Supporting Information

Micellisation of binary mixtures of surfactants Na-deoxycholate – Na-decyl sulphate and Na-hyodeoxycholate – Na-decyl sulphate in water solutions: rational development of the thermodynamic model for the excess Gibbs energy (G^E)

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Table S1. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture $\alpha_B = 0.1$; $T = 295.15$ K; $p = 1.00 \times 10^5$ Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.25	1.773
0.50	1.794
1.00	1.784
2.00	1.743
2.50	1.740
3.00	1.520
3.50	1.270
4.00	1.050
5.00	1.024
6.00	1.025
7.00	1.036
8.00	1.030
9.00	1.025
10.00	1.027

Standard uncertainties u, are: $u(c_t) = 0.01$ mmol kg^{-1}, $u(I_1/I_3) = 0.01$, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S2. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture $\alpha_B = 0.2$; $T = 295.15$ K; $p = 1.00 \times 10^5$ Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.25	1.260
0.50	1.264
1.00	1.243
2.00	1.248
2.50	1.190
3.00	1.077
3.50	1.007
4.00	0.985
5.00	0.964
6.00	0.958
7.00	0.983
8.00	0.957
9.00	0.976
10.00	0.964

Standard uncertainties u, are: $u(c_t) = 0.01$ mmol kg^{-1}, $u(I_1/I_3) = 0.01$, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S3. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture α_B = 0.3; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.25	1.788
0.50	1.779
1.00	1.773
2.00	1.750
2.50	1.470
3.00	1.196
3.50	1.026
4.00	1.011
5.00	1.005
6.00	1.004
7.00	1.022
8.00	1.013
9.00	1.018
10.00	1.020

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa, $u(\alpha)$ = 0.01.

*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S4. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture α_B = 0.4; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.25	1.806
0.50	1.783
1.00	1.800
2.00	1.743
2.50	1.401
3.00	1.139
3.50	1.035
4.00	0.992
5.00	0.996
6.00	0.985
7.00	1.006
8.00	0.980
9.00	0.983
10.00	0.980

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa, $u(\alpha)$ = 0.01.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S5. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture $\alpha_B = 0.5$; $T = 295.15$ K; $p = 1.00 \times 10^5$ Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.25	1.765
0.50	1.770
1.00	1.796
2.00	1.691
2.50	1.369
3.00	1.040
3.50	0.977
4.00	0.963
5.00	0.981
6.00	0.968
7.00	0.969
8.00	0.979
9.00	0.956
10.00	0.970

Standard uncertainties u, are: $u(c_t) = 0.01$ mmol kg^{-1}, $u(I_1/I_3) = 0.01$, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S6. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture α_B = 0.6; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.25	1.796
0.50	1.798
1.00	1.778
2.00	1.612
2.50	1.246
3.00	0.930
3.50	0.918
4.00	0.871
5.00	0.910
6.00	0.904
7.00	0.885
8.00	0.874
9.00	0.904
10.00	0.899

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa, $u(\alpha)$ = 0.01.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S7. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture $\alpha_B = 0.7$; $T = 295.15$ K; $p = 1.00 \times 10^5$ Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.25	1.746
0.50	1.749
1.00	1.568
2.00	1.159
2.50	1.026
3.00	0.832
3.50	0.816
4.00	0.833
5.00	0.813
6.00	0.825
7.00	0.808
8.00	0.800
9.00	0.820
10.00	0.798

Standard uncertainties u, are: $u(c_t) = 0.01$ mmol kg^{-1}, $u(I_1/I_3) = 0.01$, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S8. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture α_B = 0.8; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.25	1.729
0.50	1.636
1.00	1.720
2.00	1.342
2.50	1.024
3.00	0.809
3.50	0.766
4.00	0.732
5.00	0.731
6.00	0.730
7.00	0.741
8.00	0.739
9.00	0.749
10.00	0.742

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa, $u(\alpha)$ = 0.01.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S9. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture α_B = 0.9; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.25	1.781
0.50	1.777
1.00	1.686
2.00	1.363
2.50	0.964
3.00	0.756
3.50	0.722
4.00	0.727
5.00	0.709
6.00	0.721
7.00	0.734
8.00	0.711
9.00	0.732
10.00	0.739

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa, $u(\alpha)$ = 0.01.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S10. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.1; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.025	1.795
0.05	1.761
0.10	1.725
0.20	1.609
0.50	1.412
1.00	1.321
2.00	1.251
4.00	1.214
5.00	1.110
6.00	1.087
7.00	1.071
8.00	1.081
9.00	1.090
10.00	1.076
12.00	1.087

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa, $u(\alpha)$ = 0.01.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S11. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.2; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.005	1.799
0.01	1.775
0.025	1.789
0.05	1.698
0.10	1.644
0.20	1.492
0.50	1.364
1.00	1.244
2.00	1.187
4.00	1.111
5.00	1.098
6.00	1.043
8.00	1.046
10.00	1.044
12.00	1.045

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa, $u(\alpha)$ = 0.01.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S12. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.3; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.005	1.793
0.025	1.726
0.05	1.652
0.10	1.541
0.20	1.508
1.00	1.208
2.00	1.163
4.00	1.189
6.00	1.076
8.00	1.049
10.00	1.050
12.00	1.044

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa, $u(\alpha)$ = 0.01.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S13. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.4; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.005	1.740
0.01	1.676
0.025	1.721
0.05	1.565
0.10	1.540
0.20	1.433
1.00	1.177
2.00	1.138
4.00	1.101
6.00	1.046
8.00	1.048
10.00	1.047

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa, $u(\alpha)$ = 0.01.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S14. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.5; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.005	1.745
0.01	1.778
0.025	1.721
0.05	1.562
0.10	1.507
0.20	1.353
1.00	1.185
2.00	1.189
4.00	1.068
6.00	1.002
8.00	1.034
10.00	1.007

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa, $u(\alpha)$ = 0.01.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S15. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.6; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.005	1.779
0.01	1.725
0.025	1.663
0.05	1.555
0.10	1.443
0.20	1.388
1.00	1.212
2.00	1.147
4.00	1.139
6.00	1.006
8.00	0.999
10.00	1.005

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa, $u(\alpha)$ = 0.01.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S16. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.7; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.005	1.768
0.01	1.754
0.025	1.699
0.05	1.499
0.10	1.498
0.20	1.382
1.00	1.259
2.00	1.235
4.00	1.189
6.00	1.034
8.00	1.036
10.00	1.031

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa, $u(\alpha)$ = 0.01.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S17. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.8; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.01	1.724
0.025	1.613
0.05	1.557
0.10	1.450
0.20	1.388
1.00	1.254
2.00	1.196
4.00	1.137
6.00	1.070
8.00	0.985
10.00	0.974
12.00	0.974

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa, $u(\alpha)$ = 0.01.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S18. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.9; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

c_t /mmol kg^{-1}	I_1/I_3
0.01	1.621
0.025	1.613
0.05	1.568
0.10	1.419
0.20	1.316
1.00	1.226
2.00	1.123
4.00	1.068
6.00	0.990
7.00	0.991
8.00	0.980
10.00	0.990

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa, $u(\alpha)$ = 0.01.
*In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S19. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of sodium deoxycholate (NaDC) in water (NaCl 0.3 mol kg^{-1}) at T = 295.15 K; p = 1.00 × 10^5 Pa*

c_t /mmol kg^{-1}	I_1/I_3
0.25	1.268
0.50	1.279
1.00	1.282
2.00	1.170
2.50	0.933
3.00	0.698
4.00	0.696
5.00	0.720
6.00	0.700
7.00	0.684
8.00	0.707
9.00	0.695
10.00	0.702

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S20. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of sodium hyodeoxycholate (NaHDC) in water (NaCl 0.3 mol kg^{-1}) at T = 295.15 K; p = 1.00 × 10^5 Pa*

c_t /mmol kg^{-1}	I_1/I_3
0.50	1.218
1.00	1.214
2.00	1.213
3.00	1.125
4.00	1.082
5.00	1.049
6.00	1.001
7.00	0.970
8.00	0.953
9.00	0.974
10.00	0.974
11.00	0.974

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S21. Experimental values of the I_1/I_3 ratio of pyrene emission spectrum as a function of logarithm of the total concentration log c_t of sodium decyl sulfate (NaDS) in water (NaCl 0.3 mol kg^{-1}) at T = 295.15 K; p = 1.00 × 10^5 Pa*

c_t /mmol kg^{-1}	I_1/I_3
0.20	1.773
0.50	1.759
1.00	1.744
2.00	1.760
4.00	1.761
5.00	1.752
6.00	1.643
7.00	1.318
8.00	1.194
9.00	1.125
10.00	1.113
11.00	1.118
12.00	1.111

Standard uncertainties u, are: $u(c_t)$ = 0.01 mmol kg^{-1}, $u(I_1/I_3)$ = 0.01, $u(T)$ = 0.1 K, $u(p)$ = 0.002 MPa. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S22. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture α_B = 0.1; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
0.301	67
0.322	66.5
0.342	66.0
0.362	65.5
0.38	64.5
0.397	63.5
0.414	62.1
0.431	61.5
0.447	61.0
0.462	60.5
0.477	60.1
0.491	59.5
0.505	59.0
0.518	58.1
0.531	57.5
0.544	57.0
0.556	56.5
0.568	56.0
0.579	55.5
0.591	55.0

0.602	54.5
0.612	53.9
0.623	53.1
0.633	53.0
0.643	53.0
0.653	52.9
0.662	52.7
0.672	52.6
0.681	52.4
0.690	52.4
0.698	52.3
0.707	52.1
0.716	52.0
0.724	51.9
0.732	52.0
0.748	52.0
0.755	51.8
0.763	51.8
0.770	51.8
0.778	51.8

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S23. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture α_B = 0.2; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
0.322	66.5
0.342	65.5
0.362	64.5
0.38	63.9
0.397	63.2
0.414	62.5
0.431	61.5
0.447	60.3
0.462	59.2
0.477	58.5
0.491	58.0
0.505	57.0
0.518	56.5
0.531	55.7
0.544	54.9
0.556	54.2
0.568	53.5
0.579	53.1
0.591	53.0
0.602	52.9

0.612	53.0
0.623	52.8
0.633	52.9
0.643	52.9
0.653	52.8
0.662	52.8
0.672	52.7
0.681	52.7
0.690	52.8
0.698	52.7
0.707	52.6
0.716	52.7
0.724	52.6
0.732	52.7

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S24. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture α_B = 0.3; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
0.301	67.0
0.322	66.5
0.342	65.3
0.362	64.2
0.38	63.3
0.397	62.5
0.414	61.0
0.431	60.2
0.447	59.0
0.462	58.0
0.477	57.1
0.491	56.0
0.505	54.9
0.518	54.5
0.531	53.3
0.544	53.2
0.556	53.0
0.568	53.1
0.579	53.0
0.591	53.0

0.602	52.9
0.612	52.9
0.623	52.8
0.633	52.9
0.643	52.9
0.653	52.8
0.662	52.7
0.672	52.7
0.681	52.7
0.690	52.8
0.698	52.8
0.707	52.7
0.716	52.6
0.724	52.6

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S25. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture α_B = 0.4; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
0.301	65.8
0.322	64.7
0.342	63.2
0.362	61.5
0.38	60.9
0.397	60
0.414	58.8
0.431	58.2
0.447	57.3
0.462	56.1
0.477	55
0.491	53.8
0.505	53
0.518	53
0.531	53
0.544	52.8
0.556	52.8
0.568	52.7
0.579	52.7
0.591	52.6

0.602	52.6
0.612	52.7
0.623	52.6
0.633	52.5
0.643	52.5
0.653	52.4
0.662	52.3
0.672	52.3
0.681	52.2
0.690	52.2

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S26. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture α_B = 0.5; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
0.301	65.5
0.322	64.7
0.342	63.0
0.362	61.8
0.38	60.9
0.397	59.7
0.414	58.5
0.431	58.0
0.447	56.9
0.462	55.5
0.477	54.5
0.491	53.5
0.505	53.1
0.518	53.0
0.531	53.0
0.544	52.8
0.556	52.8
0.568	52.7
0.579	52.7
0.591	52.6

0.602	52.6
0.612	52.6
0.623	52.6
0.633	52.5
0.643	52.4
0.653	52.3
0.662	52.3
0.672	52.3
0.681	52.2
0.69	52.2

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S27. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture α_B = 0.6; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
0.301	67.0
0.322	66.0
0.342	64.2
0.362	62.4
0.38	61.0
0.397	60.2
0.414	59.0
0.431	58.0
0.447	57.0
0.462	56.0
0.477	55.0
0.491	54.5
0.505	54.1
0.518	54.0
0.531	53.5
0.544	53.5
0.556	53.5
0.568	53.3
0.579	53.3
0.591	53.0

0.602	53.1
0.612	53.0
0.623	52.9
0.633	52.9
0.643	52.8
0.653	52.9
0.662	52.7
0.672	52.6
0.681	52.4
0.690	52.4

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S28. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture α_B = 0.7; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
0.301	65.5
0.322	64.5
0.342	63.2
0.362	62.0
0.38	61.0
0.397	59.2
0.414	58
0.431	56.8
0.447	55.8
0.462	54.5
0.477	53.5
0.491	53.0
0.505	53.0
0.518	52.8
0.531	52.7
0.544	52.8
0.556	52.7
0.568	52.7
0.579	52.6
0.591	52.6

0.602	52.6
0.612	52.6
0.623	52.5
0.633	52.5
0.643	52.4
0.653	52.3
0.662	52.3
0.672	52.2
0.681	52.1
0.690	52.1

Standard uncertainties u, are: u(log ct) = 0.005, u(γ) = 0.5 mN m^{-1}, u(T) = 0.1 K, u(p) = 0.002 MPa, u(α) = 0.01. *In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S29. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture α_B = 0.8; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
0.301	66
0.322	64.5
0.342	63
0.362	62
0.38	61
0.397	59.5
0.414	57.8
0.431	56.5
0.447	55.6
0.462	54.5
0.477	54
0.491	53.2
0.505	53
0.518	52.8
0.531	52.7
0.544	52.8
0.556	52.7
0.568	52.7
0.579	52.6
0.591	52.6

0.602	52.6
0.612	52.6
0.623	52.5
0.633	52.5
0.643	52.4
0.653	52.3
0.662	52.3
0.672	52.2
0.681	52.1
0.69	52.1

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S30. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium deoxycholate (NaDC) – sodium decyl sulfate (NaDS); NaDC molar ratio in the binary mixture α_B = 0.9; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
0.301	66.0
0.322	64.5
0.342	63.0
0.362	62.0
0.38	61.0
0.397	59.8
0.414	59.2
0.431	58.0
0.447	57.0
0.462	56.0
0.477	55
0.491	54.1
0.505	53.5
0.518	53
0.531	52.9
0.544	53.0
0.556	53.0
0.568	52.9
0.579	52.9
0.591	52.8

0.602	52.7
0.612	52.7
0.623	52.6
0.633	52.5
0.643	52.3
0.653	52.3
0.662	52.3
0.672	52.2
0.681	52.1
0.69	52.1

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S31. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.1; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
-0.522	69.0
-0.397	67.0
-0.301	65.2
-0.22	62.8
-0.154	61.0
-0.09	60.2
-0.04	58.5
0	57.1
0.04	56.5
0.079	56.0
0.113	54.5
0.146	54
0.176	53.5
0.204	52.8
0.230	52.9
0.255	52.8
0.278	52.8
0.301	52.7
0.322	52.7
0.342	52.6

0.361	52.6
0.380	52.6
0.397	52.6

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S32. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.2; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
-0.397	65.5
-0.301	63.5
-0.22	61.8
-0.154	60.2
-0.09	59.1
-0.04	57.5
0	56.6
0.04	55.5
0.079	54.5
0.113	54.0
0.146	53.1
0.176	53.1
0.204	53.0
0.23	53.0
0.255	53.0
0.278	52.8
0.301	52.8
0.322	52.8
0.342	52.7
0.361	52.7

0.38	52.6
0.397	52.6
0.414	52.6

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S33. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.3; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
-0.397	64.4
-0.301	62.8
-0.22	61.0
-0.154	59.3
-0.09	58.0
-0.04	57.0
0	55.8
0.04	55.1
0.079	54.0
0.113	53.0
0.146	53.0
0.176	53.0
0.204	53.0
0.23	53.0
0.255	53.0
0.278	52.7
0.301	52.8
0.322	52.8
0.342	52.7
0.361	52.6

0.38	52.6
0.397	52.6
0.414	52.6

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S34. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.4; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
-0.397	65.7
-0.301	63.9
-0.22	61.4
-0.154	60.0
-0.09	58.0
-0.04	57.0
0	55.8
0.04	55.1
0.079	54
0.113	53.5
0.146	53.0
0.176	53.0
0.204	53.0
0.23	52.9
0.255	52.8
0.278	52.8
0.301	52.8
0.322	52.8
0.342	52.7
0.361	52.6

0.38	52.6
0.397	52.6
0.414	52.6

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S35. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.5; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
-0.397	66.0
-0.301	64.0
-0.22	61.5
-0.154	60.0
-0.09	59.0
-0.04	57.5
0	56.5
0.04	55.5
0.079	54.4
0.113	53.8
0.146	53.2
0.176	53.0
0.204	53.0
0.23	53.0
0.255	52.7
0.278	52.7
0.301	52.8
0.322	52.7
0.342	52.7
0.361	52.6

0.38	52.6
0.397	52.6
0.414	52.6

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S36. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.6; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
-0.397	68.0
-0.301	66.0
-0.22	64.5
-0.154	61.5
-0.09	60.5
-0.04	59.5
0	57.5
0.04	56.8
0.079	55.5
0.113	54.5
0.146	54.0
0.176	53.7
0.204	53.2
0.23	53.1
0.255	52.9
0.278	52.9
0.301	52.8
0.322	52.7
0.342	52.7
0.361	52.6

0.38	52.6
0.397	52.6
0.414	52.6

Standard uncertainties u, are: $u(\log c_i) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S37. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.7; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
-0.397	68.0
-0.301	66.0
-0.22	64.0
-0.154	62.0
-0.09	60.5
-0.04	59
0	57.2
0.04	56.5
0.079	55.5
0.113	54.5
0.146	53.6
0.176	53.2
0.204	53.2
0.23	53.1
0.255	52.9
0.278	52.9
0.301	52.8
0.322	52.7
0.342	52.7
0.361	52.6

0.38	52.6
0.397	52.6
0.414	52.6

Standard uncertainties u, are: $u(\log c_i) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S38. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.8; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
-0.301	68
-0.22	66
-0.154	63.5
-0.09	62.2
-0.04	61
0	59.8
0.04	58.5
0.079	57.3
0.113	56.2
0.146	55.3
0.176	54.6
0.204	54
0.23	53
0.255	52.9
0.278	52.9
0.301	52.8
0.322	52.7
0.342	52.7
0.361	52.6
0.38	52.6

0.397	52.6
0.414	52.6
0.431	52.6
0.447	52.6
0.467	52.5
0.477	52.6

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S39. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of the surfactants binary mixture {sodium hyodeoxycholate (NaHDC) – sodium decyl sulfate (NaDS); NaHDC molar ratio in the binary mixture α_B = 0.9; T = 295.15 K; p = 1.00 × 10^5 Pa} in water (NaCl 0.3 mol kg^{-1})*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
-0.301	68
-0.22	66
-0.154	63.8
-0.09	62.3
-0.04	61.5
0	59.5
0.04	58.2
0.079	57
0.113	56
0.146	55
0.176	54.5
0.204	53.4
0.23	53.2
0.255	52.9
0.278	52.8
0.301	52.8
0.322	52.7
0.342	52.7
0.361	52.6
0.38	52.6

0.397	52.6
0.414	52.6
0.431	52.6
0.447	52.6
0.467	52.6
0.477	52.6

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa, $u(\alpha) = 0.01$. ∗In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S40. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of sodium hyodeoxycholate (NaHDC) in water (NaCl 0.3 mol kg^{-1}) at T = 295.15 K; p = 1.00 × 10^5 Pa*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
0.74	69.1
0.748	67.5
0.755	66.5
0.763	64.5
0.77	63.2
0.778	62
0.785	60.3
0.792	59.2
0.799	57.8
0.804	56.6
0.812	55.4
0.819	54.2
0.826	53.6
0.832	53
0.838	52.9
0.845	52.8
0.851	52.8
0.857	52.7
0.863	52.7
0.869	52.6
0.875	52.6

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa. ∗ In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S41. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of sodium deoxycholate (NaDC) in water (NaCl 0.3 mol kg^{-1}) at T = 295.15 K; p = 1.00 × 10^5 Pa*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
0.301	66.0
0.322	64.5
0.342	63.0
0.362	61.5
0.38	59.7
0.397	58.8
0.414	57.5
0.431	56.0
0.447	55.1
0.462	54.0
0.477	53.0
0.491	53.0
0.505	53.0
0.518	53.0
0.531	53.0
0.544	52.9
0.556	52.8
0.568	52.9
0.579	52.9
0.591	52.8
0.602	52.7

0.612	52.7
0.623	52.6
0.633	52.5
0.643	52.3
0.653	52.2
0.662	52.2
0.672	52.2
0.681	52.1
0.690	52.1

Standard uncertainties u, are: $u(\log c_t) = 0.005$, $u(\gamma) = 0.5$ mN m^{-1}, $u(T) = 0.1$ K, $u(p) = 0.002$ MPa. ∗ In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.

Table S42. Experimental values of surface tension γ as a function of logarithm of the total concentration log c_t of sodium decyl sulfate (NaDS) in water (NaCl 0.3 mol kg^{-1}) at T = 295.15 K; p = 1.00 × 10^5 Pa*

log c_t / mmol kg^{-1}	γ/ mN m^{-1}
0.812	65.5
0.819	64
0.826	62.5
0.832	61.5
0.838	60
0.845	59.5
0.851	58
0.857	56.5
0.863	56
0.869	54.2
0.875	53.8
0.880	53
0.892	53
0.897	53
0.903	53
0.908	52.9
0.913	52.8
0.919	52.9
0.924	52.9
0.929	52.8

Standard uncertainties u, are: u(log c_t) = 0.005, u(γ) = 0.5 mN m^{-1}, u(T) = 0.1 K, u(p) = 0.002 MPa. ∗ In each measuring unit of a molality (mol per kilogram of solvent) the mass of a solvent refers to the pure water.