0.0000	0.0000
2	1	1	2	2.4142	78.7025	0.3506	-8.8640	-6.9283	0.0000	0.0000
1	1	1	3	-0.7104	22.6038	0.5309	-2.0000	-0.6614	0.0000	0.0000
2	1	1	3	1.9323	52.9368	0.6554	-8.8118	-3.9854	0.0000	0.0000
3	1	1	3	-1.2500	1.1248	-0.1230	-9.9453	-3.9000	0.0000	0.0000
1	1	3	1	-0.6848	56.7751	-1.2733	-2.2937	-4.0000	0.0000	0.0000
1	1	3	2	-1.4557	78.6279	0.9945	-3.2742	-2.4240	0.0000	0.0000
2	1	3	1	0.6928	78.1546	0.5608	-3.1713	-3.7301	0.0000	0.0000
2	1	3	2	-1.4343	77.0699	0.9875	-3.4139	-1.4053	0.0000	0.0000
1	1	3	3	0.5153	2.1584	0.2000	-6.5859	-3.0000	0.0000	0.0000
2	1	3	3	0.2018	80.0000	0.3778	-2.5000	-2.8750	0.0000	0.0000
3	1	3	1	-1.9875	79.2591	1.0000	-2.4206	-3.9342	0.0000	0.0000
3	1	3	2	-1.1000	78.8002	-1.0000	-2.6282	-4.0000	0.0000	0.0000
3	1	3	3	-1.0000	83.5323	4.3660	-2.6805	-1.2938	0.0000	0.0000
1	3	3	1	3.4682	0.0781	0.9887	-6.1195	-0.5004	0.0000	0.0000
1	3	3	2	1.0000	16.5478	-1.0313	-2.0000	-2.6888	0.0000	0.0000
2	3	3	2	4.0818	-3.2744	-0.9664	-7.1634	-3.0000	0.0000	0.0000
1	3	3	3	4.2014	-10.0642	1.8690	-2.4805	-2.5000	0.0000	0.0000
2	3	3	3	1.0000	-10.0500	-1.0000	-2.1946	-0.5300	0.0000	0.0000
3	3	3	3	1.0000	1.6871	3.0000	-6.2660	-0.5500	0.0000	0.0000
1	! Nr of hydrogen bonds. at1;at2;at3;r(hb);p(hb1);p(hb2);p(hb3)									
3	2	3	1.8130	-3.5409	2.3815	21.9463				

Table S1 Main initial reaction pathways of ethanol in additive-fuel combustion simulations at 3000 K.

Reactions	k' (relative rate constant)
$\text{CH}_3\text{CH}_2\text{OH} \rightarrow \bullet\text{H} + \bullet\text{CH}_2\text{CH}_2\text{OH}$ (reference reaction)	1
$\text{CH}_3\text{CH}_2\text{OH} + \bullet\text{OH} \rightarrow \text{CH}_2\text{O} + \bullet\text{CH}_3 + \text{H}_2\text{O}$	1.62
$\text{CH}_3\text{CH}_2\text{OH} + \bullet\text{OH} \rightarrow \text{H}_2\text{O} + \text{CH}_3\text{CH}_2\text{O}\bullet$	5.68
$\text{CH}_3\text{CH}_2\text{OH} \rightarrow \bullet\text{CH}_3 + \bullet\text{CH}_2\text{OH}$	8.03
$\bullet\text{CH}_2\text{CH}_2\text{OH} \rightarrow \text{C}_2\text{H}_4 + \bullet\text{OH}$	3.20e+03
$\text{CH}_3\text{CH}_2\text{O}\bullet \rightarrow \bullet\text{CH}_3 + \text{CH}_2\text{O}$	6.41e+03

Table S2 Main initial reaction pathways of 2-butanol in additive-fuel combustion simulations at 3000 K.

Reactions	k' (relative rate constant)
	13.23
	1.38e+02
	6.61
	1.16e+02
	4.56e+02
	75.98
	2.99
	2.06e+03
	9.06e+02
	1.87e+03
	7.37e+02

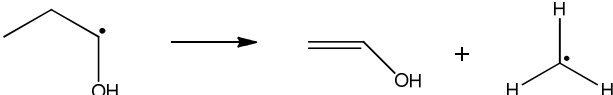
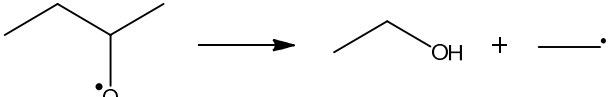
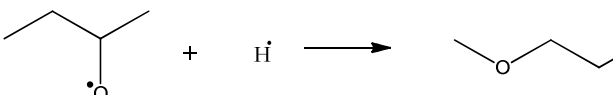
	2.42e+03
	2.89e+04
	5.25e+03

Table S3 Main initial reaction pathways of DMC in additive-fuel combustion simulations at 3000 K.

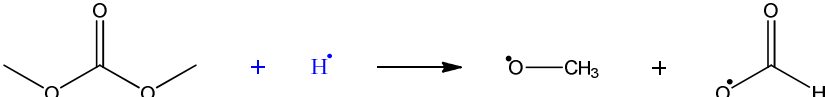
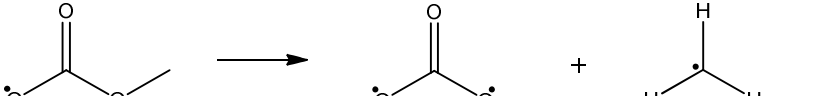
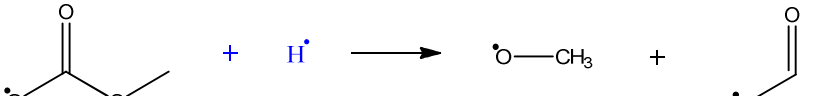
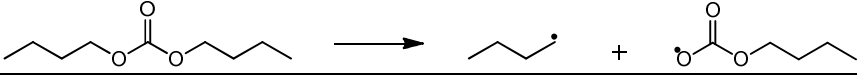
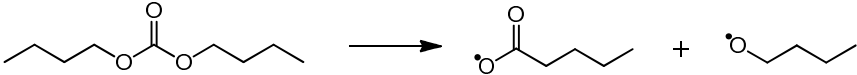
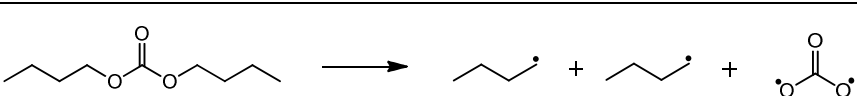
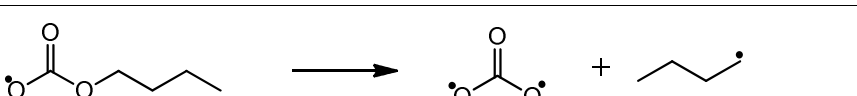
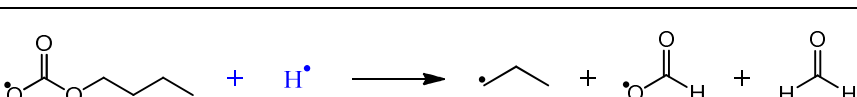
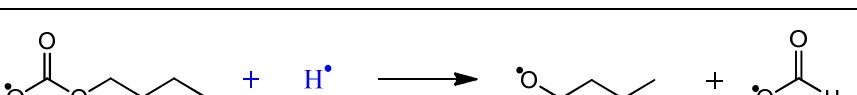
Reactions	k' (relative rate constant)
	4.26e+02
	1.17e+04
	2.60e+03
	1.37e+04

Table S4 Main initial reaction pathways of DBC in additive-fuel combustion simulations at 3000 K.

Reactions	k' (relative rate constant)
	2.00e+04
	8.90e+04
	8.90e+04
	4.06e+03
	8.11e+02
	1.87e+04

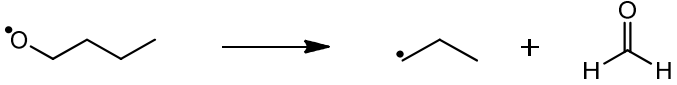
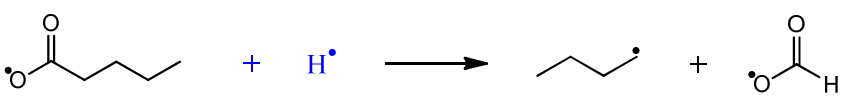
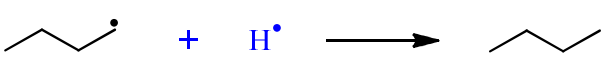
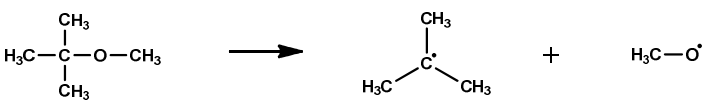
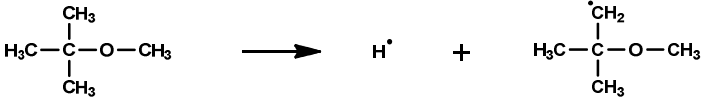
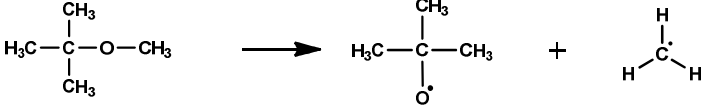
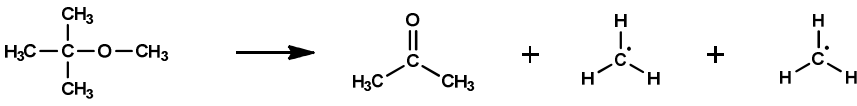
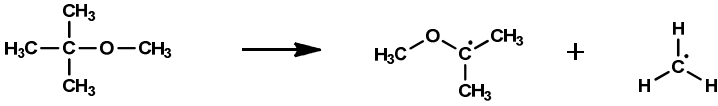
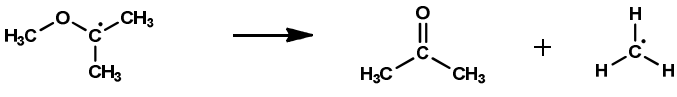
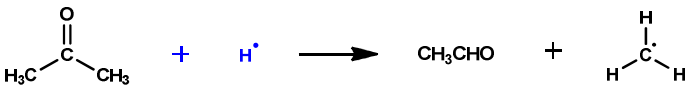
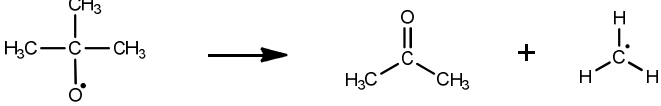
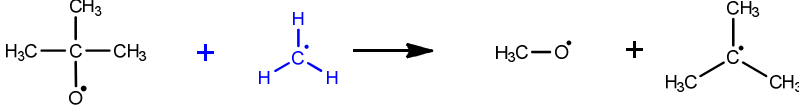
	1.29e+04
	2.30e+03
	4.72e+02
	1.96e+03

Table S5 Main initial reaction pathways of MTBE in additive-fuel combustion simulations at 3000 K.

Reactions	k' (relative rate constant)
	7.59e+02
	26.85
	1.49e+03
	80.31
	1.00e+02
	2.49e+04
	68.43
	2.32e+04
	10.47

$\text{H}_3\text{C}-\overset{\text{CH}_3}{\underset{\text{CH}_3}{\text{C}}}\cdot \longrightarrow \text{H}_2\text{C}^{\cdot}-\overset{\text{CH}_3}{\underset{\text{H}}{\text{C}}}-\text{CH}_3$	1.14e+03
$\text{H}_3\text{C}-\overset{\text{CH}_3}{\underset{\text{CH}_3}{\text{C}}}\cdot \longrightarrow \text{H}_3\text{C}-\overset{\text{CH}_2}{\underset{\text{CH}_3}{\text{C}}}=\text{CH}_2 + \text{H}^{\cdot}$	1.05e+03
$\text{H}_3\text{C}-\overset{\cdot\text{CH}_2}{\underset{\text{CH}_3}{\text{C}}}-\text{O}-\text{CH}_3 \longrightarrow \text{H}_3\text{C}-\overset{\text{CH}_2}{\underset{\text{CH}_3}{\text{C}}}=\text{CH}_2 + \text{H}_3\text{C}-\text{O}^{\cdot}$	1.83e+04
$\text{H}_3\text{C}-\text{O}^{\cdot} \longrightarrow \text{H}_3\text{C}-\text{O}^{\cdot}$	1.65e+02
$\text{H}_3\text{C}-\text{O}^{\cdot} \longrightarrow \text{CH}_2\text{O} + \text{H}^{\cdot}$	1.65e+03

Table S6 Main initial reaction pathways of TTAG in additive-fuel combustion simulations at 3000 K.

Reactions	k' (relative rate constant)
$\begin{array}{c} \text{C}_8\text{H}_{11} \\ \\ \text{O} \\ \\ \text{H}_2\text{C}-\text{C}-\text{H} \\ \quad \\ \text{O}-\text{C}-\text{O} \\ \quad \\ \text{C}_8\text{H}_{11} \end{array} \longrightarrow \begin{array}{c} \text{H}_2\text{C}-\text{O}-\text{C}_8\text{H}_{11} \\ \\ \text{HC}^{\cdot}-\text{CH}_2-\text{O}-\text{C}_8\text{H}_{11} \end{array} + \begin{array}{c} \text{H}_2\text{C}-\text{CH}_3 \\ \\ \text{H}_3\text{C}-\text{C}=\text{O} \end{array} + \begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}^{\cdot}-\text{H} \end{array}$	2.02e+02
$\begin{array}{c} \text{C}_8\text{H}_{11} \\ \\ \text{O} \\ \\ \text{H}_2\text{C}-\text{C}-\text{H} \\ \quad \\ \text{O}-\text{C}-\text{O} \\ \quad \\ \text{C}_8\text{H}_{11} \end{array} \longrightarrow \begin{array}{c} \text{H}_3\text{C}-\text{CH}_2-\text{C}^{\cdot}-\text{CH}_3 \\ \\ \text{H}_3\text{C}-\text{CH}_2 \end{array} + \begin{array}{c} \text{C}_8\text{H}_{11}-\text{O}-\text{C}-\text{CH}_2 \\ \quad \\ \text{H}_2\text{C}-\text{C}-\text{O}^{\cdot} \\ \\ \text{CH}_3 \end{array} + \begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}^{\cdot}-\text{H} \end{array}$	7.39e+02
$\begin{array}{c} \text{C}_8\text{H}_{11} \\ \\ \text{O} \\ \\ \text{H}_2\text{C}-\text{C}-\text{H} \\ \quad \\ \text{O}-\text{C}-\text{O} \\ \quad \\ \text{C}_8\text{H}_{11} \end{array} \longrightarrow \begin{array}{c} \text{H}_2\text{C}=\text{CH}-\text{O}-\text{C}_8\text{H}_{11} \\ \\ \text{HC}=\text{CH}_2 \end{array} + \begin{array}{c} \text{H}_3\text{C}-\text{O}^{\cdot} \\ \\ \text{H}_2\text{C}-\text{C}-\text{CH}_3 \\ \\ \text{CH}_3 \end{array} + \begin{array}{c} \text{H}_2\text{C}-\text{CH}_3 \\ \\ \text{H}_3\text{C}-\text{C}=\text{O} \end{array} + \begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}^{\cdot}-\text{H} \end{array}$	2.02e+02
$\begin{array}{c} \text{C}_8\text{H}_{11} \\ \\ \text{O} \\ \\ \text{H}_2\text{C}-\text{C}-\text{H} \\ \quad \\ \text{O}-\text{C}-\text{O} \\ \quad \\ \text{C}_8\text{H}_{11} \end{array} \longrightarrow \begin{array}{c} \text{H}_2\text{C}-\text{O}-\text{C}_8\text{H}_{11} \\ \\ \text{C}_8\text{H}_{11}-\text{O}-\text{C}-\text{CH}_2 \\ \\ \text{CH}_2 \end{array} + \begin{array}{c} \text{H}_3\text{C}-\text{O}^{\cdot} \\ \\ \text{H}_2\text{C}-\text{C}-\text{CH}_3 \\ \\ \text{CH}_3 \end{array}$	3.29e+02
$\begin{array}{c} \text{C}_8\text{H}_{11} \\ \\ \text{O} \\ \\ \text{H}_2\text{C}-\text{C}-\text{H} \\ \quad \\ \text{O}-\text{C}-\text{O} \\ \quad \\ \text{C}_8\text{H}_{11} \end{array} \longrightarrow \begin{array}{c} \text{H}_2\text{C}-\text{O}-\text{C}_8\text{H}_{11} \\ \\ \text{C}_8\text{H}_{11}-\text{O}-\text{C}-\text{CH}_2 \\ \\ \text{CH}_2 \end{array} + \begin{array}{c} \text{H}_3\text{C}-\text{O}^{\cdot} \\ \\ \text{H}_2\text{C}-\text{C}-\text{CH}_3 \\ \\ \text{CH}_3 \end{array}$	2.76e+02
$\begin{array}{c} \text{C}_8\text{H}_{11} \\ \\ \text{O} \\ \\ \text{H}_2\text{C}-\text{C}-\text{H} \\ \quad \\ \text{O}-\text{C}-\text{O} \\ \quad \\ \text{C}_8\text{H}_{11} \end{array} \longrightarrow \begin{array}{c} \text{H}_2\text{C}=\text{CH}-\text{O}-\text{C}_8\text{H}_{11} \\ \\ \text{HC}=\text{CH}_2 \end{array} + \begin{array}{c} \text{H}_3\text{C}-\text{O}^{\cdot} \\ \\ \text{H}_2\text{C}-\text{C}-\text{CH}_3 \\ \\ \text{CH}_3 \end{array} + \begin{array}{c} \text{H}_3\text{C}-\text{O}^{\cdot} \\ \\ \text{H}_2\text{C}-\text{C}-\text{CH}_3 \\ \\ \text{CH}_3 \end{array}$	7.39e+02

The time evolution of temperature in each system

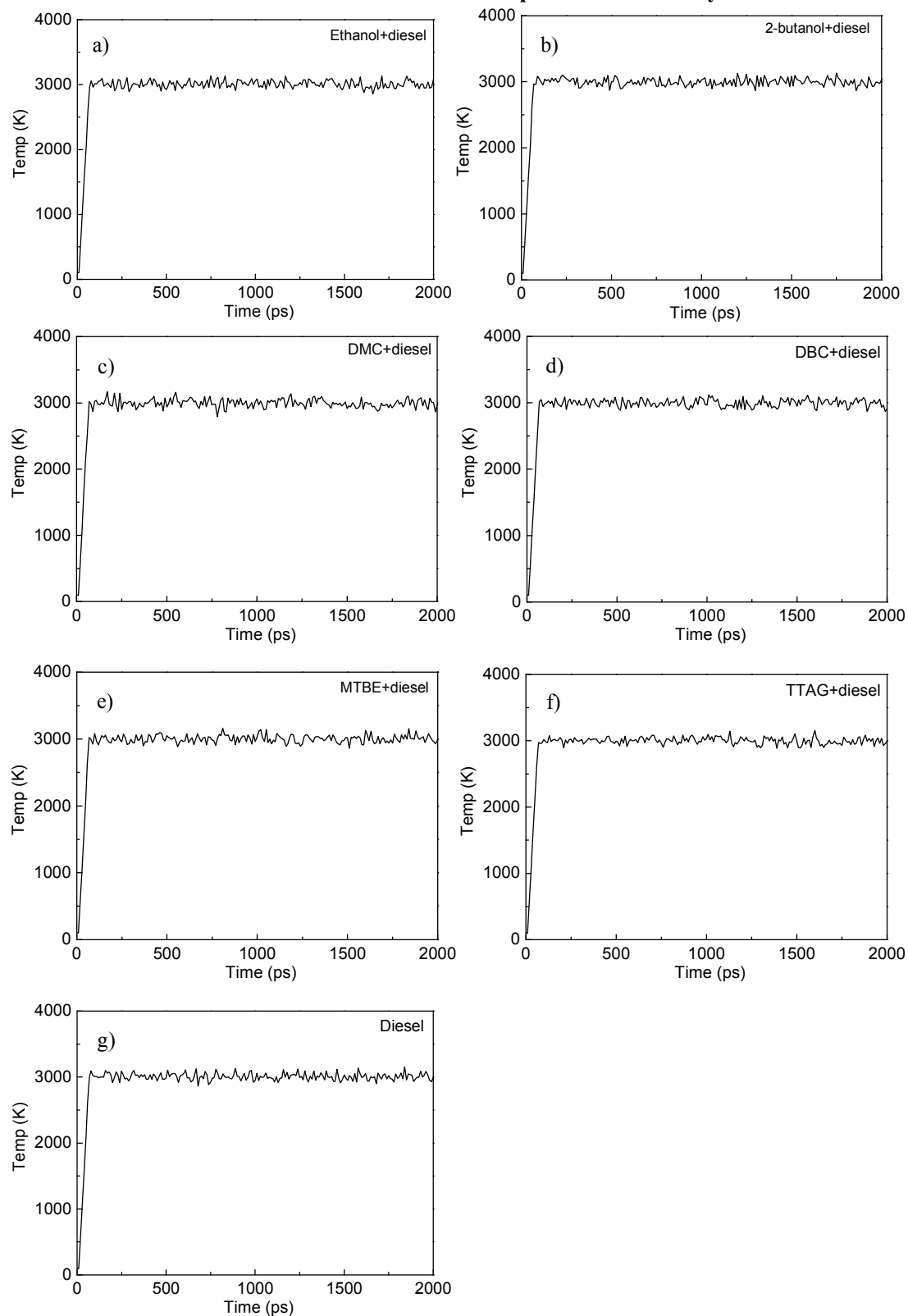
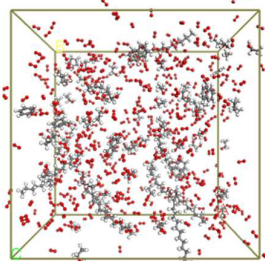
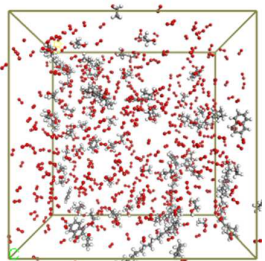
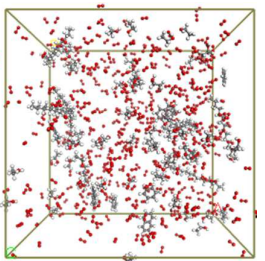
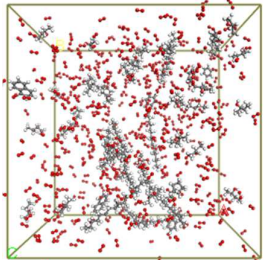
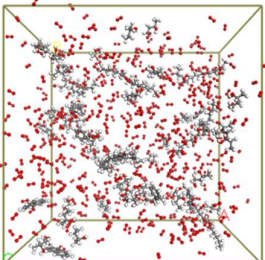
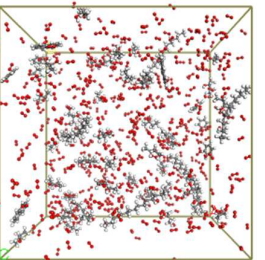
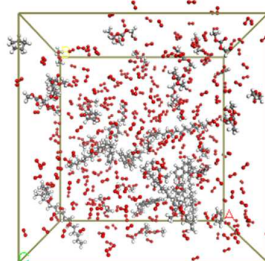
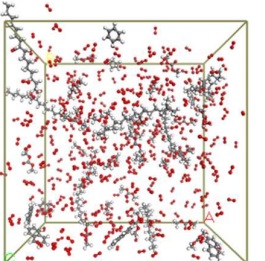
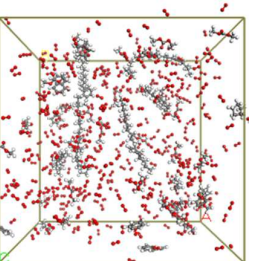
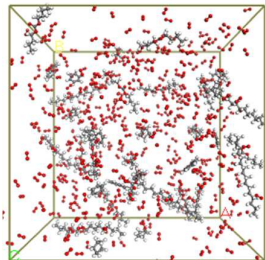
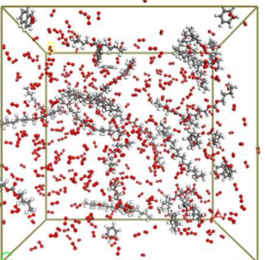
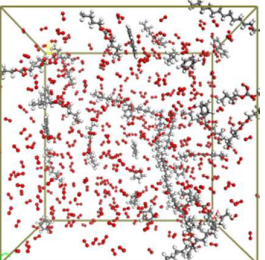
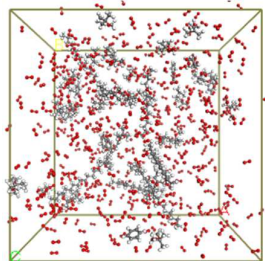
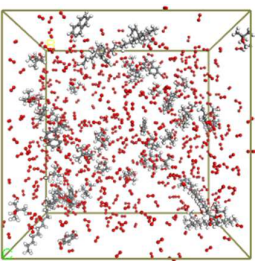
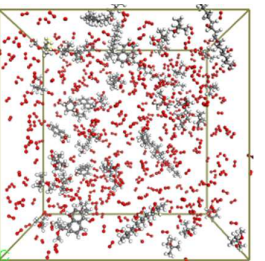


Figure S1 Time evolution of temperature in a). ethanol+diesel, b). 2-butanol+diesel, c). DMC+diesel, d). DBC+diesel, e). MTBE+diesel, f). TTAG+diesel, g). diesel systems.

The construction procedure and random settings

Table S7 The different configurations used in this work:

Systems	Configuration 1	Configuration 2	Configuration 3
Ethanol+diesel+O ₂			
2-butanol+diesel+O ₂			
DMC+diesel+O ₂			
DBC+diesel+O ₂			
MTBE+diesel+O ₂			

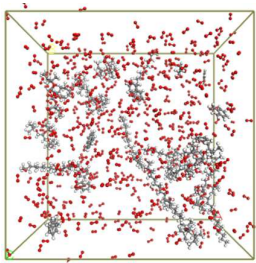
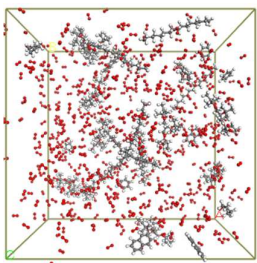
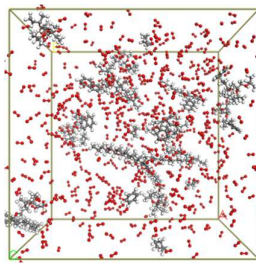
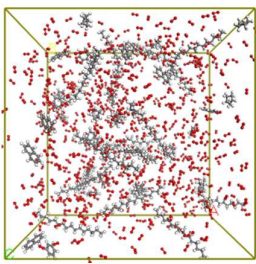
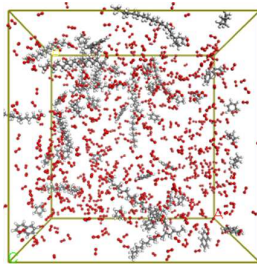
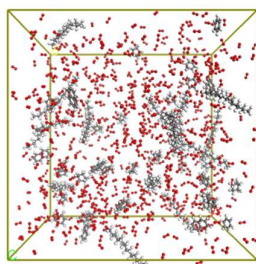
TTAG+diesel+O ₂			
Diesel+O ₂			

Table S8 Randomization settings and explanations

Steps	Explanations
1. Constructing three simulation boxes for each system (shown in Table S7).	The molecules were randomly placed and thus could ensure the different configurations between the simulation boxes.
2. Random seed (including the absolute value and direction) was set and adopted on each atom for the first simulation step.	The initial rate (including the absolute value and direction) is the basis of later integrals. Random initial rate will ensure the randomization of subsequent rates.
3. All the species underwent the low-temperature simulation firstly before the temperature was increased.	Enough low-temperature simulation allowed bond torsion and any other bond variation except for reactions, and thus ensure the diversity of configurations.
4. Heating the systems to target temperature.	Then reactions may occur.